

Acta Crystallographica Section E

Structure Reports
Online

ISSN 1600-5368

trans-Bis[2-(piperazin-1-yl)ethanamine]-bis(saccharinato)cobalt(II)

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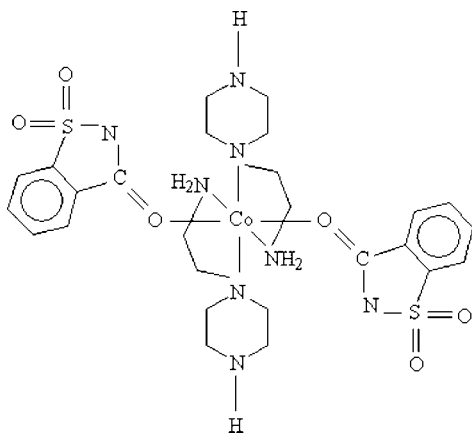
Received 19 November 2007; accepted 23 November 2007

Key indicators: single-crystal X-ray study; $T = 293$ K; mean $\sigma(\text{C}-\text{C}) = 0.004$ Å; R factor = 0.033; wR factor = 0.094; data-to-parameter ratio = 14.0.

In the centrosymmetric title complex, $[\text{Co}(\text{C}_7\text{H}_4\text{NO}_3\text{S})_2(\text{C}_6\text{H}_{15}\text{N}_3)_2]$, the Co^{II} ion is coordinated by two saccharinate (sac) anions and two neutral 2-piperazin-1-ylethanamine (ppzea) ligands, showing a distorted octahedral coordination. Sac is O-bonded *via* the carbonyl group, while ppzea acts as an N,N' -bidentate chelating ligand. The molecules are connected by $\text{N}-\text{H}\cdots\text{N}$ and $\text{N}-\text{H}\cdots\text{O}$ hydrogen bonds, forming a linear chain running parallel to the crystallographic a axis. The compound is isostructural with the reported Ni, Zn, and Cd analogues.

Related literature

For the structures of the analogous Ni, Zn, and Cd complexes, see: Guney *et al.* (2005); Yilmaz *et al.* (2005). For a review of saccharinate complexes, see: Baran & Yilmaz (2006).


Experimental
Crystal data

$[\text{Co}(\text{C}_7\text{H}_4\text{NO}_3\text{S})_2(\text{C}_6\text{H}_{15}\text{N}_3)_2]$
 $M_r = 681.69$
 Triclinic, $P\bar{1}$
 $a = 8.4294$ (7) Å
 $b = 9.3742$ (10) Å
 $c = 11.5618$ (10) Å
 $\alpha = 93.651$ (8)°
 $\beta = 110.473$ (6)°

$\gamma = 116.486$ (13)°
 $V = 739.31$ (17) Å³
 $Z = 1$
 Mo $K\alpha$ radiation
 $\mu = 0.78$ mm⁻¹
 $T = 293$ (2) K
 $0.58 \times 0.41 \times 0.25$ mm

Data collection

Stoe IPDS 2 diffractometer
 Absorption correction: integration
 (*X-RED*; Stoe & Cie, 2002)
 $T_{\text{min}} = 0.721$, $T_{\text{max}} = 0.865$

9652 measured reflections
 2913 independent reflections
 2635 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.103$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.033$
 $wR(F^2) = 0.094$
 $S = 1.07$
 2913 reflections
 208 parameters
 1 restraint

H atoms treated by a mixture of independent and constrained refinement
 $\Delta\rho_{\text{max}} = 0.47$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.64$ e Å⁻³

Table 1

Selected geometric parameters (Å, °).

Co1—O1	2.0897 (13)	Co1—N3	2.3589 (16)
Co1—N2	2.1101 (16)		
O1—Co1—N2 ⁱ	91.53 (6)	N2—Co1—N3	80.62 (6)
O1—Co1—N2	88.47 (6)	O1—Co1—N3 ⁱ	92.94 (6)
O1—Co1—N3	87.06 (6)	N2—Co1—N3 ⁱ	99.38 (6)

 Symmetry code: (i) $-x + 1, -y + 1, -z + 1$.

Table 2

Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
N2—H2B \cdots N4 ⁱⁱ	0.85 (3)	2.38 (3)	3.172 (3)	154 (2)
N2—H2A \cdots N1 ⁱ	0.91 (3)	2.25 (3)	2.982 (2)	137 (2)
N4—H4A \cdots O3 ⁱⁱⁱ	0.838 (17)	2.221 (18)	3.054 (3)	173 (3)

 Symmetry codes: (i) $-x + 1, -y + 1, -z + 1$; (ii) $x - 1, y, z$; (iii) $x + 1, y + 1, z$.

Data collection: *X-AREA* (Stoe & Cie, 2002); cell refinement: *X-AREA*; data reduction: *X-RED* (Stoe & Cie, 2002); program(s) used to solve structure: *SHELXS97* (Sheldrick, 1997); program(s) used to refine structure: *SHELXL97* (Sheldrick, 1997); molecular graphics: *ORTEP-3* (Farrugia, 1997); software used to prepare material for publication: *WinGX* (Farrugia, 1999).

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: SQ2006).

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supporting information

Acta Cryst. (2008). E64, m59–m60 [https://doi.org/10.1107/S1600536807062745]

trans*-Bis[2-(piperazin-1-yl)ethanamine]bis(saccharinato)cobalt(II)*Veysel T. Yilmaz, Serkan Guney and Canan Kazak****S1. Comment**

The saccharinate (sac) anion is formed by the deprotonation of saccharin and coordinates to various metal ions rather easily (Baran & Yilmaz 2006). In the course of the synthesis and structural characterization of mixed ligand–metal complexes of sac, recently we reported nickel(II), zinc(II) and cadmium(II) complexes of sac with 2-piperazin-1-ylethanamine (ppzea) (Guney *et al.*, 2005; Yilmaz *et al.*, 2005). In this paper, the crystal and molecular structure of the isomorphous sac complex of cobalt(II) with ppzea (I) is reported.

The title complex (I) is isostructural with the nickel(II), zinc(II) and cadmium(II) complexes of the same ligands (Guney *et al.*, 2005; Yilmaz *et al.*, 2005) and shows similar structural characteristics. In these isostructural complexes, the M^{II} ions show an elongated octahedral geometry, possibly due to a poor overlap of the sp^3 lone pair on the N atom of ppz with the valence orbitals of the metal ions. As shown in Fig. 1, (I) is a mononuclear Co^{II} complex, in which the Co^{II} ion lies on a centre of inversion and also exhibits an elongated distorted octahedral geometry with two neutral bidentate (N,N') ppzea ligands and two anionic sac ligands. In spite of the common N-coordination mode, sac coordinates to Co^{II} through the carbonyl O atom. The puckering parameters of the ppz ring system in (I) are $q = 0.538(2)\text{\AA}$ and $\Theta = 5.4(2)^\circ$, suggesting that the ppz rings exhibit a typical (*e.g.*, cyclohexane-like) chair conformation.

The amine hydrogen atoms of ppzea form intramolecular hydrogen bonds with the negatively charged N atom of sac. The individual molecules are linked by N—H \cdots N and N—H \cdots O hydrogen bonds, involving the amine H atoms of ppzea and the ring N atom of ppzea and the sulfonyl O atoms of sac, forming a linear chain running parallel to the crystallographic *a* axis.

S2. Experimental

A 20 ml ethanol solution containing ppzea (0.26 g, 2 mmol) and sacH (0.36 g, 2 mmol) was mixed with a 20 ml ethanol solution of $Co(OAc)_2 \cdot 4H_2O$ (0.25 g, 1 mmol). The reaction solution was stirred for 1 h at room temperature. Red-brown prisms were obtained after 5 days by slow evaporation of the solution at room temperature.

S3. Refinement

All N-bonded H atoms were refined freely, while C-bonded H atoms were placed in idealized locations (C—H = 0.95 Å) and included as riding atoms with $U_{iso}(H) = 1.2 * U_{eq}(C)$. The instruction *DFIX* was applied to the N4—H4A bond to increase its length to a reasonable value.

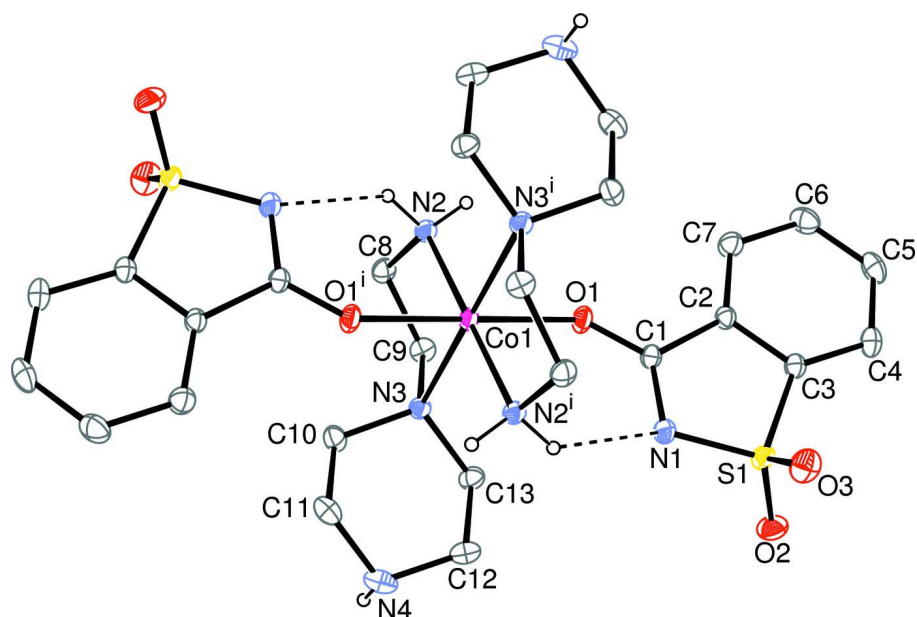


Figure 1

The molecular structure of (I) showing 30% displacement ellipsoids (arbitrary spheres for the H atoms). Symmetry code: (i) $-x + 1, -y + 1, -z + 1$. The intramolecular N—H···N hydrogen bonds are indicated by dashed lines.

***trans*-Bis[2-(piperazin-1-yl)ethanamine]bis(saccharinato)cobalt(II)**

Crystal data

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$M_r = 681.69$

Triclinic, $P\bar{1}$

Hall symbol: $-P\ 1$

$a = 8.4294$ (7) Å

$b = 9.3742$ (10) Å

$c = 11.5618$ (10) Å

$\alpha = 93.651$ (8)°

$\beta = 110.473$ (6)°

$\gamma = 116.486$ (13)°

$V = 739.31$ (17) Å³

$Z = 1$

$F(000) = 357$

$D_x = 1.531$ Mg m⁻³

Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å

Cell parameters from 16784 reflections

$\theta = 2.0$ – 28.0 °

$\mu = 0.78$ mm⁻¹

$T = 293$ K

Prism, light brown

$0.58 \times 0.41 \times 0.25$ mm

Data collection

Stoe IPDS 2

diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

Detector resolution: 6.67 pixels mm⁻¹

rotation method scans

Absorption correction: integration

(*X-RED*; Stoe & Cie, 2002)

$T_{\min} = 0.721$, $T_{\max} = 0.865$

9652 measured reflections

2913 independent reflections

2635 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.103$

$\theta_{\max} = 26.0$ °, $\theta_{\min} = 2.0$ °

$h = -10 \rightarrow 10$

$k = -10 \rightarrow 11$

$l = -14 \rightarrow 14$

Refinement

Refinement on F^2
 Least-squares matrix: full
 $R[F^2 > 2\sigma(F^2)] = 0.033$
 $wR(F^2) = 0.094$
 $S = 1.07$
 2913 reflections
 208 parameters
 1 restraint
 Primary atom site location: structure-invariant
 direct methods

Secondary atom site location: difference Fourier
 map
 Hydrogen site location: inferred from
 neighbouring sites
 H atoms treated by a mixture of independent
 and constrained refinement
 $w = 1/[\sigma^2(F_o^2) + (0.0416P)^2 + 0.1493P]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\max} < 0.001$
 $\Delta\rho_{\max} = 0.47 \text{ e } \text{Å}^{-3}$
 $\Delta\rho_{\min} = -0.64 \text{ e } \text{Å}^{-3}$

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) etc. and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
Co1	0.5000	0.5000	0.5000	0.02842 (12)
S1	0.60105 (7)	0.20581 (6)	0.86106 (4)	0.03467 (14)
O1	0.4721 (2)	0.46565 (17)	0.67018 (13)	0.0386 (3)
O2	0.7830 (2)	0.2754 (2)	0.97136 (15)	0.0540 (4)
O3	0.5203 (3)	0.03573 (19)	0.79884 (17)	0.0557 (4)
N1	0.6138 (2)	0.3163 (2)	0.75875 (16)	0.0379 (4)
N2	0.4126 (2)	0.6783 (2)	0.50483 (16)	0.0324 (3)
N3	0.8077 (2)	0.73527 (18)	0.62659 (14)	0.0312 (3)
N4	1.1486 (2)	0.7105 (2)	0.6325 (2)	0.0498 (5)
C1	0.4967 (3)	0.3785 (2)	0.74533 (16)	0.0313 (4)
C2	0.3873 (3)	0.3351 (2)	0.82672 (17)	0.0311 (4)
C3	0.4282 (3)	0.2328 (2)	0.89749 (17)	0.0332 (4)
C4	0.3404 (3)	0.1714 (3)	0.9774 (2)	0.0441 (5)
H4	0.3693	0.1030	1.0251	0.053*
C5	0.2076 (3)	0.2162 (3)	0.9835 (2)	0.0501 (5)
H5	0.1437	0.1752	1.0352	0.060*
C6	0.1672 (3)	0.3207 (3)	0.9146 (2)	0.0508 (5)
H6	0.0790	0.3508	0.9219	0.061*
C7	0.2570 (3)	0.3805 (3)	0.8352 (2)	0.0436 (5)
H7	0.2298	0.4502	0.7884	0.052*
C8	0.5815 (3)	0.8468 (2)	0.5672 (2)	0.0379 (4)
H8A	0.5428	0.9166	0.6017	0.046*
H8B	0.6277	0.8934	0.5052	0.046*
C9	0.7420 (3)	0.8403 (2)	0.67326 (19)	0.0381 (4)

H9A	0.8524	0.9514	0.7150	0.046*
H9B	0.6957	0.7976	0.7363	0.046*
C10	0.9221 (3)	0.8211 (2)	0.5561 (2)	0.0405 (4)
H10A	1.0226	0.9321	0.6082	0.049*
H10B	0.8365	0.8297	0.4785	0.049*
C11	1.0174 (3)	0.7314 (3)	0.5215 (2)	0.0475 (5)
H11A	0.9159	0.6234	0.4642	0.057*
H11B	1.0910	0.7925	0.4759	0.057*
C12	1.0468 (3)	0.6362 (3)	0.7102 (2)	0.0448 (5)
H12A	1.1410	0.6398	0.7898	0.054*
H12B	0.9525	0.5210	0.6657	0.054*
C13	0.9411 (3)	0.7169 (3)	0.74108 (18)	0.0390 (4)
H13A	0.8658	0.6515	0.7842	0.047*
H13B	1.0371	0.8253	0.7997	0.047*
H2A	0.348 (3)	0.681 (3)	0.424 (2)	0.043 (6)*
H2B	0.337 (3)	0.655 (3)	0.543 (2)	0.040 (6)*
H4A	1.244 (3)	0.804 (2)	0.677 (3)	0.064 (8)*

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Co1	0.0362 (2)	0.03155 (19)	0.02755 (19)	0.02107 (15)	0.01746 (14)	0.01371 (14)
S1	0.0412 (3)	0.0402 (3)	0.0304 (2)	0.0253 (2)	0.0159 (2)	0.01366 (19)
O1	0.0519 (8)	0.0450 (7)	0.0335 (7)	0.0294 (6)	0.0246 (6)	0.0217 (6)
O2	0.0454 (8)	0.0778 (11)	0.0397 (8)	0.0367 (8)	0.0107 (7)	0.0189 (8)
O3	0.0727 (10)	0.0430 (8)	0.0596 (10)	0.0346 (8)	0.0296 (8)	0.0114 (7)
N1	0.0430 (8)	0.0491 (9)	0.0348 (8)	0.0276 (7)	0.0228 (7)	0.0197 (7)
N2	0.0360 (8)	0.0385 (8)	0.0332 (8)	0.0238 (7)	0.0178 (7)	0.0141 (7)
N3	0.0333 (7)	0.0337 (7)	0.0310 (7)	0.0190 (6)	0.0150 (6)	0.0102 (6)
N4	0.0321 (9)	0.0534 (11)	0.0560 (12)	0.0198 (8)	0.0152 (8)	-0.0008 (9)
C1	0.0357 (9)	0.0347 (8)	0.0249 (8)	0.0176 (7)	0.0141 (7)	0.0094 (7)
C2	0.0333 (8)	0.0336 (8)	0.0267 (8)	0.0158 (7)	0.0139 (7)	0.0096 (7)
C3	0.0369 (9)	0.0344 (9)	0.0288 (9)	0.0169 (7)	0.0154 (7)	0.0104 (7)
C4	0.0558 (12)	0.0434 (10)	0.0360 (10)	0.0218 (10)	0.0251 (9)	0.0181 (9)
C5	0.0530 (12)	0.0553 (12)	0.0423 (11)	0.0178 (10)	0.0330 (10)	0.0128 (10)
C6	0.0481 (12)	0.0639 (14)	0.0501 (13)	0.0302 (11)	0.0289 (10)	0.0108 (11)
C7	0.0504 (11)	0.0519 (11)	0.0425 (11)	0.0327 (10)	0.0239 (9)	0.0181 (9)
C8	0.0467 (10)	0.0328 (9)	0.0447 (11)	0.0256 (8)	0.0222 (9)	0.0134 (8)
C9	0.0427 (10)	0.0347 (9)	0.0366 (10)	0.0208 (8)	0.0155 (8)	0.0037 (8)
C10	0.0393 (10)	0.0368 (9)	0.0458 (11)	0.0156 (8)	0.0226 (8)	0.0144 (8)
C11	0.0425 (11)	0.0496 (12)	0.0512 (12)	0.0181 (9)	0.0280 (9)	0.0087 (10)
C12	0.0377 (10)	0.0491 (11)	0.0437 (11)	0.0267 (9)	0.0077 (8)	0.0057 (9)
C13	0.0378 (10)	0.0454 (10)	0.0320 (9)	0.0229 (8)	0.0104 (8)	0.0086 (8)

Geometric parameters (Å, °)

Co1—O1 ⁱ	2.0897 (13)	C3—C4	1.378 (3)
Co1—O1	2.0897 (13)	C4—C5	1.380 (3)

Co1—N2 ⁱ	2.1101 (16)	C4—H4	0.9300
Co1—N2	2.1101 (16)	C5—C6	1.385 (4)
Co1—N3	2.3589 (16)	C5—H5	0.9300
Co1—N3 ⁱ	2.3589 (16)	C6—C7	1.379 (3)
S1—O2	1.4274 (15)	C6—H6	0.9300
S1—O3	1.4376 (16)	C7—H7	0.9300
S1—N1	1.6214 (16)	C8—C9	1.504 (3)
S1—C3	1.763 (2)	C8—H8A	0.9700
O1—C1	1.257 (2)	C8—H8B	0.9700
N1—C1	1.326 (3)	C9—H9A	0.9700
N2—C8	1.474 (2)	C9—H9B	0.9700
N2—H2A	0.91 (3)	C10—C11	1.515 (3)
N2—H2B	0.85 (3)	C10—H10A	0.9700
N3—C10	1.478 (3)	C10—H10B	0.9700
N3—C9	1.482 (2)	C11—H11A	0.9700
N3—C13	1.485 (2)	C11—H11B	0.9700
N4—C12	1.454 (3)	C12—C13	1.512 (3)
N4—C11	1.463 (3)	C12—H12A	0.9700
N4—H4A	0.838 (17)	C12—H12B	0.9700
C1—C2	1.490 (3)	C13—H13A	0.9700
C2—C7	1.373 (3)	C13—H13B	0.9700
C2—C3	1.380 (2)		
O1 ⁱ —Co1—O1	180.000 (1)	C3—C4—H4	121.5
O1 ⁱ —Co1—N2 ⁱ	88.47 (6)	C5—C4—H4	121.5
O1—Co1—N2 ⁱ	91.53 (6)	C4—C5—C6	121.7 (2)
O1 ⁱ —Co1—N2	91.53 (6)	C4—C5—H5	119.2
O1—Co1—N2	88.47 (6)	C6—C5—H5	119.2
N2 ⁱ —Co1—N2	180.000 (1)	C7—C6—C5	120.3 (2)
O1 ⁱ —Co1—N3	92.94 (6)	C7—C6—H6	119.9
O1—Co1—N3	87.06 (6)	C5—C6—H6	119.9
N2 ⁱ —Co1—N3	99.38 (6)	C2—C7—C6	118.75 (19)
N2—Co1—N3	80.62 (6)	C2—C7—H7	120.6
O1 ⁱ —Co1—N3 ⁱ	87.06 (5)	C6—C7—H7	120.6
O1—Co1—N3 ⁱ	92.94 (6)	N2—C8—C9	109.12 (15)
N2 ⁱ —Co1—N3 ⁱ	80.62 (6)	N2—C8—H8A	109.9
N2—Co1—N3 ⁱ	99.38 (6)	C9—C8—H8A	109.9
N3—Co1—N3 ⁱ	180.0	N2—C8—H8B	109.9
O2—S1—O3	114.88 (11)	C9—C8—H8B	109.9
O2—S1—N1	111.71 (10)	H8A—C8—H8B	108.3
O3—S1—N1	110.32 (10)	N3—C9—C8	112.44 (16)
O2—S1—C3	110.86 (10)	N3—C9—H9A	109.1
O3—S1—C3	110.47 (10)	C8—C9—H9A	109.1
N1—S1—C3	97.23 (9)	N3—C9—H9B	109.1
C1—O1—Co1	136.88 (13)	C8—C9—H9B	109.1
C1—N1—S1	110.72 (14)	H9A—C9—H9B	107.8
C8—N2—Co1	112.04 (11)	N3—C10—C11	111.99 (17)
C8—N2—H2A	105.0 (14)	N3—C10—H10A	109.2

Co1—N2—H2A	110.7 (16)	C11—C10—H10A	109.2
C8—N2—H2B	110.3 (16)	N3—C10—H10B	109.2
Co1—N2—H2B	110.0 (16)	C11—C10—H10B	109.2
H2A—N2—H2B	109 (2)	H10A—C10—H10B	107.9
C10—N3—C9	109.30 (15)	N4—C11—C10	113.47 (19)
C10—N3—C13	107.64 (15)	N4—C11—H11A	108.9
C9—N3—C13	107.08 (15)	C10—C11—H11A	108.9
C10—N3—Co1	115.28 (12)	N4—C11—H11B	108.9
C9—N3—Co1	99.17 (11)	C10—C11—H11B	108.9
C13—N3—Co1	117.56 (11)	H11A—C11—H11B	107.7
C12—N4—C11	109.61 (16)	N4—C12—C13	115.08 (18)
C12—N4—H4A	108 (2)	N4—C12—H12A	108.5
C11—N4—H4A	109 (2)	C13—C12—H12A	108.5
O1—C1—N1	125.08 (18)	N4—C12—H12B	108.5
O1—C1—C2	120.27 (17)	C13—C12—H12B	108.5
N1—C1—C2	114.64 (15)	H12A—C12—H12B	107.5
C7—C2—C3	120.25 (19)	N3—C13—C12	113.48 (17)
C7—C2—C1	128.66 (17)	N3—C13—H13A	108.9
C3—C2—C1	111.08 (17)	C12—C13—H13A	108.9
C4—C3—C2	122.14 (19)	N3—C13—H13B	108.9
C4—C3—S1	131.55 (16)	C12—C13—H13B	108.9
C2—C3—S1	106.31 (14)	H13A—C13—H13B	107.7
C3—C4—C5	116.93 (19)		
N2 ⁱ —Co1—O1—C1	0.70 (18)	C1—C2—C3—C4	-178.03 (17)
N2—Co1—O1—C1	-179.30 (18)	C7—C2—C3—S1	-179.55 (15)
N3—Co1—O1—C1	-98.62 (18)	C1—C2—C3—S1	1.80 (18)
N3 ⁱ —Co1—O1—C1	81.38 (18)	O2—S1—C3—C4	-64.8 (2)
O2—S1—N1—C1	-115.72 (15)	O3—S1—C3—C4	63.7 (2)
O3—S1—N1—C1	115.20 (15)	N1—S1—C3—C4	178.58 (19)
C3—S1—N1—C1	0.18 (15)	O2—S1—C3—C2	115.35 (14)
O1 ⁱ —Co1—N2—C8	-86.23 (13)	O3—S1—C3—C2	-116.12 (14)
O1—Co1—N2—C8	93.77 (13)	N1—S1—C3—C2	-1.22 (14)
N3—Co1—N2—C8	6.49 (13)	C2—C3—C4—C5	0.3 (3)
N3 ⁱ —Co1—N2—C8	-173.51 (13)	S1—C3—C4—C5	-179.46 (17)
O1 ⁱ —Co1—N3—C10	-3.65 (13)	C3—C4—C5—C6	-1.3 (3)
O1—Co1—N3—C10	176.35 (13)	C4—C5—C6—C7	1.3 (4)
N2 ⁱ —Co1—N3—C10	85.29 (13)	C3—C2—C7—C6	-0.6 (3)
N2—Co1—N3—C10	-94.71 (13)	C1—C2—C7—C6	177.80 (19)
O1 ⁱ —Co1—N3—C9	112.87 (11)	C5—C6—C7—C2	-0.4 (3)
O1—Co1—N3—C9	-67.13 (11)	Co1—N2—C8—C9	-34.5 (2)
N2 ⁱ —Co1—N3—C9	-158.19 (11)	C10—N3—C9—C8	72.7 (2)
N2—Co1—N3—C9	21.81 (11)	C13—N3—C9—C8	-170.98 (16)
O1 ⁱ —Co1—N3—C13	-132.30 (13)	Co1—N3—C9—C8	-48.30 (17)
O1—Co1—N3—C13	47.70 (13)	N2—C8—C9—N3	59.5 (2)
N2 ⁱ —Co1—N3—C13	-43.36 (14)	C9—N3—C10—C11	172.09 (16)
N2—Co1—N3—C13	136.64 (14)	C13—N3—C10—C11	56.1 (2)
Co1—O1—C1—N1	24.4 (3)	Co1—N3—C10—C11	-77.30 (18)

Co1—O1—C1—C2	-154.52 (14)	C12—N4—C11—C10	52.0 (2)
S1—N1—C1—O1	-178.10 (15)	N3—C10—C11—N4	-58.3 (2)
S1—N1—C1—C2	0.9 (2)	C11—N4—C12—C13	-49.1 (2)
O1—C1—C2—C7	-1.3 (3)	C10—N3—C13—C12	-53.1 (2)
N1—C1—C2—C7	179.63 (19)	C9—N3—C13—C12	-170.55 (17)
O1—C1—C2—C3	177.19 (16)	Co1—N3—C13—C12	79.06 (19)
N1—C1—C2—C3	-1.9 (2)	N4—C12—C13—N3	52.0 (2)
C7—C2—C3—C4	0.6 (3)		

Symmetry code: (i) $-x+1, -y+1, -z+1$.

Hydrogen-bond geometry (Å, °)

<i>D</i> —H \cdots <i>A</i>	<i>D</i> —H	H \cdots <i>A</i>	<i>D</i> \cdots <i>A</i>	<i>D</i> —H \cdots <i>A</i>
N2—H2B \cdots N4 ⁱⁱ	0.85 (3)	2.38 (3)	3.172 (3)	154 (2)
N2—H2A \cdots N1 ⁱ	0.91 (3)	2.25 (3)	2.982 (2)	137 (2)
N4—H4A \cdots O3 ⁱⁱⁱ	0.84 (2)	2.22 (2)	3.054 (3)	173 (3)

Symmetry codes: (i) $-x+1, -y+1, -z+1$; (ii) $x-1, y, z$; (iii) $x+1, y+1, z$.