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# Tetrakis(1,3,4,6,7,8-hexahydro-2H-pyrimido[1,2-a]pyrimidin-9-ido- $\kappa^2N^1, N^9$ )niobium(V) hexafluorido-phosphate

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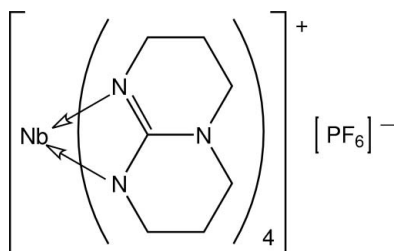
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 Key indicators: single-crystal X-ray study;  $T = 213$  K; mean  $\sigma(C-C) = 0.008$  Å;  $R$  factor = 0.050;  $wR$  factor = 0.145; data-to-parameter ratio = 14.9.

The title complex,  $[Nb(C_7H_{12}N_3)_4]PF_6$ , features chelating hpp anions (hpp is 1,3,4,6,7,8-hexahydro-2H-pyrimido[1,2-a]-pyrimidine) that define a distorted dodecahedral coordination geometry based on an  $N_8$  donor set. The Nb atom is situated on a site of symmetry  $\bar{4}$ , and the  $PF_6^-$  anion has crystallographic fourfold symmetry.

## Related literature

For background literature, see: Cotton *et al.* (1998, 2005). For related structures, see: Cotton *et al.* (2000); Coles & Hitchcock (2001).



## Experimental

## Crystal data

$[Nb(C_7H_{12}N_3)_4]PF_6$   
 $M_r = 790.66$   
 Tetragonal,  $P4/n$   
 $a = 13.531(6)$  Å  
 $c = 9.159(4)$  Å  
 $V = 1676.9(13)$  Å<sup>3</sup>

$Z = 2$   
 Mo  $K\alpha$  radiation  
 $\mu = 0.48$  mm<sup>-1</sup>  
 $T = 213(2)$  K  
 $0.20 \times 0.15 \times 0.10$  mm

## Data collection

Bruker SMART 1K CCD area-detector diffractometer  
 Absorption correction: multi-scan (SADABS; Sheldrick, 2004)  
 $T_{min} = 0.910$ ,  $T_{max} = 0.953$

10356 measured reflections  
 1655 independent reflections  
 1381 reflections with  $I > 2\sigma(I)$

## Refinement

$R[F^2 > 2\sigma(F^2)] = 0.050$   
 $wR(F^2) = 0.144$   
 $S = 1.05$   
 1655 reflections

111 parameters  
 H-atom parameters constrained  
 $\Delta\rho_{max} = 0.69$  e Å<sup>-3</sup>  
 $\Delta\rho_{min} = -0.39$  e Å<sup>-3</sup>

Data collection: SMART (Bruker, 2001); cell refinement: SAINT (Bruker, 2001); data reduction: SAINT; program(s) used to solve structure: SHELXTL (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: ORTEPII (Johnson, 1976); software used to prepare material for publication: SHELXL97.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: LX2067).

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**supplementary materials**

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## Tetrakis(1,3,4,6,7,8-hexahydro-2*H*-pyrimido[1,2-*a*]pyrimidin-9-ido- $\kappa^2N^1,N^9$ )niobium(V) hexafluoridophosphate

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### Comment

The title complex,  $[\text{Nb}(\text{hpp})_4][\text{PF}_6]$  (I), features a  $[\text{Nb}(\text{hpp})_4]^+$  cation, with the Nb atom located on a site of symmetry  $\bar{4}$ , and a  $[\text{PF}_6]^-$  anion, with fourfold symmetry; where hpp is 1,3,4,6,7,8-hexahydro-2*H*-pyrimido(1,2-*a*)pyrimidine. The Nb atom is chelated four hpp ligands and the  $N_8$  donor set defines an approximate dodecahedral coordination environment (Fig. 1).

The conformations of the N1- and N-2 containing six-membered rings is twisted chair. Such a binding mode as observed in (I) is uncommon for the hpp ligand, which normally acts as a bridging group in various paddlewheel complexes (Cotton *et al.*, 2005). A related example of hpp acting as a chelating ligand is  $[\text{Ta}(\text{hpp})_4][\text{Ta}(\text{CO})_6]$  (Cotton *et al.*, 2000). Both complexes were obtained by oxidizing the precursors  $\text{Nb}_2(\text{hpp})_4$  and  $[\text{Et}_4\text{N}][\text{Ta}(\text{CO})_6]$ , respectively. The chelating mode of hpp is also found in some Ti complexes (Coles & Hitchcock, 2001).

### Experimental

The title complex (I) was obtained unintentionally in an attempt to oxidize the paddlewheel complex  $\text{Nb}_2(\text{hpp})_4$  with  $[\text{Cp}_2\text{Fe}][\text{PF}_6]$  in  $\text{CH}_2\text{Cl}_2$ . X-ray quality crystals were obtained by slow diffusion of hexanes into a  $\text{CH}_2\text{Cl}_2$  solution of (I) at room temperature.

### Refinement

The H atoms were geometrically placed ( $\text{C}-\text{H} = 0.98 \text{ \AA}$ ) and refined as riding with  $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$ .

### Figures

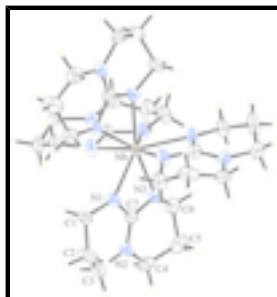


Fig. 1. Molecular structure of the cation in (I) showing the crystallographic numbering scheme. Displacement ellipsoids are shown at the 35% probability level. The Nb atom is located on a site of symmetry  $\bar{4}$ .

## Tetrakis(1,3,4,6,7,8-hexahydro-2H-pyrimido[1,2-a]pyrimidin-9-ido- $\kappa^2N^1,N^9$ )niobium(V) hexafluoridophosphate

### Crystal data

$[\text{Nb}(\text{C}_7\text{H}_{12}\text{N}_3)_4]\text{PF}_6$	$Z = 2$
$M_r = 790.66$	$F_{000} = 820$
Tetragonal, $P4/n$	$D_x = 1.566 \text{ Mg m}^{-3}$
Hall symbol: -P 4a	Mo $K\alpha$ radiation
$a = 13.531 (6) \text{ \AA}$	$\lambda = 0.71069 \text{ \AA}$
$b = 13.531 (6) \text{ \AA}$	Cell parameters from 10356 reflections
$c = 9.159 (4) \text{ \AA}$	$\theta = 2.1\text{--}27.5^\circ$
$\alpha = 90^\circ$	$\mu = 0.48 \text{ mm}^{-1}$
$\beta = 90^\circ$	$T = 213 (2) \text{ K}$
$\gamma = 90^\circ$	Block, yellow
$V = 1676.9 (13) \text{ \AA}^3$	$0.20 \times 0.15 \times 0.10 \text{ mm}$

### Data collection

Bruker SMART 1K CCD area-detector diffractometer	1655 independent reflections
Radiation source: fine-focus sealed tube	1381 reflections with $I > 2\sigma(I)$
Monochromator: graphite	$R_{\text{int}} = 0.025$
Detector resolution: $10 \text{ pixels mm}^{-1}$	$\theta_{\text{max}} = 26.0^\circ$
$T = 213(2) \text{ K}$	$\theta_{\text{min}} = 2.1^\circ$
$\omega$ and $\varphi$ scans	$h = -17 \rightarrow 12$
Absorption correction: multi-scan (SADABS; Sheldrick, 2004)	$k = -17 \rightarrow 16$
$T_{\text{min}} = 0.910$ , $T_{\text{max}} = 0.953$	$l = -10 \rightarrow 11$
10356 measured reflections	

### Refinement

Refinement on $F^2$	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.050$	H-atom parameters constrained
$wR(F^2) = 0.144$	$w = 1/[\sigma^2(F_o^2) + (0.067P)^2 + 3.4248P]$
$S = 1.05$	where $P = (F_o^2 + 2F_c^2)/3$
1655 reflections	$(\Delta/\sigma)_{\text{max}} < 0.001$
111 parameters	$\Delta\rho_{\text{max}} = 0.69 \text{ e \AA}^{-3}$
Primary atom site location: structure-invariant direct methods	$\Delta\rho_{\text{min}} = -0.39 \text{ e \AA}^{-3}$
	Extinction correction: none

*Special details*

**Geometry.** All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted R-factor wR and goodness of fit S are based on  $F^2$ , conventional R-factors R are based on F, with F set to zero for negative  $F^2$ . The threshold expression of  $F^2 > 2\text{sigma}(F^2)$  is used only for calculating R-factors(gt) etc. and is not relevant to the choice of reflections for refinement. R-factors based on  $F^2$  are statistically about twice as large as those based on F, and R- factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
Nb	0.7500	0.2500	0.5000	0.0388 (2)
N1	0.6548 (3)	0.1964 (3)	0.3195 (4)	0.0670 (11)
N2	0.6131 (3)	0.1782 (3)	0.5468 (4)	0.0677 (11)
N3	0.4885 (3)	0.1508 (3)	0.3731 (4)	0.0598 (9)
C1	0.6352 (4)	0.1994 (4)	0.1664 (5)	0.0730 (14)
H1A	0.6931	0.1747	0.1137	0.088*
H1B	0.6245	0.2682	0.1370	0.088*
C2	0.5483 (5)	0.1405 (7)	0.1244 (6)	0.116 (3)
H2A	0.5295	0.1587	0.0247	0.139*
H2B	0.5675	0.0706	0.1231	0.139*
C3	0.4618 (4)	0.1511 (4)	0.2179 (6)	0.0783 (15)
H3A	0.4158	0.0967	0.1985	0.094*
H3B	0.4281	0.2132	0.1945	0.094*
C4	0.4136 (3)	0.1333 (4)	0.4827 (6)	0.0694 (14)
H4A	0.3720	0.0777	0.4523	0.083*
H4B	0.3713	0.1919	0.4909	0.083*
C5	0.4580 (4)	0.1111 (4)	0.6279 (6)	0.0839 (18)
H5A	0.4830	0.0431	0.6279	0.101*
H5B	0.4067	0.1159	0.7032	0.101*
C6	0.5406 (4)	0.1802 (4)	0.6648 (5)	0.0745 (14)
H6A	0.5149	0.2474	0.6772	0.089*
H6B	0.5718	0.1599	0.7566	0.089*
C7	0.5805 (3)	0.1735 (3)	0.4103 (5)	0.0549 (10)
P1	0.2500	0.2500	0.9214 (2)	0.0475 (5)
F1	0.13862 (19)	0.2132 (2)	0.9219 (3)	0.0772 (9)
F2	0.2500	0.2500	0.7468 (5)	0.0695 (14)
F3	0.2500	0.2500	1.0963 (5)	0.0562 (11)

*Atomic displacement parameters ( $\text{\AA}^2$ )*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
Nb	0.0438 (3)	0.0438 (3)	0.0288 (4)	0.000	0.000	0.000

## supplementary materials

N1	0.056 (2)	0.107 (3)	0.0386 (18)	-0.013 (2)	-0.0009 (16)	-0.0025 (19)
N2	0.061 (2)	0.095 (3)	0.0472 (19)	-0.019 (2)	0.0110 (18)	-0.003 (2)
N3	0.051 (2)	0.067 (2)	0.062 (2)	0.0073 (17)	-0.0030 (17)	0.0021 (19)
C1	0.077 (3)	0.099 (4)	0.043 (2)	-0.014 (3)	-0.009 (2)	0.005 (2)
C2	0.084 (4)	0.217 (8)	0.047 (3)	-0.053 (5)	-0.014 (3)	0.001 (4)
C3	0.061 (3)	0.096 (4)	0.078 (3)	-0.001 (3)	-0.023 (3)	0.010 (3)
C4	0.048 (2)	0.055 (3)	0.105 (4)	-0.0021 (19)	0.012 (3)	-0.006 (3)
C5	0.090 (4)	0.072 (3)	0.089 (4)	-0.004 (3)	0.054 (3)	0.000 (3)
C6	0.076 (3)	0.101 (4)	0.047 (2)	-0.013 (3)	0.018 (2)	0.000 (3)
C7	0.055 (2)	0.065 (3)	0.045 (2)	-0.005 (2)	0.0064 (18)	0.0035 (19)
P1	0.0530 (7)	0.0530 (7)	0.0365 (10)	0.000	0.000	0.000
F1	0.0582 (15)	0.112 (2)	0.0619 (17)	-0.0192 (15)	0.0028 (13)	-0.0233 (16)
F2	0.087 (2)	0.087 (2)	0.035 (2)	0.000	0.000	0.000
F3	0.0660 (18)	0.0660 (18)	0.036 (2)	0.000	0.000	0.000

### Geometric parameters ( $\text{\AA}$ , $^\circ$ )

Nb—N2	2.135 (4)	C2—H2B	0.9800
Nb—N2 <sup>i</sup>	2.135 (4)	C3—H3A	0.9800
Nb—N1	2.218 (4)	C3—H3B	0.9800
Nb—N1 <sup>i</sup>	2.218 (4)	C4—C5	1.490 (8)
Nb—C7	2.648 (4)	C4—H4A	0.9800
Nb—C7 <sup>i</sup>	2.648 (4)	C4—H4B	0.9800
N1—C7	1.341 (5)	C5—C6	1.496 (8)
N1—C1	1.428 (5)	C5—H5A	0.9800
N2—C7	1.328 (6)	C5—H5B	0.9800
N2—C6	1.460 (6)	C6—H6A	0.9800
N3—C7	1.326 (5)	C6—H6B	0.9800
N3—C4	1.447 (6)	P1—F1	1.587 (3)
N3—C3	1.466 (6)	P1—F1 <sup>ii</sup>	1.587 (3)
C1—C2	1.472 (7)	P1—F1 <sup>iii</sup>	1.587 (3)
C1—H1A	0.9800	P1—F1 <sup>iv</sup>	1.587 (3)
C1—H1B	0.9800	P1—F2	1.600 (5)
C2—C3	1.457 (8)	P1—F3	1.602 (5)
C2—H2A	0.9800		
N2—Nb—N2 <sup>i</sup>	156.8 (2)	C7—N1—C1	118.4 (4)
N2—Nb—N2 <sup>v</sup>	92.31 (4)	C7—N1—Nb	92.8 (3)
N2 <sup>i</sup> —Nb—N2 <sup>v</sup>	92.31 (4)	C1—N1—Nb	146.2 (3)
N2—Nb—N2 <sup>vi</sup>	92.31 (4)	C7—N2—C6	118.3 (4)
N2 <sup>i</sup> —Nb—N2 <sup>vi</sup>	92.31 (4)	C7—N2—Nb	97.0 (3)
N2 <sup>v</sup> —Nb—N2 <sup>vi</sup>	156.8 (2)	C6—N2—Nb	136.4 (3)
N2—Nb—N1	59.80 (14)	C7—N3—C4	121.1 (4)
N2 <sup>i</sup> —Nb—N1	143.38 (14)	C7—N3—C3	118.7 (4)
N2 <sup>v</sup> —Nb—N1	80.23 (16)	C4—N3—C3	120.0 (4)
N2 <sup>vi</sup> —Nb—N1	82.54 (17)	N1—C1—C2	112.9 (4)
N2—Nb—N1 <sup>v</sup>	82.54 (17)	N1—C1—H1A	109.0

N2 <sup>i</sup> —Nb—N1 <sup>v</sup>	80.23 (16)	C2—C1—H1A	109.0
N2 <sup>v</sup> —Nb—N1 <sup>v</sup>	59.80 (14)	N1—C1—H1B	109.0
N2 <sup>vi</sup> —Nb—N1 <sup>v</sup>	143.38 (14)	C2—C1—H1B	109.0
N1—Nb—N1 <sup>v</sup>	123.75 (12)	H1A—C1—H1B	107.8
N2—Nb—N1 <sup>vi</sup>	80.23 (16)	C3—C2—C1	115.7 (6)
N2 <sup>i</sup> —Nb—N1 <sup>vi</sup>	82.54 (17)	C3—C2—H2A	108.3
N2 <sup>v</sup> —Nb—N1 <sup>vi</sup>	143.38 (14)	C1—C2—H2A	108.3
N2 <sup>vi</sup> —Nb—N1 <sup>vi</sup>	59.80 (14)	C3—C2—H2B	108.3
N1—Nb—N1 <sup>vi</sup>	123.75 (12)	C1—C2—H2B	108.3
N1 <sup>v</sup> —Nb—N1 <sup>vi</sup>	83.6 (2)	H2A—C2—H2B	107.4
N2—Nb—N1 <sup>i</sup>	143.38 (14)	C2—C3—N3	111.8 (4)
N2 <sup>i</sup> —Nb—N1 <sup>i</sup>	59.80 (14)	C2—C3—H3A	109.3
N2 <sup>v</sup> —Nb—N1 <sup>i</sup>	82.54 (17)	N3—C3—H3A	109.3
N2 <sup>vi</sup> —Nb—N1 <sup>i</sup>	80.23 (16)	C2—C3—H3B	109.3
N1—Nb—N1 <sup>i</sup>	83.6 (2)	N3—C3—H3B	109.3
N1 <sup>v</sup> —Nb—N1 <sup>i</sup>	123.75 (12)	H3A—C3—H3B	107.9
N1 <sup>vi</sup> —Nb—N1 <sup>i</sup>	123.75 (12)	N3—C4—C5	111.7 (4)
N2—Nb—C7	29.86 (14)	N3—C4—H4A	109.3
N2 <sup>i</sup> —Nb—C7	172.73 (14)	C5—C4—H4A	109.3
N2 <sup>v</sup> —Nb—C7	89.57 (16)	N3—C4—H4B	109.3
N2 <sup>vi</sup> —Nb—C7	83.26 (16)	C5—C4—H4B	109.3
N1—Nb—C7	30.39 (13)	H4A—C4—H4B	107.9
N1 <sup>v</sup> —Nb—C7	106.74 (15)	C4—C5—C6	112.1 (4)
N1 <sup>vi</sup> —Nb—C7	100.07 (15)	C4—C5—H5A	109.2
N1 <sup>i</sup> —Nb—C7	113.57 (13)	C6—C5—H5A	109.2
N2—Nb—C7 <sup>vi</sup>	89.57 (15)	C4—C5—H5B	109.2
N2 <sup>i</sup> —Nb—C7 <sup>vi</sup>	83.26 (16)	C6—C5—H5B	109.2
N2 <sup>v</sup> —Nb—C7 <sup>vi</sup>	172.73 (14)	H5A—C5—H5B	107.9
N2 <sup>vi</sup> —Nb—C7 <sup>vi</sup>	29.86 (14)	N2—C6—C5	108.9 (4)
N1—Nb—C7 <sup>vi</sup>	106.74 (15)	N2—C6—H6A	109.9
N1 <sup>v</sup> —Nb—C7 <sup>vi</sup>	113.57 (13)	C5—C6—H6A	109.9
N1 <sup>vi</sup> —Nb—C7 <sup>vi</sup>	30.39 (13)	N2—C6—H6B	109.9
N1 <sup>i</sup> —Nb—C7 <sup>vi</sup>	100.07 (15)	C5—C6—H6B	109.9
C7—Nb—C7 <sup>vi</sup>	95.53 (5)	H6A—C6—H6B	108.3
N2—Nb—C7 <sup>v</sup>	83.26 (16)	N3—C7—N2	124.4 (4)
N2 <sup>i</sup> —Nb—C7 <sup>v</sup>	89.57 (15)	N3—C7—N1	126.7 (4)
N2 <sup>v</sup> —Nb—C7 <sup>v</sup>	29.86 (14)	N2—C7—N1	108.8 (4)
N2 <sup>vi</sup> —Nb—C7 <sup>v</sup>	172.73 (14)	N3—C7—Nb	169.6 (3)
N1—Nb—C7 <sup>v</sup>	100.07 (15)	N2—C7—Nb	53.2 (2)
N1 <sup>v</sup> —Nb—C7 <sup>v</sup>	30.39 (13)	N1—C7—Nb	56.8 (2)
N1 <sup>vi</sup> —Nb—C7 <sup>v</sup>	113.57 (13)	F1—P1—F1 <sup>ii</sup>	90.000 (2)

## supplementary materials

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N1 <sup>i</sup> —Nb—C7 <sup>v</sup>	106.74 (15)	F1—P1—F1 <sup>iii</sup>	179.7 (2)
C7—Nb—C7 <sup>v</sup>	95.53 (5)	F1 <sup>ii</sup> —P1—F1 <sup>iii</sup>	90.000 (2)
C7 <sup>vi</sup> —Nb—C7 <sup>v</sup>	143.83 (18)	F1—P1—F1 <sup>iv</sup>	90.000 (1)
N2—Nb—C7 <sup>i</sup>	172.73 (14)	F1 <sup>ii</sup> —P1—F1 <sup>iv</sup>	179.7 (2)
N2 <sup>i</sup> —Nb—C7 <sup>i</sup>	29.86 (14)	F1 <sup>iii</sup> —P1—F1 <sup>iv</sup>	90.000 (2)
N2 <sup>v</sup> —Nb—C7 <sup>i</sup>	83.26 (16)	F1—P1—F2	90.16 (12)
N2 <sup>vi</sup> —Nb—C7 <sup>i</sup>	89.57 (16)	F1 <sup>ii</sup> —P1—F2	90.16 (12)
N1—Nb—C7 <sup>i</sup>	113.57 (13)	F1 <sup>iii</sup> —P1—F2	90.16 (12)
N1 <sup>v</sup> —Nb—C7 <sup>i</sup>	100.07 (15)	F1 <sup>iv</sup> —P1—F2	90.16 (12)
N1 <sup>vi</sup> —Nb—C7 <sup>i</sup>	106.74 (15)	F1—P1—F3	89.84 (12)
N1 <sup>i</sup> —Nb—C7 <sup>i</sup>	30.39 (13)	F1 <sup>ii</sup> —P1—F3	89.84 (12)
C7—Nb—C7 <sup>i</sup>	143.83 (18)	F1 <sup>iii</sup> —P1—F3	89.84 (12)
C7 <sup>vi</sup> —Nb—C7 <sup>i</sup>	95.53 (5)	F1 <sup>iv</sup> —P1—F3	89.84 (12)
C7 <sup>v</sup> —Nb—C7 <sup>i</sup>	95.53 (5)	F2—P1—F3	180.000 (2)
N2—Nb—N1—C7	7.7 (3)	N3—C4—C5—C6	45.0 (6)
N2 <sup>i</sup> —Nb—N1—C7	-173.2 (3)	C7—N2—C6—C5	39.2 (7)
N2 <sup>v</sup> —Nb—N1—C7	106.2 (3)	Nb—N2—C6—C5	179.1 (4)
N2 <sup>vi</sup> —Nb—N1—C7	-89.4 (3)	C4—C5—C6—N2	-54.6 (6)
N1 <sup>v</sup> —Nb—N1—C7	63.0 (4)	C4—N3—C7—N2	3.0 (7)
N1 <sup>vi</sup> —Nb—N1—C7	-43.6 (4)	C3—N3—C7—N2	178.2 (5)
N1 <sup>i</sup> —Nb—N1—C7	-170.3 (4)	C4—N3—C7—N1	-176.2 (5)
C7 <sup>vi</sup> —Nb—N1—C7	-71.7 (3)	C3—N3—C7—N1	-1.0 (7)
C7 <sup>v</sup> —Nb—N1—C7	83.7 (2)	C4—N3—C7—Nb	-70.1 (19)
C7 <sup>i</sup> —Nb—N1—C7	-175.65 (18)	C3—N3—C7—Nb	105.1 (17)
N2—Nb—N1—C1	165.9 (7)	C6—N2—C7—N3	-14.1 (7)
N2 <sup>i</sup> —Nb—N1—C1	-15.0 (8)	Nb—N2—C7—N3	-167.5 (4)
N2 <sup>v</sup> —Nb—N1—C1	-95.6 (7)	C6—N2—C7—N1	165.2 (5)
N2 <sup>vi</sup> —Nb—N1—C1	68.8 (7)	Nb—N2—C7—N1	11.8 (4)
N1 <sup>v</sup> —Nb—N1—C1	-138.8 (6)	C6—N2—C7—Nb	153.4 (5)
N1 <sup>vi</sup> —Nb—N1—C1	114.6 (6)	C1—N1—C7—N3	1.6 (8)
N1 <sup>i</sup> —Nb—N1—C1	-12.1 (6)	Nb—N1—C7—N3	168.0 (4)
C7—Nb—N1—C1	158.2 (9)	C1—N1—C7—N2	-177.7 (5)
C7 <sup>vi</sup> —Nb—N1—C1	86.5 (7)	Nb—N1—C7—N2	-11.3 (4)
C7 <sup>v</sup> —Nb—N1—C1	-118.0 (7)	C1—N1—C7—Nb	-166.4 (5)
C7 <sup>i</sup> —Nb—N1—C1	-17.4 (7)	N2—Nb—C7—N3	80.4 (17)
N2 <sup>i</sup> —Nb—N2—C7	173.5 (3)	N2 <sup>v</sup> —Nb—C7—N3	175.8 (17)
N2 <sup>v</sup> —Nb—N2—C7	-85.1 (3)	N2 <sup>vi</sup> —Nb—C7—N3	-26.3 (17)
N2 <sup>vi</sup> —Nb—N2—C7	72.2 (3)	N1—Nb—C7—N3	-113.0 (18)
N1—Nb—N2—C7	-7.8 (3)	N1 <sup>v</sup> —Nb—C7—N3	117.6 (17)
N1 <sup>v</sup> —Nb—N2—C7	-144.2 (3)	N1 <sup>vi</sup> —Nb—C7—N3	31.3 (18)
N1 <sup>vi</sup> —Nb—N2—C7	131.0 (3)	N1 <sup>i</sup> —Nb—C7—N3	-102.5 (17)

N1 <sup>i</sup> —Nb—N2—C7	-4.5 (5)	C7 <sup>vi</sup> —Nb—C7—N3	1.0 (17)
C7 <sup>vi</sup> —Nb—N2—C7	101.9 (3)	C7 <sup>v</sup> —Nb—C7—N3	146.5 (18)
C7 <sup>v</sup> —Nb—N2—C7	-113.6 (3)	C7 <sup>i</sup> —Nb—C7—N3	-106.3 (18)
N2 <sup>i</sup> —Nb—N2—C6	28.4 (5)	N2 <sup>v</sup> —Nb—C7—N2	95.4 (3)
N2 <sup>v</sup> —Nb—N2—C6	129.7 (6)	N2 <sup>vi</sup> —Nb—C7—N2	-106.7 (3)
N2 <sup>vi</sup> —Nb—N2—C6	-73.0 (5)	N1—Nb—C7—N2	166.6 (5)
N1—Nb—N2—C6	-153.0 (6)	N1 <sup>v</sup> —Nb—C7—N2	37.2 (3)
N1 <sup>v</sup> —Nb—N2—C6	70.6 (5)	N1 <sup>vi</sup> —Nb—C7—N2	-49.1 (3)
N1 <sup>vi</sup> —Nb—N2—C6	-14.2 (5)	N1 <sup>i</sup> —Nb—C7—N2	177.1 (3)
N1 <sup>i</sup> —Nb—N2—C6	-149.7 (5)	C7 <sup>v</sup> —Nb—C7—N2	66.1 (3)
C7—Nb—N2—C6	-145.2 (7)	C7 <sup>i</sup> —Nb—C7—N2	173.3 (3)
C7 <sup>vi</sup> —Nb—N2—C6	-43.3 (5)	N2—Nb—C7—N1	-166.6 (5)
C7 <sup>v</sup> —Nb—N2—C6	101.2 (5)	N2 <sup>v</sup> —Nb—C7—N1	-71.2 (3)
C7—N1—C1—C2	-22.7 (8)	N2 <sup>vi</sup> —Nb—C7—N1	86.7 (3)
Nb—N1—C1—C2	-177.8 (5)	N1 <sup>v</sup> —Nb—C7—N1	-129.3 (3)
N1—C1—C2—C3	44.3 (9)	N1 <sup>vi</sup> —Nb—C7—N1	144.4 (3)
C1—C2—C3—N3	-43.2 (8)	N1 <sup>i</sup> —Nb—C7—N1	10.5 (4)
C7—N3—C3—C2	21.6 (8)	C7 <sup>vi</sup> —Nb—C7—N1	114.0 (3)
C4—N3—C3—C2	-163.1 (5)	C7 <sup>v</sup> —Nb—C7—N1	-100.5 (3)
C7—N3—C4—C5	-19.0 (6)	C7 <sup>i</sup> —Nb—C7—N1	6.8 (3)
C3—N3—C4—C5	165.8 (5)		

Symmetry codes: (i)  $-x+3/2, -y+1/2, z$ ; (ii)  $-y+1/2, x, z$ ; (iii)  $-x+1/2, -y+1/2, z$ ; (iv)  $y, -x+1/2, z$ ; (v)  $-y+1, x-1/2, -z+1$ ; (vi)  $y+1/2, -x+1, -z+1$ .

