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Exhaustive symmetry mode searches: phase transitions in pyrochlore Bi₂Sn₂O₇

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Bi2Sn2O7 pyrochlore has been shown to undergo phase transitions from a cubic parent γ phase to β and a phases on cooling. Until recently, the only reliable structural model for the low temperature phases has been an a phase model reported by Evans et al. [1]. This was derived by testing a limited number of candidate structures based on experimental information from the literature and assuming group-subgroup relationships.

We have now developed an exhaustive approach for searching all possible subgroups between a parent structure (here the γ phase) and a child structure with sufficient degrees of freedom to explain all the peaks seen in powder diffraction data from any form of the material (here a P1 structure with a cell parameters of root2ap x root2ap x 2ap, where ap is the cell parameter of the parent cubic γ phase). Our method uses ISODISTORT [2] to produce a 547-membered subgroup tree and Topas Academic to automatically test each of the candidate structures. Using this approach we have determined the first definitive model for the β -phase (Aba2, a = 7.571833(8)Å, b = 21.41262(2)Å, c = 15.132459(14)Å) and a much simpler model than previously reported for the a-phase (Cc, a = 13.15493(6)Å, b = 7.54118(4)Å, c = 15.07672(7)Å, β = 125.0120(3)°) [3].

By using a symmetry mode basis we can describe the principal distortions in each phase in terms of coupled rotations of the cristobalite-like Bi2O' framework that allow Bi to adopt lone-pair cation preferred low-symmetry sites. In the β -phase we find that Bi is displaced towards an edge of its O6 hexagonal coordination environment whereas in the a-phase the displacement is towards an apex.

The use of symmetry modes enables an exhaustive approach to symmetry determination in a modest time period (~48 hours on an i7, 3.4GHz desktop PC in this work) that is applicable across a range of structure determination problems where only powder diffraction data is available.

[1] Evans, I. R., Howard, J. A. K. and Evans, J. S. O. J. Mater. Chem. 2003, 13, 2098

[2] Campbell, B. J.; Stokes, H. T.; Tanner, D. E.; Hatch, D. M. Journal of Applied Crystallography 2006, 39, 607

[3] Lewis, J.W., Payne J.L., Evans I.R., Stokes H.T., Campbell B.J., and Evans J.S.O.. Journal of the American Chemical Society 2016, 138, 25, 8031-8042.

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