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## Structure Reports

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# (Z)-Methyl 4-(1,3-benzothiazol-2-ylsulfanyl)-2-(methoxyimino)-3-oxobutanoate

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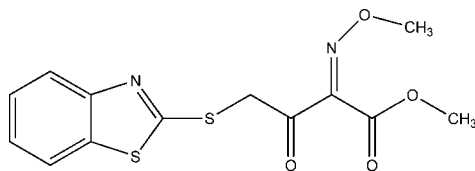
Received 1 December 2008; accepted 3 December 2008

Key indicators: single-crystal X-ray study;  $T = 293$  K; mean  $\sigma(\text{C}-\text{C}) = 0.004$  Å;  $R$  factor = 0.034;  $wR$  factor = 0.101; data-to-parameter ratio = 13.8.

In the molecular structure of the title compound,  $\text{C}_{13}\text{H}_{12}\text{N}_2\text{O}_4\text{S}_2$ , there is a dihedral angle of  $0.41$  ( $13$ )° between the benzene and thiazole rings. In the crystal, inversion dimers linked by two  $\text{C}-\text{H}\cdots\text{O}$  interactions together with  $\pi-\pi$  stacking between the parallel benzene rings of adjacent molecules [centroid-centroid distance =  $3.673$  ( $2$ ) Å].

## Related literature

For general background to benzothiazole derivatives and their biological activities, see: Bradshaw *et al.* (2008); Moharram (1990); Spillane *et al.* (2007).



## Experimental

## Crystal data

 $\text{C}_{13}\text{H}_{12}\text{N}_2\text{O}_4\text{S}_2$   
 $M_r = 324.39$ 

 Triclinic,  $P\bar{1}$   
 $a = 8.019$  ( $3$ ) Å

 $b = 10.037$  ( $4$ ) Å  
 $c = 10.662$  ( $5$ ) Å  
 $\alpha = 76.44$  ( $2$ )°  
 $\beta = 67.997$  ( $14$ )°  
 $\gamma = 74.964$  ( $15$ )°  
 $V = 759.3$  ( $6$ ) Å<sup>3</sup>
 $Z = 2$   
 Mo  $K\alpha$  radiation  
 $\mu = 0.37$  mm<sup>-1</sup>  
 $T = 293$  ( $2$ ) K  
 $0.22 \times 0.19 \times 0.18$  mm

## Data collection

 Bruker SMART CCD area-detector diffractometer  
 Absorption correction: multi-scan (*SADABS*; Bruker, 2002)  
 $T_{\min} = 0.913$ ,  $T_{\max} = 0.939$ 

 7739 measured reflections  
 2620 independent reflections  
 2278 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.020$ 

## Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.034$   
 $wR(F^2) = 0.101$   
 $S = 1.10$   
 2620 reflections

 190 parameters  
 H-atom parameters constrained  
 $\Delta\rho_{\text{max}} = 0.23$  e Å<sup>-3</sup>  
 $\Delta\rho_{\text{min}} = -0.27$  e Å<sup>-3</sup>

Table 1

Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{C3}-\text{H3}\cdots\text{O1}^i$	0.93	2.55	3.398 (3)	152

 Symmetry code: (i)  $-x + 1, -y, -z + 1$ .

Data collection: *SMART* (Bruker, 2002); cell refinement: *SAINTE* (Bruker, 2002); data reduction: *SAINTE*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEP-3 for Windows* (Farrugia, 1997); software used to prepare material for publication: *WinGX* (Farrugia, 1999).

The authors acknowledge the National Key Technologies R&D Program of China (2006BAE01A01-13) for supporting this work.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: AT2688).

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**supplementary materials**

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**(Z)-Methyl 4-(1,3-benzothiazol-2-ylsulfanyl)-2-(methoxyimino)-3-oxobutanoate**

**Q.-Z. Li, B.-A. Song, S. Yang, Y.-G. Zheng and Q.-Q. Guo**

**Comment**

A wide range of biological activities have been attributed to compounds containing benzothiazole moiety such as anticancer (Bradshaw *et al.*, 2008), antibacterial (Moharram *et al.*, 1990) and cytotoxic activity (Spillane *et al.*, 2007), Herein we present the crystal structure of the title benzothiazole derivative (I).

The crystal structure of the title compound (I) is represented in Fig. 1. There is a dihedral angle of 0.41 (13)° between the benzene ring and thiazole ring. In the crystal structure, weak intermolecular C—H···O interactions (Table 1), together with  $\pi$ - $\pi$  stacking between parallel benzene rings of adjacent molecules stabilize the packing, the centroid-to-centroid distance of two benzene rings is 3.673 (2) Å (symmetry codes: 1 - x, -y, 1 - z).

**Experimental**

A solution of (Z)-methyl-4-bromo-2-(methoxyimino)-3-oxobutanoate 30 ml (2.38 g, 0.01 mol) in methanol (10 ml) was added dropwise to a stirred solution of benzothiazole-2-thiol (1.67 g, 0.01 mol) and sodium hydroxide (0.40 g, 0.01 mol) in water (25 ml). The resulting mixture was stirred at room temperature for 3 h, the mixture was filtered and the residue was dissolved in 30 ml ethanol, Single crystals of (I) were obtained after several days.

**Refinement**

H atoms were placed in calculated positions with C—H = 0.93–0.97 Å, and refined in riding mode with  $U_{\text{iso}}(\text{H}) = 1.2$ – $1.5 U_{\text{eq}}(\text{C})$ .

**Figures**

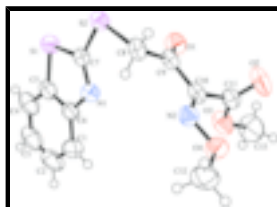


Fig. 1. The molecular structure of (I) showing the atom-labelling scheme. Displacement ellipsoids are drawn at the 50% probability level.

**(Z)-Methyl 4-(1,3-benzothiazol-2-ylsulfanyl)-2-(methoxyimino)-3-oxobutanoate**

*Crystal data*

$\text{C}_{13}\text{H}_{12}\text{N}_2\text{O}_4\text{S}_2$

$M_r = 324.39$

Triclinic, *PT*

$Z = 2$

$F_{000} = 336$

$D_x = 1.419 \text{ Mg m}^{-3}$

# supplementary materials

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Hall symbol: -P 1

$a = 8.019$  (3) Å

$b = 10.037$  (4) Å

$c = 10.662$  (5) Å

$\alpha = 76.44$  (2)°

$\beta = 67.997$  (14)°

$\gamma = 74.964$  (15)°

$V = 759.3$  (6) Å<sup>3</sup>

Mo  $K\alpha$  radiation

$\lambda = 0.71073$  Å

Cell parameters from 2679 reflections

$\theta = 2.1$ – $25.0$ °

$\mu = 0.37$  mm<sup>-1</sup>

$T = 293$  (2) K

Block, colourless

$0.22 \times 0.19 \times 0.18$  mm

## Data collection

Bruker SMART CCD area-detector  
diffractometer

Radiation source: fine-focus sealed tube

Monochromator: graphite

$T = 293$ (2) K

$\varphi$  and  $\omega$  scans

Absorption correction: multi-scan  
(SADABS; Bruker, 2002)

$T_{\min} = 0.913$ ,  $T_{\max} = 0.939$

7739 measured reflections

2620 independent reflections

2278 reflections with  $I > 2\sigma(I)$

$R_{\text{int}} = 0.020$

$\theta_{\max} = 25.0$ °

$\theta_{\min} = 2.1$ °

$h = -9 \rightarrow 9$

$k = -11 \rightarrow 11$

$l = -12 \rightarrow 12$

## Refinement

Refinement on  $F^2$

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.034$

$wR(F^2) = 0.101$

$S = 1.10$

2620 reflections

190 parameters

Primary atom site location: structure-invariant direct  
methods

Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring  
sites

H-atom parameters constrained

$w = 1/[\sigma^2(F_o^2) + (0.0551P)^2 + 0.1857P]$

where  $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{\max} < 0.001$

$\Delta\rho_{\max} = 0.23$  e Å<sup>-3</sup>

$\Delta\rho_{\min} = -0.26$  e Å<sup>-3</sup>

Extinction correction: none

## Special details

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
C1	0.2301 (4)	0.0381 (3)	0.4741 (3)	0.0665 (7)
H1	0.1206	0.0298	0.5455	0.080*
C2	0.3264 (5)	-0.0698 (2)	0.3972 (3)	0.0788 (9)
H2	0.2803	-0.1514	0.4185	0.095*
C3	0.4890 (4)	-0.0581 (3)	0.2898 (3)	0.0735 (8)
H3	0.5497	-0.1320	0.2407	0.088*
C4	0.5617 (4)	0.0593 (3)	0.2549 (3)	0.0635 (6)
H4	0.6707	0.0670	0.1829	0.076*
C5	0.4664 (3)	0.1679 (2)	0.3314 (2)	0.0472 (5)
C6	0.3029 (3)	0.1583 (2)	0.4407 (2)	0.0452 (5)
C7	0.3251 (3)	0.3681 (2)	0.45184 (18)	0.0387 (4)
C8	0.0771 (3)	0.5086 (2)	0.65173 (19)	0.0436 (5)
H8A	0.0138	0.5988	0.6798	0.052*
H8B	-0.0068	0.4706	0.6302	0.052*
C9	0.1277 (3)	0.41269 (19)	0.76915 (19)	0.0399 (4)
C10	-0.0218 (3)	0.3515 (2)	0.88308 (19)	0.0403 (4)
C11	0.0216 (3)	0.2738 (2)	1.0097 (2)	0.0443 (5)
C12	-0.4696 (3)	0.3219 (3)	0.9528 (3)	0.0694 (7)
H12A	-0.5552	0.2765	1.0303	0.104*
H12B	-0.4454	0.2809	0.8734	0.104*
H12C	-0.5204	0.4195	0.9369	0.104*
C13	0.1207 (5)	0.0538 (3)	1.1217 (3)	0.0890 (9)
H13A	0.1554	-0.0424	1.1090	0.134*
H13B	0.0174	0.0646	1.2040	0.134*
H13C	0.2215	0.0828	1.1287	0.134*
N1	0.2241 (2)	0.27441 (17)	0.50850 (16)	0.0448 (4)
N2	-0.1769 (2)	0.36936 (18)	0.86710 (16)	0.0451 (4)
O1	0.28034 (19)	0.38332 (16)	0.77643 (15)	0.0530 (4)
O2	0.0143 (3)	0.32938 (19)	1.09837 (17)	0.0774 (5)
O3	0.0721 (2)	0.13890 (16)	1.00578 (16)	0.0638 (4)
O4	-0.30078 (19)	0.30477 (17)	0.97942 (15)	0.0568 (4)
S1	0.52463 (7)	0.32799 (6)	0.31355 (5)	0.05263 (18)
S2	0.27235 (7)	0.53273 (5)	0.50010 (5)	0.04578 (17)

Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
C1	0.0938 (19)	0.0544 (14)	0.0641 (14)	-0.0322 (13)	-0.0349 (13)	0.0011 (11)
C2	0.138 (3)	0.0375 (13)	0.093 (2)	-0.0260 (15)	-0.074 (2)	0.0004 (13)
C3	0.103 (2)	0.0500 (14)	0.0869 (19)	0.0087 (14)	-0.0594 (18)	-0.0267 (13)
C4	0.0734 (16)	0.0570 (14)	0.0702 (15)	0.0041 (12)	-0.0340 (13)	-0.0302 (12)
C5	0.0562 (12)	0.0444 (11)	0.0480 (11)	-0.0042 (9)	-0.0252 (10)	-0.0131 (9)
C6	0.0596 (12)	0.0385 (11)	0.0474 (11)	-0.0108 (9)	-0.0292 (10)	-0.0048 (8)
C7	0.0414 (10)	0.0418 (10)	0.0358 (9)	-0.0071 (8)	-0.0154 (8)	-0.0082 (8)

## supplementary materials

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C8	0.0432 (11)	0.0490 (11)	0.0406 (10)	0.0008 (9)	-0.0174 (8)	-0.0155 (8)
C9	0.0420 (11)	0.0403 (10)	0.0438 (10)	0.0002 (8)	-0.0203 (8)	-0.0174 (8)
C10	0.0438 (10)	0.0413 (10)	0.0424 (10)	-0.0005 (8)	-0.0213 (8)	-0.0155 (8)
C11	0.0440 (11)	0.0488 (12)	0.0439 (10)	-0.0041 (9)	-0.0188 (9)	-0.0127 (9)
C12	0.0513 (13)	0.0842 (18)	0.0843 (17)	-0.0193 (12)	-0.0342 (13)	-0.0082 (14)
C13	0.127 (3)	0.0679 (18)	0.0844 (19)	-0.0117 (17)	-0.068 (2)	0.0130 (15)
N1	0.0484 (9)	0.0457 (10)	0.0433 (9)	-0.0136 (8)	-0.0150 (7)	-0.0077 (7)
N2	0.0446 (9)	0.0497 (10)	0.0449 (9)	-0.0089 (7)	-0.0185 (7)	-0.0088 (7)
O1	0.0438 (8)	0.0620 (9)	0.0594 (9)	-0.0054 (7)	-0.0272 (7)	-0.0094 (7)
O2	0.1180 (15)	0.0696 (11)	0.0605 (10)	0.0034 (10)	-0.0516 (10)	-0.0270 (9)
O3	0.0931 (12)	0.0467 (9)	0.0642 (10)	-0.0046 (8)	-0.0473 (9)	-0.0069 (7)
O4	0.0463 (8)	0.0745 (11)	0.0546 (9)	-0.0184 (7)	-0.0238 (7)	0.0000 (7)
S1	0.0523 (3)	0.0552 (4)	0.0485 (3)	-0.0167 (3)	-0.0037 (2)	-0.0193 (2)
S2	0.0515 (3)	0.0413 (3)	0.0472 (3)	-0.0109 (2)	-0.0144 (2)	-0.0131 (2)

### *Geometric parameters (Å, °)*

C1—C6	1.389 (3)	C8—H8A	0.9700
C1—C2	1.397 (4)	C8—H8B	0.9700
C1—H1	0.9300	C9—O1	1.210 (2)
C2—C3	1.387 (4)	C9—C10	1.493 (3)
C2—H2	0.9300	C10—N2	1.279 (2)
C3—C4	1.362 (4)	C10—C11	1.506 (3)
C3—H3	0.9300	C11—O2	1.182 (2)
C4—C5	1.398 (3)	C11—O3	1.315 (3)
C4—H4	0.9300	C12—O4	1.445 (3)
C5—C6	1.398 (3)	C12—H12A	0.9600
C5—S1	1.736 (2)	C12—H12B	0.9600
C6—N1	1.397 (3)	C12—H12C	0.9600
C7—N1	1.291 (2)	C13—O3	1.450 (3)
C7—S2	1.745 (2)	C13—H13A	0.9600
C7—S1	1.753 (2)	C13—H13B	0.9600
C8—C9	1.507 (3)	C13—H13C	0.9600
C8—S2	1.800 (2)	N2—O4	1.386 (2)
C6—C1—C2	118.0 (3)	O1—C9—C10	118.71 (17)
C6—C1—H1	121.0	O1—C9—C8	124.10 (19)
C2—C1—H1	121.0	C10—C9—C8	117.19 (16)
C3—C2—C1	121.5 (2)	N2—C10—C9	118.18 (17)
C3—C2—H2	119.3	N2—C10—C11	124.72 (18)
C1—C2—H2	119.3	C9—C10—C11	117.11 (16)
C4—C3—C2	121.3 (2)	O2—C11—O3	125.40 (19)
C4—C3—H3	119.3	O2—C11—C10	123.41 (19)
C2—C3—H3	119.3	O3—C11—C10	111.17 (16)
C3—C4—C5	117.6 (3)	O4—C12—H12A	109.5
C3—C4—H4	121.2	O4—C12—H12B	109.5
C5—C4—H4	121.2	H12A—C12—H12B	109.5
C6—C5—C4	122.2 (2)	O4—C12—H12C	109.5
C6—C5—S1	109.43 (15)	H12A—C12—H12C	109.5
C4—C5—S1	128.4 (2)	H12B—C12—H12C	109.5

C1—C6—N1	124.8 (2)	O3—C13—H13A	109.5
C1—C6—C5	119.4 (2)	O3—C13—H13B	109.5
N1—C6—C5	115.76 (18)	H13A—C13—H13B	109.5
N1—C7—S2	124.48 (15)	O3—C13—H13C	109.5
N1—C7—S1	117.24 (15)	H13A—C13—H13C	109.5
S2—C7—S1	118.25 (11)	H13B—C13—H13C	109.5
C9—C8—S2	113.07 (14)	C7—N1—C6	109.26 (17)
C9—C8—H8A	109.0	C10—N2—O4	111.87 (15)
S2—C8—H8A	109.0	C11—O3—C13	115.94 (19)
C9—C8—H8B	109.0	N2—O4—C12	109.46 (16)
S2—C8—H8B	109.0	C5—S1—C7	88.31 (10)
H8A—C8—H8B	107.8	C7—S2—C8	98.95 (10)

*Hydrogen-bond geometry (Å, °)*

<i>D</i> —H $\cdots$ <i>A</i>	<i>D</i> —H	H $\cdots$ <i>A</i>	<i>D</i> $\cdots$ <i>A</i>	<i>D</i> —H $\cdots$ <i>A</i>
C3—H3 $\cdots$ O1 <sup>i</sup>	0.93	2.55	3.398 (3)	152

Symmetry codes: (i)  $-x+1, -y, -z+1$ .

Fig. 1

