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## But-2-enal 2,4-dinitrophenylhydrazone

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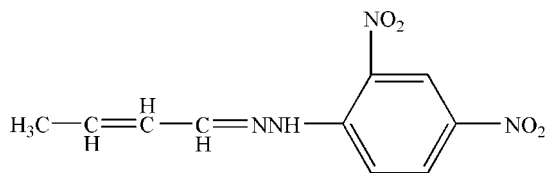
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 Key indicators: single-crystal X-ray study;  $T = 293$  K; mean  $\sigma(\text{C}-\text{C}) = 0.003$  Å;  $R$  factor = 0.042;  $wR$  factor = 0.111; data-to-parameter ratio = 15.0.

In the title compound,  $\text{C}_{10}\text{H}_{10}\text{N}_4\text{O}_4$ , the but-2-enal chain is almost planar, the largest deviation from the mean plane being 0.013 (1) Å, and this plane makes a dihedral angle of 9.95 (24)° with the benzene ring. Of the two nitro groups, one is twisted with respect to the benzene ring, making a dihedral angle of 5.7 (1)°, whereas the other is nearly in the plane of the benzene ring, with a twist angle of only 0.7 (1)°. This difference is related to the occurrence of an intramolecular  $\text{N}-\text{H}\cdots\text{O}$  hydrogen bond with the O atom of the less twisted nitro group. The NH group is also involved in a weak interaction with the same O atom of a symmetry-related molecule, thus forming a pseudo inversion dimer.

## Related literature

For general background, see: Okabe *et al.* (1993). For related structures, see: Bolte & Dill (1998); Ohba (1996).



## Experimental

## Crystal data

 $\text{C}_{10}\text{H}_{10}\text{N}_4\text{O}_4$   
 $M_r = 250.22$ 

 Monoclinic,  $P2_1/c$   
 $a = 4.6699$  (8) Å

 $b = 13.188$  (2) Å  
 $c = 18.880$  (3) Å  
 $\beta = 92.565$  (3)°  
 $V = 1161.6$  (3) Å<sup>3</sup>  
 $Z = 4$ 

 Mo  $K\alpha$  radiation  
 $\mu = 0.11$  mm<sup>-1</sup>  
 $T = 293$  (2) K  
 $0.27 \times 0.23 \times 0.23$  mm

## Data collection

 Bruker SMART APEX CCD area-detector diffractometer  
 Absorption correction: multi-scan (SADABS; Bruker, 1998)  
 $T_{\min} = 0.972$ ,  $T_{\max} = 0.976$ 

 9152 measured reflections  
 2458 independent reflections  
 1326 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.039$ 

## Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.042$   
 $wR(F^2) = 0.111$   
 $S = 0.95$   
 2458 reflections

 164 parameters  
 H-atom parameters constrained  
 $\Delta\rho_{\max} = 0.16$  e Å<sup>-3</sup>  
 $\Delta\rho_{\min} = -0.11$  e Å<sup>-3</sup>

Table 1

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{N2}-\text{H2}\cdots\text{O1}$	0.86	2.02	2.6314 (18)	127
$\text{N2}-\text{H2}\cdots\text{O1}^i$	0.86	2.52	3.331 (2)	159

 Symmetry code: (i)  $-x + 1, -y + 1, -z$ .

Data collection: SMART (Bruker, 1998); cell refinement: SAINT (Bruker, 1998); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: ORTEPIII (Burnett & Johnson, 1996), ORTEP-3 for Windows (Farrugia, 1997) and PLATON (Spek, 2003); software used to prepare material for publication: SHELXL97.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: DN2390).

## References

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**supplementary materials**

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## But-2-enal 2,4-dinitrophenylhydrazone

Z.-G. Yin, H.-Y. Qian, Y.-Z. Chen and J. Hu

### Comment

2,4-Dinitrophenylhydrazine has applications in organic synthesis and some of its derivatives have been shown to be potentially DNA-damaging and mutagenic agents (Okabe *et al.*, 1993). Some phenylhydrazone derivatives have been synthesized in our laboratory. As part of our work, we report the synthesis and crystal structure of the title compound(I).

In the title compound, the but-2-enal chain is planar with the largest deviation from the mean plane being 0.013 (1)Å at C1. This plane makes a dihedral angle of 9.95 (24)° with the benzene ring, so the whole molecule is roughly planar (Fig. 1). Of the the two nitro groups, one is twisted with respect to the benzene ring making a dihedral angle of 5.7 (1)° whereas the other is nearly in the plane of the benzene ring with a twist angle of only 0.7 (1)°. This difference is related to the occurrence of an intramolecular hydrogen N-H...bond with the O atom of the less twisted nitro group (Table 1). The NH is also in weak intermolecular interaction with the same O atom of a symmetry related molecule building a pseudo dimer (Table 1, Fig. 2).

Bond lengths and bond angles are consistent with those of other dinitrophenylhydrazone derivatives(Ohba,1996; Bolte & Dill, 1998)

### Experimental

2,4-Dinitrophenylhydrazine (1 mmol, 0.198 g) was dissolved in anhydrous methanol, H<sub>2</sub>SO<sub>4</sub> (98% 0.5 ml) was added to this, the mixture was stirred for several minutes at 351 K, but-2-enal (1 mmol 0.070 g) in methanol (8 ml) was added dropwise and the mixture was stirred at refluxing temperature for 2 h. The product was isolated and recrystallized in methanol, red single crystals of (I) was obtained after two weeks.

### Refinement

H atoms were placed in calculated position and treated as riding with C—H= 0.93Å (aromatic), 0.96Å(methyl) and N-H= 0.86Å with  $U_{iso}(H)=1.2U_{eq}(C,N)$  or  $U_{iso}(H)=1.5U_{eq}(methyl)$ .

### Figures

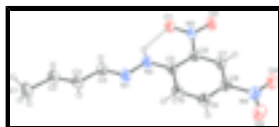


Fig. 1. Molecular view of (I) with the atom-labeling scheme. Displacement ellipsoids are drawn at the 30% probability level. H atoms are presented as small spheres of arbitrary radii. Intramolecular H bond is shown as dashed line.

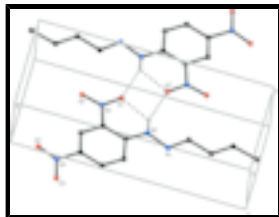


Fig. 2. Partial packing view showing the formation of pseudo dimer through weak intermolecular N-H...O hydrogen bonds which are shown as dashed lines. H atoms not involved in hydrogen bondings have been omitted for clarity. [Symmetry code: (i) -x+1, -y+1, -z]

## But-2-enal 2,4-dinitrophenylhydrazone

### Crystal data

$C_{10}H_{10}N_4O_4$

$M_r = 250.22$

Monoclinic,  $P2_1/c$

Hall symbol: -P 2ybc

$a = 4.6699$  (8) Å

$b = 13.188$  (2) Å

$c = 18.880$  (3) Å

$\beta = 92.565$  (3)°

$V = 1161.6$  (3) Å<sup>3</sup>

$Z = 4$

$F_{000} = 520$

$D_x = 1.432$  Mg m<sup>-3</sup>

Mo  $K\alpha$  radiation

$\lambda = 0.71073$  Å

Cell parameters from 1673 reflections

$\theta = 2.1$ – $25.5$ °

$\mu = 0.11$  mm<sup>-1</sup>

$T = 293$  (2) K

Block, red

$0.27 \times 0.23 \times 0.23$  mm

### Data collection

Bruker SMART APEX CCD area-detector diffractometer

Radiation source: sealed tube

Monochromator: graphite

$T = 293$ (2) K

$\varphi$  and  $\omega$  scans

Absorption correction: multi-scan (SADABS; Bruker, 1998)

$T_{\min} = 0.972$ ,  $T_{\max} = 0.976$

9152 measured reflections

2458 independent reflections

1326 reflections with  $I > 2\sigma(I)$

$R_{\text{int}} = 0.039$

$\theta_{\max} = 27.0$ °

$\theta_{\min} = 2.7$ °

$h = -5 \rightarrow 5$

$k = -16 \rightarrow 16$

$l = -24 \rightarrow 24$

### Refinement

Refinement on  $F^2$

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.042$

$wR(F^2) = 0.111$

$S = 0.95$

2458 reflections

164 parameters

Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring sites

H-atom parameters constrained

$w = 1/[\sigma^2(F_o^2) + (0.0552P)^2]$

where  $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{\max} < 0.001$

$\Delta\rho_{\max} = 0.16$  e Å<sup>-3</sup>

$\Delta\rho_{\min} = -0.11$  e Å<sup>-3</sup>

Primary atom site location: structure-invariant direct methods Extinction correction: none

*Special details*

**Geometry.** All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
N1	0.0380 (3)	0.35052 (10)	0.11439 (8)	0.0607 (4)
N2	0.1292 (3)	0.44685 (9)	0.09740 (8)	0.0563 (4)
H2	0.2613	0.4552	0.0676	0.068*
N3	0.2785 (3)	0.65667 (11)	0.05938 (8)	0.0554 (4)
N4	-0.3915 (4)	0.77803 (13)	0.22849 (9)	0.0673 (5)
O1	0.4042 (3)	0.58766 (9)	0.02881 (7)	0.0693 (4)
O2	0.3266 (3)	0.74610 (9)	0.04720 (8)	0.0761 (4)
O3	-0.5851 (3)	0.75891 (12)	0.26884 (9)	0.0872 (5)
O4	-0.3125 (3)	0.86398 (12)	0.21507 (9)	0.0909 (5)
C1	0.1822 (6)	-0.01443 (14)	0.09371 (14)	0.1054 (9)
H1A	0.0157	-0.0191	0.1216	0.158*
H1B	0.3428	-0.0449	0.1191	0.158*
H1C	0.1469	-0.0492	0.0495	0.158*
C2	0.2468 (5)	0.09525 (15)	0.07954 (12)	0.0770 (6)
H22	0.4003	0.1089	0.0513	0.092*
C3	0.1060 (5)	0.17380 (13)	0.10341 (10)	0.0651 (5)
H3	-0.0512	0.1615	0.1307	0.078*
C4	0.1811 (4)	0.27691 (12)	0.08969 (10)	0.0585 (5)
H4	0.3370	0.2906	0.0623	0.070*
C5	0.0079 (4)	0.52752 (12)	0.12821 (9)	0.0497 (4)
C6	0.0733 (4)	0.63020 (12)	0.11150 (9)	0.0506 (4)
C7	-0.0580 (4)	0.71065 (13)	0.14479 (10)	0.0539 (5)
H7	-0.0131	0.7771	0.1330	0.065*
C8	-0.2534 (4)	0.69186 (13)	0.19495 (10)	0.0556 (5)
C9	-0.3232 (4)	0.59241 (15)	0.21334 (10)	0.0603 (5)
H9	-0.4562	0.5806	0.2476	0.072*
C10	-0.1960 (4)	0.51299 (13)	0.18104 (10)	0.0582 (5)
H10	-0.2439	0.4473	0.1939	0.070*

## supplementary materials

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### Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
N1	0.0727 (10)	0.0384 (8)	0.0716 (10)	-0.0054 (7)	0.0084 (8)	0.0014 (7)
N2	0.0651 (9)	0.0405 (8)	0.0639 (9)	-0.0042 (7)	0.0116 (7)	0.0008 (7)
N3	0.0681 (10)	0.0380 (8)	0.0602 (9)	0.0009 (7)	0.0032 (8)	0.0008 (7)
N4	0.0677 (11)	0.0668 (12)	0.0670 (11)	0.0122 (9)	-0.0015 (9)	-0.0095 (9)
O1	0.0902 (10)	0.0448 (7)	0.0747 (9)	0.0027 (7)	0.0259 (7)	-0.0002 (6)
O2	0.1022 (11)	0.0386 (7)	0.0894 (10)	-0.0035 (7)	0.0259 (8)	0.0066 (6)
O3	0.0782 (10)	0.0913 (11)	0.0937 (11)	0.0151 (8)	0.0195 (9)	-0.0164 (8)
O4	0.1150 (13)	0.0545 (9)	0.1047 (12)	0.0163 (9)	0.0220 (10)	-0.0058 (8)
C1	0.162 (2)	0.0418 (11)	0.110 (2)	0.0019 (14)	-0.0186 (17)	0.0004 (12)
C2	0.1028 (17)	0.0469 (11)	0.0809 (14)	0.0004 (11)	0.0013 (12)	-0.0004 (10)
C3	0.0838 (14)	0.0428 (10)	0.0687 (12)	-0.0059 (9)	0.0057 (10)	0.0026 (9)
C4	0.0724 (12)	0.0427 (10)	0.0606 (11)	-0.0040 (9)	0.0060 (9)	0.0011 (8)
C5	0.0527 (10)	0.0394 (9)	0.0564 (11)	-0.0010 (8)	-0.0037 (8)	-0.0010 (7)
C6	0.0528 (10)	0.0445 (9)	0.0543 (10)	-0.0006 (8)	-0.0001 (9)	0.0012 (8)
C7	0.0572 (11)	0.0406 (9)	0.0632 (11)	0.0014 (8)	-0.0062 (9)	-0.0002 (8)
C8	0.0548 (11)	0.0519 (11)	0.0596 (11)	0.0098 (9)	-0.0014 (9)	-0.0042 (9)
C9	0.0545 (11)	0.0661 (12)	0.0605 (12)	0.0025 (9)	0.0051 (9)	-0.0002 (9)
C10	0.0622 (11)	0.0474 (10)	0.0649 (12)	-0.0042 (9)	0.0039 (9)	0.0056 (9)

### Geometric parameters ( $\text{\AA}$ , $^\circ$ )

N1—C4	1.278 (2)	C2—C3	1.317 (3)
N1—N2	1.3821 (17)	C2—H22	0.9300
N2—C5	1.349 (2)	C3—C4	1.431 (2)
N2—H2	0.8600	C3—H3	0.9300
N3—O2	1.2245 (17)	C4—H4	0.9300
N3—O1	1.2396 (17)	C5—C10	1.422 (3)
N3—C6	1.446 (2)	C5—C6	1.427 (2)
N4—O4	1.2221 (19)	C6—C7	1.390 (2)
N4—O3	1.234 (2)	C7—C8	1.367 (3)
N4—C8	1.465 (2)	C7—H7	0.9300
C1—C2	1.504 (3)	C8—C9	1.399 (3)
C1—H1A	0.9600	C9—C10	1.362 (2)
C1—H1B	0.9600	C9—H9	0.9300
C1—H1C	0.9600	C10—H10	0.9300
C4—N1—N2	116.23 (15)	N1—C4—C3	121.30 (18)
C5—N2—N1	119.02 (14)	N1—C4—H4	119.4
C5—N2—H2	120.5	C3—C4—H4	119.4
N1—N2—H2	120.5	N2—C5—C10	120.21 (15)
O2—N3—O1	121.65 (15)	N2—C5—C6	123.70 (16)
O2—N3—C6	119.54 (15)	C10—C5—C6	116.08 (16)
O1—N3—C6	118.80 (14)	C7—C6—C5	121.43 (16)
O4—N4—O3	123.62 (17)	C7—C6—N3	116.26 (15)
O4—N4—C8	119.16 (19)	C5—C6—N3	122.30 (15)

O3—N4—C8	117.21 (17)	C8—C7—C6	119.78 (16)
C2—C1—H1A	109.5	C8—C7—H7	120.1
C2—C1—H1B	109.5	C6—C7—H7	120.1
H1A—C1—H1B	109.5	C7—C8—C9	120.80 (17)
C2—C1—H1C	109.5	C7—C8—N4	118.64 (17)
H1A—C1—H1C	109.5	C9—C8—N4	120.55 (18)
H1B—C1—H1C	109.5	C10—C9—C8	119.93 (18)
C3—C2—C1	126.1 (2)	C10—C9—H9	120.0
C3—C2—H22	117.0	C8—C9—H9	120.0
C1—C2—H22	117.0	C9—C10—C5	121.97 (16)
C2—C3—C4	123.7 (2)	C9—C10—H10	119.0
C2—C3—H3	118.1	C5—C10—H10	119.0
C4—C3—H3	118.1		
C5—N2—N1—C4	-172.68 (16)	N1—C4—C3—C2	-179.7 (2)
N2—N1—C4—C3	-179.67 (15)	C4—C3—C2—C1	178.41 (19)

*Hydrogen-bond geometry (Å, °)*

<i>D</i> —H $\cdots$ <i>A</i>	<i>D</i> —H	H $\cdots$ <i>A</i>	<i>D</i> $\cdots$ <i>A</i>	<i>D</i> —H $\cdots$ <i>A</i>
N2—H2 $\cdots$ O1	0.86	2.02	2.6314 (18)	127
N2—H2 $\cdots$ O1 <sup>i</sup>	0.86	2.52	3.331 (2)	159

Symmetry codes: (i)  $-x+1, -y+1, -z$ .

Fig. 1

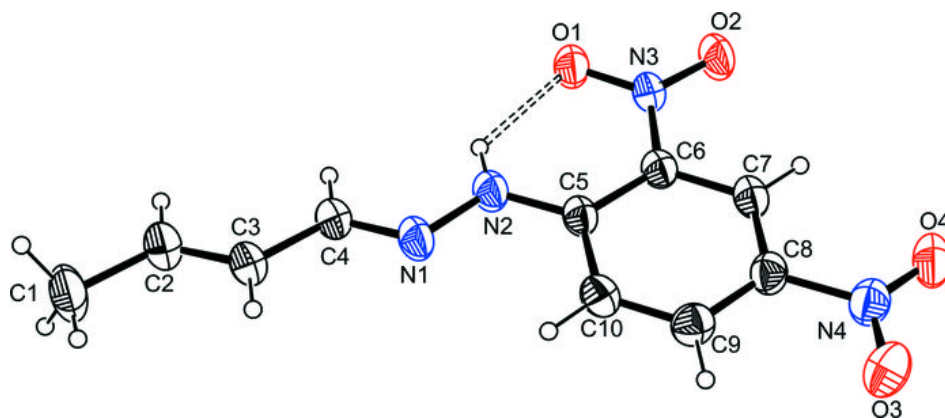


Fig. 2

