

Acta Crystallographica Section E

Structure Reports

Online

ISSN 1600-5368

2-(4-Iodophenyl)-5,7-dimethyl-3-methylsulfanyl-1-benzofuran

Hong Dae Choi,^a Pil Ja Seo,^a Byung Wha Son^b and Uk Lee^{b*}

^aDepartment of Chemistry, Donggeui University, San 24 Kaya-dong, Busanjin-gu, Busan 614-714, Republic of Korea, and ^bDepartment of Chemistry, Pukyong National University, 599-1 Daeyeon 3-dong, Nam-gu, Busan 608-737, Republic of Korea

Correspondence e-mail: uklee@pknu.ac.kr

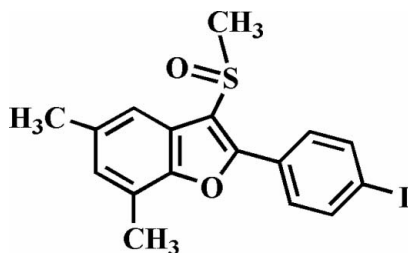
Received 8 May 2008; accepted 11 May 2008

Key indicators: single-crystal X-ray study; $T = 293$ K; mean $\sigma(\text{C}-\text{C}) = 0.005$ Å; R factor = 0.026; wR factor = 0.080; data-to-parameter ratio = 17.8.

The title compound, $\text{C}_{17}\text{H}_{15}\text{IO}_2\text{S}$, was prepared by the oxidation of 2-(4-iodophenyl)-5,7-dimethyl-3-methylsulfanyl-1-benzofuran using 3-chloroperoxybenzoic acid. The 4-iodophenyl ring makes a dihedral angle of $26.0(1)^\circ$ with the plane of the benzofuran fragment, and the O atom and the methyl group of the methylsulfanyl substituent lie on opposite sides of this plane. The crystal structure is stabilized by inter- and intramolecular $\text{C}-\text{H}\cdots\text{O}$ hydrogen bonds, and by an $\text{I}\cdots\text{O}$ halogen bond with an $\text{I}\cdots\text{O}$ distance of $3.145(2)$ Å and a nearly linear $\text{C}-\text{I}\cdots\text{O}$ angle of $164.01(9)^\circ$.

Related literature

For the crystal structures of similar 2-aryl-3-methylsulfanyl-1-benzofuran compounds, see: Choi *et al.* (2007a,b). For a review of halogen bonding, see: Politzer *et al.* (2007).



Experimental

Crystal data

$\text{C}_{17}\text{H}_{15}\text{IO}_2\text{S}$
 $M_r = 410.25$
 Triclinic, $P\bar{1}$
 $a = 8.6320(9)$ Å
 $b = 8.917(1)$ Å
 $c = 11.638(1)$ Å
 $\alpha = 94.580(2)^\circ$
 $\beta = 100.949(2)^\circ$
 $\gamma = 113.725(2)^\circ$
 $V = 792.90(14)$ Å³
 $Z = 2$
 Mo $K\alpha$ radiation
 $\mu = 2.15$ mm⁻¹
 $T = 293(2)$ K
 $0.40 \times 0.20 \times 0.20$ mm

Data collection

Bruker SMART CCD diffractometer
 Absorption correction: multi-scan (SADABS; Sheldrick, 2000)
 $T_{\min} = 0.594$, $T_{\max} = 0.647$
 6882 measured reflections
 3408 independent reflections
 3214 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.029$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.025$
 $wR(F^2) = 0.080$
 $S = 1.22$
 3408 reflections
 192 parameters
 H-atom parameters constrained
 $\Delta\rho_{\text{max}} = 0.50$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.66$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{C16}-\text{H16B}\cdots\text{O1}$	0.96	2.55	2.975 (4)	107
$\text{C16}-\text{H16A}\cdots\text{O2}^{\text{i}}$	0.96	2.39	3.288 (4)	156
$\text{C17}-\text{H17B}\cdots\text{O1}^{\text{ii}}$	0.96	2.51	3.422 (4)	159

Symmetry codes: (i) $x, y + 1, z$; (ii) $-x, -y + 1, -z + 1$.

Data collection: SMART (Bruker, 2001); cell refinement: SAINT (Bruker, 2001); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: ORTEP-3 (Farrugia, 1997) and DIAMOND (Brandenburg, 1998); software used to prepare material for publication: SHELXL97.

This work was supported by a grant from Donggeui University (2008 A A098).

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: ZL2115).

References

- Brandenburg, K. (1998). DIAMOND. Crystal Impact GbR, Bonn, Germany.
 Bruker (2001). SAINT and SMART. Bruker AXS Inc., Madison, Wisconsin, USA.
 Choi, H. D., Seo, P. J., Son, B. W. & Lee, U. (2007a). Acta Cryst. E63, o3295.
 Choi, H. D., Seo, P. J., Son, B. W. & Lee, U. (2007b). Acta Cryst. E63, o4282.
 Farrugia, L. J. (1997). J. Appl. Cryst. 30, 565.
 Politzer, P., Lane, P., Concha, M. C., Ma, Y. & Murray, J. S. (2007). J. Mol. Model. 13, 305–311.
 Sheldrick, G. M. (2000). SADABS. University of Göttingen, Germany.
 Sheldrick, G. M. (2008). Acta Cryst. A64, 112–122.

supplementary materials

Acta Cryst. (2008). E64, o1088 [doi:10.1107/S1600536808014104]

2-(4-Iodophenyl)-5,7-dimethyl-3-methylsulfinyl-1-benzofuran

H. D. Choi, P. J. Seo, B. W. Son and U. Lee

Comment

As a part of our ongoing studies on the synthesis and structure of 2-aryl-3-methylsulfinyl-1-benzofuran analogues, the crystal structure of 2-(4-bromophenyl)-5-methyl-3-methylsulfinyl-1-benzofuran (Choi *et al.*, 2007*a*) and 2-(4-bromophenyl)-5,7-dimethyl-3-methylsulfinyl-1-benzofuran (Choi *et al.*, 2007*b*) have been described in the literature. Here we report the crystal structure of the title compound, 2-(4-iodophenyl)-5,7-dimethyl-3-methylsulfinyl-1-benzofuran (Fig. 1).

The benzofuran unit is essentially planar, with a mean deviation of 0.01 Å from the least-squares plane defined by the nine constituent atoms. The molecular packing (Fig. 2) is stabilized by three different C—H \cdots O hydrogen bonds; one between a methyl H atom and the furan O atom, *i.e.* C16—H16B \cdots O1, and a second between a methyl H atom and the oxygen of a neighbouring S=O unit, *i.e.* C16—H16A \cdots O2ⁱ, and a third between a methyl H atom of the methylsulfinyl substituent and the furan O atom of neighbouring molecules, *i.e.* C17—H17B \cdots O1ⁱⁱ, (Fig. 2 and Table 1; symmetry code as in Fig. 2). Further stabilization of the structure comes from a weak I \cdots O halogen bond (Fig. 2) (Politzer *et al.*, 2007) between the iodine atom and the oxygen of a neighbouring S=O unit, with an I \cdots O2ⁱⁱⁱ distance of 3.145 (2) Å (Symmetry code as in Fig. 2).

Experimental

77% 3-chloroperoxybenzoic acid (359 mg, 1.60 mmol) was added in small portions to a stirred solution of 2-(4-iodophenyl)-5,7-dimethyl-3-methylsulfinyl-1-benzofuran (591 mg, 1.50 mmol) in dichloromethane (30 ml) at 273 K. After being stirred at room temperature for 2 h, the mixture was washed with saturated sodium bicarbonate solution and the organic layer was separated, dried over magnesium sulfate, filtered and concentrated in vacuum. The residue was purified by column chromatography (ethyl acetate) to afford the title compound as a colorless solid [yield 80%, m.p. 450–451 K; R_f = 0.57 (ethyl acetate)]. Single crystals suitable for X-ray diffraction were prepared by slow evaporation of a solution of the title compound in tetrahydrofuran at room temperature. Spectroscopic analysis: ¹H NMR (CDCl₃, 400 MHz) δ 2.44 (s, 3H), 2.53 (s, 3H), 3.10 (s, 3H), 7.03 (s, 1H), 7.59 (d, J = 8.44 Hz, 2H), 7.80 (s, 1H), 7.84 (d, J = 8.44 Hz, 2H); EI—MS 410 [M^+].

Refinement

All H atoms were geometrically located in ideal positions and refined using a riding model, with C—H = 0.95 Å for aromatic H atoms and 0.98 Å for methyl H atoms, and with $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$ for aromatic H atoms, and $1.5U_{\text{eq}}(\text{C})$ for methyl H atoms.

Figures

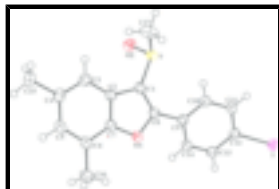


Fig. 1. The molecular structure of the title compound, showing displacement ellipsoids drawn at the 50% probability level.

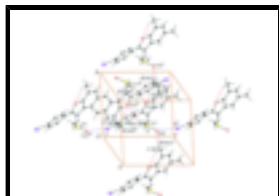


Fig. 2. C—H...O hydrogen bond and I...O halogen bond (dotted lines) in the title compound. [Symmetry codes: (i) $x, y + 1, z$; (ii) $-x, -y + 1, -z + 1$; (iii) $x, y, z - 1$; (iv) $x, y, z + 1$; (v) $x, y - 1, z$.]

2-(4-Iodophenyl)-5,7-dimethyl-3-methylsulfinyl-1-benzofuran

Crystal data

$C_{17}H_{15}IO_2S$

$M_r = 410.25$

Triclinic, $P\bar{1}$

Hall symbol: $-p_1$

$a = 8.6320$ (9) Å

$b = 8.917$ (1) Å

$c = 11.638$ (1) Å

$\alpha = 94.580$ (2)°

$\beta = 100.949$ (2)°

$\gamma = 113.725$ (2)°

$V = 792.90$ (14) Å³

$Z = 2$

$F_{000} = 404$

$D_x = 1.718$ Mg m⁻³

Melting point = 450–451 K

Mo $K\alpha$ radiation

$\lambda = 0.71069$ Å

Cell parameters from 5631 reflections

$\theta = 2.5$ – 28.3 °

$\mu = 2.15$ mm⁻¹

$T = 293$ (2) K

Block, colorless

$0.40 \times 0.20 \times 0.20$ mm

Data collection

Bruker SMART CCD
diffractometer

Radiation source: fine-focus sealed tube

Monochromator: graphite

Detector resolution: 10.0 pixels mm⁻¹

$T = 293$ (2) K

φ and ω scans

Absorption correction: multi-scan
(SADABS; Sheldrick, 2000)

$T_{\min} = 0.594$, $T_{\max} = 0.647$

6882 measured reflections

3408 independent reflections

3214 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.029$

$\theta_{\text{max}} = 27.0$ °

$\theta_{\text{min}} = 1.8$ °

$h = -11 \rightarrow 10$

$k = -11 \rightarrow 11$

$l = -14 \rightarrow 14$

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.025$	H-atom parameters constrained
$wR(F^2) = 0.080$	$w = 1/[\sigma^2(F_o^2) + (0.0374P)^2 + 0.4132P]$
$S = 1.22$	where $P = (F_o^2 + 2F_c^2)/3$
3408 reflections	$(\Delta/\sigma)_{\max} = 0.001$
192 parameters	$\Delta\rho_{\max} = 0.50 \text{ e } \text{\AA}^{-3}$
Primary atom site location: structure-invariant direct methods	$\Delta\rho_{\min} = -0.66 \text{ e } \text{\AA}^{-3}$
	Extinction correction: none

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
I	0.26443 (2)	0.23987 (2)	-0.006964 (16)	0.03233 (9)
S	0.10120 (10)	0.19886 (8)	0.60394 (7)	0.02931 (16)
O1	0.3173 (3)	0.6586 (2)	0.54519 (18)	0.0262 (4)
O2	0.2055 (3)	0.1813 (3)	0.7150 (2)	0.0374 (5)
C1	0.1717 (4)	0.4134 (3)	0.6002 (3)	0.0262 (6)
C2	0.2089 (4)	0.5443 (3)	0.6979 (3)	0.0259 (5)
C3	0.1802 (4)	0.5533 (4)	0.8114 (3)	0.0294 (6)
H3	0.1197	0.4569	0.8393	0.035*
C4	0.2430 (4)	0.7078 (4)	0.8822 (3)	0.0309 (6)
C5	0.3335 (4)	0.8521 (4)	0.8378 (3)	0.0307 (6)
H5	0.3738	0.9548	0.8862	0.037*
C6	0.3655 (4)	0.8489 (4)	0.7256 (3)	0.0282 (6)
C7	0.2995 (4)	0.6914 (3)	0.6586 (3)	0.0256 (6)
C8	0.2386 (4)	0.4876 (3)	0.5117 (3)	0.0256 (5)
C9	0.2460 (4)	0.4269 (3)	0.3936 (2)	0.0247 (5)
C10	0.1189 (4)	0.2744 (4)	0.3281 (3)	0.0286 (6)
H10	0.0280	0.2098	0.3599	0.034*
C11	0.1279 (4)	0.2191 (4)	0.2158 (3)	0.0292 (6)

supplementary materials

H11	0.0438	0.1169	0.1728	0.035*
C12	0.2623 (4)	0.3162 (4)	0.1674 (3)	0.0268 (6)
C13	0.3902 (4)	0.4699 (4)	0.2319 (3)	0.0289 (6)
H13	0.4803	0.5348	0.1997	0.035*
C14	0.3811 (4)	0.5239 (3)	0.3437 (3)	0.0274 (6)
H14	0.4656	0.6259	0.3866	0.033*
C15	0.2162 (5)	0.7226 (5)	1.0062 (3)	0.0416 (8)
H15A	0.0937	0.6706	1.0029	0.062*
H15B	0.2633	0.8380	1.0409	0.062*
H15C	0.2745	0.6687	1.0538	0.062*
C16	0.4651 (4)	1.0036 (4)	0.6797 (3)	0.0387 (7)
H16A	0.3979	1.0671	0.6682	0.058*
H16B	0.4867	0.9730	0.6054	0.058*
H16C	0.5742	1.0693	0.7362	0.058*
C17	-0.1093 (4)	0.1531 (4)	0.6283 (4)	0.0471 (9)
H17A	-0.1675	0.0368	0.6314	0.071*
H17B	-0.1764	0.1801	0.5646	0.071*
H17C	-0.0976	0.2176	0.7022	0.071*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
I	0.03823 (13)	0.03425 (13)	0.02494 (12)	0.01616 (9)	0.00863 (8)	0.00135 (8)
S	0.0364 (4)	0.0219 (3)	0.0299 (4)	0.0120 (3)	0.0091 (3)	0.0058 (3)
O1	0.0313 (10)	0.0209 (9)	0.0256 (10)	0.0103 (8)	0.0077 (8)	0.0028 (8)
O2	0.0405 (12)	0.0360 (12)	0.0384 (13)	0.0193 (10)	0.0058 (10)	0.0135 (10)
C1	0.0298 (13)	0.0227 (13)	0.0263 (14)	0.0123 (11)	0.0052 (11)	0.0037 (11)
C2	0.0287 (13)	0.0234 (13)	0.0264 (14)	0.0126 (11)	0.0052 (11)	0.0045 (11)
C3	0.0312 (14)	0.0304 (14)	0.0289 (15)	0.0145 (12)	0.0085 (12)	0.0075 (12)
C4	0.0295 (14)	0.0385 (16)	0.0265 (15)	0.0166 (13)	0.0069 (11)	0.0027 (12)
C5	0.0298 (14)	0.0269 (14)	0.0318 (15)	0.0109 (12)	0.0050 (12)	-0.0023 (12)
C6	0.0240 (13)	0.0253 (13)	0.0339 (15)	0.0100 (11)	0.0068 (11)	0.0011 (11)
C7	0.0254 (13)	0.0250 (13)	0.0264 (14)	0.0117 (11)	0.0049 (11)	0.0031 (11)
C8	0.0260 (13)	0.0219 (13)	0.0269 (14)	0.0097 (10)	0.0032 (11)	0.0036 (11)
C9	0.0271 (13)	0.0237 (13)	0.0224 (13)	0.0114 (11)	0.0031 (10)	0.0042 (10)
C10	0.0275 (13)	0.0271 (14)	0.0272 (14)	0.0077 (11)	0.0060 (11)	0.0051 (11)
C11	0.0294 (14)	0.0248 (13)	0.0257 (14)	0.0067 (11)	0.0017 (11)	0.0000 (11)
C12	0.0299 (14)	0.0283 (14)	0.0233 (13)	0.0150 (11)	0.0038 (11)	0.0031 (11)
C13	0.0293 (14)	0.0277 (14)	0.0296 (15)	0.0109 (11)	0.0091 (12)	0.0069 (11)
C14	0.0276 (13)	0.0227 (13)	0.0271 (14)	0.0078 (11)	0.0037 (11)	0.0017 (11)
C15	0.0498 (19)	0.0454 (18)	0.0279 (17)	0.0184 (15)	0.0118 (15)	0.0014 (14)
C16	0.0398 (17)	0.0259 (15)	0.0469 (19)	0.0079 (13)	0.0182 (15)	0.0024 (14)
C17	0.0318 (16)	0.0336 (17)	0.076 (3)	0.0111 (14)	0.0151 (17)	0.0201 (18)

Geometric parameters (\AA , $^\circ$)

I—C12	2.094 (3)	C9—C10	1.396 (4)
I—O2 ⁱ	3.145 (2)	C9—C14	1.404 (4)

S—O2	1.486 (2)	C10—C11	1.387 (4)
S—C1	1.766 (3)	C10—H10	0.9300
S—C17	1.780 (4)	C11—C12	1.390 (4)
O1—C7	1.379 (3)	C11—H11	0.9300
O1—C8	1.382 (3)	C12—C13	1.401 (4)
C1—C8	1.364 (4)	C13—C14	1.378 (4)
C1—C2	1.452 (4)	C13—H13	0.9300
C2—C3	1.391 (4)	C14—H14	0.9300
C2—C7	1.396 (4)	C15—H15A	0.9600
C3—C4	1.385 (4)	C15—H15B	0.9600
C3—H3	0.9300	C15—H15C	0.9600
C4—C5	1.409 (4)	C16—H16A	0.9600
C4—C15	1.508 (4)	C16—H16B	0.9600
C5—C6	1.385 (4)	C16—H16C	0.9600
C5—H5	0.9300	C17—H17A	0.9600
C6—C7	1.385 (4)	C17—H17B	0.9600
C6—C16	1.505 (4)	C17—H17C	0.9600
C8—C9	1.459 (4)		
C12—I—O2 ⁱ	164.01 (9)	C11—C10—H10	119.9
O2—S—C1	107.81 (13)	C9—C10—H10	119.9
O2—S—C17	105.98 (17)	C10—C11—C12	120.1 (3)
C1—S—C17	99.04 (15)	C10—C11—H11	119.9
C7—O1—C8	106.5 (2)	C12—C11—H11	119.9
C8—C1—C2	107.5 (2)	C11—C12—C13	120.3 (3)
C8—C1—S	123.8 (2)	C11—C12—I	119.8 (2)
C2—C1—S	127.1 (2)	C13—C12—I	119.8 (2)
C3—C2—C7	119.0 (3)	C14—C13—C12	119.4 (3)
C3—C2—C1	136.5 (3)	C14—C13—H13	120.3
C7—C2—C1	104.5 (2)	C12—C13—H13	120.3
C4—C3—C2	119.1 (3)	C13—C14—C9	120.9 (3)
C4—C3—H3	120.5	C13—C14—H14	119.5
C2—C3—H3	120.5	C9—C14—H14	119.5
C3—C4—C5	119.4 (3)	C4—C15—H15A	109.5
C3—C4—C15	120.7 (3)	C4—C15—H15B	109.5
C5—C4—C15	119.9 (3)	H15A—C15—H15B	109.5
C6—C5—C4	123.4 (3)	C4—C15—H15C	109.5
C6—C5—H5	118.3	H15A—C15—H15C	109.5
C4—C5—H5	118.3	H15B—C15—H15C	109.5
C7—C6—C5	114.7 (3)	C6—C16—H16A	109.5
C7—C6—C16	122.2 (3)	C6—C16—H16B	109.5
C5—C6—C16	123.0 (3)	H16A—C16—H16B	109.5
O1—C7—C6	124.7 (3)	C6—C16—H16C	109.5
O1—C7—C2	110.9 (2)	H16A—C16—H16C	109.5
C6—C7—C2	124.4 (3)	H16B—C16—H16C	109.5
C1—C8—O1	110.5 (2)	S—C17—H17A	109.5
C1—C8—C9	134.5 (3)	S—C17—H17B	109.5
O1—C8—C9	115.0 (2)	H17A—C17—H17B	109.5
C10—C9—C14	119.2 (3)	S—C17—H17C	109.5

supplementary materials

C10—C9—C8	121.0 (3)	H17A—C17—H17C	109.5
C14—C9—C8	119.8 (2)	H17B—C17—H17C	109.5
C11—C10—C9	120.1 (3)		
O2—S—C1—C8	120.8 (3)	C1—C2—C7—O1	0.7 (3)
C17—S—C1—C8	-129.1 (3)	C3—C2—C7—C6	0.5 (4)
O2—S—C1—C2	-43.3 (3)	C1—C2—C7—C6	-178.1 (3)
C17—S—C1—C2	66.8 (3)	C2—C1—C8—O1	0.1 (3)
C8—C1—C2—C3	-178.7 (3)	S—C1—C8—O1	-166.67 (19)
S—C1—C2—C3	-12.5 (5)	C2—C1—C8—C9	179.4 (3)
C8—C1—C2—C7	-0.5 (3)	S—C1—C8—C9	12.6 (5)
S—C1—C2—C7	165.7 (2)	C7—O1—C8—C1	0.4 (3)
C7—C2—C3—C4	-0.3 (4)	C7—O1—C8—C9	-179.1 (2)
C1—C2—C3—C4	177.8 (3)	C1—C8—C9—C10	27.6 (5)
C2—C3—C4—C5	0.3 (4)	O1—C8—C9—C10	-153.1 (3)
C2—C3—C4—C15	-179.4 (3)	C1—C8—C9—C14	-153.4 (3)
C3—C4—C5—C6	-0.6 (5)	O1—C8—C9—C14	25.9 (4)
C15—C4—C5—C6	179.1 (3)	C14—C9—C10—C11	0.8 (4)
C4—C5—C6—C7	0.7 (4)	C8—C9—C10—C11	179.8 (3)
C4—C5—C6—C16	-178.9 (3)	C9—C10—C11—C12	-0.8 (4)
C8—O1—C7—C6	178.1 (3)	C10—C11—C12—C13	0.4 (4)
C8—O1—C7—C2	-0.7 (3)	C10—C11—C12—I	-175.5 (2)
C5—C6—C7—O1	-179.4 (3)	C11—C12—C13—C14	-0.1 (4)
C16—C6—C7—O1	0.2 (5)	I—C12—C13—C14	175.8 (2)
C5—C6—C7—C2	-0.7 (4)	C12—C13—C14—C9	0.2 (4)
C16—C6—C7—C2	178.9 (3)	C10—C9—C14—C13	-0.5 (4)
C3—C2—C7—O1	179.3 (2)	C8—C9—C14—C13	-179.5 (3)

Symmetry codes: (i) $x, y, z-1$.

Hydrogen-bond geometry (\AA , $^\circ$)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
C16—H16B \cdots O1	0.96	2.55	2.975 (4)	107
C16—H16A \cdots O2 ⁱⁱ	0.96	2.39	3.288 (4)	156
C17—H17B \cdots O1 ⁱⁱⁱ	0.96	2.51	3.422 (4)	159

Symmetry codes: (ii) $x, y+1, z$; (iii) $-x, -y+1, -z+1$.

Fig. 1

