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3-[(Methylcarbamoyl)amino]-1*H*-isoindolium chlorideBushra Maliha,^a Muhammad Ilyas Tariq,^b M. Nawaz Tahir,^{c*} Ishtiaq Hussain^d and Hamid Latif Siddiqui^a^aUniversity of the Punjab, Institute of Chemistry, Lahore 54590, Pakistan, ^bUniversity of Sargodha, Department of Chemistry, Sargodha, Pakistan, ^cUniversity of Sargodha, Department of Physics, Sargodha, Pakistan, and ^dUniversity of the Punjab, Institute of Chemistry, Lahore-54590, Pakistan

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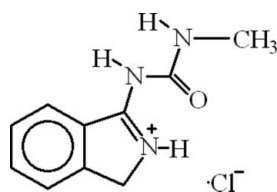
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Key indicators: single-crystal X-ray study; $T = 296$ K; mean $\sigma(\text{C}-\text{C}) = 0.002$ Å; R factor = 0.032; wR factor = 0.101; data-to-parameter ratio = 16.2.

The title compound, $\text{C}_{10}\text{H}_{12}\text{N}_3\text{O}^+\cdot\text{Cl}^-$, is a derivative of *o*-phthaldehyde and methylthiourea. The molecules form dimers through intra- and intermolecular $\text{N}-\text{H}\cdots\text{O}$ hydrogen bonds. The dimers are further linked into chains through one $\text{C}-\text{H}\cdots\text{Cl}$ and two $\text{N}-\text{H}\cdots\text{Cl}$ hydrogen bonds.

Related literature

For applications of iminium salts, see: Page *et al.* (2008); Skalkos *et al.* (1994) Tariq *et al.* (2008). For the formation of derivatives of *o*-phthaldehyde with different ureas, see: Maliha, Tariq, Tahir, Hussain & Ali (2009); Maliha, Tariq, Tahir, Hussain & Siddiqui (2009); Maliha *et al.* (2008). For a related structure, see: Arfan *et al.* (2008).



Experimental

Crystal data

 $\text{C}_{10}\text{H}_{12}\text{N}_3\text{O}^+\cdot\text{Cl}^-$ $M_r = 225.68$ Triclinic, $P\bar{1}$ $a = 7.1171$ (5) Å $b = 7.7900$ (6) Å $c = 10.3033$ (8) Å $\alpha = 89.484$ (3)° $\beta = 69.997$ (2)° $\gamma = 74.613$ (4)° $V = 515.43$ (7) Å³ $Z = 2$ Mo $K\alpha$ radiation $\mu = 0.35$ mm⁻¹ $T = 296$ (2) K

0.30 × 0.10 × 0.06 mm

Data collection

Bruker Kappa APEXII CCD

diffractometer

Absorption correction: multi-scan

(SADABS; Bruker, 2005)

 $T_{\min} = 0.982$, $T_{\max} = 0.989$

8853 measured reflections

2369 independent reflections

2210 reflections with $I > 2\sigma(I)$ $R_{\text{int}} = 0.020$

Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.032$ $wR(F^2) = 0.101$ $S = 1.01$

2369 reflections

146 parameters

H atoms treated by a mixture of independent and constrained refinement

 $\Delta\rho_{\text{max}} = 0.69$ e Å⁻³ $\Delta\rho_{\text{min}} = -0.22$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{N1}-\text{H1N}\cdots\text{O1}$	0.81 (2)	2.22 (2)	2.7097 (16)	119.9 (18)
$\text{N1}-\text{H1N}\cdots\text{O1}^{\text{i}}$	0.81 (2)	2.15 (2)	2.8760 (17)	150 (2)
$\text{N2}-\text{H2N}\cdots\text{Cl1}^{\text{ii}}$	0.90 (2)	2.23 (2)	3.0969 (13)	160.5 (18)
$\text{N3}-\text{H3N}\cdots\text{Cl1}^{\text{ii}}$	0.86 (2)	2.40 (2)	3.2082 (13)	157.0 (17)
$\text{C1}-\text{H1B}\cdots\text{Cl1}^{\text{iii}}$	0.97	2.74	3.6755 (16)	162

Symmetry codes: (i) $-x + 1, -y, -z + 1$; (ii) $-x, -y + 1, -z + 1$; (iii) $x + 1, y, z$.

Data collection: APEX2 (Bruker, 2007); cell refinement: SAINT (Bruker, 2007); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: ORTEP-3 for Windows (Farrugia, 1997) and PLATON (Spek, 2003); software used to prepare material for publication: WinGX (Farrugia, 1999) and PLATON.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: E22156).

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supplementary materials

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3-[(Methylcarbamoyl)amino]-1*H*-isoindolium chloride

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Comment

Iminium salts are a class of organic compounds acting as reactive intermediates. These are being extensively synthesized and further reacted in an efficient and convenient manner for the preparation of multidrug-resistance reversal agents and pesticides and for metal complexation to form new photosensitizers (Page *et al.*, 2008; Skalkos *et al.*, 1994; Tariq *et al.*, 2008). The present paper relates to the continuation of our studies regarding the formation of derivatives of *O*-phthaldehyde with different ureas (Maliha *et al.*, 2008; Maliha, Tariq, Tahir, Hussain & Ali, 2009; Maliha, Tariq, Tahir, Hussain & Siddiqui, 2009).

The molecule of the title compound (I; Fig 1), is almost planar. It has three N—H bonds in the asymmetric unit. The N—H of the pyrrole ring forms an intramolecular as well as an intermolecular N—H \cdots O H-bond. The title compound forms a dimer (Fig 2) with a central four membered O \cdots H \cdots O \cdots H unit. The other two N—H groups in the methyl urea moiety are involved in intermolecular H-bonding with the Cl $^-$. The Cl $^-$ also forms H-bonds with the methylene group of the pyrrole ring. Consequently, the chlorine anion makes three H-bonds. Similar Cl $^-$ bonding behavior has been reported by Arfan *et al.* (2008). No strong π interactions are observed.

Experimental

O-phthaldehyde (200 mmol), methylthiourea (200 mmol) and a few drops of 2*M* HCl were mixed and ground in mortar and pestle. The product obtained was washed sequentially with hexane, ether, ethanol and water. The precipitate was dried and recrystallized from a mixture of methanol:acetone (9:1), by slow evaporation at room temperature.

Figures

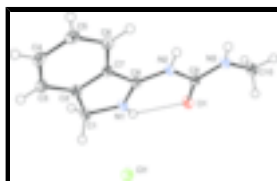


Fig. 1. The title compound (I) with displacement ellipsoids drawn at 50% probability level. Intramolecular H-bonds are indicated by broken lines.

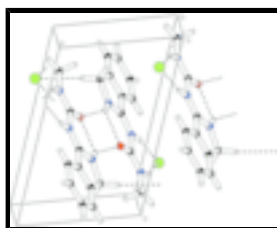


Fig. 2. The packing of I (*PLATON*: Spek, 2003) showing the dimers and H-bonds to the chloride anion. H-bonds are indicated by broken lines.

3-[(Methylcarbamoyl)amino]-1*H*-isoindolium chloride

Crystal data

$C_{10}H_{12}N_3O^+ \cdot Cl^-$	$Z = 2$
$M_r = 225.68$	$F_{000} = 236$
Triclinic, $P\bar{1}$	$D_x = 1.454 \text{ Mg m}^{-3}$
Hall symbol: -P 1	Mo $K\alpha$ radiation
$a = 7.1171 (5) \text{ \AA}$	$\lambda = 0.71073 \text{ \AA}$
$b = 7.7900 (6) \text{ \AA}$	Cell parameters from 2369 reflections
$c = 10.3033 (8) \text{ \AA}$	$\theta = 3.1\text{--}27.5^\circ$
$\alpha = 89.484 (3)^\circ$	$\mu = 0.35 \text{ mm}^{-1}$
$\beta = 69.997 (2)^\circ$	$T = 296 \text{ K}$
$\gamma = 74.613 (4)^\circ$	Needle, light yellow
$V = 515.43 (7) \text{ \AA}^3$	$0.30 \times 0.10 \times 0.06 \text{ mm}$

Data collection

Bruker Kappa APEXII CCD diffractometer	2369 independent reflections
Radiation source: fine-focus sealed tube	2210 reflections with $I > 2\sigma(I)$
Monochromator: graphite	$R_{\text{int}} = 0.020$
Detector resolution: $7.50 \text{ pixels mm}^{-1}$	$\theta_{\text{max}} = 27.5^\circ$
$T = 296 \text{ K}$	$\theta_{\text{min}} = 3.1^\circ$
ω scans	$h = -9 \rightarrow 9$
Absorption correction: multi-scan (SADABS; Bruker, 2005)	$k = -10 \rightarrow 10$
$T_{\text{min}} = 0.982$, $T_{\text{max}} = 0.989$	$l = -13 \rightarrow 13$
8853 measured reflections	

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: mixed
$R[F^2 > 2\sigma(F^2)] = 0.032$	H atoms treated by a mixture of independent and constrained refinement
$wR(F^2) = 0.101$	$w = 1/[\sigma^2(F_o^2) + (0.0646P)^2 + 0.2853P]$
$S = 1.01$	where $P = (F_o^2 + 2F_c^2)/3$
2369 reflections	$(\Delta/\sigma)_{\text{max}} < 0.001$
146 parameters	$\Delta\rho_{\text{max}} = 0.69 \text{ e \AA}^{-3}$
Primary atom site location: structure-invariant direct methods	$\Delta\rho_{\text{min}} = -0.21 \text{ e \AA}^{-3}$
	Extinction correction: none

Special details

Geometry. Bond distances, angles etc. have been calculated using the rounded fractional coordinates. All su's are estimated from the variances of the (full) variance-covariance matrix. The cell esds are taken into account in the estimation of distances, angles and torsion angles

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) etc. and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
O1	0.29475 (16)	0.09740 (13)	0.60677 (11)	0.0195 (3)
N1	0.46180 (18)	0.25619 (17)	0.37719 (12)	0.0161 (3)
N2	0.19866 (18)	0.40058 (16)	0.58681 (12)	0.0155 (3)
N3	0.05849 (19)	0.28726 (17)	0.79177 (13)	0.0177 (3)
C1	0.5740 (2)	0.29979 (19)	0.23859 (15)	0.0175 (4)
C2	0.4836 (2)	0.49861 (19)	0.24873 (15)	0.0165 (4)
C3	0.5278 (2)	0.6215 (2)	0.15233 (15)	0.0201 (4)
C4	0.4209 (2)	0.8007 (2)	0.19298 (17)	0.0222 (4)
C5	0.2750 (2)	0.8576 (2)	0.32626 (17)	0.0214 (4)
C6	0.2298 (2)	0.73562 (19)	0.42329 (15)	0.0183 (4)
C7	0.3359 (2)	0.55629 (19)	0.38103 (14)	0.0149 (4)
C8	0.3263 (2)	0.39698 (18)	0.45638 (14)	0.0147 (3)
C9	0.1900 (2)	0.24711 (19)	0.66151 (14)	0.0155 (3)
C10	0.0312 (3)	0.1475 (2)	0.88541 (15)	0.0218 (4)
Cl1	0.11104 (5)	0.28646 (4)	0.19897 (3)	0.0186 (1)
H1A	0.54729	0.23997	0.16735	0.0210*
H1B	0.72287	0.26795	0.21941	0.0210*
H1N	0.487 (3)	0.156 (3)	0.401 (2)	0.0194*
H2N	0.113 (3)	0.507 (3)	0.631 (2)	0.0186*
H3	0.62547	0.58510	0.06360	0.0242*
H3N	-0.007 (3)	0.397 (3)	0.819 (2)	0.0212*
H4	0.44729	0.88501	0.12976	0.0266*
H5	0.20729	0.97864	0.35036	0.0256*
H6	0.13317	0.77212	0.51234	0.0220*
H10A	0.01212	0.04982	0.83996	0.0327*
H10B	-0.08889	0.19448	0.96722	0.0327*
H10C	0.15238	0.10608	0.91081	0.0327*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
O1	0.0212 (5)	0.0138 (5)	0.0177 (5)	-0.0003 (4)	-0.0033 (4)	0.0024 (4)
N1	0.0167 (5)	0.0144 (6)	0.0145 (6)	-0.0019 (4)	-0.0039 (4)	0.0029 (4)

supplementary materials

N2	0.0166 (5)	0.0126 (6)	0.0143 (6)	-0.0020 (4)	-0.0032 (4)	0.0016 (4)
N3	0.0208 (6)	0.0145 (6)	0.0145 (6)	-0.0031 (5)	-0.0036 (5)	0.0021 (4)
C1	0.0162 (6)	0.0182 (7)	0.0149 (6)	-0.0030 (5)	-0.0029 (5)	0.0021 (5)
C2	0.0146 (6)	0.0192 (7)	0.0165 (7)	-0.0058 (5)	-0.0057 (5)	0.0024 (5)
C3	0.0193 (7)	0.0251 (8)	0.0169 (7)	-0.0097 (6)	-0.0049 (5)	0.0049 (6)
C4	0.0251 (7)	0.0221 (8)	0.0249 (7)	-0.0130 (6)	-0.0111 (6)	0.0101 (6)
C5	0.0245 (7)	0.0158 (7)	0.0264 (8)	-0.0074 (6)	-0.0109 (6)	0.0047 (6)
C6	0.0181 (6)	0.0176 (7)	0.0200 (7)	-0.0053 (5)	-0.0074 (5)	0.0016 (5)
C7	0.0142 (6)	0.0160 (6)	0.0164 (7)	-0.0055 (5)	-0.0067 (5)	0.0037 (5)
C8	0.0140 (6)	0.0156 (6)	0.0158 (6)	-0.0040 (5)	-0.0071 (5)	0.0018 (5)
C9	0.0160 (6)	0.0156 (6)	0.0155 (6)	-0.0048 (5)	-0.0062 (5)	0.0033 (5)
C10	0.0280 (7)	0.0197 (7)	0.0156 (7)	-0.0068 (6)	-0.0050 (6)	0.0060 (5)
C11	0.0195 (2)	0.0144 (2)	0.0192 (2)	-0.0041 (1)	-0.0040 (1)	0.0002 (1)

Geometric parameters (Å, °)

O1—C9	1.2234 (17)	C4—C5	1.397 (2)
N1—C1	1.4654 (19)	C5—C6	1.390 (2)
N1—C8	1.3096 (19)	C6—C7	1.392 (2)
N2—C8	1.3356 (18)	C7—C8	1.463 (2)
N2—C9	1.4197 (19)	C1—H1A	0.9700
N3—C9	1.3296 (19)	C1—H1B	0.9700
N3—C10	1.456 (2)	C3—H3	0.9300
N1—H1N	0.81 (2)	C4—H4	0.9300
N2—H2N	0.90 (2)	C5—H5	0.9300
N3—H3N	0.86 (2)	C6—H6	0.9300
C1—C2	1.501 (2)	C10—H10A	0.9600
C2—C3	1.390 (2)	C10—H10B	0.9600
C2—C7	1.395 (2)	C10—H10C	0.9600
C3—C4	1.390 (2)		
C11...C1	3.4838 (16)	C8...C7 ^{vii}	3.457 (2)
C11...C2	3.6463 (16)	C9...C3 ^{vii}	3.539 (2)
C11...C5 ⁱ	3.6252 (16)	C9...C6 ⁱⁱ	3.358 (2)
C11...N2 ⁱⁱ	3.0969 (13)	C9...C4 ^{vii}	3.513 (2)
C11...N3 ⁱⁱ	3.2082 (13)	C9...C5 ⁱⁱ	3.575 (2)
C11...H5 ⁱ	2.9100	C10...C4 ⁱⁱ	3.500 (3)
C11...H1A	2.9300	C10...C5 ⁱⁱ	3.579 (3)
C11...H1B ⁱⁱⁱ	2.7400	C10...C10 ^{ix}	3.283 (2)
C11...H2N ⁱⁱ	2.23 (2)	C1...H10B ^x	2.9400
C11...H3 ^{iv}	3.0500	C5...H10A ⁱⁱ	3.0400
C11...H3N ⁱⁱ	2.40 (2)	C6...H2N	2.82 (2)
C11...H6 ⁱⁱ	2.9900	C9...H1N	2.75 (2)
C11...H10A ^v	3.0400	C10...H10B ^{ix}	3.0600
O1...O1 ^{vi}	2.9982 (16)	C10...H10C ^{ix}	3.0700
O1...N1 ^{vi}	2.8760 (17)	H1A...C11	2.9300
O1...N1	2.7097 (16)	H1B...C11 ^{xi}	2.7400

O1...H1N ^{vi}	2.15 (2)	H1B...H10B ^x	2.4700
O1...H10A	2.6500	H1N...O1 ^{vi}	2.15 (2)
O1...H1N	2.22 (2)	H1N...C9	2.75 (2)
N1...O1	2.7097 (16)	H1N...O1	2.22 (2)
N1...O1 ^{vi}	2.8760 (17)	H2N...H6	2.4000
N2...C11 ⁱⁱ	3.0969 (13)	H2N...H3N	2.11 (3)
N3...C5 ⁱⁱ	3.436 (2)	H2N...C6	2.82 (2)
N3...C3 ^{vii}	3.433 (2)	H2N...C11 ⁱⁱ	2.23 (2)
N3...C11 ⁱⁱ	3.2082 (13)	H3...C11 ^{iv}	3.0500
N2...H6	2.9400	H3N...H2N	2.11 (3)
C1...C11	3.4838 (16)	H3N...C11 ⁱⁱ	2.40 (2)
C2...C11	3.6463 (16)	H5...C11 ^{viii}	2.9100
C3...N3 ^{vii}	3.433 (2)	H6...H2N	2.4000
C3...C9 ^{vii}	3.539 (2)	H6...N2	2.9400
C4...C9 ^{vii}	3.513 (2)	H6...C11 ⁱⁱ	2.9900
C4...C10 ⁱⁱ	3.500 (3)	H10A...O1	2.6500
C5...N3 ⁱⁱ	3.436 (2)	H10A...C11 ^v	3.0400
C5...C9 ⁱⁱ	3.575 (2)	H10A...C5 ⁱⁱ	3.0400
C5...C10 ⁱⁱ	3.579 (3)	H10B...C10 ^{ix}	3.0600
C5...C11 ^{viii}	3.6252 (16)	H10B...H1B ^{xii}	2.4700
C6...C9 ⁱⁱ	3.358 (2)	H10B...C1 ^{xii}	2.9400
C7...C8 ^{vii}	3.457 (2)	H10C...C10 ^{ix}	3.0700
C1—N1—C8	112.52 (12)	N2—C9—N3	112.47 (12)
C8—N2—C9	124.09 (12)	O1—C9—N2	121.32 (12)
C9—N3—C10	120.48 (13)	O1—C9—N3	126.20 (13)
C8—N1—H1N	125.2 (14)	N1—C1—H1A	111.00
C1—N1—H1N	122.2 (14)	N1—C1—H1B	111.00
C8—N2—H2N	118.4 (13)	C2—C1—H1A	111.00
C9—N2—H2N	117.5 (13)	C2—C1—H1B	111.00
C10—N3—H3N	121.3 (13)	H1A—C1—H1B	109.00
C9—N3—H3N	118.2 (13)	C2—C3—H3	121.00
N1—C1—C2	101.98 (12)	C4—C3—H3	121.00
C1—C2—C3	130.68 (13)	C3—C4—H4	119.00
C3—C2—C7	120.12 (13)	C5—C4—H4	119.00
C1—C2—C7	109.19 (12)	C4—C5—H5	120.00
C2—C3—C4	117.74 (14)	C6—C5—H5	120.00
C3—C4—C5	121.78 (14)	C5—C6—H6	122.00
C4—C5—C6	120.86 (14)	C7—C6—H6	122.00
C5—C6—C7	116.91 (13)	N3—C10—H10A	109.00
C2—C7—C8	106.74 (12)	N3—C10—H10B	109.00
C2—C7—C6	122.59 (13)	N3—C10—H10C	109.00
C6—C7—C8	130.66 (13)	H10A—C10—H10B	109.00
N1—C8—C7	109.54 (12)	H10A—C10—H10C	109.00
N1—C8—N2	126.82 (13)	H10B—C10—H10C	109.00
N2—C8—C7	123.63 (12)		

supplementary materials

C8—N1—C1—C2	1.26 (17)	C1—C2—C7—C6	177.86 (14)
C1—N1—C8—C7	-1.95 (18)	C1—C2—C7—C8	-1.00 (17)
C1—N1—C8—N2	178.99 (15)	C3—C2—C7—C6	-1.0 (2)
C9—N2—C8—N1	-0.2 (3)	C3—C2—C7—C8	-179.83 (14)
C8—N2—C9—N3	176.55 (15)	C2—C3—C4—C5	0.6 (2)
C9—N2—C8—C7	-179.15 (14)	C3—C4—C5—C6	-0.6 (2)
C8—N2—C9—O1	-3.7 (2)	C4—C5—C6—C7	-0.1 (2)
C10—N3—C9—O1	1.4 (3)	C5—C6—C7—C8	179.48 (16)
C10—N3—C9—N2	-178.92 (15)	C5—C6—C7—C2	0.9 (2)
N1—C1—C2—C3	178.61 (16)	C2—C7—C8—N1	1.83 (18)
N1—C1—C2—C7	-0.06 (17)	C6—C7—C8—N1	-176.90 (16)
C1—C2—C3—C4	-178.36 (16)	C6—C7—C8—N2	2.2 (3)
C7—C2—C3—C4	0.2 (2)	C2—C7—C8—N2	-179.07 (15)

Symmetry codes: (i) $x, y-1, z$; (ii) $-x, -y+1, -z+1$; (iii) $x-1, y, z$; (iv) $-x+1, -y+1, -z$; (v) $-x, -y, -z+1$; (vi) $-x+1, -y, -z+1$; (vii) $-x+1, -y+1, -z+1$; (viii) $x, y+1, z$; (ix) $-x, -y, -z+2$; (x) $x+1, y, z-1$; (xi) $x+1, y, z$; (xii) $x-1, y, z+1$.

Hydrogen-bond geometry ($\text{\AA}, ^\circ$)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
N1—H1N \cdots O1	0.81 (2)	2.22 (2)	2.7097 (16)	119.9 (18)
N1—H1N \cdots O1 ^{vi}	0.81 (2)	2.15 (2)	2.8760 (17)	150 (2)
N2—H2N \cdots C11 ⁱⁱ	0.90 (2)	2.23 (2)	3.0969 (13)	160.5 (18)
N3—H3N \cdots C11 ⁱⁱ	0.86 (2)	2.40 (2)	3.2082 (13)	157.0 (17)
C1—H1B \cdots C11 ^{xi}	0.9700	2.7400	3.6755 (16)	162.00

Symmetry codes: (vi) $-x+1, -y, -z+1$; (ii) $-x, -y+1, -z+1$; (xi) $x+1, y, z$.

Fig. 1

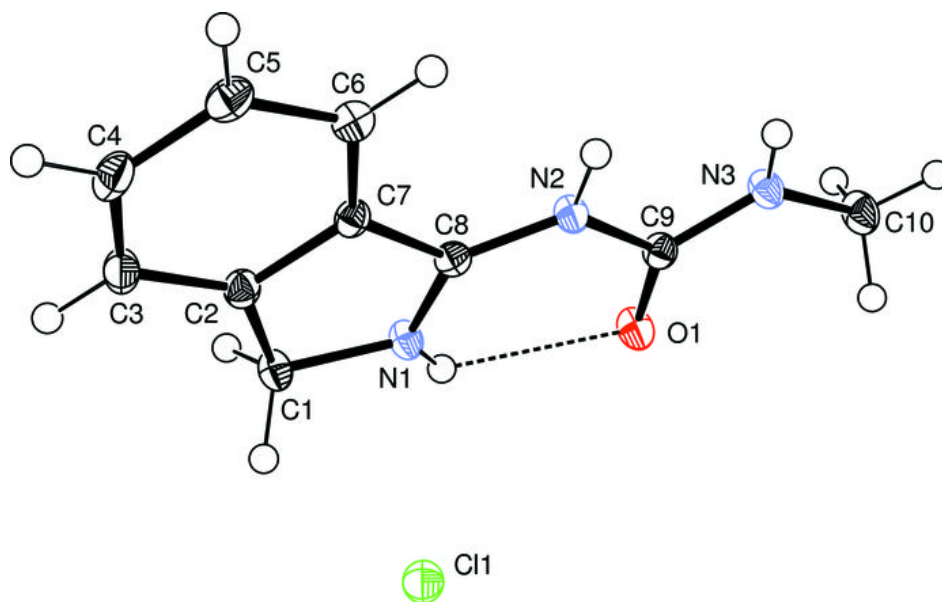


Fig. 2

