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## Structure Reports

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## Methyl 2,5-dichlorobenzoate

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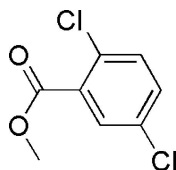
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Key indicators: single-crystal X-ray study;  $T = 150$  K; mean  $\sigma(\text{C}-\text{C}) = 0.005$  Å;  $R$  factor = 0.054;  $wR$  factor = 0.146; data-to-parameter ratio = 16.9.

In the molecule of the title compound,  $\text{C}_8\text{H}_6\text{Cl}_2\text{O}_2$ , the benzene ring is oriented with respect to the planar ester group at a dihedral angle of  $39.22(3)^\circ$ .

## Related literature

For general background, see: Zheng *et al.* (2003); Al-Talib *et al.* (1990); Yousif *et al.* (1986); Ahmad *et al.* (2001); Al-Soud *et al.* (2004); El-Emam *et al.* (2004); Weinstock *et al.* (1991). For a description of the Cambridge Structural Database, see: Allen (2002); and of *MOGUL*, see: Bruno *et al.* (2004).



## Experimental

## Crystal data

$\text{C}_8\text{H}_6\text{Cl}_2\text{O}_2$   
 $M_r = 205.03$   
 Triclinic,  $P\bar{1}$   
 $a = 3.8452(3)$  Å  
 $b = 7.0158(4)$  Å  
 $c = 15.8510(10)$  Å  
 $\alpha = 77.189(6)^\circ$   
 $\beta = 89.130(7)^\circ$

$\gamma = 83.741(5)^\circ$   
 $V = 414.46(5)$  Å<sup>3</sup>  
 $Z = 2$   
 Mo  $K\alpha$  radiation  
 $\mu = 0.73$  mm<sup>-1</sup>  
 $T = 150(1)$  K  
 $0.68 \times 0.11 \times 0.06$  mm

## Data collection

Bruker–Nonius Kappa CCD area-detector diffractometer  
 Absorption correction: Gaussian (Coppens, 1970)  
 $T_{\min} = 0.864$ ,  $T_{\max} = 0.971$

5966 measured reflections  
 1840 independent reflections  
 1455 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.110$

## Refinement

$R[F^2 > 2\sigma(F^2)] = 0.053$   
 $wR(F^2) = 0.146$   
 $S = 1.09$   
 1840 reflections

109 parameters  
 H-atom parameters constrained  
 $\Delta\rho_{\max} = 0.44$  e Å<sup>-3</sup>  
 $\Delta\rho_{\min} = -0.57$  e Å<sup>-3</sup>

Data collection: *COLLECT* (Hooft, 1998) and *DENZO* (Otwinowski & Minor, 1997); cell refinement: *DIRAX/LSQ* (Duisenberg, 1992); data reduction: *EvalCCD* (Duisenberg, 1992); program(s) used to solve structure: *SIR92* (Altomare *et al.*, 1994); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *PLATON* (Spek, 2003); software used to prepare material for publication: *SHELXL97*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HK2533).

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**supplementary materials**

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## Methyl 2,5-dichlorobenzoate

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### Comment

The title compound is a lachrymator and a drug intermediate. Methyl 2,5-dichlorobenzoate is widely employed in synthetic organic chemistry for example, 2,5-dichlorobenzohydrazide, 2,5-disubstituted-1,3,4-oxadiazoles (Zheng *et al.*, 2003; Al-Talib *et al.*, 1990) and 5-substituted-2-mercapto-1,3,4-oxadiazoles (Yousif *et al.*, 1986; Ahmad *et al.*, 2001; Al-Soud *et al.*, 2004; El-Emam *et al.*, 2004). In addition, methyl 4-(bromomethyl)benzoate has been used in the synthesis of 1-(carboxybenzyl)imidazole-5-acrylic acids, which are potent and selective angiotensin II receptor antagonists (Weinstock *et al.*, 1991). In view of the versatility of these compounds, we have synthesized the title compound, and report herein its crystal structure.

In the molecule of the title compound, (Fig. 1), the bond lengths and angles are generally within normal ranges (Cambridge Structural Database, Version 5.28, November 2006; Mogul Version 1.1; Allen, 2002, Bruno *et al.*, 2004). The benzene ring (C1-C6) is oriented with respect to the planar ester group (O1/O2/C1/C7/C8) at a dihedral angle of 39.22 (3)°.

### Experimental

For the preparation of the title compound, the mixture of 2,5-dichlorobenzoic acid (2.05 g, 10 mmol) and absolute methanol (50 ml) in the presence of a few drops of sulphuric acid was refluxed for 5 h. The excess of solvent was removed by distillation. The solid residue was filtered off, washed with water and recrystallized from ethanol (30%) to give the title compound (yield; 88%, m.p. 319-321 K). Suitable single crystals of the title compound were obtained by slow evaporation of an ethanol solution at room temperature.

### Refinement

H atoms were positioned geometrically, with C-H = 0.93 and 0.96 Å for aromatic and methyl H, respectively, and constrained to ride on their parent atoms with  $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$ .

### Figures

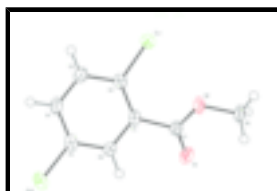


Fig. 1. The molecular structure of the title molecule, with the atom-numbering scheme. Displacement ellipsoids are drawn at the 50% probability level.



Fig. 2. Reaction scheme.

## methyl 2,5-dichlorobenzoate

### Crystal data

$C_8H_6Cl_2O_2$	$Z = 2$
$M_r = 205.03$	$F_{000} = 208$
Triclinic, $P\bar{1}$	$D_x = 1.643 \text{ Mg m}^{-3}$
Hall symbol: -P 1	Melting point: 319(2) K
$a = 3.8452 (3) \text{ \AA}$	Mo $K\alpha$ radiation
$b = 7.0158 (4) \text{ \AA}$	$\lambda = 0.71073 \text{ \AA}$
$c = 15.8510 (10) \text{ \AA}$	Cell parameters from 6024 reflections
$\alpha = 77.189 (6)^\circ$	$\theta = 1\text{--}27.5^\circ$
$\beta = 89.130 (7)^\circ$	$\mu = 0.73 \text{ mm}^{-1}$
$\gamma = 83.741 (5)^\circ$	$T = 150 (1) \text{ K}$
$V = 414.46 (5) \text{ \AA}^3$	Needle, colorless
	$0.68 \times 0.11 \times 0.06 \text{ mm}$

### Data collection

Bruker–Nonius KappaCCD area-detector diffractometer	1840 independent reflections
Radiation source: fine-focus sealed tube	1455 reflections with $I > 2\sigma(I)$
Monochromator: graphite	$R_{\text{int}} = 0.110$
Detector resolution: 9.091 pixels $\text{mm}^{-1}$	$\theta_{\text{max}} = 27.5^\circ$
$T = 150(1) \text{ K}$	$\theta_{\text{min}} = 2.6^\circ$
$\varphi$ and $\omega$ scans	$h = -4 \rightarrow 4$
Absorption correction: Gaussian (Coppens, 1970)	$k = -9 \rightarrow 8$
$T_{\text{min}} = 0.864$ , $T_{\text{max}} = 0.971$	$l = -20 \rightarrow 20$
5966 measured reflections	

### Refinement

Refinement on $F^2$	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.053$	H-atom parameters constrained
$wR(F^2) = 0.146$	$w = 1/[\sigma^2(F_o^2) + (0.0527P)^2 + 0.5264P]$
$S = 1.10$	where $P = (F_o^2 + 2F_c^2)/3$
1840 reflections	$(\Delta/\sigma)_{\text{max}} < 0.001$
109 parameters	$\Delta\rho_{\text{max}} = 0.44 \text{ e \AA}^{-3}$
Primary atom site location: structure-invariant direct methods	$\Delta\rho_{\text{min}} = -0.57 \text{ e \AA}^{-3}$
	Extinction correction: none

*Special details*

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
Cl1	0.6965 (2)	0.52269 (12)	0.70713 (5)	0.0309 (2)
Cl2	1.3585 (2)	-0.28182 (13)	0.92310 (5)	0.0361 (3)
O1	0.6439 (7)	-0.0150 (4)	0.62680 (14)	0.0347 (6)
O2	0.8542 (6)	0.2744 (4)	0.57949 (13)	0.0288 (5)
C1	0.9066 (8)	0.1286 (5)	0.72871 (18)	0.0220 (6)
C2	0.8841 (8)	0.2952 (5)	0.76262 (19)	0.0235 (6)
C3	1.0100 (9)	0.2857 (5)	0.8458 (2)	0.0278 (7)
H3	0.9954	0.3983	0.8680	0.033*
C4	1.1574 (9)	0.1080 (5)	0.89485 (19)	0.0286 (7)
H4	1.2435	0.1006	0.9501	0.034*
C5	1.1755 (8)	-0.0580 (5)	0.8613 (2)	0.0257 (7)
C6	1.0489 (8)	-0.0510 (5)	0.77933 (19)	0.0245 (6)
H6	1.0581	-0.1648	0.7581	0.029*
C7	0.7825 (8)	0.1206 (5)	0.64006 (18)	0.0229 (6)
C8	0.7421 (9)	0.2749 (5)	0.49289 (19)	0.0294 (7)
H8A	0.8509	0.1593	0.4761	0.035*
H8B	0.8089	0.3898	0.4536	0.035*
H8C	0.4924	0.2757	0.4914	0.035*

*Atomic displacement parameters ( $\text{\AA}^2$ )*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
Cl1	0.0411 (5)	0.0249 (4)	0.0250 (4)	0.0029 (3)	-0.0044 (3)	-0.0047 (3)
Cl2	0.0429 (5)	0.0328 (5)	0.0276 (4)	0.0027 (4)	-0.0115 (3)	0.0014 (3)
O1	0.0496 (16)	0.0344 (14)	0.0231 (11)	-0.0124 (11)	-0.0066 (10)	-0.0089 (10)
O2	0.0368 (13)	0.0354 (13)	0.0137 (10)	-0.0082 (10)	-0.0056 (8)	-0.0017 (9)
C1	0.0214 (15)	0.0287 (16)	0.0156 (13)	-0.0036 (12)	-0.0005 (10)	-0.0039 (11)
C2	0.0233 (15)	0.0256 (16)	0.0205 (14)	-0.0043 (12)	-0.0010 (11)	-0.0021 (12)
C3	0.0331 (18)	0.0295 (18)	0.0230 (15)	-0.0050 (13)	-0.0007 (12)	-0.0096 (13)
C4	0.0336 (18)	0.0348 (18)	0.0172 (14)	-0.0055 (14)	-0.0045 (12)	-0.0042 (12)
C5	0.0233 (16)	0.0303 (17)	0.0206 (14)	-0.0028 (12)	-0.0014 (11)	0.0003 (12)
C6	0.0302 (17)	0.0236 (16)	0.0205 (14)	-0.0008 (12)	-0.0026 (11)	-0.0073 (12)

## supplementary materials

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C7	0.0257 (15)	0.0265 (16)	0.0166 (13)	0.0004 (12)	-0.0013 (11)	-0.0062 (11)
C8	0.0367 (19)	0.0368 (19)	0.0141 (13)	0.0003 (14)	-0.0047 (12)	-0.0061 (12)

### *Geometric parameters (Å, °)*

C11—C2	1.728 (3)	C3—H3	0.9300
C12—C5	1.737 (3)	C4—C3	1.382 (5)
O1—C7	1.199 (4)	C4—C5	1.378 (5)
O2—C7	1.326 (4)	C4—H4	0.9300
O2—C8	1.444 (3)	C5—C6	1.384 (4)
C1—C2	1.385 (5)	C6—H6	0.9301
C1—C6	1.395 (4)	C8—H8A	0.9600
C1—C7	1.505 (4)	C8—H8B	0.9600
C2—C3	1.396 (4)	C8—H8C	0.9600
C7—O2—C8	115.5 (3)	C4—C5—C12	119.7 (2)
C2—C1—C6	119.3 (3)	C6—C5—C12	118.9 (3)
C2—C1—C7	125.6 (3)	C5—C6—C1	119.3 (3)
C6—C1—C7	115.1 (3)	C5—C6—H6	120.4
C1—C2—C11	123.1 (2)	C1—C6—H6	120.2
C1—C2—C3	120.7 (3)	O1—C7—O2	124.7 (3)
C3—C2—C11	116.3 (3)	O1—C7—C1	122.4 (3)
C4—C3—C2	119.7 (3)	O2—C7—C1	112.9 (3)
C4—C3—H3	120.1	O2—C8—H8A	109.4
C2—C3—H3	120.3	O2—C8—H8B	109.5
C5—C4—C3	119.5 (3)	H8A—C8—H8B	109.5
C5—C4—H4	120.3	O2—C8—H8C	109.5
C3—C4—H4	120.2	H8A—C8—H8C	109.5
C4—C5—C6	121.4 (3)	H8B—C8—H8C	109.5

Fig. 1

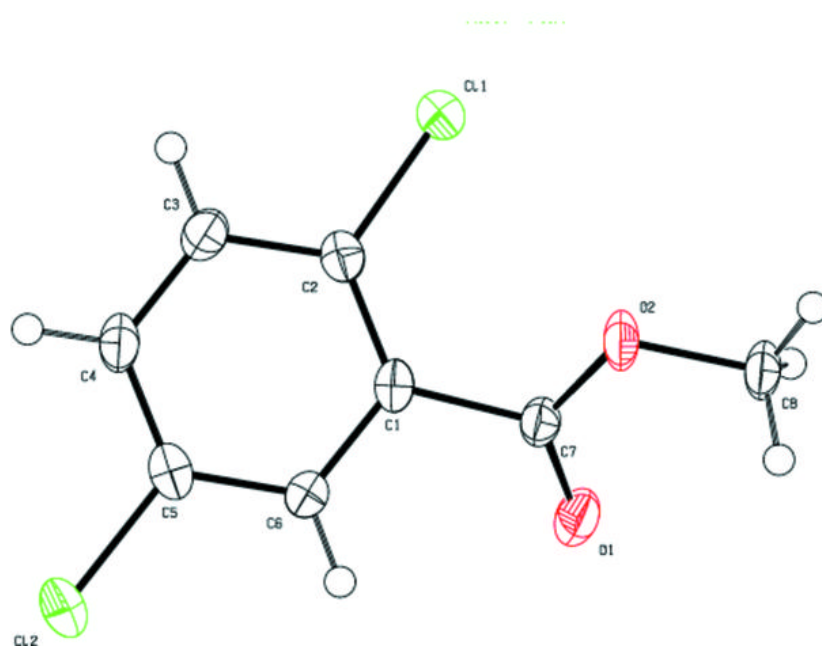


Fig. 2

