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## Structure Reports

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## 3-Nitrobenzaldehyde thiosemicarbazone

De-Hong Wu,\* Zhu-Feng Li and You-Hong Zhang

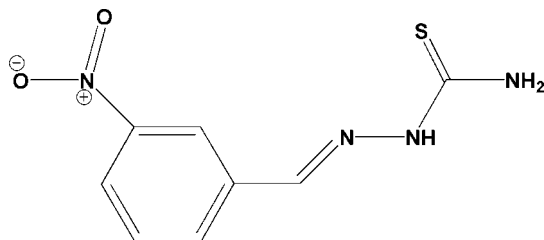
 Ordered Matter Science Research Center, College of Chemistry and Chemical Engineering, Southeast University, Nanjing 210096, People's Republic of China  
 Correspondence e-mail: wudh1971@sohu.com

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 Key indicators: single-crystal X-ray study;  $T = 291$  K; mean  $\sigma(\text{C}-\text{C}) = 0.005$  Å;  $R$  factor = 0.069;  $wR$  factor = 0.117; data-to-parameter ratio = 15.2.

 The molecule of the title compound,  $\text{C}_8\text{H}_8\text{N}_4\text{O}_2\text{S}$ , adopts an *E* configuration about both the C–N bonds. In the crystal structure, adjacent molecules are linked by intermolecular N–H...S hydrogen-bonding interactions, forming chains running parallel to the *b* axis.

## Related literature

 For general background to thiosemicarbazone compounds, see: Casas *et al.* (2000); Tarafder *et al.* (2000); Deschamps *et al.* (2003); Liu *et al.* (1999); Wu *et al.* (2000). For similar structures, see: Sutton (1965).


## Experimental

## Crystal data

 $\text{C}_8\text{H}_8\text{N}_4\text{O}_2\text{S}$   
 $M_r = 224.25$   
 Monoclinic,  $P2_1/c$   
 $a = 13.276$  (3) Å  
 $b = 8.225$  (7) Å

 $c = 10.491$  (4) Å  
 $\beta = 112.78$  (5)°  
 $V = 1056.2$  (11) Å<sup>3</sup>  
 $Z = 4$   
 Mo  $K\alpha$  radiation

 $\mu = 0.29$  mm<sup>-1</sup>  
 $T = 291$  (2) K

 $0.20 \times 0.20 \times 0.20$  mm

## Data collection

 Rigaku Mercury2 diffractometer  
 Absorption correction: multi-scan  
 (*CrystalClear*; Rigaku, 2005)  
 $T_{\text{min}} = 0.930$ ,  $T_{\text{max}} = 0.940$ 

 9317 measured reflections  
 2071 independent reflections  
 1343 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.081$ 

## Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.069$   
 $wR(F^2) = 0.117$   
 $S = 1.01$   
 2071 reflections

 136 parameters  
 H-atom parameters constrained  
 $\Delta\rho_{\text{max}} = 0.23$  e Å<sup>-3</sup>  
 $\Delta\rho_{\text{min}} = -0.18$  e Å<sup>-3</sup>

Table 1

Hydrogen-bond geometry (Å, °).

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
N1—H1B...S1 <sup>i</sup>	0.86	2.56	3.369 (4)	158
N2—H2B...S1 <sup>ii</sup>	0.86	2.56	3.394 (4)	165

 Symmetry codes: (i)  $-x + 2, y + \frac{1}{2}, -z + \frac{5}{2}$ ; (ii)  $-x + 2, y - \frac{1}{2}, -z + \frac{5}{2}$ .

 Data collection: *CrystalClear* (Rigaku, 2005); cell refinement: *CrystalClear*; data reduction: *CrystalClear*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: RZ2279).

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## supporting information

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## 3-Nitrobenzaldehyde thiosemicarbazone

De-Hong Wu, Zhu-Feng Li and You-Hong Zhang

### S1. Comment

Thiosemicarbazones constitute an important class of N, S donor ligands with their propensity to react with a wide range of metals (Casas *et al.*, 2000). Schiff bases show potential activity as antimicrobial and anticancer agents (Taraferder *et al.*, 2000; Deschamps *et al.*, 2003) and so have biochemical and pharmacological applications. It has been postulated that extensive electron delocalization in the thiosemicarbazone moiety helps the free thiosemicarbazone ligands and their metal complexes to exhibit SHG (second harmonic generation) efficiency (Liu *et al.*, 1999; Wu *et al.*, 2000). As part of a research on non-linear optical materials, specifically thiosemicarbazones and their metal complexes, we report here the crystal structure of a new Schiff base compound derived from thiosemicarbazide and 3-nitrobenzaldehyde.

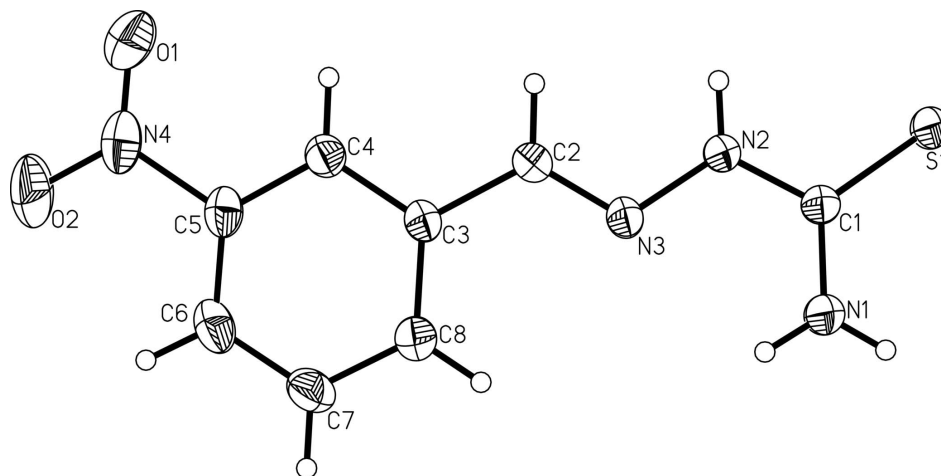
In the title compound (Fig. 1), the thiosemicarbazone moiety is nearly planar (maximum deviation 0.077 (3) Å for atom N2) and shows an *E* configuration about the C2—N3 bond. The molecule is not strictly planar, the dihedral angle between the thiosemicarbazone moiety and the phenyl ring being 13.45 (12)°. The C—S bond distance of 1.695 (3) Å agrees well with similar bonds in related compounds, being intermediate between 1.82 Å for a C—S single bond and 1.56 Å for a C=S double bond (Sutton, 1965). The C1—N2 bond distance (1.346 (4) Å) is indicative of some double-bond character, suggesting extensive electron delocalization in the whole molecule. In the crystal packing, adjacent molecules are linked by intermolecular N—H···S hydrogen bonds (Table 1) to form chains running parallel to the *b* axis.

### S2. Experimental

The title compound was synthesized by refluxing 3-nitrobenzaldehyde (6.04 g, 4 mmol) and thiosemicarbazide (0.37 g, 4 mmol) in absolute ethanol (30 ml) for 8 h. After cooling to room temperature, the yellow solid formed was isolated and dried under vacuum (0.76 g, yield 85%). Single crystals suitable for X-ray structure analysis were obtained by slow evaporation of an ethanol solution in air.

### S3. Refinement

H atoms were placed at calculated positions (N—H = 0.86 Å, C—H = 0.93 Å), and refined using the riding model approximation, with  $U_{\text{iso}}(\text{H}) = 1.2 U_{\text{eq}}(\text{C}, \text{N})$ .



**Figure 1**

The molecular structure of the title compound, showing the atomic numbering scheme. Displacement ellipsoids are drawn at the 30% probability level.

### 3-Nitrobenzaldehyde thiosemicarbazone

#### Crystal data

$C_8H_8N_4O_2S$

$M_r = 224.25$

Monoclinic,  $P2_1/c$

Hall symbol:  $-P\ 2_1/c$

$a = 13.276\ (3)\ \text{\AA}$

$b = 8.225\ (7)\ \text{\AA}$

$c = 10.491\ (4)\ \text{\AA}$

$\beta = 112.78\ (5)^\circ$

$V = 1056.2\ (11)\ \text{\AA}^3$

$Z = 4$

$F(000) = 464$

$D_x = 1.410\ \text{Mg m}^{-3}$

Mo  $K\alpha$  radiation,  $\lambda = 0.71073\ \text{\AA}$

Cell parameters from 1712 reflections

$\theta = 3.0\text{--}27.4^\circ$

$\mu = 0.29\ \text{mm}^{-1}$

$T = 291\ \text{K}$

Block, yellow

$0.20 \times 0.20 \times 0.20\ \text{mm}$

#### Data collection

Rigaku Mercury2  
diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

Detector resolution:  $13.6612\ \text{pixels mm}^{-1}$

CCD\_Profile\_fitting scans

Absorption correction: multi-scan  
(*CrystalClear*; Rigaku, 2005)

$T_{\min} = 0.930$ ,  $T_{\max} = 0.940$

9317 measured reflections

2071 independent reflections

1343 reflections with  $I > 2\sigma(I)$

$R_{\text{int}} = 0.081$

$\theta_{\max} = 26.0^\circ$ ,  $\theta_{\min} = 3.0^\circ$

$h = -16 \rightarrow 16$

$k = -10 \rightarrow 10$

$l = -12 \rightarrow 12$

#### Refinement

Refinement on  $F^2$

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.069$

$wR(F^2) = 0.117$

$S = 1.01$

2071 reflections

136 parameters

0 restraints

Primary atom site location: structure-invariant  
direct methods

Secondary atom site location: difference Fourier  
map

Hydrogen site location: inferred from  
neighbouring sites

H-atom parameters constrained

$$w = 1/[\sigma^2(F_o^2) + (0.0122P)^2 + 1.5479P]$$

where  $P = (F_o^2 + 2F_c^2)/3$   
 $(\Delta/\sigma)_{\max} < 0.001$

$$\Delta\rho_{\max} = 0.23 \text{ e } \text{\AA}^{-3}$$

$$\Delta\rho_{\min} = -0.18 \text{ e } \text{\AA}^{-3}$$

*Special details*

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
C1	0.9248 (3)	1.1141 (4)	1.1340 (3)	0.0416 (8)
C2	0.8132 (3)	0.8365 (4)	0.8634 (3)	0.0444 (9)
H2A	0.8406	0.7429	0.9149	0.053*
C3	0.7415 (3)	0.8218 (4)	0.7163 (3)	0.0412 (9)
C4	0.7111 (3)	0.6666 (4)	0.6615 (3)	0.0441 (9)
H4A	0.7375	0.5743	0.7153	0.053*
C5	0.6401 (3)	0.6545 (5)	0.5237 (4)	0.0476 (10)
C6	0.5984 (3)	0.7864 (5)	0.4403 (4)	0.0584 (11)
H6A	0.5501	0.7733	0.3490	0.070*
C7	0.6301 (3)	0.9401 (5)	0.4955 (4)	0.0603 (12)
H7A	0.6036	1.0314	0.4404	0.072*
C8	0.7017 (3)	0.9588 (5)	0.6334 (4)	0.0512 (10)
H8A	0.7229	1.0622	0.6697	0.061*
N1	0.8999 (3)	1.2497 (4)	1.0619 (3)	0.0572 (10)
H1A	0.8691	1.2460	0.9732	0.069*
H1B	0.9145	1.3420	1.1036	0.069*
N2	0.9000 (2)	0.9732 (3)	1.0633 (3)	0.0425 (8)
H2B	0.9222	0.8823	1.1055	0.051*
N3	0.8385 (2)	0.9751 (3)	0.9216 (3)	0.0411 (7)
N4	0.6100 (3)	0.4887 (5)	0.4659 (4)	0.0656 (10)
O1	0.6466 (3)	0.3718 (4)	0.5411 (4)	0.0905 (11)
O2	0.5489 (3)	0.4772 (5)	0.3440 (3)	0.1039 (13)
S1	0.98388 (9)	1.10945 (12)	1.30902 (9)	0.0529 (3)

*Atomic displacement parameters ( $\text{\AA}^2$ )*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
C1	0.046 (2)	0.038 (2)	0.0381 (19)	-0.0034 (18)	0.0136 (17)	0.0019 (17)
C2	0.048 (2)	0.041 (2)	0.0373 (19)	0.0029 (18)	0.0093 (17)	0.0012 (17)
C3	0.045 (2)	0.044 (2)	0.0326 (18)	-0.0016 (17)	0.0121 (17)	-0.0025 (16)
C4	0.043 (2)	0.047 (2)	0.041 (2)	0.0013 (17)	0.0138 (18)	-0.0016 (17)
C5	0.038 (2)	0.056 (3)	0.047 (2)	-0.0044 (18)	0.0144 (18)	-0.0142 (19)

C6	0.048 (3)	0.077 (3)	0.040 (2)	0.003 (2)	0.007 (2)	-0.008 (2)
C7	0.063 (3)	0.061 (3)	0.046 (2)	0.016 (2)	0.010 (2)	0.009 (2)
C8	0.054 (3)	0.047 (2)	0.045 (2)	0.0008 (19)	0.0120 (19)	-0.0025 (18)
N1	0.084 (3)	0.0394 (19)	0.0371 (17)	-0.0042 (17)	0.0120 (18)	-0.0012 (14)
N2	0.054 (2)	0.0345 (17)	0.0332 (16)	-0.0005 (14)	0.0102 (15)	-0.0003 (13)
N3	0.0443 (19)	0.0418 (18)	0.0335 (15)	-0.0040 (14)	0.0110 (14)	-0.0023 (13)
N4	0.056 (3)	0.076 (3)	0.063 (3)	-0.017 (2)	0.022 (2)	-0.029 (2)
O1	0.107 (3)	0.056 (2)	0.095 (3)	-0.016 (2)	0.024 (2)	-0.0196 (19)
O2	0.104 (3)	0.110 (3)	0.069 (2)	-0.028 (2)	0.003 (2)	-0.042 (2)
S1	0.0735 (7)	0.0436 (5)	0.0345 (5)	-0.0028 (5)	0.0130 (5)	-0.0035 (4)

*Geometric parameters (Å, °)*

C1—N1	1.316 (4)	C6—C7	1.386 (5)
C1—N2	1.346 (4)	C6—H6A	0.9300
C1—S1	1.695 (3)	C7—C8	1.398 (5)
C2—N3	1.276 (4)	C7—H7A	0.9300
C2—C3	1.471 (4)	C8—H8A	0.9300
C2—H2A	0.9300	N1—H1A	0.8600
C3—C4	1.394 (5)	N1—H1B	0.8600
C3—C8	1.396 (5)	N2—N3	1.392 (4)
C4—C5	1.391 (5)	N2—H2B	0.8600
C4—H4A	0.9300	N4—O1	1.219 (5)
C5—C6	1.369 (5)	N4—O2	1.225 (4)
C5—N4	1.483 (5)		
N1—C1—N2	117.4 (3)	C7—C6—H6A	120.9
N1—C1—S1	123.3 (3)	C6—C7—C8	120.5 (4)
N2—C1—S1	119.3 (3)	C6—C7—H7A	119.8
N3—C2—C3	121.3 (3)	C8—C7—H7A	119.8
N3—C2—H2A	119.4	C3—C8—C7	119.9 (4)
C3—C2—H2A	119.4	C3—C8—H8A	120.1
C4—C3—C8	120.2 (3)	C7—C8—H8A	120.1
C4—C3—C2	118.3 (3)	C1—N1—H1A	120.0
C8—C3—C2	121.5 (3)	C1—N1—H1B	120.0
C5—C4—C3	117.8 (3)	H1A—N1—H1B	120.0
C5—C4—H4A	121.1	C1—N2—N3	119.7 (3)
C3—C4—H4A	121.1	C1—N2—H2B	120.1
C6—C5—C4	123.5 (4)	N3—N2—H2B	120.1
C6—C5—N4	119.3 (3)	C2—N3—N2	116.0 (3)
C4—C5—N4	117.2 (4)	O1—N4—O2	123.5 (4)
C5—C6—C7	118.3 (4)	O1—N4—C5	118.9 (3)
C5—C6—H6A	120.9	O2—N4—C5	117.6 (4)
N3—C2—C3—C4	-175.8 (4)	C2—C3—C8—C7	-177.4 (4)
N3—C2—C3—C8	2.5 (6)	C6—C7—C8—C3	-0.1 (6)
C8—C3—C4—C5	-0.6 (5)	N1—C1—N2—N3	-7.7 (5)
C2—C3—C4—C5	177.7 (3)	S1—C1—N2—N3	171.4 (2)

C3—C4—C5—C6	-0.4 (6)	C3—C2—N3—N2	175.3 (3)
C3—C4—C5—N4	178.9 (3)	C1—N2—N3—C2	-176.1 (3)
C4—C5—C6—C7	1.1 (6)	C6—C5—N4—O1	-179.3 (4)
N4—C5—C6—C7	-178.2 (4)	C4—C5—N4—O1	1.3 (6)
C5—C6—C7—C8	-0.8 (6)	C6—C5—N4—O2	0.5 (6)
C4—C3—C8—C7	0.9 (6)	C4—C5—N4—O2	-178.9 (4)

*Hydrogen-bond geometry (Å, °)*

<i>D</i> —H $\cdots$ <i>A</i>	<i>D</i> —H	H $\cdots$ <i>A</i>	<i>D</i> $\cdots$ <i>A</i>	<i>D</i> —H $\cdots$ <i>A</i>
N1—H1B $\cdots$ S1 <sup>i</sup>	0.86	2.56	3.369 (4)	158
N2—H2B $\cdots$ S1 <sup>ii</sup>	0.86	2.56	3.394 (4)	165

Symmetry codes: (i)  $-x+2, y+1/2, -z+5/2$ ; (ii)  $-x+2, y-1/2, -z+5/2$ .