

Acta Crystallographica Section E

Structure Reports

Online

ISSN 1600-5368

N-Phenylpyridine-2-carbamide

Yu-Guo Zhuang,* Hua-Jiang Jiang, Zhi Hong and Fang-Li Qiu

School of Pharmaceutical and Chemical Engineering, Taizhou University, Linhai 317000, People's Republic of China

Correspondence e-mail: hieagle@126.com

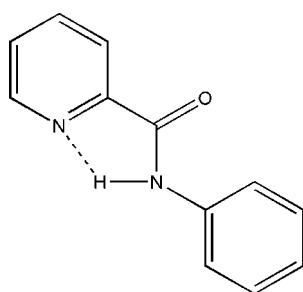
Received 1 September 2008; accepted 4 September 2008

Key indicators: single-crystal X-ray study; $T = 296$ K; mean $\sigma(\text{C}-\text{C}) = 0.003$ Å; R factor = 0.030; wR factor = 0.075; data-to-parameter ratio = 8.5.

In the title compound, $\text{C}_{12}\text{H}_{10}\text{N}_2\text{O}$, the dihedral angle between the pyridine ring system and the phenyl ring is $1.8(1)^\circ$. There is an intramolecular $\text{N}-\text{H}\cdots\text{N}$ hydrogen bond between the pyridine N atom and the amide NH function.

Related literature

For general background, see: Sousa & Filgueiras (1990); Gomes *et al.* (2007); Morsali *et al.* (2003); Jacob & Mukherjee (2006); Marumoto *et al.* (1981); Piatnitski & Kiselyov (2004). For related structures, see: Qi *et al.* (2003); Zhang *et al.* (2006); Yin *et al.* (2007). For the synthesis, see: Chan *et al.* (2004).



Experimental

Crystal data

 $\text{C}_{12}\text{H}_{10}\text{N}_2\text{O}$ $M_r = 198.22$ Monoclinic, Pn $a = 5.7469(2)$ Å $b = 6.2382(2)$ Å $c = 14.0158(3)$ Å $\beta = 94.752(2)^\circ$ $V = 500.74(3)$ Å³ $Z = 2$ Mo $K\alpha$ radiation $\mu = 0.09$ mm⁻¹ $T = 296(2)$ K $0.15 \times 0.14 \times 0.09$ mm

Data collection

Bruker APEXII CCD diffractometer
Absorption correction: multi-scan (SADABS; Bruker, 2004)
 $T_{\min} = 0.968$, $T_{\max} = 0.992$

5000 measured reflections
1162 independent reflections
1036 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.017$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.030$
 $wR(F^2) = 0.075$
 $S = 1.00$
1162 reflections

137 parameters
H-atom parameters constrained
 $\Delta\rho_{\text{max}} = 0.11$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.12$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{N1}-\text{H101}\cdots\text{N2}$	0.86	2.28	2.697 (2)	110

Data collection: APEX2 (Bruker, 2004); cell refinement: SAINT-Plus (Bruker, 2004); data reduction: SAINT-Plus; program(s) used to solve structure: SHELXL97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: ORTEP-3 for Windows (Farrugia, 1997); software used to prepare material for publication: WinGX (Farrugia, 1999).

The authors are grateful for financial support from the Key Discipline Open Foundation of Zhejiang University of Technology (grant No. 20080604). The authors thank Mr Jian-Ming Gu (Testing and Analysis Center, Zhejiang University) for guidance in the structure analysis.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: IM2083).

References

- Bruker (2004). APEX2, SAINT-Plus and SADABS. Bruker AXS Inc., Madison, Wisconsin, USA.
- Chan, A. S. C., Qi, J. Y., Pai, C. C., Li, X. J., Deng, L. S., Li, W. Z. & Hu, J. Y. (2004). US Patent 6 680 385.
- Farrugia, L. J. (1997). *J. Appl. Cryst.* **30**, 565.
- Farrugia, L. J. (1999). *J. Appl. Cryst.* **32**, 837–838.
- Gomes, L., Low, J. N., Valente, M. A. D. C., Freire, C. & Castro, B. (2007). *Acta Cryst.* **C63**, m293–m296.
- Jacob, W. & Mukherjee, R. (2006). *Inorg. Chim. Acta*, **359**, 4565–4573.
- Marumoto, R., Shunsuke, S. & Masao, T. (1981). Eur. Patent EP 38 161.
- Morsali, A., Ramazani, A. & Mahjoub, A. R. (2003). *J. Coord. Chem.* **56**, 1555–1566.
- Piatnitski, E. & Kiselyov, A. S. (2004). US Patent 0 017 248.
- Qi, J. Y., Yang, Q. Y., Lam, K. H., Zhou, Z. Y. & Chan, A. S. C. (2003). *Acta Cryst.* **E59**, o374–o375.
- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
- Sousa, G. F. & Filgueiras, C. A. L. (1990). *Transition Met. Chem.* **15**, 286–289.
- Yin, X.-H., Zhao, K., Feng, Y. & Zhu, J. (2007). *Acta Cryst.* **E63**, o4617.
- Zhang, Q., Zhang, S.-P. & Shao, S.-C. (2006). *Acta Cryst.* **E62**, o4695–o4696.

supplementary materials

Acta Cryst. (2008). E64, o1904 [doi:10.1107/S1600536808028274]

***N*-Phenylpyridine-2-carbamide**

Y.-G. Zhuang, H.-J. Jiang, Z. Hong and F.-L. Qiu

Comment

Pyridine-containing amides continue to attract considerable interest as ligands for metals (Sousa & Filgueiras, 1990; Gomes *et al.*, 2007; Jacob & Mukherjee, 2006), building blocks in organic synthesis (Marumoto *et al.*, 1981) and physiologically active compounds (Piatnitski & Kiselyov, 2004). As part of our studies on the synthesis and characterization of these compounds, we report here the crystal structure of the title compound.

The C7—O1 [1.223 (3) Å], N1—C7 [1.344 (3) Å] and N1—C6 [1.410 (2) Å] bond lengths indicate extensive electron delocalization in the amide linkage. The pyridyl and phenyl rings of the title compound are almost coplanar, forming a dihedral angle of 1.8 (1)°. In the crystal structure, there is an intramolecular hydrogen bond (N1—H101...N2) and no intermolecular hydrogen bonds are observed (Table 1).

The reported monoclinic space-group is in a non-standard setting (*Pn*, #7). There is a strong feature ($h + 1 = 2n$) in *hkl* data. Setting up the space group as *Pc* results in a β angle of 23° or 157°, respectively. Obviously such an unit-cell division is inappropriate. Therefore, the non-standard setting *Pn* was chosen.

Experimental

The title compound was synthesized from pyridine-2-carboxylic acid and aniline according to the procedure of Chan *et al.* (2004). The crystal used for data collection was obtained by slow evaporation from a saturated ethanol/water solution at room temperature.

Refinement

All H atoms were placed in geometrically idealized positions and constrained to ride on their parent atoms, with N—H = 0.86 Å, C—H = 0.93 Å, and with $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{parent atom})$.

Figures

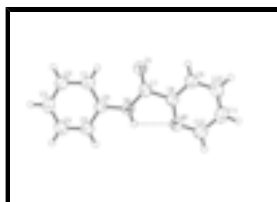


Fig. 1. The molecular structure of the title compound, shown with 50% probability displacement ellipsoids.

N-Phenylpyridine-2-carbamide

Crystal data

$C_{12}H_{10}N_2O$	$F_{000} = 208$
$M_r = 198.22$	$D_x = 1.315 \text{ Mg m}^{-3}$
Monoclinic, Pn	Mo $K\alpha$ radiation
Hall symbol: P -2yac	$\lambda = 0.71073 \text{ \AA}$
$a = 5.7469 (2) \text{ \AA}$	Cell parameters from 2226 reflections
$b = 6.2382 (2) \text{ \AA}$	$\theta = 2.9\text{--}25.1^\circ$
$c = 14.0158 (3) \text{ \AA}$	$\mu = 0.09 \text{ mm}^{-1}$
$\beta = 94.752 (2)^\circ$	$T = 296 (2) \text{ K}$
$V = 500.74 (3) \text{ \AA}^3$	Block, colourless
$Z = 2$	$0.15 \times 0.14 \times 0.09 \text{ mm}$

Data collection

Bruker APEXII CCD diffractometer	1162 independent reflections
Radiation source: fine-focus sealed tube	1036 reflections with $I > 2\sigma(I)$
Monochromator: graphite	$R_{\text{int}} = 0.017$
Detector resolution: 10 pixels mm^{-1}	$\theta_{\text{max}} = 27.7^\circ$
$T = 296(2) \text{ K}$	$\theta_{\text{min}} = 2.9^\circ$
ω scans	$h = -7 \rightarrow 7$
Absorption correction: multi-scan (SADABS; Bruker, 2004)	$k = -8 \rightarrow 5$
$T_{\text{min}} = 0.968, T_{\text{max}} = 0.992$	$l = -18 \rightarrow 18$
5000 measured reflections	

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.030$	H-atom parameters constrained
$wR(F^2) = 0.075$	$w = 1/[\sigma^2(F_o^2) + (0.0365P)^2 + 0.0548P]$
$S = 1.00$	where $P = (F_o^2 + 2F_c^2)/3$
1162 reflections	$(\Delta/\sigma)_{\text{max}} < 0.001$
137 parameters	$\Delta\rho_{\text{max}} = 0.11 \text{ e \AA}^{-3}$
Primary atom site location: structure-invariant direct methods	$\Delta\rho_{\text{min}} = -0.12 \text{ e \AA}^{-3}$
	Extinction correction: none

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
N1	-0.0173 (3)	0.2086 (3)	0.59409 (11)	0.0452 (4)
H101	-0.1420	0.2510	0.6183	0.054*
C8	0.1414 (3)	0.5113 (3)	0.68241 (13)	0.0435 (4)
C6	-0.0401 (3)	0.0235 (3)	0.53631 (12)	0.0412 (4)
C1	0.1246 (4)	-0.0375 (4)	0.47407 (15)	0.0509 (5)
H1	0.2570	0.0459	0.4685	0.061*
C7	0.1751 (4)	0.3270 (3)	0.61595 (15)	0.0475 (5)
C2	0.0901 (4)	-0.2230 (4)	0.42049 (16)	0.0569 (5)
H2	0.1999	-0.2635	0.3787	0.068*
N2	-0.0537 (3)	0.5139 (3)	0.72828 (12)	0.0526 (4)
C5	-0.2373 (4)	-0.1016 (3)	0.54213 (15)	0.0476 (5)
H5	-0.3507	-0.0595	0.5819	0.057*
O1	0.3656 (3)	0.2937 (3)	0.58575 (15)	0.0753 (6)
C10	0.2767 (5)	0.8384 (3)	0.75234 (17)	0.0575 (5)
H10	0.3863	0.9481	0.7597	0.069*
C11	0.0794 (4)	0.8441 (4)	0.80026 (16)	0.0612 (6)
H11	0.0523	0.9575	0.8410	0.073*
C3	-0.1044 (4)	-0.3483 (4)	0.42826 (16)	0.0553 (5)
H3	-0.1255	-0.4731	0.3923	0.066*
C12	-0.0787 (4)	0.6788 (4)	0.78711 (17)	0.0629 (6)
H12	-0.2107	0.6823	0.8213	0.076*
C9	0.3106 (4)	0.6677 (3)	0.69302 (15)	0.0516 (5)
H9	0.4451	0.6581	0.6607	0.062*
C4	-0.2676 (4)	-0.2879 (4)	0.48961 (17)	0.0553 (5)
H4	-0.3986	-0.3728	0.4956	0.066*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
N1	0.0458 (9)	0.0397 (9)	0.0515 (9)	-0.0007 (7)	0.0121 (7)	-0.0031 (7)
C8	0.0478 (10)	0.0408 (10)	0.0422 (10)	0.0013 (8)	0.0051 (8)	0.0021 (8)
C6	0.0458 (10)	0.0369 (10)	0.0408 (10)	0.0030 (8)	0.0042 (8)	0.0021 (8)

supplementary materials

C1	0.0504 (11)	0.0521 (12)	0.0517 (11)	-0.0041 (10)	0.0120 (9)	-0.0028 (10)
C7	0.0497 (11)	0.0398 (11)	0.0539 (12)	-0.0030 (9)	0.0092 (9)	-0.0007 (9)
C2	0.0594 (13)	0.0609 (14)	0.0515 (12)	0.0060 (11)	0.0117 (9)	-0.0103 (10)
N2	0.0477 (9)	0.0540 (10)	0.0566 (10)	-0.0015 (8)	0.0067 (8)	-0.0096 (8)
C5	0.0464 (10)	0.0458 (11)	0.0515 (10)	-0.0004 (9)	0.0088 (8)	-0.0024 (9)
O1	0.0531 (9)	0.0669 (11)	0.1095 (15)	-0.0111 (9)	0.0283 (9)	-0.0330 (10)
C10	0.0699 (14)	0.0471 (11)	0.0550 (11)	-0.0102 (12)	0.0014 (10)	-0.0007 (11)
C11	0.0714 (16)	0.0553 (14)	0.0559 (13)	0.0065 (12)	-0.0002 (11)	-0.0156 (11)
C3	0.0652 (13)	0.0459 (11)	0.0537 (11)	0.0014 (11)	-0.0023 (10)	-0.0119 (10)
C12	0.0561 (13)	0.0705 (15)	0.0635 (14)	0.0047 (11)	0.0122 (11)	-0.0168 (12)
C9	0.0559 (11)	0.0486 (12)	0.0513 (11)	-0.0086 (10)	0.0116 (9)	-0.0018 (9)
C4	0.0550 (12)	0.0506 (13)	0.0598 (12)	-0.0070 (10)	0.0027 (10)	-0.0027 (11)

Geometric parameters (Å, °)

N1—C7	1.344 (3)	N2—C12	1.333 (3)
N1—C6	1.410 (2)	C5—C4	1.379 (3)
N1—H101	0.8600	C5—H5	0.9300
C8—N2	1.338 (2)	C10—C11	1.365 (4)
C8—C9	1.377 (3)	C10—C9	1.375 (3)
C8—C7	1.502 (3)	C10—H10	0.9300
C6—C5	1.384 (3)	C11—C12	1.377 (3)
C6—C1	1.392 (3)	C11—H11	0.9300
C1—C2	1.385 (3)	C3—C4	1.377 (3)
C1—H1	0.9300	C3—H3	0.9300
C7—O1	1.223 (3)	C12—H12	0.9300
C2—C3	1.376 (4)	C9—H9	0.9300
C2—H2	0.9300	C4—H4	0.9300
C7—N1—C6	128.07 (17)	C4—C5—H5	119.6
C7—N1—H101	116.0	C6—C5—H5	119.6
C6—N1—H101	116.0	C11—C10—C9	118.9 (2)
N2—C8—C9	123.48 (19)	C11—C10—H10	120.6
N2—C8—C7	117.62 (17)	C9—C10—H10	120.6
C9—C8—C7	118.89 (18)	C10—C11—C12	118.7 (2)
C5—C6—C1	119.06 (17)	C10—C11—H11	120.7
C5—C6—N1	117.74 (17)	C12—C11—H11	120.7
C1—C6—N1	123.19 (17)	C2—C3—C4	119.6 (2)
C2—C1—C6	119.57 (19)	C2—C3—H3	120.2
C2—C1—H1	120.2	C4—C3—H3	120.2
C6—C1—H1	120.2	N2—C12—C11	123.8 (2)
O1—C7—N1	124.81 (19)	N2—C12—H12	118.1
O1—C7—C8	120.61 (18)	C11—C12—H12	118.1
N1—C7—C8	114.58 (18)	C10—C9—C8	118.7 (2)
C3—C2—C1	120.8 (2)	C10—C9—H9	120.6
C3—C2—H2	119.6	C8—C9—H9	120.6
C1—C2—H2	119.6	C3—C4—C5	120.1 (2)
C12—N2—C8	116.38 (18)	C3—C4—H4	119.9
C4—C5—C6	120.77 (19)	C5—C4—H4	119.9

Hydrogen-bond geometry (Å, °)

<i>D—H···A</i>	<i>D—H</i>	<i>H···A</i>	<i>D···A</i>	<i>D—H···A</i>
N1—H101···N2	0.86	2.28	2.697 (2)	110

Fig. 1

