

Acta Crystallographica Section E

Structure Reports

Online

ISSN 1600-5368

2-Methyl-2-(4-nitrophenoxy)propanoic acid

 Gabriel Navarrete-Vázquez,^a Hector Torres-Gómez,^a Sergio Hidalgo-Figueroa^a and Hugo Tlahuext^{b*}

^aFacultad de Farmacia, Universidad Autónoma del Estado de Morelos, Av. Universidad 1001 Col Chamilpa CP 62100, Cuernavaca Mor., México, and ^bCentro de Investigaciones Químicas, Universidad Autónoma del Estado de Morelos, Av. Universidad 1001 Col. Chamilpa, CP 62100, Cuernavaca Mor., México
Correspondence e-mail: tlahuext@ciq.uaem.mx

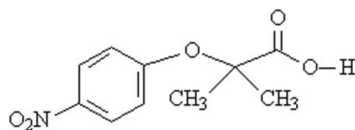
Received 17 October 2008; accepted 29 October 2008

Key indicators: single-crystal X-ray study; $T = 294$ K; mean $\sigma(\text{C}-\text{C}) = 0.004$ Å; R factor = 0.065; wR factor = 0.163; data-to-parameter ratio = 12.6.

The title compound, $\text{C}_{10}\text{H}_{11}\text{NO}_5$, is of interest with respect to its antidyslipidemic activity. It was prepared by reaction of 4-nitrophenol with ethyl 2-bromo-2-methylpropionate followed by ethyl ester hydrolysis. In the crystal, molecules are linked into centrosymmetric dimers by intermolecular $\text{O}-\text{H}\cdots\text{O}$ hydrogen bonds and the dimers are connected into chains by weak $\text{C}-\text{H}\cdots\text{O}$ interactions. The packing is further stabilized by offset $\pi-\pi$ interactions between adjacent benzene rings with a centroid-centroid distance of 3.8643 (17) Å.

Related literature

For related literature on fibrate structures and hypolipidemic activity, see: Navarrete-Vázquez *et al.* (2008); Henry *et al.* (2003); Rath *et al.* (2005); Djinović *et al.* (1989); Thorp (1962); Thorp & Waring (1962); Miller & Spence (1998); Forcheron *et al.* (2002). For details of the graph-set analysis of hydrogen-bonding patterns, see: Bernstein *et al.* (1995)



Experimental

Crystal data

$\text{C}_{10}\text{H}_{11}\text{NO}_5$	$c = 14.518$ (2) Å
$M_r = 225.20$	$\beta = 93.794$ (2)°
Monoclinic, $C2/c$	$V = 2170.2$ (5) Å ³
$a = 21.296$ (3) Å	$Z = 8$
$b = 7.0348$ (9) Å	Mo $K\alpha$ radiation

$\mu = 0.11$ mm ⁻¹	$0.32 \times 0.25 \times 0.20$ mm
$T = 294$ (2) K	

Data collection

Bruker SMART APEX CCD area-detector diffractometer	10072 measured reflections
Absorption correction: multi-scan (SADABS; Sheldrick, 1996)	1920 independent reflections
$T_{\min} = 0.763$, $T_{\max} = 0.978$	1745 reflections with $I > 2\sigma(I)$
	$R_{\text{int}} = 0.035$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.065$	H atoms treated by a mixture of independent and constrained refinement
$wR(F^2) = 0.163$	$\Delta\rho_{\text{max}} = 0.23$ e Å ⁻³
$S = 1.12$	$\Delta\rho_{\text{min}} = -0.22$ e Å ⁻³
1920 reflections	
152 parameters	

Table 1
Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{O5}-\text{H5A}\cdots\text{O4}^i$	0.92 (4)	1.75 (4)	2.659 (3)	173 (3)
$\text{C6}-\text{H6}\cdots\text{O1}^{ii}$	0.93	2.34	3.165 (4)	147

Symmetry codes: (i) $-x + \frac{1}{2}, -y + \frac{3}{2}, -z + 1$; (ii) $x, y + 1, z$.

Data collection: SMART (Bruker, 2000); cell refinement: SAINT-Plus-NT (Bruker, 2001); data reduction: SAINT-Plus-NT; program(s) used to solve structure: SHELXTL-NT (Sheldrick, 2008); program(s) used to refine structure: SHELXTL-NT; molecular graphics: SHELXTL-NT (Sheldrick, 2008); software used to prepare material for publication: PLATON (Spek, 2003) and publCIF (Westrip, 2008).

This work was supported by the Consejo Nacional de Ciencia y Tecnología (CONACyT) under grants Nos. 55591 and 3562P-E.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: SJ2550).

References

- Bernstein, J., Davis, R. E., Shimoni, L. & Chang, N.-L. (1995). *Angew. Chem. Int. Ed. Engl.* **34**, 1555-1573.
- Bruker (2000). SMART. Bruker AXS Inc., Madison, Wisconsin, USA.
- Bruker (2001). SAINT-Plus-NT. Bruker AXS Inc., Madison, Wisconsin, USA.
- Djinović, K., Globokar, M. & Zupet, P. (1989). *Acta Cryst.* **C45**, 772-775.
- Forcheron, F., Cachefo, A., Thevenon, S., Pinteur, C. & Beylot, M. (2002). *Diabetes*, **51**, 3486-3491.
- Henry, R. F., Zhang, G. Z., Gao, Y. & Buckner, I. S. (2003). *Acta Cryst.* **E59**, o699-o700.
- Miller, D. B. & Spence, J. D. (1998). *Clin. Pharmacokinet.* **34**, 155-162.
- Navarrete-Vázquez, G., Villalobos-Molina, R., Estrada-Soto, S., Ortiz-Andrade, R. & Tlahuext, H. (2008). *Acta Cryst.* **E64**, o91.
- Rath, N. P., Haq, W. & Balendiran, G. K. (2005). *Acta Cryst.* **C61**, o81-o84.
- Sheldrick, G. M. (1996). SADABS. University of Göttingen, Germany.
- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112-122.
- Spek, A. L. (2003). *J. Appl. Cryst.* **36**, 7-13.
- Thorp, J. M. (1962). *Lancet*, **1**, 1323-1326.
- Thorp, J. M. & Waring, W. S. (1962). *Nature (London)*, **194**, 948-949.
- Westrip, S. P. (2008). publCIF. In preparation.

supplementary materials

Acta Cryst. (2008). E64, o2261 [doi:10.1107/S1600536808035411]

2-Methyl-2-(4-nitrophenoxy)propanoic acid

G. Navarrete-Vázquez, H. Torres-Gómez, S. Hidalgo-Figueroa and H. Tlahuext

Comment

The fibrates belong to a class of lipid-modifying agents that decrease plasma triglycerides (Thorp, 1962; Miller & Spence, 1998; Forcheron *et al.*, 2002). These compounds are used as therapeutic agents in the treatment of dyslipidemia, heart disease and diabetic complications in humans. The fibric acid pharmacophore has been of interest to medicinal chemists ever since the initial discovery that ethyl chlorophenoxyisobutyrate possessed hypolipidemic properties (Thorp & Waring, 1962).

In order to assist our knowledge about the electronic and steric requirements to shown antihyperlipidemic activity, we have determined the crystal structure of the title compound, (I), Fig 1, which is a bioisoster of clofibrac acid, with a nitro group instead of chlorine atom. The crystal structure is stabilized by strong O5—H5A···O4 hydrogen-bonding interactions, forming $R_2^2(8)$ motifs (Bernstein *et al.*, 1995) (Fig. 1; Table 2). In the crystal packing there are also weak C6—H6···O1 hydrogen bonding interactions that interconnect molecules into chains running along the *b* axis. The crystal structure is also stabilized by offset π – π interactions between two adjacent molecules, with a distance between centroids of the C1—C6 benzene rings [$Cg1, Cg1'$ (Symmetry code: $-x, y, -z + 1/2$)] of 3.8643 (17) Å. (Fig. 2; Table 1).

Experimental

A mixture of 4-nitrophenol (1.0 g, 4.44 mmol), potassium carbonate (1.22 g, 8.88 mmol) in acetonitrile, was added dropwise to 1.04 ml of ethyl 2-bromo-2-methylpropionate (1.29 g, 6.66 mmol). The mixture was stirred and heated under reflux for 8 h then poured onto cold water. The resulting oil was treated with a mixture of THF/MeOH/H₂O (3:2:1), and LiOH (5 equiv) was added. The mixture stirred at room temperature for 3 h, 10% HCl solution added, and most of the organic solvents removed *in vacuo*. The partly solid residue was extracted with CH₂Cl₂ (3 x 10 ml), dried with Na₂SO₄, filtered, and concentrated in vacuo to give a yellow solid (m.p. 396 K). Single crystals of (I) were obtained from chloroform.

Refinement

Aromatic and methyl H atoms were positioned geometrically, were constrained to the riding-model approximation [$C-H_{\text{aryl}} = 0.93$ Å, $U_{\text{iso}}(H_{\text{aryl}}) = 1.2 U_{\text{eq}}(C_{\text{aryl}})$; $C-H_{\text{methyl}} = 0.96$ Å, $U_{\text{iso}}(H_{\text{methyl}}) = 1.5 U_{\text{eq}}(C_{\text{methyl}})$]. Atom H5A, which is involved in hydrogen-bonding interaction, was located in a difference Fourier map and refined freely with isotropic displacement parameters.

Figures

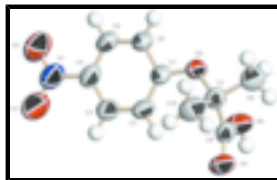


Fig. 1. The molecular structure of I showing 50% probability displacement ellipsoids and the atomic numbering. H atoms are shown as small spheres of arbitrary radius.

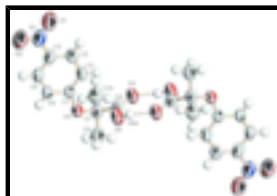


Fig. 2. Centrosymmetric dimers generated by O4—H4...O3 intermolecular hydrogen bonds (dotted lines) forming an $R_2^2(8)$ motif.

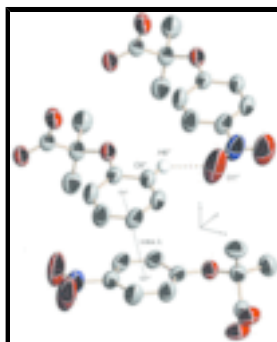


Fig. 3. View of the offset π - π interaction and C—H...O hydrogen bonds between adjacent molecules. The dashed line indicates the vector between the centroids (Cg1, Cg1''). The hydrogen bond is represented by a dotted line and H atoms not involved in hydrogen bonding have been omitted for clarity.

2-Methyl-2-(4-nitrophenoxy)propanoic acid

Crystal data

$C_{10}H_{11}NO_5$

$M_r = 225.20$

Monoclinic, $C2/c$

Hall symbol: $-C 2yc$

$a = 21.296 (3) \text{ \AA}$

$b = 7.0348 (9) \text{ \AA}$

$c = 14.518 (2) \text{ \AA}$

$\beta = 93.794 (2)^\circ$

$V = 2170.2 (5) \text{ \AA}^3$

$Z = 8$

$F_{000} = 944$

$D_x = 1.378 \text{ Mg m}^{-3}$

Melting point: 396 K

Mo $K\alpha$ radiation

$\lambda = 0.71073 \text{ \AA}$

Cell parameters from 5496 reflections

$\theta = 2.5\text{--}2.9^\circ$

$\mu = 0.11 \text{ mm}^{-1}$

$T = 294 (2) \text{ K}$

Prism, colourless

$0.32 \times 0.25 \times 0.20 \text{ mm}$

Data collection

Bruker SMART APEX CCD area-detector diffractometer

Radiation source: fine-focus sealed tube

Monochromator: graphite

$T = 294(2) \text{ K}$

1920 independent reflections

1745 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.035$

$\theta_{\text{max}} = 25.0^\circ$

φ and ω scans	$\theta_{\min} = 1.9^\circ$
Absorption correction: multi-scan (SADABS; Sheldrick, 1996)	$h = -25 \rightarrow 25$
$T_{\min} = 0.763$, $T_{\max} = 0.978$	$k = -8 \rightarrow 8$
10072 measured reflections	$l = -17 \rightarrow 17$

Refinement

Refinement on F^2	Hydrogen site location: inferred from neighbouring sites
Least-squares matrix: full	H atoms treated by a mixture of independent and constrained refinement
$R[F^2 > 2\sigma(F^2)] = 0.065$	$w = 1/[\sigma^2(F_o^2) + (0.0689P)^2 + 2.0562P]$ where $P = (F_o^2 + 2F_c^2)/3$
$wR(F^2) = 0.163$	$(\Delta/\sigma)_{\max} < 0.001$
$S = 1.12$	$\Delta\rho_{\max} = 0.23 \text{ e } \text{\AA}^{-3}$
1920 reflections	$\Delta\rho_{\min} = -0.22 \text{ e } \text{\AA}^{-3}$
152 parameters	Extinction correction: SHELXTL-NT (Sheldrick, 2008), $F_c^* = kFc[1 + 0.001 \times Fc^2 \lambda^3 / \sin(2\theta)]^{-1/4}$
Primary atom site location: structure-invariant direct methods	Extinction coefficient: 0.020 (2)
Secondary atom site location: difference Fourier map	

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
C1	0.43849 (11)	0.9257 (3)	0.38604 (15)	0.0513 (6)
C2	0.42396 (12)	0.7351 (4)	0.3768 (2)	0.0696 (8)
H2	0.3822	0.6954	0.3747	0.084*
C3	0.47162 (13)	0.6040 (4)	0.3706 (2)	0.0783 (9)
H3	0.4623	0.4754	0.3641	0.094*
C4	0.53299 (12)	0.6654 (4)	0.3741 (2)	0.0667 (7)
C5	0.54791 (12)	0.8533 (4)	0.38383 (19)	0.0661 (7)
H5	0.5897	0.8926	0.3866	0.079*
C6	0.50077 (11)	0.9819 (4)	0.38940 (18)	0.0613 (7)
H6	0.5106	1.1103	0.3956	0.074*

supplementary materials

C7	0.33120 (11)	1.0562 (4)	0.36820 (17)	0.0584 (7)
C8	0.29840 (10)	0.9120 (3)	0.42697 (16)	0.0532 (6)
C9	0.30573 (14)	1.2523 (4)	0.3905 (2)	0.0801 (9)
H9A	0.3183	1.2844	0.4533	0.120*
H9B	0.2606	1.2512	0.3823	0.120*
H9C	0.3222	1.3448	0.3499	0.120*
C10	0.32173 (14)	1.0132 (5)	0.26611 (19)	0.0772 (8)
H10A	0.3449	1.1033	0.2320	0.116*
H10B	0.2778	1.0221	0.2471	0.116*
H10C	0.3365	0.8871	0.2544	0.116*
H5A	0.2918 (16)	0.838 (5)	0.547 (3)	0.101 (11)*
N1	0.58387 (13)	0.5287 (4)	0.3648 (2)	0.0966 (9)
O1	0.57049 (14)	0.3662 (4)	0.3469 (3)	0.1713 (19)
O2	0.63689 (12)	0.5842 (4)	0.3687 (3)	0.1449 (13)
O3	0.39722 (7)	1.0706 (2)	0.39675 (12)	0.0594 (5)
O4	0.25911 (8)	0.8033 (3)	0.39279 (12)	0.0711 (6)
O5	0.31446 (10)	0.9212 (3)	0.51375 (13)	0.0770 (7)

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
C1	0.0472 (12)	0.0552 (14)	0.0519 (13)	-0.0041 (10)	0.0076 (9)	0.0040 (10)
C2	0.0477 (13)	0.0595 (16)	0.103 (2)	-0.0087 (12)	0.0140 (13)	0.0028 (14)
C3	0.0658 (17)	0.0506 (15)	0.121 (2)	-0.0042 (13)	0.0250 (16)	0.0062 (15)
C4	0.0515 (14)	0.0686 (17)	0.0816 (18)	0.0059 (13)	0.0163 (12)	0.0123 (14)
C5	0.0453 (13)	0.0758 (18)	0.0784 (17)	-0.0102 (13)	0.0142 (12)	-0.0012 (14)
C6	0.0536 (14)	0.0583 (14)	0.0736 (16)	-0.0112 (12)	0.0151 (11)	-0.0034 (12)
C7	0.0481 (13)	0.0609 (15)	0.0659 (15)	0.0001 (11)	0.0021 (10)	0.0060 (12)
C8	0.0418 (11)	0.0595 (14)	0.0583 (14)	0.0012 (10)	0.0025 (10)	-0.0016 (11)
C9	0.0691 (17)	0.0653 (17)	0.106 (2)	0.0078 (14)	0.0085 (15)	0.0093 (16)
C10	0.0730 (18)	0.092 (2)	0.0657 (17)	-0.0098 (16)	0.0012 (13)	0.0144 (15)
N1	0.0659 (17)	0.0787 (19)	0.148 (3)	0.0083 (14)	0.0311 (16)	0.0179 (18)
O1	0.105 (2)	0.0664 (17)	0.353 (6)	0.0126 (15)	0.091 (3)	0.017 (2)
O2	0.0574 (15)	0.127 (2)	0.252 (4)	0.0171 (15)	0.0225 (18)	-0.013 (2)
O3	0.0483 (9)	0.0532 (10)	0.0770 (12)	-0.0036 (7)	0.0049 (8)	-0.0015 (8)
O4	0.0619 (11)	0.0864 (14)	0.0648 (11)	-0.0253 (10)	0.0017 (8)	-0.0029 (9)
O5	0.0797 (13)	0.0928 (15)	0.0581 (11)	-0.0337 (11)	0.0008 (9)	0.0035 (10)

Geometric parameters (\AA , $^\circ$)

C1—O3	1.361 (3)	C7—C8	1.524 (3)
C1—C2	1.380 (4)	C7—C9	1.525 (4)
C1—C6	1.382 (3)	C8—O4	1.215 (3)
C2—C3	1.379 (4)	C8—O5	1.285 (3)
C2—H2	0.9300	C9—H9A	0.9600
C3—C4	1.374 (4)	C9—H9B	0.9600
C3—H3	0.9300	C9—H9C	0.9600
C4—C5	1.365 (4)	C10—H10A	0.9600
C4—N1	1.462 (4)	C10—H10B	0.9600

C5—C6	1.358 (4)	C10—H10C	0.9600
C5—H5	0.9300	N1—O2	1.193 (4)
C6—H6	0.9300	N1—O1	1.202 (4)
C7—O3	1.443 (3)	O5—H5A	0.92 (4)
C7—C10	1.513 (4)		
O3—C1—C2	126.6 (2)	C10—C7—C9	111.2 (2)
O3—C1—C6	114.0 (2)	C8—C7—C9	107.5 (2)
C2—C1—C6	119.3 (2)	O4—C8—O5	124.3 (2)
C3—C2—C1	119.6 (2)	O4—C8—C7	121.3 (2)
C3—C2—H2	120.2	O5—C8—C7	114.4 (2)
C1—C2—H2	120.2	C7—C9—H9A	109.5
C4—C3—C2	119.3 (3)	C7—C9—H9B	109.5
C4—C3—H3	120.3	H9A—C9—H9B	109.5
C2—C3—H3	120.3	C7—C9—H9C	109.5
C5—C4—C3	121.5 (2)	H9A—C9—H9C	109.5
C5—C4—N1	118.6 (2)	H9B—C9—H9C	109.5
C3—C4—N1	119.9 (3)	C7—C10—H10A	109.5
C6—C5—C4	118.9 (2)	C7—C10—H10B	109.5
C6—C5—H5	120.5	H10A—C10—H10B	109.5
C4—C5—H5	120.5	C7—C10—H10C	109.5
C5—C6—C1	121.2 (2)	H10A—C10—H10C	109.5
C5—C6—H6	119.4	H10B—C10—H10C	109.5
C1—C6—H6	119.4	O2—N1—O1	122.1 (3)
O3—C7—C10	111.1 (2)	O2—N1—C4	119.0 (3)
O3—C7—C8	111.08 (19)	O1—N1—C4	118.6 (3)
C10—C7—C8	112.3 (2)	C1—O3—C7	122.60 (18)
O3—C7—C9	103.2 (2)	C8—O5—H5A	111 (2)
O3—C1—C2—C3	-177.1 (3)	O3—C7—C8—O5	42.7 (3)
C6—C1—C2—C3	-0.3 (4)	C10—C7—C8—O5	167.7 (2)
C1—C2—C3—C4	0.2 (5)	C9—C7—C8—O5	-69.6 (3)
C2—C3—C4—C5	0.2 (5)	C5—C4—N1—O2	2.2 (5)
C2—C3—C4—N1	-178.2 (3)	C3—C4—N1—O2	-179.4 (4)
C3—C4—C5—C6	-0.6 (4)	C5—C4—N1—O1	-172.6 (4)
N1—C4—C5—C6	177.8 (3)	C3—C4—N1—O1	5.9 (5)
C4—C5—C6—C1	0.5 (4)	C2—C1—O3—C7	-22.5 (4)
O3—C1—C6—C5	177.1 (2)	C6—C1—O3—C7	160.6 (2)
C2—C1—C6—C5	-0.1 (4)	C10—C7—O3—C1	-58.6 (3)
O3—C7—C8—O4	-139.5 (2)	C8—C7—O3—C1	67.2 (3)
C10—C7—C8—O4	-14.5 (3)	C9—C7—O3—C1	-177.9 (2)
C9—C7—C8—O4	108.2 (3)		

Hydrogen-bond geometry (\AA , $^\circ$)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
O5—H5A \cdots O4 ⁱ	0.92 (4)	1.75 (4)	2.659 (3)	173 (3)
C6—H6 \cdots O1 ⁱⁱ	0.93	2.34	3.165 (4)	147

Symmetry codes: (i) $-x+1/2, -y+3/2, -z+1$; (ii) $x, y+1, z$.

Fig. 1

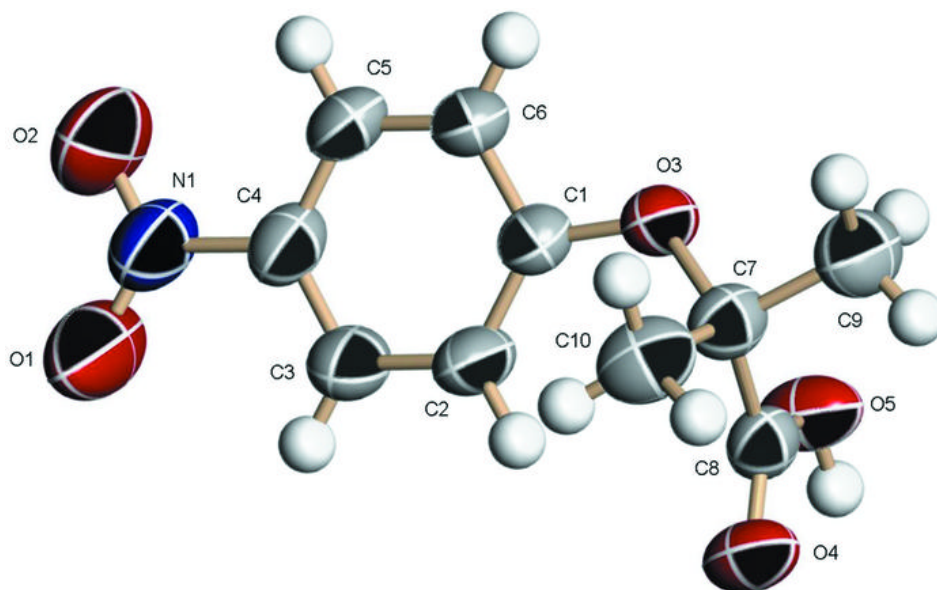


Fig. 2

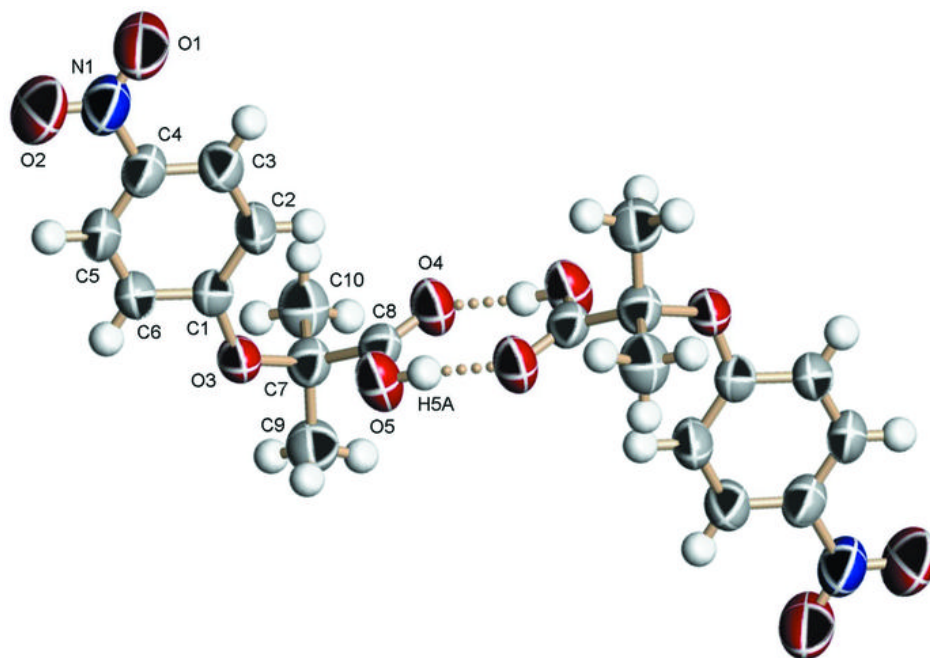


Fig. 3

