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Chlorido(pyridine-2-carbaldehyde oximato- κ^2N,N')(pyridine-2-carbaldehyde oxime- κ^2N,N')copper(II)

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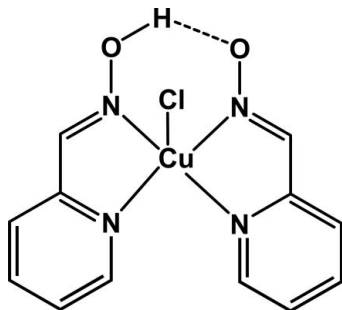
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Key indicators: single-crystal X-ray study; $T = 293$ K; mean $\sigma(\text{C}-\text{C}) = 0.006$ Å; R factor = 0.030; wR factor = 0.110; data-to-parameter ratio = 12.8.

In the title compound, $[\text{Cu}(\text{C}_6\text{H}_5\text{N}_2\text{O})\text{Cl}(\text{C}_6\text{H}_6\text{N}_2\text{O})]$, the Cu atom is coordinated by one neutral and one deprotonated pyridine-2-carboxaldehyde oxime (pco) ligand, resulting in the formation of two five-membered CuN_2C_2 rings. Together with the additional coordinating chloride anion, the coordination polyhedron of copper is best described as a distorted square-pyramid, the distortion parameter being 0.288. The two organic ligands are linked by an intramolecular $\text{O}-\text{H}\cdots\text{O}$ hydrogen bond.

Related literature

For related literature, see: Addison *et al.* (1984); Afrati *et al.* (2005); Korpi *et al.* (2005); Pearse *et al.* (1989); Stamatatos *et al.* (2006).



Experimental

Crystal data

 $[\text{Cu}(\text{C}_6\text{H}_5\text{N}_2\text{O})\text{Cl}(\text{C}_6\text{H}_6\text{N}_2\text{O})]$ $M_r = 342.24$ Monoclinic, $C2/c$ $a = 16.686$ (2) Å $b = 12.064$ (2) Å $c = 13.805$ (1) Å $\beta = 109.02$ (1)° $V = 2627.3$ (5) Å³ $Z = 8$ Mo $K\alpha$ radiation $\mu = 1.87$ mm⁻¹ $T = 293$ (2) K $0.22 \times 0.18 \times 0.15$ mm

Data collection

Bruker SMART CCD area-detector diffractometer

Absorption correction: multi-scan (*SHELXTL*; Sheldrick, 2008) $T_{\min} = 0.488$, $T_{\max} = 0.594$

(expected range = 0.620–0.755)

6487 measured reflections

2318 independent reflections

1788 reflections with $I > 2\sigma(I)$ $R_{\text{int}} = 0.034$

Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.029$ $wR(F^2) = 0.110$ $S = 1.01$

2318 reflections

181 parameters

H-atom parameters constrained

 $\Delta\rho_{\max} = 0.40$ e Å⁻³ $\Delta\rho_{\min} = -0.39$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{O1}-\text{H1B}\cdots\text{O2}$	0.82	1.70	2.488 (5)	162

Data collection: *SMART* (Bruker, 1997); cell refinement: *SAINTE* (Bruker, 1997); data reduction: *SAINTE*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: IM2064).

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supplementary materials

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Chlorido(pyridine-2-carbaldehyde oximato- κ^2N,N')(pyridine-2-carbaldehyde oxime- κ^2N,N')copper(II)

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Comment

Pyridine-2-carbaldehyde oxime ligands usually bind to metals in a bidentate fashion, either chelating one metal center or bridging two metals. Their complexes find application in diverse areas such as functional supramolecular design, magnetic materials and catalysis (Korpi *et al.*, 2005; Pearse *et al.*, 1989; Afrati *et al.*, 2005; Stamatatos *et al.*, 2006). The title compound is a new copper complex from the reaction of CuCl_2 with pyridine-2-carbaldehyde oxime (pco). The compound consists of two N,N-chelating ligands and one chloride anion. The two pco ligands are coordinated to copper to form two five-membered CuC_2N_2 rings. The copper atom adopts a distorted 4 + 1 square-pyramidal coordination mode with the distortion parameter being 0.288 (Addison *et al.*, 1984) and the angles around copper ion ranging from $79.07(1)^\circ$ for N3—Cu1—N4 to $168.37(1)^\circ$ for N2—Cu1—N3 . From the viewpoint of charge balance, it is presumed there exists one deprotonated and one protonated oxime ligand with a strong intramolecular hydrogen bond between the OH group and the negatively charged oxygen of the other ligand ($\text{O1}\cdots\text{O2} = 2.488 \text{ \AA}$) which would also give an explanation for the rather unusual cis-arrangement of the ligands (Scheme 1, Figure 1.).

Experimental

A methanolic solution (15 ml) containing pco (0.1 mmol, 0.012 g) was added to an methanolic solution (10 ml) containing $\text{CuCl}_2 \times 2 \text{ H}_2\text{O}$ (0.1 mmol, 0.017 g). After stirring for 2 h, the solution was filtered. Dark green needle-like crystals suitable for single-crystal X-ray diffraction were obtained by evaporating the resulting filtrate in air for several days (yield 65.6% based on the ligand).

Refinement

H atoms were placed geometrically and allowed to ride during refinement with $\text{C—H} = 0.93\text{--}0.96 \text{ \AA}$ and $\text{O—H} = 0.82 \text{ \AA}$ with $U_{\text{iso}}(\text{H}) = 1.2$ or $1.5U_{\text{eq}}(\text{C or O})$.

Figures

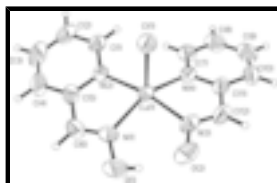


Fig. 1. The molecular structure of the title compound with the atom-numbering scheme. Displacement ellipsoids were drawn at the 50% probability level.

Chlorido(pyridine-2-carbaldehyde oximato- $\kappa^2\text{N},\text{N}'$)(pyridine-2-carbaldehyde oxime- $\kappa^2\text{N},\text{N}'$)copper(II)

Crystal data

[Cu(C ₆ H ₅ N ₂ O)Cl(C ₆ H ₆ N ₂ O)]	$F_{000} = 1384$
$M_r = 342.24$	$D_x = 1.730 \text{ Mg m}^{-3}$
Monoclinic, $C2/c$	Mo $K\alpha$ radiation
Hall symbol: -C 2yc	$\lambda = 0.71073 \text{ \AA}$
$a = 16.686 (2) \text{ \AA}$	Cell parameters from 2343 reflections
$b = 12.064 (2) \text{ \AA}$	$\theta = 2.4\text{--}26.6^\circ$
$c = 13.805 (1) \text{ \AA}$	$\mu = 1.87 \text{ mm}^{-1}$
$\beta = 109.02 (1)^\circ$	$T = 293 (2) \text{ K}$
$V = 2627.3 (5) \text{ \AA}^3$	Block, dark green
$Z = 8$	$0.22 \times 0.18 \times 0.15 \text{ mm}$

Data collection

Bruker SMART CCD area-detector diffractometer	2318 independent reflections
Radiation source: fine-focus sealed tube	1788 reflections with $I > 2\sigma(I)$
Monochromator: graphite	$R_{\text{int}} = 0.034$
$T = 293(2) \text{ K}$	$\theta_{\text{max}} = 25.0^\circ$
φ and ω scans	$\theta_{\text{min}} = 2.1^\circ$
Absorption correction: multi-scan (SHELXTL; Sheldrick, 2008)	$h = -14 \rightarrow 19$
$T_{\text{min}} = 0.488$, $T_{\text{max}} = 0.594$	$k = -14 \rightarrow 13$
6487 measured reflections	$l = -16 \rightarrow 16$

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.029$	H-atom parameters constrained
$wR(F^2) = 0.110$	$w = 1/[\sigma^2(F_o^2) + (0.065P)^2 + 1.2P]$
$S = 1.01$	where $P = (F_o^2 + 2F_c^2)/3$
2318 reflections	$(\Delta/\sigma)_{\text{max}} = 0.001$
181 parameters	$\Delta\rho_{\text{max}} = 0.40 \text{ e \AA}^{-3}$
Primary atom site location: structure-invariant direct methods	$\Delta\rho_{\text{min}} = -0.39 \text{ e \AA}^{-3}$
	Extinction correction: none

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
Cu1	0.45583 (2)	0.75857 (3)	0.08754 (3)	0.03490 (18)
Cl1	0.39133 (6)	0.74459 (6)	-0.09737 (7)	0.0452 (3)
N2	0.39301 (17)	0.8985 (2)	0.1014 (2)	0.0365 (6)
N3	0.53622 (18)	0.6345 (2)	0.0978 (2)	0.0432 (7)
N4	0.39418 (17)	0.6311 (2)	0.1352 (2)	0.0386 (7)
O2	0.60997 (16)	0.6484 (2)	0.0801 (2)	0.0594 (7)
C1	0.3107 (2)	0.9115 (3)	0.0892 (3)	0.0445 (9)
H1A	0.2761	0.8491	0.0764	0.053*
N1	0.54953 (17)	0.8713 (2)	0.1156 (2)	0.0431 (7)
C11	0.4316 (2)	0.5328 (3)	0.1314 (3)	0.0427 (9)
C5	0.4419 (2)	0.9892 (3)	0.1166 (3)	0.0433 (8)
O1	0.63020 (15)	0.8491 (3)	0.1224 (2)	0.0666 (8)
H1B	0.6346	0.7834	0.1100	0.100*
C2	0.2750 (3)	1.0126 (4)	0.0946 (3)	0.0585 (11)
H2A	0.2177	1.0187	0.0866	0.070*
C9	0.3236 (3)	0.4362 (4)	0.1762 (3)	0.0647 (12)
H9A	0.3002	0.3707	0.1903	0.078*
C7	0.3232 (2)	0.6295 (3)	0.1599 (3)	0.0457 (9)
H7A	0.2973	0.6968	0.1638	0.055*
C12	0.5115 (2)	0.5390 (3)	0.1109 (3)	0.0463 (9)
H12A	0.5426	0.4760	0.1076	0.056*
C6	0.5299 (2)	0.9703 (3)	0.1257 (3)	0.0483 (9)
H6A	0.5692	1.0276	0.1381	0.058*
C10	0.3975 (3)	0.4349 (3)	0.1509 (3)	0.0579 (11)
H10A	0.4239	0.3680	0.1471	0.069*
C8	0.2865 (3)	0.5352 (4)	0.1798 (3)	0.0593 (11)
H8A	0.2365	0.5388	0.1957	0.071*
C4	0.4094 (3)	1.0927 (3)	0.1215 (3)	0.0636 (12)
H4A	0.4440	1.1550	0.1313	0.076*
C3	0.3253 (3)	1.1034 (3)	0.1118 (4)	0.0718 (13)
H3A	0.3029	1.1729	0.1171	0.086*

supplementary materials

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Cu1	0.0277 (3)	0.0340 (3)	0.0441 (3)	0.00547 (15)	0.0132 (2)	-0.00158 (17)
Cl1	0.0482 (6)	0.0432 (5)	0.0407 (5)	0.0035 (4)	0.0094 (4)	-0.0023 (4)
N2	0.0355 (16)	0.0317 (15)	0.0418 (16)	0.0041 (12)	0.0118 (13)	-0.0026 (12)
N3	0.0398 (17)	0.0480 (19)	0.0432 (17)	0.0156 (14)	0.0152 (14)	0.0003 (14)
N4	0.0385 (16)	0.0366 (16)	0.0392 (16)	0.0031 (12)	0.0107 (13)	0.0023 (12)
O2	0.0440 (16)	0.0712 (19)	0.0731 (19)	0.0207 (14)	0.0327 (14)	0.0047 (15)
C1	0.040 (2)	0.044 (2)	0.049 (2)	0.0096 (16)	0.0142 (17)	0.0009 (16)
N1	0.0286 (16)	0.0516 (19)	0.0505 (18)	-0.0040 (13)	0.0147 (14)	-0.0005 (14)
C11	0.052 (2)	0.0372 (19)	0.032 (2)	0.0057 (16)	0.0044 (17)	-0.0001 (15)
C5	0.052 (2)	0.038 (2)	0.040 (2)	0.0022 (16)	0.0157 (18)	-0.0024 (16)
O1	0.0326 (15)	0.077 (2)	0.092 (2)	-0.0025 (13)	0.0241 (15)	-0.0042 (17)
C2	0.051 (2)	0.066 (3)	0.059 (3)	0.029 (2)	0.019 (2)	0.002 (2)
C9	0.069 (3)	0.058 (3)	0.058 (3)	-0.024 (2)	0.009 (2)	0.009 (2)
C7	0.041 (2)	0.051 (2)	0.047 (2)	-0.0006 (17)	0.0166 (17)	0.0022 (17)
C12	0.052 (2)	0.044 (2)	0.040 (2)	0.0195 (18)	0.0113 (18)	-0.0006 (17)
C6	0.045 (2)	0.045 (2)	0.056 (2)	-0.0115 (17)	0.0177 (18)	-0.0050 (18)
C10	0.071 (3)	0.036 (2)	0.053 (2)	-0.0020 (19)	0.002 (2)	0.0001 (18)
C8	0.052 (3)	0.067 (3)	0.057 (3)	-0.014 (2)	0.016 (2)	0.006 (2)
C4	0.080 (3)	0.031 (2)	0.079 (3)	0.0027 (19)	0.025 (3)	-0.0044 (19)
C3	0.088 (4)	0.044 (3)	0.084 (3)	0.033 (2)	0.029 (3)	-0.004 (2)

Geometric parameters (\AA , $^\circ$)

Cu1—N3	1.984 (3)	C5—C4	1.371 (5)
Cu1—N1	2.012 (3)	C5—C6	1.451 (5)
Cu1—N2	2.029 (3)	O1—H1B	0.8200
Cu1—N4	2.072 (3)	C2—C3	1.352 (6)
Cu1—Cl1	2.4316 (10)	C2—H2A	0.9300
N2—C1	1.338 (4)	C9—C8	1.354 (6)
N2—C5	1.340 (4)	C9—C10	1.385 (6)
N3—C12	1.256 (5)	C9—H9A	0.9300
N3—O2	1.341 (3)	C7—C8	1.361 (5)
N4—C7	1.335 (4)	C7—H7A	0.9300
N4—C11	1.350 (4)	C12—H12A	0.9300
C1—C2	1.370 (5)	C6—H6A	0.9300
C1—H1A	0.9300	C10—H10A	0.9300
N1—C6	1.258 (4)	C8—H8A	0.9300
N1—O1	1.345 (3)	C4—C3	1.371 (6)
C11—C10	1.375 (5)	C4—H4A	0.9300
C11—C12	1.453 (5)	C3—H3A	0.9300
N3—Cu1—N1	91.79 (14)	C4—C5—C6	123.0 (4)
N3—Cu1—N2	168.29 (12)	N1—O1—H1B	109.5
N1—Cu1—N2	79.19 (11)	C3—C2—C1	118.4 (4)
N3—Cu1—N4	79.15 (12)	C3—C2—H2A	120.8

N1—Cu1—N4	151.07 (12)	C1—C2—H2A	120.8
N2—Cu1—N4	105.21 (11)	C8—C9—C10	118.3 (4)
N3—Cu1—C11	94.60 (9)	C8—C9—H9A	120.9
N1—Cu1—C11	107.40 (9)	C10—C9—H9A	120.9
N2—Cu1—C11	95.22 (8)	N4—C7—C8	124.0 (4)
N4—Cu1—C11	100.71 (8)	N4—C7—H7A	118.0
C1—N2—C5	118.1 (3)	C8—C7—H7A	118.0
C1—N2—Cu1	128.8 (2)	N3—C12—C11	116.1 (3)
C5—N2—Cu1	112.8 (2)	N3—C12—H12A	121.9
C12—N3—O2	120.3 (3)	C11—C12—H12A	121.9
C12—N3—Cu1	117.1 (2)	N1—C6—C5	115.7 (3)
O2—N3—Cu1	122.1 (2)	N1—C6—H6A	122.2
C7—N4—C11	117.2 (3)	C5—C6—H6A	122.2
C7—N4—Cu1	131.7 (2)	C11—C10—C9	119.9 (4)
C11—N4—Cu1	110.9 (2)	C11—C10—H10A	120.0
N2—C1—C2	122.9 (4)	C9—C10—H10A	120.0
N2—C1—H1A	118.5	C9—C8—C7	119.2 (4)
C2—C1—H1A	118.5	C9—C8—H8A	120.4
C6—N1—O1	118.2 (3)	C7—C8—H8A	120.4
C6—N1—Cu1	116.6 (2)	C5—C4—C3	119.3 (4)
O1—N1—Cu1	125.2 (2)	C5—C4—H4A	120.4
N4—C11—C10	121.4 (4)	C3—C4—H4A	120.4
N4—C11—C12	115.3 (3)	C2—C3—C4	119.8 (4)
C10—C11—C12	123.2 (3)	C2—C3—H3A	120.1
N2—C5—C4	121.4 (4)	C4—C3—H3A	120.1
N2—C5—C6	115.6 (3)		

Hydrogen-bond geometry (Å, °)

<i>D—H...A</i>	<i>D—H</i>	<i>H...A</i>	<i>D...A</i>	<i>D—H...A</i>
O1—H1B...O2	0.82	1.70	2.488 (5)	162

Fig. 1

