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catena-Poly[[diaqua(2,2′-bipyridine-κ^2N,N')nickel(II)]-μ-biphenyl-2,2′-dicarboxylato-$\kappa^2O:O'$]

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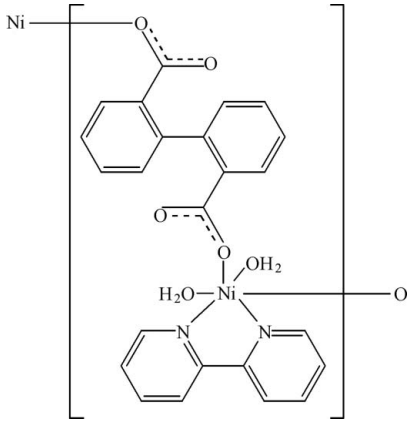
Key indicators: single-crystal X-ray study; T = 296 K; mean σ(C–C) = 0.005 Å; R factor = 0.034; wR factor = 0.088; data-to-parameter ratio = 12.8.

In the title compound, $[Ni(C_{14}H_8O_4)(C_{10}H_8N_2)(H_2O)_2]_n$, the Ni^{II} atom is coordinated in a slightly distorted octahedral geometry by two water molecules, two N atoms from a 2,2′-bipyridine ligand and two O atoms from the carboxylate groups of two 2,2′-biphenyldicarboxylate (2,2′-dpa) ligands. The 2,2′-dpa ligand acts as a bridge between neighbouring Ni^{II} atoms, forming one-dimensional coordination polymers along [100]. The coordinated water molecules form hydrogen bonds to the carboxylate O atoms of 2,2′-dpa within the same coordination polymer, and one O—H$\cdots\pi$ interaction is also formed to 2,2′-dpa.

Related literature

For other metal–organic frameworks containing 2,2′-dpa, see: Rueff *et al.* (2003); Wang *et al.* (2006); Xu *et al.* (2006).



Experimental

Crystal data

$[Ni(C_{14}H_8O_4)(C_{10}H_8N_2)(H_2O)_2]$
M_r = 491.11
Orthorhombic, $P2_12_12_1$
a = 10.9087 (15) Å
b = 11.214 (2) Å
c = 18.129 (3) Å
V = 2217.6 (6) Å^3
Z = 4
Mo $K\alpha$ radiation
μ = 0.92 mm^{-1}
T = 296 (2) K
0.42 × 0.27 × 0.19 mm

Data collection

Bruker APEXII CCD diffractometer
Absorption correction: multi-scan (*SADABS*; Bruker, 2001)
T_{min} = 0.699, T_{max} = 0.845
11746 measured reflections
3953 independent reflections
3351 reflections with $I > 2\sigma(I)$
R_{int} = 0.033

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.034
$wR(F^2)$ = 0.088
S = 1.00
3953 reflections
310 parameters
H atoms treated by a mixture of independent and constrained refinement
$\Delta\rho_{max}$ = 0.21 e Å^{-3}
$\Delta\rho_{min}$ = −0.30 e Å^{-3}
Absolute structure: Flack (1983), 1694 Friedel pairs
Flack parameter: 0.042 (16)

Table 1
Hydrogen-bond geometry (Å, °).

D—H$\cdots A$	D—H	H$\cdots A$	$D\cdots A$	D—H$\cdots A$
O5—H1$W\cdots Cg1$	0.86 (2)	2.91	3.741 (3)	163
O5—H2$W\cdots$O2^{i}	0.85 (2)	1.90 (2)	2.740 (3)	169 (3)
O6—H3$W\cdots$O4^{i}	0.81 (2)	1.91 (2)	2.676 (4)	158 (4)
O6—H4$W\cdots$O2	0.82 (2)	1.98 (2)	2.790 (3)	167 (4)

Symmetry code: (i) $x-\frac{1}{2}, -y+\frac{3}{2}, -z+2$. $Cg1$ is the centroid of the C2–C7 benzene ring.

Data collection: *APEX2* (Bruker, 2004); cell refinement: *SAINT-Plus* (Bruker, 2001); data reduction: *SAINT-Plus*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL*.

The authors acknowledge financial support from Maoming University.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: BI2317).

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Wang, R.-H., Gong, Y.-Q., Han, L., Yuan, D.-Q., Lou, B.-Y., Wu, B.-L. & Hong, M.-C. (2006). *J. Mol. Struct.* **784**, 1–6.
Xu, X.-X., Lu, Y., Wang, E.-B., Ma, Y. & Bai, X.-L. (2006). *Cryst. Growth Des.* **6**, 2029–2035.

supporting information

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Zhe An and Xian-Chun Niu

S1. Comment

2,2′-Biphenyldicarboxylic acid (H_2dpa) has been demonstrated to be a useful ligand for constructing metal-organic frameworks (Rueff *et al.,* 2003; Wang *et al.,* 2006; Xu *et al.,* 2006). The title compound is a Ni^{II} coordination polymer in which 2,2′-biphenyldicarboxylate (2,2′-dpa) acts as a bridging ligand.

The asymmetric unit (Fig. 1) contains one Ni^{II} atom coordinated by one 2,2′-bipyridine ligand, 2,2′-dpa and two water molecules. The Ni^{II} atom is hexacoordinated in a slightly distorted octahedral geometry by two water molecules, two N atoms from 2,2′-bipyridine, and two O atoms from carboxylate groups of two 2,2′-dpa. The 2,2′-dpa ligand acts as a bridge to link two neighboring Ni^{II} atoms, forming a 1-D coordination polymer along [100] (Fig. 2). Hydrogen bonds from the coordinated water molecules and the O atoms of the carboxylate groups are formed within the same coordination polymer (Fig. 3). One water molecule also forms an O—H···πi interaction to the neighbouring benzene ring of 2,2′-dpa.

S2. Experimental

A mixture of nickel(II) chloride hexahydrate (0.1 mmol), 2,2′-bipyridine (0.1 mmol), biphenyl-2,2′-dicarboxylic acid (0.2 mmol) and H_2O (16 ml) in a 25 ml Teflon-lined stainless steel autoclave was kept at 463 K for five days. Green crystals were obtained after cooling to room temperature with a yield of 12%. Elemental analysis calculated: C 58.64, H 4.89, N 5.70%; found: C 58.62, H 4.86, N 5.65%.

S3. Refinement

H atoms of the water molecules were located from difference Fourier maps and refined freely with U_{iso}(H) = 1.2U_{eq}(O). All other H atoms were placed in calculated positions with C—H = 0.93 Å and allowed to ride with U_{iso}(H) = 1.2U_{eq}(C).

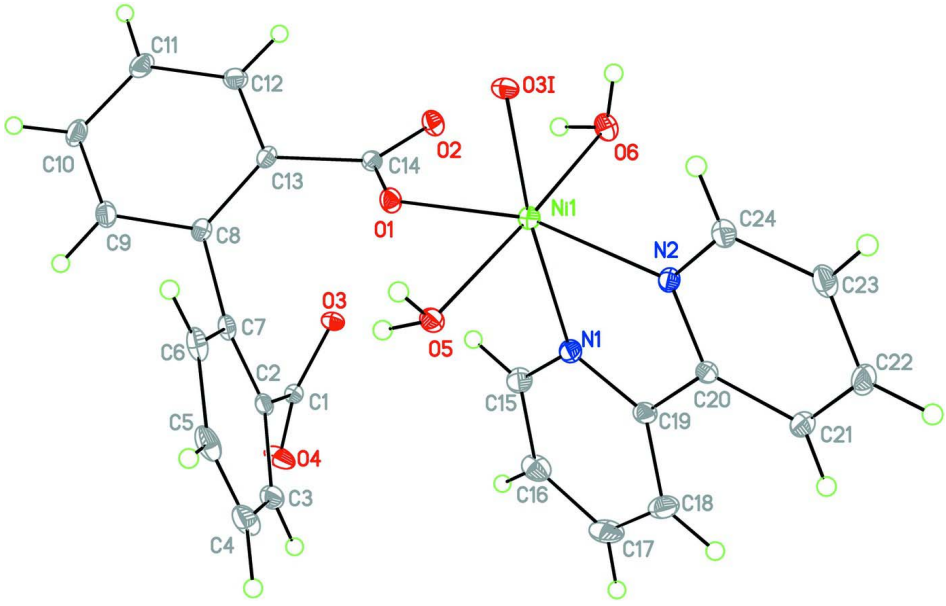

Figure 1

Asymmetric unit of the title compound showing displacement ellipsoids at 30% for non-H atoms.

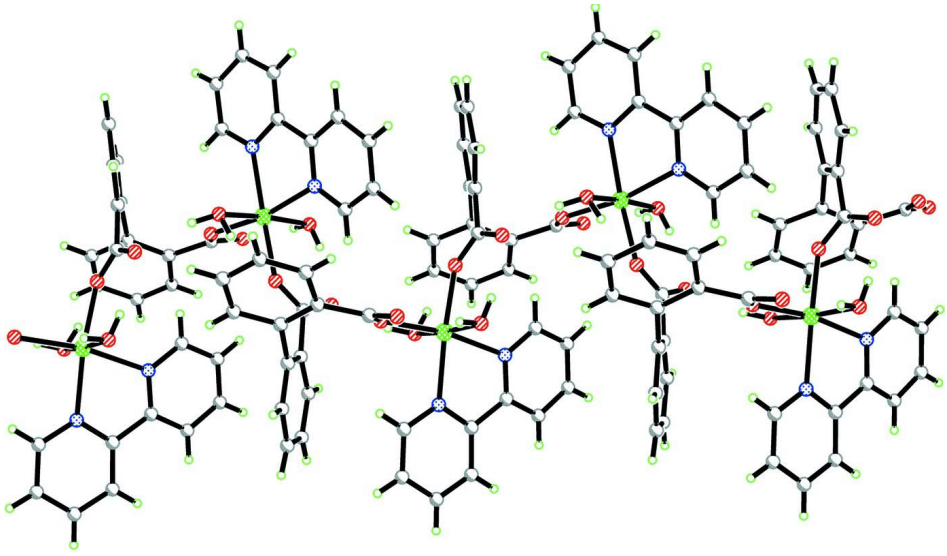

Figure 2

1-D coordination polymer running along the [100] direction.

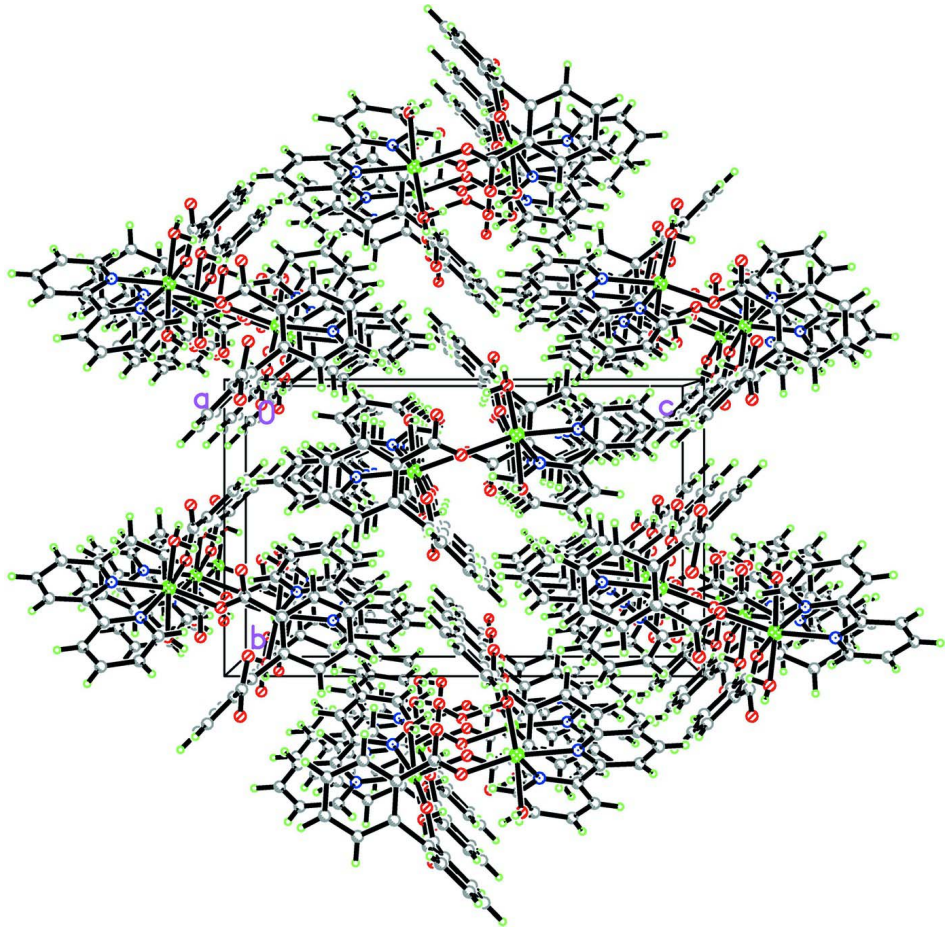

Figure 3

View of the packing along the *a* axis.

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Crystal data

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$M_r = 491.11$
Orthorhombic, $P2_12_12_1$
Hall symbol: P 2ac 2ab
$a = 10.9087\ (15)$ Å
$b = 11.214\ (2)$ Å
$c = 18.129\ (3)$ Å
$V = 2217.6\ (6)$ Å^3
$Z = 4$
$F(000) = 1016$
$D_x = 1.471$ Mg m^{-3}
Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å
Cell parameters from 3953 reflections
$\theta = 2.1–25.1°$
$\mu = 0.92$ mm^{-1}
$T = 296$ K
Block, green
0.42 × 0.27 × 0.19 mm

Data collection

Bruker APEXII CCD diffractometer
Radiation source: fine-focus sealed tube
Graphite monochromator
φ and ω scans
Absorption correction: multi-scan (*SADABS*; Bruker, 2001)
$T_{min} = 0.699$, $T_{max} = 0.845$
11746 measured reflections
3953 independent reflections
3351 reflections with $I > 2\sigma(I)$

R_{int} = 0.033
θ_{max} = 25.1°, θ_{min} = 2.1°
h = −13→12
k = −9→13
l = −20→21

Refinement

Refinement on F^2
Least-squares matrix: full
$R[F^2 > 2\sigma(F^2)]$ = 0.034
$wR(F^2)$ = 0.088
S = 1.00
3953 reflections
310 parameters
0 restraints
Primary atom site location: structure-invariant direct methods
Secondary atom site location: difference Fourier map
Hydrogen site location: inferred from neighbouring sites
H atoms treated by a mixture of independent and constrained refinement
$w = 1/[\sigma^2(F_o^2) + (0.051P)^2]$ where $P = (F_o^2 + 2F_c^2)/3$
$(\Delta/\sigma)_{max} < 0.001$
$\Delta\rho_{max}$ = 0.21 e Å^{-3}
$\Delta\rho_{min}$ = −0.30 e Å^{-3}
Absolute structure: Flack (1983), 1694 Friedel pairs
Absolute structure parameter: 0.042 (16)

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R-factor wR and goodness of fit S are based on F^2, conventional R-factors R are based on F, with F set to zero for negative F^2. The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R-factors(gt) *etc*. and is not relevant to the choice of reflections for refinement. R-factors based on F^2 are statistically about twice as large as those based on F, and R- factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å²)

	x	y	z	U_{iso}*/U_{eq}
Ni1	0.24756 (4)	0.69154 (3)	0.898676 (19)	0.03527 (12)
C1	0.5625 (3)	1.0010 (3)	1.06148 (16)	0.0367 (7)
C2	0.4305 (3)	1.0354 (3)	1.03948 (16)	0.0377 (8)
C3	0.4167 (4)	1.1232 (3)	0.98610 (18)	0.0525 (9)
H3	0.4864	1.1582	0.9661	0.063*
C4	0.3031 (5)	1.1602 (4)	0.9617 (2)	0.0673 (14)
H4	0.2977	1.2188	0.9255	0.081*
C5	0.1971 (4)	1.1108 (4)	0.9905 (2)	0.0616 (12)
H5	0.1208	1.1353	0.9735	0.074*
C6	0.2059 (3)	1.0243 (3)	1.04525 (19)	0.0515 (10)
H6	0.1351	0.9918	1.0656	0.062*
C7	0.3233 (3)	0.9853 (3)	1.07010 (16)	0.0386 (8)
C8	0.3257 (3)	0.9030 (3)	1.13631 (16)	0.0360 (7)
C9	0.3102 (3)	0.9525 (3)	1.20745 (18)	0.0492 (9)
H9	0.2988	1.0345	1.2110	0.059*
C10	0.3108 (4)	0.8874 (4)	1.27158 (19)	0.0552 (10)
H10	0.2954	0.9239	1.3167	0.066*
C11	0.3343 (4)	0.7687 (4)	1.26763 (18)	0.0553 (10)
H11	0.3383	0.7222	1.3101	0.066*
C12	0.3524 (3)	0.7181 (3)	1.19772 (17)	0.0469 (8)

H12	0.3706	0.6372	1.1950	0.056*
C13	0.3445 (3)	0.7820 (3)	1.13252 (16)	0.0356 (7)
C14	0.3500 (3)	0.7168 (3)	1.05824 (16)	0.0342 (7)
C15	0.4937 (3)	0.8346 (3)	0.8927 (2)	0.0509 (9)
H15	0.4958	0.8224	0.9435	0.061*
C16	0.5799 (3)	0.9085 (3)	0.8606 (2)	0.0608 (10)
H16	0.6392	0.9457	0.8894	0.073*
C17	0.5771 (4)	0.9267 (4)	0.7847 (3)	0.0678 (12)
H17	0.6337	0.9772	0.7625	0.081*
C18	0.4900 (4)	0.8695 (4)	0.7424 (2)	0.0596 (10)
H18	0.4877	0.8802	0.6915	0.071*
C19	0.4047 (3)	0.7949 (3)	0.77769 (18)	0.0401 (7)
C20	0.3066 (3)	0.7286 (3)	0.73743 (17)	0.0383 (8)
C21	0.3005 (4)	0.7230 (3)	0.65934 (18)	0.0527 (10)
H21	0.3595	0.7609	0.6306	0.063*
C22	0.2058 (4)	0.6605 (3)	0.6264 (2)	0.0587 (11)
H22	0.2007	0.6555	0.5753	0.070*
C23	0.1192 (4)	0.6058 (3)	0.66986 (18)	0.0525 (10)
H23	0.0547	0.5638	0.6486	0.063*
C24	0.1297 (3)	0.6143 (3)	0.74713 (18)	0.0451 (8)
H24	0.0705	0.5780	0.7764	0.054*
N1	0.4065 (2)	0.7797 (2)	0.85219 (14)	0.0402 (7)
N2	0.2231 (2)	0.6733 (2)	0.78055 (13)	0.0363 (6)
O1	0.28142 (19)	0.75413 (19)	1.00612 (11)	0.0407 (6)
O2	0.4185 (2)	0.6272 (2)	1.05422 (12)	0.0451 (6)
O3	0.58196 (19)	0.89368 (18)	1.07528 (12)	0.0385 (5)
O4	0.6422 (2)	1.0789 (2)	1.06222 (17)	0.0670 (8)
O5	0.1519 (2)	0.85857 (18)	0.89069 (12)	0.0394 (5)
H1W	0.187 (2)	0.895 (3)	0.9264 (15)	0.047*
H2W	0.0764 (16)	0.858 (3)	0.9025 (17)	0.047*
O6	0.3552 (2)	0.5345 (2)	0.91659 (13)	0.0466 (6)
H3W	0.302 (3)	0.484 (3)	0.9225 (18)	0.056*
H4W	0.385 (3)	0.557 (3)	0.9558 (13)	0.056*

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Ni1	0.0390 (2)	0.0344 (2)	0.0325 (2)	−0.0010 (2)	0.0015 (2)	−0.00163 (16)
C1	0.047 (2)	0.0325 (18)	0.0310 (16)	0.0033 (16)	0.0056 (14)	0.0009 (14)
C2	0.058 (2)	0.0275 (16)	0.0275 (15)	0.0093 (16)	−0.0028 (15)	−0.0034 (13)
C3	0.080 (3)	0.0377 (19)	0.0401 (19)	0.006 (2)	−0.0049 (19)	0.0003 (17)
C4	0.114 (4)	0.042 (2)	0.045 (2)	0.026 (2)	−0.033 (2)	−0.0035 (18)
C5	0.080 (3)	0.049 (2)	0.056 (2)	0.034 (2)	−0.032 (2)	−0.015 (2)
C6	0.054 (2)	0.052 (2)	0.048 (2)	0.0185 (18)	−0.0131 (16)	−0.0215 (18)
C7	0.049 (2)	0.0350 (18)	0.0319 (16)	0.0132 (15)	−0.0035 (15)	−0.0093 (14)
C8	0.0320 (17)	0.0416 (18)	0.0343 (16)	0.0010 (14)	0.0026 (14)	−0.0022 (14)
C9	0.056 (2)	0.046 (2)	0.045 (2)	0.0093 (17)	0.0060 (17)	−0.0082 (17)
C10	0.065 (2)	0.068 (3)	0.0335 (18)	−0.004 (2)	0.0110 (17)	−0.0084 (18)

C11	0.071 (3)	0.065 (3)	0.0296 (17)	−0.011 (2)	0.0051 (18)	0.0021 (17)
C12	0.059 (2)	0.0417 (19)	0.0402 (18)	−0.0088 (17)	−0.0019 (17)	0.0041 (16)
C13	0.0349 (17)	0.0436 (19)	0.0283 (15)	−0.0034 (15)	0.0040 (13)	−0.0020 (14)
C14	0.0370 (17)	0.0334 (18)	0.0321 (16)	−0.0062 (15)	0.0009 (14)	−0.0021 (14)
C15	0.0413 (19)	0.054 (2)	0.057 (2)	−0.0117 (16)	0.0082 (18)	−0.0048 (19)
C16	0.043 (2)	0.053 (2)	0.087 (3)	−0.0107 (19)	0.007 (2)	−0.002 (2)
C17	0.048 (2)	0.054 (2)	0.102 (3)	−0.013 (2)	0.015 (2)	0.019 (2)
C18	0.061 (2)	0.056 (2)	0.062 (2)	−0.002 (2)	0.016 (2)	0.018 (2)
C19	0.0407 (17)	0.0313 (17)	0.0482 (19)	0.0054 (15)	0.0106 (15)	0.0042 (15)
C20	0.0513 (19)	0.0280 (16)	0.0356 (17)	0.0079 (15)	0.0065 (15)	0.0024 (14)
C21	0.078 (3)	0.043 (2)	0.0376 (18)	0.0082 (19)	0.0099 (18)	0.0045 (16)
C22	0.096 (3)	0.050 (2)	0.0305 (17)	0.018 (2)	−0.0062 (19)	0.0007 (17)
C23	0.072 (3)	0.041 (2)	0.044 (2)	0.0062 (19)	−0.0159 (18)	−0.0114 (18)
C24	0.053 (2)	0.042 (2)	0.0401 (19)	−0.0008 (17)	−0.0012 (16)	−0.0061 (16)
N1	0.0388 (15)	0.0395 (15)	0.0423 (16)	−0.0044 (13)	0.0074 (12)	−0.0037 (13)
N2	0.0432 (16)	0.0361 (14)	0.0295 (12)	0.0032 (12)	0.0015 (11)	−0.0020 (11)
O1	0.0520 (14)	0.0380 (12)	0.0322 (11)	0.0053 (10)	−0.0061 (10)	−0.0038 (9)
O2	0.0518 (14)	0.0400 (13)	0.0436 (13)	0.0029 (12)	−0.0086 (11)	−0.0076 (11)
O3	0.0403 (12)	0.0297 (12)	0.0456 (12)	0.0018 (10)	−0.0019 (10)	0.0082 (10)
O4	0.0539 (15)	0.0333 (14)	0.114 (2)	−0.0040 (13)	0.0030 (16)	0.0071 (15)
O5	0.0418 (12)	0.0328 (12)	0.0436 (13)	0.0019 (10)	0.0014 (11)	−0.0006 (10)
O6	0.0465 (15)	0.0415 (14)	0.0519 (15)	0.0074 (11)	−0.0036 (12)	−0.0072 (12)

Geometric parameters (Å, º)

Ni1—O3^{i}	2.098 (2)	C13—C14	1.533 (4)
Ni1—O1	2.103 (2)	C14—O2	1.255 (4)
Ni1—O6	2.142 (2)	C14—O1	1.276 (3)
Ni1—O5	2.149 (2)	C15—N1	1.351 (4)
Ni1—N1	2.166 (3)	C15—C16	1.383 (5)
Ni1—N2	2.168 (2)	C15—H15	0.930
C1—O4	1.233 (4)	C16—C17	1.391 (6)
C1—O3	1.247 (4)	C16—H16	0.930
C1—C2	1.543 (5)	C17—C18	1.379 (6)
C2—C3	1.388 (5)	C17—H17	0.930
C2—C7	1.412 (5)	C18—C19	1.406 (5)
C3—C4	1.380 (6)	C18—H18	0.930
C3—H3	0.930	C19—N1	1.361 (4)
C4—C5	1.385 (6)	C19—C20	1.493 (5)
C4—H4	0.930	C20—N2	1.352 (4)
C5—C6	1.391 (5)	C20—C21	1.419 (4)
C5—H5	0.930	C21—C22	1.385 (6)
C6—C7	1.427 (4)	C21—H21	0.930
C6—H6	0.930	C22—C23	1.374 (5)
C7—C8	1.514 (4)	C22—H22	0.930
C8—C13	1.374 (5)	C23—C24	1.409 (5)
C8—C9	1.414 (4)	C23—H23	0.930
C9—C10	1.373 (5)	C24—N2	1.357 (4)

C9—H9	0.930	C24—H24	0.930
C10—C11	1.358 (6)	O3—Ni1^{ii}	2.098 (2)
C10—H10	0.930	O5—H1W	0.86 (2)
C11—C12	1.402 (5)	O5—H2W	0.85 (2)
C11—H11	0.930	O6—H3W	0.81 (2)
C12—C13	1.385 (4)	O6—H4W	0.82 (2)
C12—H12	0.930		

O3^{i}—Ni1—O1	95.44 (8)	C8—C13—C12	118.5 (3)
O3^{i}—Ni1—O6	93.63 (9)	C8—C13—C14	121.3 (3)
O1—Ni1—O6	92.16 (9)	C12—C13—C14	120.1 (3)
O3^{i}—Ni1—O5	89.65 (8)	O2—C14—O1	124.7 (3)
O1—Ni1—O5	81.77 (8)	O2—C14—C13	117.1 (3)
O6—Ni1—O5	173.36 (9)	O1—C14—C13	118.1 (3)
O3^{i}—Ni1—N1	169.97 (9)	N1—C15—C16	121.5 (4)
O1—Ni1—N1	93.86 (9)	N1—C15—H15	119.2
O6—Ni1—N1	89.75 (10)	C16—C15—H15	119.2
O5—Ni1—N1	87.97 (9)	C15—C16—C17	119.3 (4)
O3^{i}—Ni1—N2	94.20 (9)	C15—C16—H16	120.3
O1—Ni1—N2	165.48 (9)	C17—C16—H16	120.4
O6—Ni1—N2	98.04 (9)	C18—C17—C16	119.8 (4)
O5—Ni1—N2	87.46 (9)	C18—C17—H17	120.1
N1—Ni1—N2	75.96 (10)	C16—C17—H17	120.1
O4—C1—O3	124.2 (3)	C17—C18—C19	118.7 (4)
O4—C1—C2	118.9 (3)	C17—C18—H18	120.7
O3—C1—C2	116.9 (3)	C19—C18—H18	120.7
C3—C2—C7	117.8 (3)	N1—C19—C18	121.1 (3)
C3—C2—C1	117.3 (3)	N1—C19—C20	115.7 (3)
C7—C2—C1	124.9 (3)	C18—C19—C20	123.2 (3)
C4—C3—C2	122.3 (4)	N2—C20—C21	121.7 (3)
C4—C3—H3	118.8	N2—C20—C19	115.4 (3)
C2—C3—H3	118.8	C21—C20—C19	122.9 (3)
C5—C4—C3	120.5 (3)	C22—C21—C20	119.2 (4)
C5—C4—H4	119.7	C22—C21—H21	120.4
C3—C4—H4	119.7	C20—C21—H21	120.4
C4—C5—C6	119.4 (3)	C23—C22—C21	119.5 (3)
C4—C5—H5	120.3	C23—C22—H22	120.3
C6—C5—H5	120.3	C21—C22—H22	120.3
C5—C6—C7	120.0 (4)	C22—C23—C24	119.0 (3)
C5—C6—H6	120.0	C22—C23—H23	120.5
C7—C6—H6	120.0	C24—C23—H23	120.5
C2—C7—C6	119.8 (3)	N2—C24—C23	122.5 (3)
C2—C7—C8	122.7 (3)	N2—C24—H24	118.7
C6—C7—C8	116.9 (3)	C23—C24—H24	118.7
C13—C8—C9	116.8 (3)	C15—N1—C19	119.5 (3)
C13—C8—C7	124.4 (3)	C15—N1—Ni1	124.1 (2)
C9—C8—C7	118.8 (3)	C19—N1—Ni1	115.6 (2)
C10—C9—C8	124.3 (3)	C20—N2—C24	118.1 (3)

C10—C9—H9	117.9	C20—N2—Ni1	116.4 (2)
C8—C9—H9	117.8	C24—N2—Ni1	125.4 (2)
C11—C10—C9	118.5 (3)	C14—O1—Ni1	132.81 (19)
C11—C10—H10	120.7	C1—O3—Ni1^{ii}	129.1 (2)
C9—C10—H10	120.8	Ni1—O5—H1W	99 (2)
C10—C11—C12	118.1 (3)	Ni1—O5—H2W	117 (2)
C10—C11—H11	121.0	H1W—O5—H2W	104 (2)
C12—C11—H11	120.9	Ni1—O6—H3W	102 (3)
C13—C12—C11	123.6 (3)	Ni1—O6—H4W	95 (3)
C13—C12—H12	118.2	H3W—O6—H4W	112 (3)
C11—C12—H12	118.2		

Symmetry codes: (i) $x-1/2, -y+3/2, -z+2$; (ii) $x+1/2, -y+3/2, -z+2$.

Hydrogen-bond geometry (Å, º)

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
O5—H1*W*···*Cg*1	0.86 (2)	2.91	3.741 (3)	163
O5—H2*W*···O2^{i}	0.85 (2)	1.90 (2)	2.740 (3)	169 (3)
O6—H3*W*···O4^{i}	0.81 (2)	1.91 (2)	2.676 (4)	158 (4)
O6—H4*W*···O2	0.82 (2)	1.98 (2)	2.790 (3)	167 (4)

Symmetry code: (i) $x-1/2, -y+3/2, -z+2$.