

Tetrapyrazineplatinum(II) bis(tetrafluoroborate) acetonitrile hemisolvate

 Paul J. Derry,^a Xiaoping Wang^b and Bradley W. Smucker^{a*}
^aDepartment of Chemistry, Austin College, 900 North Grand, Sherman, TX 75090-4400, USA, and ^bOak Ridge National Laboratory, PO Box 2008 MS6460, Oak Ridge, TN 37831-6460, USA

Correspondence e-mail: bsmucker@austincollege.edu

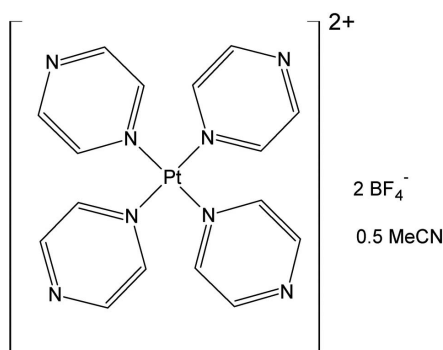
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 Key indicators: single-crystal X-ray study; $T = 296$ K; mean $\sigma(\text{C}-\text{C}) = 0.013$ Å; disorder in solvent or counterion; R factor = 0.033; wR factor = 0.114; data-to-parameter ratio = 16.4.

The improved synthesis and characterization of tetrapyrazineplatinum(II) bis(tetrafluoroborate) acetonitrile hemisolvate, $[\text{Pt}(\text{C}_4\text{H}_4\text{N}_2)_4](\text{BF}_4)_2 \cdot 0.5\text{CH}_3\text{CN}$, is reported. The unit cell contains a half equivalent of an acetonitrile solvent molecule per tetrapyrazineplatinum(II) ion. The coordination geometry of the Pt^{II} ion is almost square-planar, with the Pt atom residing on an inversion center. The BF_4^- counter-anion, located at a general position, has an idealized tetrahedral geometry and an acetonitrile solvent molecule, the methyl group of which is disordered over two equal positions, sits on a twofold rotation axis.

Related literature

For general background, see: Derossi *et al.* (2007); Klika *et al.* (2007); Pearson *et al.* (1960); Schweiger *et al.* (2001); Wendt *et al.* (1997); Willermann *et al.* (2006). For related structures, see: Wei *et al.* (1989).



Experimental

Crystal data

$[\text{Pt}(\text{C}_4\text{H}_4\text{N}_2)_4](\text{BF}_4)_2 \cdot 0.5\text{C}_2\text{H}_3\text{N}$	$V = 2587.8$ (13) Å ³
$M_r = 709.6$	$Z = 4$
Monoclinic, $C2/c$	Mo $K\alpha$ radiation
$a = 13.862$ (4) Å	$\mu = 5.50$ mm ⁻¹
$b = 10.819$ (3) Å	$T = 296$ (2) K
$c = 17.262$ (5) Å	$0.27 \times 0.21 \times 0.19$ mm
$\beta = 91.607$ (3)°	

Data collection

Bruker APEXII CCD diffractometer	11212 measured reflections
Absorption correction: multi-scan (SADABS; Bruker, 2004)	2736 independent reflections
$T_{\text{min}} = 0.246$, $T_{\text{max}} = 0.352$	1846 reflections with $I > 2\sigma(I)$
	$R_{\text{int}} = 0.039$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.033$	2 restraints
$wR(F^2) = 0.114$	H-atom parameters constrained
$S = 1.20$	$\Delta\rho_{\text{max}} = 1.58$ e Å ⁻³
2736 reflections	$\Delta\rho_{\text{min}} = -0.45$ e Å ⁻³
167 parameters	

Data collection: APEX2 (Bruker, 2006); cell refinement: APEX2; data reduction: APEX2; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: XP in SHELXTL (Sheldrick, 2008) and PLATON (Spek, 2003); software used to prepare material for publication: SHELXL97 and WinGX (Farrugia, 1999).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: PK2119).

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Tetrapyrazineplatinum(II) bis(tetrafluoroborate) acetonitrile hemisolvate

P. J. Derry, X. Wang and B. W. Smucker

Comment

The synthesis of the nitrate salt of tetrapyrazineplatinum(II), which has use as a precursor to other platinum compounds, was recently reported in low yield (Klika *et al.*, 2007). In our pursuit of utilizing pyrazine as a bridging ligand to form compounds which add to the growing number of supramolecular metallacycles such as those by Schweiger *et al.* (2001), Willermann *et al.* (2006), and Derossi *et al.* (2007), to name but a few, we sought a more favorable yield for this tetrapyrazineplatinum(II) precursor. We settled on a synthetic method that generates a high yield of tetrapyrazineplatinum(II) tetrafluoroborate following the use of a large excess of pyrazine and nitromethane, a non-coordinating solvent which Pearson *et al.* (1960) reported as increasing the lability of a ligand bound to Pt^{II}.

The structure of the cation of tetrapyrazineplatinum(II) tetrafluoroborate (Figure 1) has a square planar conformation. The Pt^{II} ion resides on an inversion center with Pt1—N1 and Pt1—N2 distances of 2.026 (6) and 2.012 (5) Å, respectively. The two unique coordinated pyrazine ligands (N1 C1 C2 C3 C4 N3, N2 C5 C6 C7 C8 N4) are slightly canted relative to the (N1—Pt—N2) plane (dihedral angles between pyrazine planes and N1—Pt—N2 plane are 84.1 (3)° and 69.3 (3)°, respectively). The canted conformation of the pyrazine ligands around the platinum atom is similar to that of pyridine in the closely-related tetrapyridine platinum(II) chloride trihydrate reported by Wei *et al.* (1989).

The tetrafluoroborate anion is positioned near three tetrapyrazineplatinum(II) cations and oriented such that every fluorine atom is between 2.1 and 2.4 Å from a hydrogen atom in the 3 or 5 position of a coordinated pyrazine. (Figure 2).

Experimental

53 mg (0.099 mmol) of [Pt(NCMe)₄](BF₄)₂, (synthesized according to Wendt *et al.* (1997)) and 344 mg (4.3 mmol) of pyrazine (Aldrich) were dissolved in 11 ml of degassed MeNO₂. The schlenk flask was covered in foil and the stirring solution was moderately heated under a nitrogen environment for 18 h. 50 ml of degassed Et₂O was added, the solution was cooled in an ice bath, and the resulting white solid was isolated by removal of the liquid *via* cannula and washing with 2x5mL degassed Et₂O. The resulting solid was vacuum dried to give 67 mg (98% yield.) of a white solid. FT—IR (nujol, CsI plates): (cm⁻¹) 1433(*s*, C—H), 1055 (*b*, B—F), 520 and 503 (*w*, Pt—N). ¹H NMR (d-MeNO₂): 9.00 (*m*) and 8.79 (*m*). ESI—M. S. 602 (Pt(C₄H₄N₂)₄)(BF₄)⁺, 258 ([Pt(C₄H₄N₂)₄]²⁺).

Pale yellow crystals were grown by liquid diffusion of diethylether into a nitromethane solution containing tetrapyrazineplatinum(II) tetrafluoroborate and excess pyrazine.

Refinement

The positions of the pyrazine and methyl H atoms were refined using a riding model, in accordance with the HFIX 43 and HFIX 137 instructions of SHELXL97, with pyrazine C—H distances of 0.93 Å, methyl C—H distance of 0.96 Å, and

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with $U_{\text{iso}}(\text{H})$ values of 1.2Ueq(C) for pyrazine and 1.5Ueq(C) for methyl group in the acetonitrile solvent molecule. The acetonitrile solvate resides on a crystallographic two-fold axis with the methyl H atoms disordered in two position. Each of the methyl H atoms were refined accordingly as half occupied. The largest positive and negative peaks in the final difference map are 1.58 and -0.45 Å, respectively, from the Pt atom.

Figures

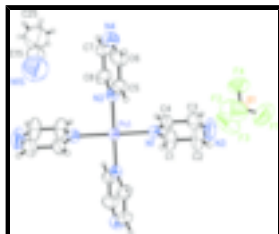


Fig. 1. Ellipsoid plot of tetrapyrazineplatinum(II) tetrafluoroborate. Unlabelled atoms are generated via inversion through Pt1, symmetry code: $-x, -y + 1, -z$ (50% probability displacement ellipsoids).

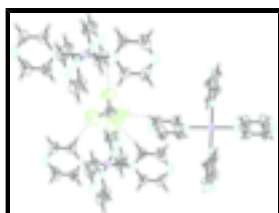


Fig. 2. View of the hydrogen bonding between the tetrafluoroborate anion and the tetrapyrazineplatinum(II) cation (50% probability displacement ellipsoids).

tetrapyrazineplatinum(II) bis(tetrafluoroborate) acetonitrile hemisolvate

Crystal data

$[\text{Pt}(\text{C}_4\text{H}_4\text{N}_2)_4](\text{BF}_4)_2 \cdot 0.5\text{C}_2\text{H}_3\text{N}$

$M_r = 709.6$

Monoclinic, $C2/c$

Hall symbol: $-C 2yc$

$a = 13.862$ (4) Å

$b = 10.819$ (3) Å

$c = 17.262$ (5) Å

$\beta = 91.607$ (3)°

$V = 2587.8$ (13) Å³

$Z = 4$

$F_{000} = 1356$

$D_x = 1.821$ Mg m⁻³

Mo $K\alpha$ radiation

$\lambda = 0.71073$ Å

Cell parameters from 8501 reflections

$\theta = 2.3$ – 27.8°

$\mu = 5.50$ mm⁻¹

$T = 296$ (2) K

Block, pale yellow

$0.27 \times 0.21 \times 0.19$ mm

Data collection

Bruker APEXII CCD
diffractometer

ω scans

Absorption correction: Multi-scan
(SADABS; Bruker, 2004)

$T_{\text{min}} = 0.246$, $T_{\text{max}} = 0.352$

11212 measured reflections

2736 independent reflections

$R_{\text{int}} = 0.039$

$\theta_{\text{max}} = 26.7^\circ$

$\theta_{\text{min}} = 2.4^\circ$

$h = -17 \rightarrow 17$

$k = -13 \rightarrow 13$

$l = -21 \rightarrow 21$

1846 reflections with $I > 2\sigma(I)$

Refinement

Refinement on F^2	2 restraints
Least-squares matrix: full	H-atom parameters constrained
$R[F^2 > 2\sigma(F^2)] = 0.033$	$w = 1/[\sigma^2(F_o^2) + (0.0518P)^2 + 10.4546P]$
$wR(F^2) = 0.114$	where $P = (F_o^2 + 2F_c^2)/3$
$S = 1.20$	$(\Delta/\sigma)_{\max} = 0.012$
2736 reflections	$\Delta\rho_{\max} = 1.58 \text{ e } \text{\AA}^{-3}$
167 parameters	$\Delta\rho_{\min} = -0.45 \text{ e } \text{\AA}^{-3}$
	Extinction correction: none

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$	Occ. (<1)
Pt1	0	0.5	0	0.04102 (14)	
N1	0.0732 (4)	0.5078 (5)	0.1031 (4)	0.0463 (13)	
N2	0.1147 (4)	0.4171 (5)	-0.0466 (3)	0.0435 (13)	
N3	0.1705 (8)	0.5173 (9)	0.2459 (5)	0.090 (3)	
N4	0.2743 (5)	0.3006 (7)	-0.1078 (4)	0.0693 (19)	
C1	0.1234 (7)	0.6071 (9)	0.1267 (5)	0.072 (3)	
H1A	0.1267	0.6756	0.0943	0.087*	
C2	0.1699 (8)	0.6096 (12)	0.1971 (6)	0.092 (3)	
H2A	0.2031	0.6811	0.2113	0.11*	
C3	0.1246 (7)	0.4181 (12)	0.2223 (5)	0.081 (3)	
H3A	0.1257	0.3491	0.2545	0.097*	
C4	0.0729 (6)	0.4100 (8)	0.1496 (5)	0.060 (2)	
H4A	0.04	0.3383	0.1352	0.072*	
C5	0.1983 (6)	0.4785 (6)	-0.0567 (5)	0.0509 (19)	
H5A	0.2027	0.5615	-0.043	0.061*	
C6	0.2764 (6)	0.4203 (9)	-0.0868 (5)	0.066 (2)	
H6A	0.3329	0.465	-0.093	0.079*	
C7	0.1921 (7)	0.2422 (8)	-0.0975 (5)	0.063 (2)	
H7A	0.1881	0.1592	-0.1112	0.076*	
C8	0.1128 (6)	0.2975 (7)	-0.0678 (5)	0.058 (2)	
H8A	0.0566	0.2518	-0.0622	0.07*	
B1	0.1048 (10)	0.1540 (11)	0.4316 (8)	0.081 (3)	
F1	0.1311 (10)	0.1855 (9)	0.5039 (6)	0.226 (6)	
F2	0.1324 (7)	0.2427 (8)	0.3805 (6)	0.163 (4)	

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F3	0.0102 (6)	0.1483 (9)	0.4290 (6)	0.172 (4)	
F4	0.1480 (6)	0.0465 (7)	0.4153 (5)	0.125 (3)	
N1S	0	0.224 (3)	-0.25	0.174 (15)*	0.5
C1S	0	0.124 (3)	-0.25	0.093 (9)*	0.5
C2S	0	-0.014 (3)	-0.25	0.16 (2)*	0.5
H2S1	0.039	-0.0435	-0.207	0.241*	0.25
H2S2	0.0259	-0.0435	-0.2976	0.241*	0.25
H2S3	-0.0649	-0.0435	-0.2454	0.241*	0.25

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Pt1	0.0407 (2)	0.0471 (2)	0.0350 (2)	0.00181 (19)	-0.00349 (13)	-0.00503 (18)
N1	0.045 (3)	0.053 (3)	0.041 (3)	0.007 (3)	-0.003 (2)	-0.006 (3)
N2	0.045 (3)	0.049 (3)	0.037 (3)	0.004 (3)	-0.003 (2)	0.002 (2)
N3	0.092 (7)	0.128 (9)	0.049 (5)	-0.012 (5)	-0.026 (4)	-0.011 (5)
N4	0.054 (4)	0.081 (5)	0.073 (5)	0.015 (4)	0.003 (3)	-0.014 (4)
C1	0.082 (6)	0.074 (6)	0.060 (5)	-0.027 (5)	-0.016 (5)	-0.004 (4)
C2	0.101 (8)	0.111 (9)	0.062 (6)	-0.038 (7)	-0.028 (6)	-0.007 (6)
C3	0.080 (6)	0.108 (9)	0.053 (5)	0.021 (6)	-0.004 (5)	0.013 (6)
C4	0.063 (5)	0.062 (5)	0.055 (5)	0.006 (4)	-0.016 (4)	0.004 (4)
C5	0.047 (4)	0.047 (5)	0.058 (5)	0.000 (3)	-0.001 (3)	-0.001 (3)
C6	0.046 (4)	0.080 (6)	0.072 (6)	0.002 (4)	0.005 (4)	0.001 (5)
C7	0.074 (6)	0.049 (5)	0.066 (5)	0.016 (4)	-0.004 (4)	-0.013 (4)
C8	0.061 (5)	0.052 (4)	0.062 (5)	-0.004 (4)	0.000 (4)	-0.012 (4)
B1	0.091 (9)	0.064 (7)	0.088 (9)	0.024 (6)	-0.018 (7)	-0.003 (6)
F1	0.329 (15)	0.179 (9)	0.166 (9)	0.089 (9)	-0.080 (10)	-0.091 (8)
F2	0.192 (9)	0.109 (6)	0.188 (9)	0.034 (6)	0.032 (7)	0.062 (6)
F3	0.088 (5)	0.200 (9)	0.227 (11)	0.021 (6)	-0.015 (6)	0.031 (7)
F4	0.172 (8)	0.073 (3)	0.131 (6)	0.050 (5)	-0.001 (5)	-0.006 (4)

Geometric parameters (\AA , $^\circ$)

Pt1—N2 ⁱ	2.012 (5)	C4—H4A	0.93
Pt1—N2	2.012 (5)	C5—C6	1.368 (11)
Pt1—N1 ⁱ	2.026 (6)	C5—H5A	0.93
Pt1—N1	2.026 (6)	C6—H6A	0.93
N1—C4	1.327 (10)	C7—C8	1.364 (11)
N1—C1	1.337 (10)	C7—H7A	0.93
N2—C8	1.345 (9)	C8—H8A	0.93
N2—C5	1.352 (9)	B1—F3	1.312 (14)
N3—C3	1.307 (14)	B1—F1	1.335 (14)
N3—C2	1.307 (13)	B1—F4	1.342 (12)
N4—C7	1.319 (11)	B1—F2	1.365 (14)
N4—C6	1.344 (11)	N1S—C1S	1.086 (14)
C1—C2	1.360 (12)	C1S—C2S	1.492 (10)
C1—H1A	0.93	C2S—H2S1	0.96
C2—H2A	0.93	C2S—H2S2	0.96

C3—C4	1.432 (12)	C2S—H2S3	0.96
C3—H3A	0.93		
N2 ⁱ —Pt1—N2	180.0 (3)	N2—C5—C6	120.9 (7)
N2 ⁱ —Pt1—N1 ⁱ	89.3 (2)	N2—C5—H5A	119.6
N2—Pt1—N1 ⁱ	90.7 (2)	C6—C5—H5A	119.6
N2 ⁱ —Pt1—N1	90.7 (2)	N4—C6—C5	122.3 (8)
N2—Pt1—N1	89.3 (2)	N4—C6—H6A	118.8
N1 ⁱ —Pt1—N1	180.00 (18)	C5—C6—H6A	118.8
C4—N1—C1	117.9 (7)	N4—C7—C8	123.3 (8)
C4—N1—Pt1	119.2 (5)	N4—C7—H7A	118.4
C1—N1—Pt1	123.0 (5)	C8—C7—H7A	118.4
C8—N2—C5	116.6 (6)	N2—C8—C7	121.0 (8)
C8—N2—Pt1	121.9 (5)	N2—C8—H8A	119.5
C5—N2—Pt1	121.5 (5)	C7—C8—H8A	119.5
C3—N3—C2	115.6 (9)	F3—B1—F1	106.8 (12)
C7—N4—C6	115.9 (7)	F3—B1—F4	113.9 (12)
N1—C1—C2	121.1 (9)	F1—B1—F4	107.8 (10)
N1—C1—H1A	119.4	F3—B1—F2	108.0 (10)
C2—C1—H1A	119.4	F1—B1—F2	110.5 (12)
N3—C2—C1	123.8 (10)	F4—B1—F2	109.8 (11)
N3—C2—H2A	118.1	N1S—C1S—C2S	180.000 (5)
C1—C2—H2A	118.1	C1S—C2S—H2S1	109.5
N3—C3—C4	123.4 (10)	C1S—C2S—H2S2	109.5
N3—C3—H3A	118.3	H2S1—C2S—H2S2	109.5
C4—C3—H3A	118.3	C1S—C2S—H2S3	109.5
N1—C4—C3	118.1 (8)	H2S1—C2S—H2S3	109.5
N1—C4—H4A	120.9	H2S2—C2S—H2S3	109.5
C3—C4—H4A	120.9		

Symmetry codes: (i) $-x, -y+1, -z$.

Fig. 1

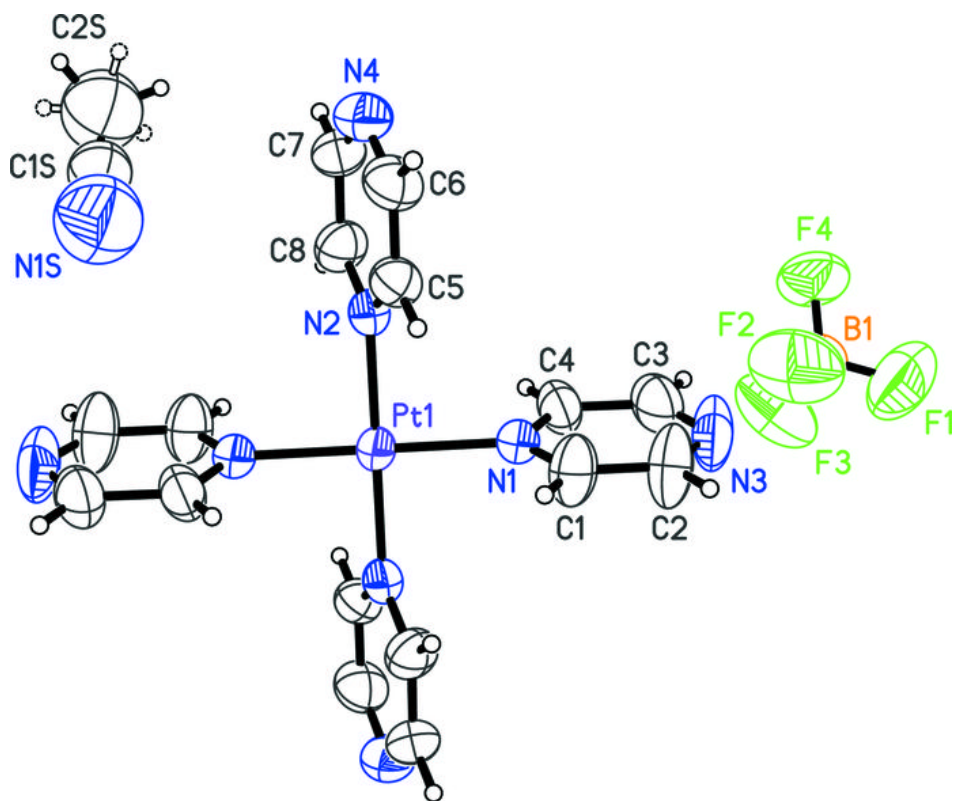


Fig. 2

