

trans-Bis(methanol- κ O)bis(quinoline-2-carboxylato- κ^2 N,O)manganese(II)

 Danuta Dobrzyńska^{a*} and Lucjan B. Jerzykiewicz^b
^aFaculty of Chemistry, Wrocław University of Technology, Wybrzeże Wyspiańskiego 27, 50-37 Wrocław, Poland, and ^bFaculty of Chemistry, University of Wrocław, Joliot-Curie 14, 50-383 Wrocław, Poland

Correspondence e-mail: danuta.dobrzynska@pwr.wroc.pl

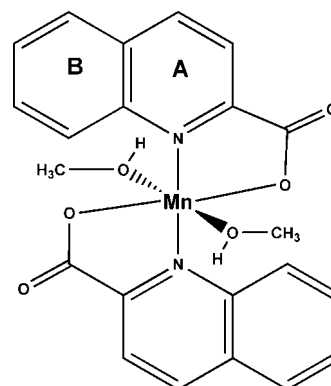
Received 18 September 2008; accepted 3 October 2008

 Key indicators: single-crystal X-ray study; $T = 100$ K; mean $\sigma(\text{C}-\text{C}) = 0.003$ Å; R factor = 0.034; wR factor = 0.089; data-to-parameter ratio = 13.2.

The title compound, $[\text{Mn}(\text{C}_{10}\text{H}_6\text{NO}_2)_2(\text{CH}_3\text{O})_2]$, was obtained unintentionally as the product of an attempt to synthesize a polynuclear carboxylate bridged manganese(III/IV) complex, using methanol to reduce the permanganate ion. The molecule is centrosymmetric; the pairs of equivalent ligands coordinate *trans* to each other in a distorted octahedral geometry. Intramolecular $\text{C}-\text{H}\cdots\text{O}$ bonds lying in the equatorial plane stabilize the molecule. In the crystal, molecules are linked by $\text{O}-\text{H}\cdots\text{O}$ and $\text{C}-\text{H}\cdots\text{O}$ hydrogen bonds, creating a three-dimensional supramolecular structure. $\pi-\pi$ and $\text{C}-\text{H}\cdots\pi$ interactions are also observed. The dihedral angle and centroid-to-centroid distance between the pyridine ring (*A*) and the benzene ring (*B*) of a symmetrically related molecule [symmetry code: (i) $-1 - x, -y, -z$] are 1.27 (11)° and 3.974 (2) Å, respectively. For the $\text{C}-\text{H}\cdots\pi$ interactions, the relevant distances and angles are: $\text{C}\cdots\text{Cg}[A^{\text{ii}}] = 3.643$ (2) Å, $\text{H}\cdots\text{Cg}[A^{\text{ii}}] = 2.750$ (2) Å and $\text{C}-\text{H}\cdots\text{Cg}[A^{\text{ii}}] = 155$ (1)° [symmetry code: (ii) $x, -1 + y, z$].

Related literature

For previously reported Mn^{II} complexes with the quinoline-2-carboxylate ligand, see: Okabe & Koizumi (1997); Goher & Mautner (1993); Haendler (1996); Dobrzyńska & Jerzykiewicz (2004); Dobrzyńska *et al.* (2005, 2006).



Experimental

Crystal data

$[\text{Mn}(\text{C}_{10}\text{H}_6\text{NO}_2)_2(\text{CH}_3\text{O})_2]$
 $M_r = 463.34$
 Monoclinic, $P2_1/n$
 $a = 10.596$ (5) Å
 $b = 7.243$ (3) Å
 $c = 13.534$ (3) Å
 $\beta = 106.59$ (4)°

$V = 995.5$ (7) Å³
 $Z = 2$
 Mo $K\alpha$ radiation
 $\mu = 0.71$ mm⁻¹
 $T = 100$ (1) K
 $0.43 \times 0.12 \times 0.09$ mm

Data collection

Kuma KM-4 CCD κ -axis diffractometer
 Absorption correction: analytical (*CrysAlis RED*; Oxford Diffraction, 2006)
 $T_{\text{min}} = 0.873$, $T_{\text{max}} = 0.902$

5405 measured reflections
 1924 independent reflections
 1475 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.031$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.034$
 $wR(F^2) = 0.089$
 $S = 0.98$
 1924 reflections
 146 parameters

H atoms treated by a mixture of independent and constrained refinement
 $\Delta\rho_{\text{max}} = 0.37$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.31$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

| $D-\text{H}\cdots A$ | $D-\text{H}$ | $\text{H}\cdots A$ | $D\cdots A$ | $D-\text{H}\cdots A$ |
|--|--------------|--------------------|-------------|----------------------|
| $\text{O3}-\text{H3}\cdots\text{O2}^{\text{i}}$ | 0.80 (3) | 1.83 (3) | 2.623 (3) | 172 (3) |
| $\text{C2}-\text{H2A}\cdots\text{O1}^{\text{ii}}$ | 0.93 | 2.58 | 3.411 (3) | 148 |
| $\text{C8}-\text{H8A}\cdots\text{O1}^{\text{iii}}$ | 0.93 | 2.36 | 3.241 (3) | 158 |

Symmetry codes: (i) $-x - \frac{1}{2}, y - \frac{1}{2}, -z - \frac{1}{2}$; (ii) $-x - \frac{1}{2}, y + \frac{1}{2}, -z - \frac{1}{2}$; (iii) $-x, -y, -z$.

Data collection: *CrysAlis CCD* (Oxford Diffraction, 2006); cell refinement: *CrysAlis RED* (Oxford Diffraction, 2006); data reduction: *CrysAlis RED*; program(s) used to solve structure: *SHELXTL-NT* (Sheldrick, 2008); program(s) used to refine structure: *SHELXTL-NT*; molecular graphics: *SHELXTL-NT*; software used to prepare material for publication: *SHELXTL-NT*.

The authors thank Wrocław University of Technology for financial support.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: SU2065).

References

- Dobrzyńska, D. & Jerzykiewicz, L. B. (2004). *J. Am. Chem. Soc.* **126**, 11118–11119.
- Dobrzyńska, D., Jerzykiewicz, L. B., Jezierska, J. & Duczmal, M. (2005). *Cryst. Growth Des.* **5**, 1945–1951.
- Dobrzyńska, D., Jerzykiewicz, L. B., Jezierska, J. & Słonec, E. (2006). *Pol. J. Chem.* **80**, 1789–1797.
- Goher, M. A. S. & Mautner, F. A. (1993). *Polyhedron*, **12**, 1863–1870.
- Haendler, H. M. (1996). *Acta Cryst.* **C52**, 801–803.
- Okabe, N. & Koizumi, M. (1997). *Acta Cryst.* **C53**, 852–854.
- Oxford Diffraction (2006). *CrysAlis CCD* and *CrysAlis RED*. Oxford Diffraction Poland, Wrocław, Poland.
- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.

supplementary materials

Acta Cryst. (2008). E64, m1383-m1384 [doi:10.1107/S1600536808031905]

trans-Bis(methanol- κ O)bis(quinoline-2-carboxylato- κ^2N,O)manganese(II)

D. Dobrzynska and LB. . Jerzykiewicz

Comment

The quinoline-2-carboxylate (quin-2-c) ion is known as an effective chelator. A few Mn(II) complexes with the quin-2-c ion and different coligands have been reported previously (Okabe *et al.*, 1997, Goher *et al.*, 1993, Haendler, 1996, Dobrzyńska *et al.*, 2004, 2005, 2006). The title complex, (I), is centrosymmetric (Fig. 1). The quin-2-c ion coordinates in a typical O,*N* chelate mode. The pairs of the equivalent ligands lie *trans* to each other in a distorted octahedral geometry. The bite angle of the chelating ligand is 74.93 (7)° and falls in the range observed for other manganese(II) complexes with the quin-2-c ion (73.1° - 78.5°; see references quoted above). The intramolecular C—H \cdots O bonds lying in the equatorial plane stabilize the molecule (Table 1).

In the crystal molecules are linked by O—H \cdots O and C—H \cdots O hydrogen bonds creating a three-dimensional supramolecular structure (see Table 1 and Fig. 2). $\pi\cdots\pi$ and C—H $\cdots\pi$ interactions are also observed. The dihedral angle and centroid-to-centroid distance between rings A [= N1,C1-C4,C9] and Bⁱ [= (C4-C9)ⁱ; symmetry code (i) -1-x, -y,-z] are 1.27° and 3.974 Å, respectively. For the C—H $\cdots\pi$ interactions the relevant distances and angles are: d(C11 \cdots Cg[Aⁱⁱ]) = 3.643 Å, d(H11A \cdots Cg[Aⁱⁱ]) = 2.750 Å with angle (C11-H11A \cdots Cg[Aⁱⁱ]) = 155° (symmetry code (ii) x, -1+y, z).

Experimental

Compound (I) was obtained unintentionally as the product of an attempt to synthesize the polynuclear, carboxylate bridged manganese(III/IV) complex mixing a methanol solution of quinoline-2-carboxylic acid and potassium permanganate at room temperature.

Refinement

The hydroxyl H-atom was located in a difference Fourier map and refined isotropically with the O-H distance restrained to 0.80 (3) Å. The C-bound H-atoms were included in calculated positions and treated as riding atoms: C-H = 0.93 - 0.96 Å with $U_{\text{iso}}(\text{H}) = 1.2$ or $1.5U_{\text{eq}}(\text{parent C atom})$.

Figures

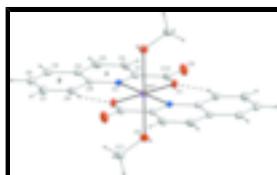


Fig. 1. The molecular structure of compound (I), with atom labels and 50% probability displacement ellipsoids for non-H atoms.

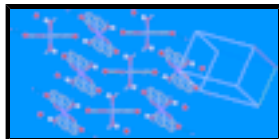


Fig. 2. The crystal packing of compound (I), showing one layer of molecules connected by O—H···O and C—H···O hydrogen bonds (dashed lines). H and O atoms participating in O—H···O bonds are shown as balls.

trans-Bis(methanol- κ O)bis(quinoline-2-carboxylato- κ^2 N,O)manganese(II)

Crystal data

[Mn(C₁₀H₆NO₂)₂(CH₄O)₂]

$M_r = 463.34$

Monoclinic, $P2_1/n$

Hall symbol: -P 2yn

$a = 10.596$ (5) Å

$b = 7.243$ (3) Å

$c = 13.534$ (3) Å

$\beta = 106.59$ (4)°

$V = 995.5$ (7) Å³

$Z = 2$

$F_{000} = 478$

$D_x = 1.546$ Mg m⁻³

Mo $K\alpha$ radiation

$\lambda = 0.71073$ Å

Cell parameters from 3842 reflections

$\theta = 3$ – 26°

$\mu = 0.71$ mm⁻¹

$T = 100$ (1) K

Block, yellow

$0.43 \times 0.12 \times 0.09$ mm

Data collection

Kuma KM-4-CCD κ -axis diffractometer

Radiation source: fine-focus sealed tube

Monochromator: graphite

$T = 100$ (1) K

ω scans

Absorption correction: analytical (CrysAlis RED; Oxford Diffraction, 2006)

$T_{\min} = 0.873$, $T_{\max} = 0.902$

5405 measured reflections

1924 independent reflections

1475 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.031$

$\theta_{\text{max}} = 26.0^\circ$

$\theta_{\text{min}} = 3.2^\circ$

$h = -13 \rightarrow 12$

$k = -8 \rightarrow 6$

$l = -16 \rightarrow 16$

Refinement

Refinement on F^2

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.034$

$wR(F^2) = 0.089$

$S = 0.98$

1924 reflections

146 parameters

Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring sites

H atoms treated by a mixture of independent and constrained refinement

$$w = 1/[\sigma^2(F_o^2) + (0.0573P)^2]$$

$$\text{where } P = (F_o^2 + 2F_c^2)/3$$

$$(\Delta/\sigma)_{\text{max}} < 0.001$$

$$\Delta\rho_{\text{max}} = 0.37 \text{ e } \text{Å}^{-3}$$

$$\Delta\rho_{\text{min}} = -0.31 \text{ e } \text{Å}^{-3}$$

Primary atom site location: structure-invariant direct methods Extinction correction: none

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

| | <i>x</i> | <i>y</i> | <i>z</i> | $U_{\text{iso}}^*/U_{\text{eq}}$ |
|------|---------------|-------------|---------------|----------------------------------|
| Mn1 | 0.0000 | 0.0000 | 0.0000 | 0.01282 (16) |
| O1 | -0.06256 (14) | 0.0841 (2) | -0.15489 (11) | 0.0151 (3) |
| O2 | -0.21857 (15) | 0.2157 (2) | -0.28166 (11) | 0.0235 (4) |
| O3 | -0.08755 (15) | -0.2745 (2) | -0.04702 (12) | 0.0179 (4) |
| N1 | -0.21067 (17) | 0.1118 (2) | -0.02417 (13) | 0.0126 (4) |
| C1 | -0.2606 (2) | 0.1766 (3) | -0.11886 (16) | 0.0137 (5) |
| C2 | -0.3844 (2) | 0.2649 (3) | -0.15224 (16) | 0.0166 (5) |
| H2A | -0.4157 | 0.3101 | -0.2191 | 0.020* |
| C3 | -0.4573 (2) | 0.2825 (3) | -0.08415 (17) | 0.0187 (5) |
| H3A | -0.5391 | 0.3400 | -0.1047 | 0.022* |
| C4 | -0.4091 (2) | 0.2138 (3) | 0.01703 (16) | 0.0156 (5) |
| C5 | -0.4793 (2) | 0.2244 (3) | 0.09220 (17) | 0.0184 (5) |
| H5A | -0.5612 | 0.2818 | 0.0756 | 0.022* |
| C6 | -0.4280 (2) | 0.1518 (3) | 0.18819 (17) | 0.0196 (5) |
| H6A | -0.4751 | 0.1600 | 0.2364 | 0.024* |
| C7 | -0.3044 (2) | 0.0644 (3) | 0.21473 (17) | 0.0184 (5) |
| H7A | -0.2713 | 0.0136 | 0.2802 | 0.022* |
| C8 | -0.2318 (2) | 0.0528 (3) | 0.14573 (16) | 0.0153 (5) |
| H8A | -0.1494 | -0.0031 | 0.1647 | 0.018* |
| C9 | -0.2833 (2) | 0.1269 (3) | 0.04511 (16) | 0.0132 (5) |
| C10 | -0.1752 (2) | 0.1571 (3) | -0.19219 (16) | 0.0145 (5) |
| C11 | -0.1393 (2) | -0.3854 (3) | 0.02034 (17) | 0.0207 (5) |
| H11A | -0.1730 | -0.4990 | -0.0136 | 0.031* |
| H11B | -0.2090 | -0.3196 | 0.0373 | 0.031* |
| H11C | -0.0705 | -0.4118 | 0.0823 | 0.031* |
| H3 | -0.141 (3) | -0.280 (4) | -0.102 (2) | 0.050 (10)* |

Atomic displacement parameters (\AA^2)

| | U^{11} | U^{22} | U^{33} | U^{12} | U^{13} | U^{23} |
|-----|------------|------------|------------|------------|--------------|------------|
| Mn1 | 0.0121 (2) | 0.0150 (3) | 0.0109 (2) | 0.0008 (2) | 0.00264 (18) | 0.0002 (2) |

supplementary materials

| | | | | | | |
|-----|-------------|-------------|-------------|--------------|-------------|--------------|
| O1 | 0.0141 (8) | 0.0196 (8) | 0.0121 (7) | 0.0013 (6) | 0.0045 (6) | 0.0006 (6) |
| O2 | 0.0180 (8) | 0.0358 (10) | 0.0161 (8) | 0.0018 (7) | 0.0040 (7) | 0.0060 (7) |
| O3 | 0.0188 (8) | 0.0180 (8) | 0.0143 (8) | -0.0024 (7) | 0.0007 (7) | 0.0003 (7) |
| N1 | 0.0132 (9) | 0.0113 (9) | 0.0131 (9) | -0.0014 (7) | 0.0035 (7) | -0.0022 (7) |
| C1 | 0.0121 (10) | 0.0118 (11) | 0.0168 (11) | -0.0023 (8) | 0.0035 (9) | -0.0017 (9) |
| C2 | 0.0149 (11) | 0.0158 (11) | 0.0179 (11) | 0.0003 (9) | 0.0027 (9) | 0.0044 (9) |
| C3 | 0.0150 (11) | 0.0148 (11) | 0.0248 (12) | 0.0041 (9) | 0.0032 (10) | 0.0033 (10) |
| C4 | 0.0154 (11) | 0.0113 (10) | 0.0200 (12) | -0.0015 (8) | 0.0049 (9) | -0.0031 (9) |
| C5 | 0.0141 (11) | 0.0157 (11) | 0.0265 (13) | 0.0004 (9) | 0.0078 (10) | -0.0037 (10) |
| C6 | 0.0197 (12) | 0.0205 (12) | 0.0224 (12) | -0.0037 (10) | 0.0120 (10) | -0.0067 (10) |
| C7 | 0.0206 (12) | 0.0200 (11) | 0.0148 (11) | -0.0023 (9) | 0.0053 (10) | -0.0010 (9) |
| C8 | 0.0144 (11) | 0.0153 (11) | 0.0158 (11) | 0.0013 (8) | 0.0037 (9) | -0.0026 (8) |
| C9 | 0.0142 (11) | 0.0095 (11) | 0.0167 (11) | -0.0022 (8) | 0.0056 (9) | -0.0018 (9) |
| C10 | 0.0144 (11) | 0.0147 (11) | 0.0119 (11) | -0.0028 (9) | -0.0001 (9) | 0.0003 (9) |
| C11 | 0.0228 (12) | 0.0167 (12) | 0.0223 (12) | -0.0006 (10) | 0.0058 (10) | 0.0002 (10) |

Geometric parameters (Å, °)

| | | | |
|-------------------------|------------|-----------|-------------|
| Mn1—O1 | 2.100 (2) | C4—C9 | 1.424 (3) |
| Mn1—O3 | 2.209 (2) | C4—C5 | 1.424 (3) |
| Mn1—N1 | 2.308 (2) | C5—C6 | 1.363 (3) |
| Mn1—O1 ⁱ | 2.100 (2) | C6—C7 | 1.406 (3) |
| Mn1—O3 ⁱ | 2.209 (2) | C7—C8 | 1.372 (3) |
| Mn1—N1 ⁱ | 2.308 (2) | C8—C9 | 1.420 (3) |
| O1—C10 | 1.271 (3) | C2—H2A | 0.93 |
| O2—C10 | 1.241 (3) | C3—H3A | 0.93 |
| O3—C11 | 1.436 (3) | C5—H5A | 0.93 |
| O3—H3 | 0.80 (3) | C6—H6A | 0.93 |
| N1—C1 | 1.325 (3) | C7—H7A | 0.93 |
| N1—C9 | 1.377 (3) | C8—H8A | 0.93 |
| C1—C10 | 1.529 (3) | C11—H11A | 0.96 |
| C1—C2 | 1.413 (3) | C11—H11B | 0.96 |
| C2—C3 | 1.367 (3) | C11—H11C | 0.96 |
| C3—C4 | 1.409 (3) | | |
| O1—Mn1—O3 | 89.23 (7) | C3—C4—C5 | 123.9 (2) |
| O1—Mn1—N1 | 74.93 (7) | C4—C5—C6 | 120.9 (2) |
| O1—Mn1—O1 ⁱ | 180.00 | C5—C6—C7 | 120.4 (2) |
| O1—Mn1—O3 ⁱ | 90.77 (7) | C6—C7—C8 | 121.1 (2) |
| O1—Mn1—N1 ⁱ | 105.07 (7) | C7—C8—C9 | 119.6 (2) |
| O3—Mn1—N1 | 88.01 (7) | N1—C9—C8 | 119.1 (2) |
| O1 ⁱ —Mn1—O3 | 90.77 (7) | C4—C9—C8 | 119.8 (2) |
| O3—Mn1—O3 ⁱ | 180.00 | N1—C9—C4 | 121.06 (19) |
| O3—Mn1—N1 ⁱ | 91.99 (7) | O2—C10—C1 | 118.6 (2) |
| O1 ⁱ —Mn1—N1 | 105.07 (7) | O1—C10—O2 | 125.0 (2) |
| O3 ⁱ —Mn1—N1 | 91.99 (7) | O1—C10—C1 | 116.33 (18) |
| N1—Mn1—N1 ⁱ | 180.00 | C1—C2—H2A | 121.00 |

| | | | |
|--------------------------------------|--------------|---------------|-------------|
| O1 ⁱ —Mn1—O3 ⁱ | 89.23 (7) | C3—C2—H2A | 121.00 |
| O1 ⁱ —Mn1—N1 ⁱ | 74.93 (7) | C2—C3—H3A | 120.00 |
| O3 ⁱ —Mn1—N1 ⁱ | 88.01 (7) | C4—C3—H3A | 120.00 |
| Mn1—O1—C10 | 120.67 (14) | C4—C5—H5A | 120.00 |
| Mn1—O3—C11 | 121.70 (13) | C6—C5—H5A | 120.00 |
| C11—O3—H3 | 105 (2) | C5—C6—H6A | 120.00 |
| Mn1—O3—H3 | 116 (2) | C7—C6—H6A | 120.00 |
| Mn1—N1—C9 | 129.59 (14) | C6—C7—H7A | 119.00 |
| C1—N1—C9 | 118.96 (19) | C8—C7—H7A | 119.00 |
| Mn1—N1—C1 | 111.36 (15) | C7—C8—H8A | 120.00 |
| C2—C1—C10 | 120.13 (19) | C9—C8—H8A | 120.00 |
| N1—C1—C2 | 123.2 (2) | O3—C11—H11A | 109.00 |
| N1—C1—C10 | 116.64 (19) | O3—C11—H11B | 109.00 |
| C1—C2—C3 | 118.6 (2) | O3—C11—H11C | 109.00 |
| C2—C3—C4 | 120.3 (2) | H11A—C11—H11B | 109.00 |
| C3—C4—C9 | 117.9 (2) | H11A—C11—H11C | 109.00 |
| C5—C4—C9 | 118.26 (19) | H11B—C11—H11C | 110.00 |
| O3—Mn1—O1—C10 | 89.82 (16) | C1—N1—C9—C4 | -1.3 (3) |
| N1—Mn1—O1—C10 | 1.67 (15) | Mn1—N1—C9—C4 | 174.99 (15) |
| O3 ⁱ —Mn1—O1—C10 | -90.18 (16) | Mn1—N1—C9—C8 | -5.6 (3) |
| N1 ⁱ —Mn1—O1—C10 | -178.33 (15) | C10—C1—C2—C3 | -179.0 (2) |
| O1—Mn1—O3—C11 | -149.62 (16) | C2—C1—C10—O1 | 176.7 (2) |
| N1—Mn1—O3—C11 | -74.68 (16) | N1—C1—C10—O1 | -1.5 (3) |
| O1 ⁱ —Mn1—O3—C11 | 30.38 (16) | N1—C1—C10—O2 | 179.83 (19) |
| N1 ⁱ —Mn1—O3—C11 | 105.32 (16) | N1—C1—C2—C3 | -0.9 (3) |
| O1—Mn1—N1—C1 | -2.35 (14) | C2—C1—C10—O2 | -1.9 (3) |
| O1—Mn1—N1—C9 | -178.82 (18) | C1—C2—C3—C4 | 0.1 (3) |
| O3—Mn1—N1—C1 | -92.09 (14) | C2—C3—C4—C9 | 0.1 (3) |
| O3—Mn1—N1—C9 | 91.44 (17) | C2—C3—C4—C5 | -179.2 (2) |
| O1 ⁱ —Mn1—N1—C1 | 177.65 (14) | C3—C4—C9—N1 | 0.5 (3) |
| O1 ⁱ —Mn1—N1—C9 | 1.18 (18) | C3—C4—C5—C6 | 178.7 (2) |
| O3 ⁱ —Mn1—N1—C1 | 87.91 (14) | C9—C4—C5—C6 | -0.6 (3) |
| O3 ⁱ —Mn1—N1—C9 | -88.56 (17) | C5—C4—C9—C8 | 0.4 (3) |
| Mn1—O1—C10—C1 | -0.8 (3) | C3—C4—C9—C8 | -178.9 (2) |
| Mn1—O1—C10—O2 | 177.74 (17) | C5—C4—C9—N1 | 179.8 (2) |
| Mn1—N1—C1—C2 | -175.44 (17) | C4—C5—C6—C7 | -0.1 (3) |
| C9—N1—C1—C2 | 1.5 (3) | C5—C6—C7—C8 | 1.1 (3) |
| C9—N1—C1—C10 | 179.63 (18) | C6—C7—C8—C9 | -1.3 (3) |
| Mn1—N1—C1—C10 | 2.7 (2) | C7—C8—C9—N1 | -178.9 (2) |
| C1—N1—C9—C8 | 178.2 (2) | C7—C8—C9—C4 | 0.5 (3) |

Symmetry codes: (i) $-x, -y, -z$.

Hydrogen-bond geometry (\AA , $^\circ$)

| $D-H\cdots A$ | $D-H$ | $H\cdots A$ | $D\cdots A$ | $D-H\cdots A$ |
|---------------------------------|----------|-------------|-------------|---------------|
| O3—H3 \cdots O2 ⁱⁱ | 0.80 (3) | 1.83 (3) | 2.623 (3) | 172 (3) |

supplementary materials

| | | | | |
|----------------------------|------|------|-----------|-----|
| C2—H2A···O1 ⁱⁱⁱ | 0.93 | 2.58 | 3.411 (3) | 148 |
| C8—H8A···O1 ⁱ | 0.93 | 2.36 | 3.241 (3) | 158 |

Symmetry codes: (ii) $-x-1/2, y-1/2, -z-1/2$; (iii) $-x-1/2, y+1/2, -z-1/2$; (i) $-x, -y, -z$.

Fig. 2

