

Bis(η^5 -cyclopentadienyl)bis(2,4,6-trimethylphenyltellurolato)zirconium(IV)

Andrew L. Hector, William Levason, Gillian Reid,
Stuart D. Reid and Michael Webster*

School of Chemistry, University of Southampton, Southampton SO17 1BJ, England
Correspondence e-mail: m.webster@soton.ac.uk

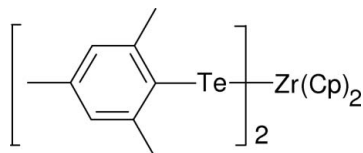
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Key indicators: single-crystal X-ray study; $T = 120$ K; mean $\sigma(\text{C}-\text{C}) = 0.011$ Å;
R factor = 0.040; wR factor = 0.087; data-to-parameter ratio = 21.0.

The structure of the title compound, $[\text{Zr}(\text{C}_5\text{H}_5)_2(\text{C}_9\text{H}_{11}\text{Te})_2]$, consists of a zirconium(IV) centre bonded to two η^5 -coordinated cyclopentadienyl groups and two mesityltellurolate ligands; the discrete molecule has crystallographic twofold rotation symmetry. The structural parameters compared with those in $[(\eta^5\text{-Me}_5\text{Cp})_2\text{Zr}(\text{TePh})_2]$ [Howard, Trnka & Parkin (1995). *Inorg. Chem.* **34**, 5900–5909] show that the greater steric demands of the bulky mesityl substituents are accommodated by widening $\text{Te}-\text{Zr}-\text{Te}$ ($\sim 8^\circ$) and by more acute $\text{Zr}-\text{Te}-\text{C}$ ($\sim 5^\circ$) angles, although the $\text{Zr}-\text{Te}$ distances are essentially the same. The crystal studied exhibited some inversion twinning.

Related literature

For a review, see: Arnold (1995). For related structures, see: Christou *et al.* (1993); Hector *et al.* (2008); Howard *et al.* (1995); Sato & Yoshida (1974).



Experimental

Crystal data

$[\text{Zr}(\text{C}_5\text{H}_5)_2(\text{C}_9\text{H}_{11}\text{Te})_2]$	$V = 2733.3$ (12) Å ³
$M_r = 714.96$	$Z = 4$
Orthorhombic, <i>Aba2</i>	Mo $K\alpha$ radiation
$a = 9.0483$ (15) Å	$\mu = 2.51$ mm ⁻¹
$b = 21.881$ (6) Å	$T = 120$ (2) K
$c = 13.806$ (4) Å	$0.20 \times 0.10 \times 0.02$ mm

Data collection

Bruker Nonius KappaCCD diffractometer	9201 measured reflections
Absorption correction: multi-scan (<i>SADABS</i> ; Sheldrick, 2007)	3046 independent reflections
$T_{\min} = 0.758$, $T_{\max} = 0.951$	2444 reflections with $I > 2\sigma(I)$
	$R_{\text{int}} = 0.059$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.040$	H-atom parameters constrained
$wR(F^2) = 0.087$	$\Delta\rho_{\max} = 1.50$ e Å ⁻³
$S = 1.05$	$\Delta\rho_{\min} = -1.16$ e Å ⁻³
3046 reflections	Absolute structure: Flack (1983),
145 parameters	with 1405 Friedel pairs
1 restraint	Flack parameter: 0.14 (5)

Table 1

Selected geometric parameters (Å, °).

Zr1—C1	2.519 (8)	Zr1—C5	2.504 (8)
Zr1—C2	2.519 (7)	Zr1—Te1	2.8694 (10)
Zr1—C3	2.455 (7)	Te1—C6	2.150 (7)
Zr1—C4	2.470 (7)		
Te1—Zr1—Te1 ⁱ	105.34 (5)	C6—Te1—Zr1	108.06 (16)
Zr1—Te1—C6—C7	77.0 (5)	Te1 ⁱ —Zr1—Te1—C6	-79.80 (19)
Zr1—Te1—C6—C11	-103.8 (5)		

Symmetry code: (i) $-x + 1, -y + 1, z$.

Data collection: *COLLECT* (Hooft, 1998) and *DENZO* (Otwinowski & Minor, 1997); cell refinement: *COLLECT* and *DENZO*; data reduction: *COLLECT* and *DENZO*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEPII* (Johnson, 1976); software used to prepare material for publication: *SHELXL97*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: SJ2480).

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supplementary materials

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Bis(η^5 -cyclopentadienyl)bis(2,4,6-trimethylphenyltellurolato)zirconium(IV)

A. L. Hector, W. Levason, G. Reid, S. D. Reid and M. Webster

Comment

Thiolate ligands (RS^-) form complexes with most metals and metalloids in the Periodic Table. In contrast, much less is known about corresponding selenolates (RSe^-), and few tellurolate (RTe^-) complexes have been characterized in detail (Arnold, 1995). The latter include $[(\eta^5\text{-Me}_5\text{Cp})_2\text{Zr}(\text{TePh})_2]$ prepared from $[(\eta^5\text{-Me}_5\text{Cp})_2\text{Zr}(\text{CO})_2]$, PhOH and Ph_2Te_2 , (Howard *et al.*, 1995) and $[(\eta^5\text{-Cp})_2\text{Zr}(\text{TePh})_2]$ prepared from $[(\eta^5\text{-Cp})_2\text{ZrCl}_2]$ and PhTeLi, (Sato & Yoshida, 1974). We have recently characterized a range of complexes of Ti, Zr and Hf (M) of type $[(\eta^5\text{-Cp})_2M(\text{Se}^t\text{Bu})_2]$ and shown that these complexes serve as single-source precursors for LPCVD (low pressure chemical vapour deposition) of $M\text{Se}_2$ films (Hector *et al.*, 2008), but that the corresponding *t*-butyltellurolates decompose to deposit elemental tellurium. During attempts to improve the stability of the tellurato-complexes, we obtained crystals of the title complex which we now report.

Red crystals of the title compound (I) were obtained in poor yield by reaction of $[(\eta^5\text{-Cp})_2\text{ZrCl}_2]$ with $(\text{Me}_3\text{C}_6\text{H}_2)\text{TeMgBr}$ in anhydrous THF solution. The discrete molecule has 2-fold crystallographic symmetry, and shows the typical metallocene geometry with η^5 -coordinated Cp rings (Zr—C 2.455 (7)–2.519 (8) Å, 2.49 (3) Å (av)) rather shorter than those in $[(\eta^5\text{-Me}_5\text{Cp})_2\text{Zr}(\text{TePh})_2]$ (2.56 (5) Å (av)) (Howard *et al.*, 1995), but similar to those in the silyltellurolate $[(\eta^5\text{-Cp})_2\text{Zr}\{\text{TeSi}(\text{SiMe}_3)_3\}_2]$ (2.50 (1) Å (av)) (Christou *et al.*, 1993). The Zr—Te distances in $[(\eta^5\text{-Cp})_2\text{Zr}\{\text{TeSi}(\text{SiMe}_3)_3\}_2]$ (2.866 (1) Å), $[(\eta^5\text{-Cp})_2\text{Zr}(\text{TeC}_6\text{H}_2\text{Me}_3)_2]$ (2.869 (1) Å), and $[(\eta^5\text{-Me}_5\text{Cp})_2\text{Zr}(\text{TePh})_2]$ (2.87 (2) Å) are very similar as are the Te—C distances in the last two compounds (2.150 (7) and 2.12 (2) Å respectively). A more notable difference is in the Te—Zr—Te and Zr—Te—C angles between $[(\eta^5\text{-Cp})_2\text{Zr}(\text{TeC}_6\text{H}_2\text{Me}_3)_2]$ and $[(\eta^5\text{-Me}_5\text{Cp})_2\text{Zr}(\text{TePh})_2]$ with Te—Zr—Te 105.34 (5) ° *versus* 97.2 (1) °, and Zr—Te—C 108.06 (16) ° *versus* 113.1 (7) °, consistent with the greater steric effects of the mesityl groups.

Experimental

To a stirred suspension of Mg turnings (66 mg, 2.72 mmol) in THF (30 ml) was added MesBr (56 mg, 2.82 mmol). The resulting mixture was stirred for 2 h, after which the Grignard was transferred by cannula to a suspension of freshly ground Te powder (300 mg, 2.35 mmol) in THF (10 ml). The mixture turned orange and was stirred for 1 h. Cp_2ZrCl_2 (345 mg, 1.18 mmol) was dissolved in THF (10 ml) and the Grignard solution added dropwise by cannula, during which time the solution turned red. The reaction was allowed to proceed overnight. The volatiles were removed *in vacuo*, the residue extracted with CH_2Cl_2 (20 ml) and filtered through celite. The solvent was removed under reduced pressure, the residue crystallized from Et_2O to produce a small number of red crystals. $^{125}\text{Te}\{^1\text{H}\}$ NMR ($\text{CH}_2\text{Cl}_2/\text{CDCl}_3$, 300 K): $\delta_{\text{Te}} = 887$ p.p.m.

Refinement

H atoms were placed in calculated positions [C—H = 0.95 (aromatic and Cp) and 0.98 Å (methyl)]. $U_{\text{iso}}(\text{H})$ values for methyl H atoms were set at $1.5U_{\text{eq}}(\text{C})$ of the bonded C, and the rest at $1.2U_{\text{eq}}(\text{C})$. Racemic twinning was allowed in the final refinement. The number of Friedel pairs measured is 1405.

Figures

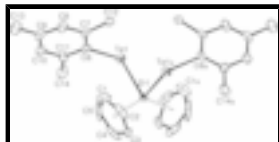


Fig. 1. The discrete molecule of (I) showing the atom-labelling scheme. Displacement ellipsoids are drawn at the 50% probability level and H atoms are omitted for clarity. Symmetry operation: $a = 1 - x, 1 - y, z$.

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Crystal data

$[\text{Zr}(\text{C}_5\text{H}_5)_2(\text{C}_9\text{H}_{11}\text{Te})_2]$

$M_r = 714.96$

Orthorhombic, *Aba2*

Hall symbol: A 2 -2ac

$a = 9.0483$ (15) Å

$b = 21.881$ (6) Å

$c = 13.806$ (4) Å

$V = 2733.3$ (12) Å³

$Z = 4$

$F_{000} = 1376$

$D_x = 1.737$ Mg m⁻³

Mo $K\alpha$ radiation

$\lambda = 0.71073$ Å

Cell parameters from 7078 reflections

$\theta = 2.9\text{--}27.5^\circ$

$\mu = 2.51$ mm⁻¹

$T = 120$ (2) K

Plate, red

$0.20 \times 0.10 \times 0.02$ mm

Data collection

Bruker Nonius KappaCCD
diffractometer

3046 independent reflections

Radiation source: Bruker-Nonius FR591 rotating anode

2444 reflections with $I > 2\sigma(I)$

Monochromator: graphite

$R_{\text{int}} = 0.059$

$T = 120$ (2) K

$\theta_{\text{max}} = 27.5^\circ$

φ and ω scans

$\theta_{\text{min}} = 3.5^\circ$

Absorption correction: multi-scan
(SADABS; Sheldrick, 2007)

$h = -9 \rightarrow 11$

$T_{\text{min}} = 0.758, T_{\text{max}} = 0.951$

$k = -28 \rightarrow 26$

9201 measured reflections

$l = -17 \rightarrow 16$

Refinement

Refinement on F^2

Hydrogen site location: inferred from neighbouring sites

Least-squares matrix: full

$$R[F^2 > 2\sigma(F^2)] = 0.040$$

$$wR(F^2) = 0.087$$

$$S = 1.05$$

3046 reflections

145 parameters

1 restraint

Primary atom site location: structure-invariant direct methods

Secondary atom site location: difference Fourier map

H-atom parameters constrained

$$w = 1/[\sigma^2(F_o^2) + (0.0276P)^2 + 1.8865P]$$

$$\text{where } P = (F_o^2 + 2F_c^2)/3$$

$$(\Delta/\sigma)_{\max} = 0.001$$

$$\Delta\rho_{\max} = 1.50 \text{ e } \text{\AA}^{-3}$$

$$\Delta\rho_{\min} = -1.16 \text{ e } \text{\AA}^{-3}$$

Extinction correction: none

Absolute structure: Flack (1983), 1405 Friedel pairs

Flack parameter: 0.14 (5)

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
Zr1	0.5000	0.5000	-0.06403 (8)	0.01613 (16)
Te1	0.74620 (4)	0.522546 (15)	0.06201 (7)	0.02158 (12)
C1	0.4387 (14)	0.6120 (4)	-0.0508 (8)	0.066 (3)
H1	0.4389	0.6355	0.0070	0.079*
C2	0.3184 (8)	0.5854 (4)	-0.0922 (6)	0.043 (2)
H2	0.2208	0.5869	-0.0670	0.052*
C3	0.3574 (7)	0.5607 (3)	-0.1807 (6)	0.0324 (16)
H3	0.2915	0.5453	-0.2284	0.039*
C4	0.5054 (9)	0.5768 (3)	-0.1952 (7)	0.051 (2)
H4	0.5578	0.5726	-0.2544	0.062*
C5	0.5545 (11)	0.6072 (4)	-0.1164 (10)	0.067 (4)
H5	0.6475	0.6271	-0.1110	0.081*
C6	0.7316 (6)	0.6154 (3)	0.1121 (5)	0.0222 (14)
C7	0.6309 (7)	0.6307 (3)	0.1845 (4)	0.0221 (16)
C8	0.6241 (7)	0.6926 (3)	0.2140 (6)	0.0278 (15)
H8	0.5545	0.7039	0.2624	0.033*
C9	0.7141 (8)	0.7370 (4)	0.1756 (6)	0.0293 (19)
C10	0.8146 (8)	0.7201 (3)	0.1041 (6)	0.0312 (17)
H10	0.8773	0.7502	0.0764	0.037*
C11	0.8248 (6)	0.6597 (3)	0.0723 (6)	0.0261 (14)
C12	0.5339 (7)	0.5842 (3)	0.2340 (5)	0.0296 (17)

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H12A	0.4787	0.5611	0.1852	0.044*
H12B	0.4644	0.6052	0.2771	0.044*
H12C	0.5955	0.5562	0.2719	0.044*
C13	0.6987 (9)	0.8020 (4)	0.2084 (7)	0.041 (2)
H13A	0.5976	0.8161	0.1961	0.062*
H13B	0.7685	0.8278	0.1727	0.062*
H13C	0.7198	0.8046	0.2779	0.062*
C14	0.9425 (8)	0.6443 (4)	-0.0014 (6)	0.0388 (19)
H14A	0.8958	0.6352	-0.0638	0.058*
H14B	0.9984	0.6085	0.0207	0.058*
H14C	1.0097	0.6791	-0.0088	0.058*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Zr1	0.0179 (3)	0.0137 (3)	0.0168 (3)	0.0028 (3)	0.000	0.000
Te1	0.02085 (18)	0.01997 (19)	0.0239 (2)	0.00485 (15)	-0.00464 (18)	-0.0039 (2)
C1	0.139 (10)	0.024 (4)	0.035 (7)	0.037 (5)	-0.038 (7)	-0.002 (4)
C2	0.033 (4)	0.041 (5)	0.055 (7)	0.022 (3)	0.023 (4)	0.028 (4)
C3	0.049 (4)	0.014 (3)	0.035 (4)	-0.002 (3)	-0.028 (4)	0.003 (3)
C4	0.065 (5)	0.039 (4)	0.050 (6)	0.032 (4)	0.032 (6)	0.034 (4)
C5	0.060 (6)	0.016 (5)	0.126 (11)	-0.014 (4)	-0.053 (7)	0.028 (6)
C6	0.022 (3)	0.022 (3)	0.022 (4)	0.006 (2)	-0.004 (3)	-0.006 (3)
C7	0.023 (3)	0.026 (4)	0.017 (4)	0.002 (3)	-0.003 (2)	-0.008 (3)
C8	0.024 (3)	0.023 (4)	0.036 (4)	0.003 (3)	-0.002 (3)	-0.009 (3)
C9	0.025 (4)	0.031 (5)	0.032 (5)	-0.002 (3)	-0.002 (3)	-0.013 (4)
C10	0.032 (4)	0.028 (4)	0.033 (5)	-0.007 (3)	0.003 (3)	-0.007 (3)
C11	0.019 (3)	0.027 (3)	0.032 (4)	-0.003 (2)	0.002 (3)	0.002 (3)
C12	0.037 (4)	0.028 (4)	0.023 (4)	-0.002 (3)	0.002 (3)	-0.004 (3)
C13	0.039 (4)	0.030 (5)	0.055 (6)	0.003 (4)	0.001 (4)	-0.013 (4)
C14	0.033 (4)	0.044 (5)	0.040 (5)	-0.010 (3)	0.016 (3)	-0.015 (4)

Geometric parameters (\AA , $^\circ$)

Zr1—C1	2.519 (8)	C5—H5	0.9500
Zr1—C2	2.519 (7)	C6—C7	1.393 (9)
Zr1—C3	2.455 (7)	C6—C11	1.398 (9)
Zr1—C4	2.470 (7)	C7—C8	1.415 (9)
Zr1—C5	2.504 (8)	C7—C12	1.507 (9)
Zr1—C3 ⁱ	2.455 (7)	C8—C9	1.375 (11)
Zr1—C4 ⁱ	2.470 (7)	C8—H8	0.9500
Zr1—C5 ⁱ	2.504 (8)	C9—C10	1.393 (11)
Zr1—C2 ⁱ	2.519 (7)	C9—C13	1.499 (12)
Zr1—C1 ⁱ	2.519 (8)	C10—C11	1.395 (9)
Zr1—Te1	2.8694 (10)	C10—H10	0.9500
Zr1—Te1 ⁱ	2.8694 (10)	C11—C14	1.512 (9)
Te1—C6	2.150 (7)	C12—H12A	0.9800

C1—C2	1.360 (13)	C12—H12B	0.9800
C1—C5	1.390 (15)	C12—H12C	0.9800
C1—H1	0.9500	C13—H13A	0.9800
C2—C3	1.382 (12)	C13—H13B	0.9800
C2—H2	0.9500	C13—H13C	0.9800
C3—C4	1.399 (10)	C14—H14A	0.9800
C3—H3	0.9500	C14—H14B	0.9800
C4—C5	1.350 (13)	C14—H14C	0.9800
C4—H4	0.9500		
C3—Zr1—C3 ⁱ	98.0 (4)	C2—C1—C5	107.3 (9)
C3—Zr1—C4 ⁱ	82.9 (3)	C2—C1—Zr1	74.3 (5)
C3 ⁱ —Zr1—C4 ⁱ	33.0 (2)	C5—C1—Zr1	73.3 (6)
C3—Zr1—C4	33.0 (2)	C2—C1—H1	125.9
C3 ⁱ —Zr1—C4	82.9 (3)	C5—C1—H1	125.9
C4 ⁱ —Zr1—C4	85.7 (5)	Zr1—C1—H1	125.9
C3—Zr1—C5 ⁱ	102.3 (4)	C1—C2—C3	109.6 (7)
C3 ⁱ —Zr1—C5 ⁱ	53.6 (2)	C1—C2—Zr1	74.3 (4)
C4 ⁱ —Zr1—C5 ⁱ	31.5 (3)	C3—C2—Zr1	71.3 (4)
C4—Zr1—C5 ⁱ	115.4 (5)	C1—C2—H2	125.1
C3—Zr1—C5	53.6 (2)	C3—C2—H2	125.1
C3 ⁱ —Zr1—C5	102.3 (4)	Zr1—C2—H2	125.1
C4 ⁱ —Zr1—C5	115.4 (5)	C2—C3—C4	105.8 (7)
C4—Zr1—C5	31.5 (3)	C2—C3—Zr1	76.4 (4)
C5 ⁱ —Zr1—C5	146.4 (6)	C4—C3—Zr1	74.1 (4)
C3—Zr1—C2 ⁱ	130.0 (3)	C2—C3—H3	126.3
C3 ⁱ —Zr1—C2 ⁱ	32.2 (3)	C4—C3—H3	126.3
C4 ⁱ —Zr1—C2 ⁱ	52.8 (2)	Zr1—C3—H3	126.3
C4—Zr1—C2 ⁱ	112.2 (3)	C5—C4—C3	108.9 (8)
C5 ⁱ —Zr1—C2 ⁱ	52.3 (3)	C5—C4—Zr1	75.6 (5)
C5—Zr1—C2 ⁱ	121.5 (3)	C3—C4—Zr1	72.9 (4)
C3—Zr1—C2	32.2 (3)	C5—C4—H4	125.2
C3 ⁱ —Zr1—C2	130.0 (3)	C3—C4—H4	125.2
C4 ⁱ —Zr1—C2	112.2 (3)	Zr1—C4—H4	125.2
C4—Zr1—C2	52.8 (2)	C4—C5—C1	108.3 (8)
C5 ⁱ —Zr1—C2	121.5 (3)	C4—C5—Zr1	72.9 (5)
C5—Zr1—C2	52.3 (3)	C1—C5—Zr1	74.5 (6)
C2 ⁱ —Zr1—C2	162.2 (4)	C4—C5—H5	125.6
C3—Zr1—C1 ⁱ	133.6 (3)	C1—C5—H5	125.6
C3 ⁱ —Zr1—C1 ⁱ	53.5 (3)	Zr1—C5—H5	125.6
C4 ⁱ —Zr1—C1 ⁱ	52.8 (3)	C7—C6—C11	120.6 (6)
C4—Zr1—C1 ⁱ	135.3 (3)	C7—C6—Te1	119.9 (5)
C5 ⁱ —Zr1—C1 ⁱ	32.1 (3)	C11—C6—Te1	119.5 (5)
C5—Zr1—C1 ⁱ	152.7 (3)	C6—C7—C8	117.7 (6)

supplementary materials

C2 ⁱ —Zr1—C1 ⁱ	31.3 (3)	C6—C7—C12	123.0 (6)
C2—Zr1—C1 ⁱ	151.2 (3)	C8—C7—C12	119.3 (6)
C3—Zr1—C1	53.5 (3)	C9—C8—C7	122.6 (7)
C3 ⁱ —Zr1—C1	133.6 (3)	C9—C8—H8	118.7
C4 ⁱ —Zr1—C1	135.3 (3)	C7—C8—H8	118.7
C4—Zr1—C1	52.8 (3)	C8—C9—C10	118.2 (7)
C5 ⁱ —Zr1—C1	152.7 (3)	C8—C9—C13	119.9 (7)
C5—Zr1—C1	32.1 (3)	C10—C9—C13	121.8 (8)
C2 ⁱ —Zr1—C1	151.2 (3)	C9—C10—C11	121.2 (7)
C2—Zr1—C1	31.3 (3)	C9—C10—H10	119.4
C1 ⁱ —Zr1—C1	171.7 (5)	C11—C10—H10	119.4
C3—Zr1—Te1	135.49 (14)	C10—C11—C6	119.6 (6)
C3 ⁱ —Zr1—Te1	94.75 (19)	C10—C11—C14	118.1 (6)
C4 ⁱ —Zr1—Te1	125.22 (16)	C6—C11—C14	122.3 (6)
C4—Zr1—Te1	108.2 (2)	C7—C12—H12A	109.5
C5 ⁱ —Zr1—Te1	119.3 (3)	C7—C12—H12B	109.5
C5—Zr1—Te1	82.0 (2)	H12A—C12—H12B	109.5
C2 ⁱ —Zr1—Te1	73.43 (16)	C7—C12—H12C	109.5
C2—Zr1—Te1	118.2 (2)	H12A—C12—H12C	109.5
C1 ⁱ —Zr1—Te1	87.3 (3)	H12B—C12—H12C	109.5
C1—Zr1—Te1	87.7 (3)	C9—C13—H13A	109.5
C3—Zr1—Te1 ⁱ	94.75 (19)	C9—C13—H13B	109.5
C3 ⁱ —Zr1—Te1 ⁱ	135.49 (14)	H13A—C13—H13B	109.5
C4 ⁱ —Zr1—Te1 ⁱ	108.2 (2)	C9—C13—H13C	109.5
C4—Zr1—Te1 ⁱ	125.22 (16)	H13A—C13—H13C	109.5
C5 ⁱ —Zr1—Te1 ⁱ	82.0 (2)	H13B—C13—H13C	109.5
C5—Zr1—Te1 ⁱ	119.3 (3)	C11—C14—H14A	109.5
C2 ⁱ —Zr1—Te1 ⁱ	118.2 (2)	C11—C14—H14B	109.5
C2—Zr1—Te1 ⁱ	73.43 (16)	H14A—C14—H14B	109.5
C1 ⁱ —Zr1—Te1 ⁱ	87.7 (3)	C11—C14—H14C	109.5
C1—Zr1—Te1 ⁱ	87.3 (3)	H14A—C14—H14C	109.5
Te1—Zr1—Te1 ⁱ	105.34 (5)	H14B—C14—H14C	109.5
C6—Te1—Zr1	108.06 (16)		
Zr1—Te1—C6—C7	77.0 (5)	Te1 ⁱ —Zr1—Te1—C6	-79.80 (19)
Zr1—Te1—C6—C11	-103.8 (5)		

Symmetry codes: (i) $-x+1, -y+1, z$.

Fig. 1

