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catena-Poly[[[diaquaterbium(III)]-μ-6-carboxynicotinato-μ-pyridine-2,5-dicarboxylato] dihydrate]

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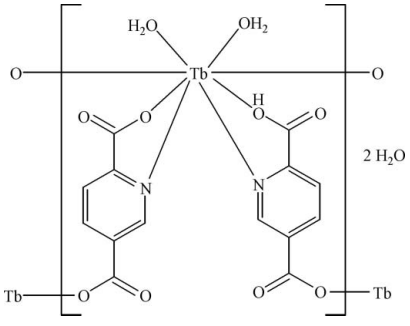
Key indicators: single-crystal X-ray study; T = 298 K; mean σ(C–C) = 0.014 Å; R factor = 0.053; wR factor = 0.131; data-to-parameter ratio = 11.4.

The title compound, $\{[Tb(C_7H_3NO_4)(C_7H_4NO_4)(H_2O)_2]\cdot 2H_2O\}_n$, is isotypic with the analogous Tm^{III} compound [Li, Zhang, Wang & Bai (2009). *Acta Cryst.* E**65**, m411]. The Tb^{III} atom is octacoordinated by two water molecules and by four carboxylate O atoms and two pyridyl N atoms from two pyridine-2,5-dicarboxylate (2,5-pydc) and two 6-carboxynicotinate (2,5-Hpydc) ligands. The 2,5-pydc and 2,5-Hpydc ligands bridge Tb^{III} atoms, generating helical coordination polymers along [001]. An extensive network of O—H···O hydrogen bonds is formed between the coordination polymers and the uncoordinated water molecules. The refined Flack parameter of 0.54 (2) suggests inversion twinning.

Related literature

For the isotypic Tm^{III} compound, see Li *et al.* (2009). For other related structures, see: Huang *et al.* (2007).



Experimental

Crystal data

$[Tb(C_7H_3NO_4)(C_7H_4NO_4)(H_2O)_2]\cdot 2H_2O$
M_r = 562.20
Tetragonal, $I\overline{4}$
a = 15.107 (2) Å
c = 14.8587 (15) Å
V = 3391.1 (7) Å^3
Z = 8
Mo $K\alpha$ radiation
μ = 4.25 mm^{-1}
T = 298 K
0.12 × 0.11 × 0.08 mm

Data collection

Bruker APEXII CCD diffractometer
Absorption correction: multi-scan (*SADABS*; Bruker, 2001)
T_{min} = 0.617, T_{max} = 0.713
6901 measured reflections
3001 independent reflections
2886 reflections with $I > 2\sigma(I)$
R_{int} = 0.072

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.053
$wR(F^2)$ = 0.131
S = 1.03
3001 reflections
263 parameters
H-atom parameters constrained
$\Delta\rho_{max}$ = 3.51 e Å^{-3}
$\Delta\rho_{min}$ = −1.17 e Å^{-3}
Absolute structure: Flack (1983), with 1387 Friedel pairs
Flack parameter: 0.54 (2)

Table 1
Hydrogen-bond geometry (Å, °).

D—H···A	D—H	H···A	D···A	D—H···A
O1—H1···O12^{i}	0.85	1.98	2.801 (12)	162
O9—H91···O4^{ii}	0.85	1.86	2.706 (10)	180
O9—H92···O4^{iii}	0.85	1.99	2.842 (11)	180
O10—H101···O7^{iv}	0.85	1.83	2.679 (11)	179
O10—H102···O9^{i}	0.85	2.15	3.000 (11)	180
O11—H111···O5	0.85	2.00	2.851 (12)	180
O11—H112···O2^{iv}	0.85	1.91	2.758 (12)	180
O12—H121···O6^{v}	0.85	2.15	3.000 (12)	180
O12—H122···O6^{vi}	0.85	2.08	2.930 (13)	180

Symmetry codes: (i) $y+\frac{1}{2}, -x+\frac{1}{2}, -z+\frac{1}{2}$; (ii) $-x+\frac{1}{2}, -y+\frac{1}{2}, z+\frac{1}{2}$; (iii) $-y+1, x, -z$; (iv) $-x+\frac{1}{2}, -y+\frac{1}{2}, z-\frac{1}{2}$; (v) $-x, -y, z$; (vi) $y, -x, -z$.

Data collection: *APEX2* (Bruker, 2004); cell refinement: *SAINT-Plus* (Bruker, 2001); data reduction: *SAINT-Plus*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: BI2349).

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Bruker (2004). *APEX2*. Bruker AXS Inc., Madison, Wisconsin, USA.
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Huang, Y. G., Wu, B. L., Yuan, D. Q., Xu, Y. Q., Jiang, F. L. & Hong, M. C. (2007). *Inorg. Chem.* **46**, 1171–1176.
Li, S., Chen, Y., He, H.-M. & Ma, Y.-F. (2009). *Acta Cryst.* E**65**, m411.
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supporting information

Acta Cryst. (2009). E**65**, m410 [doi:10.1107/S1600536809008824]

catena-Poly[[[diaquaterbium(III)]-μ-6-carboxynicotinato-μ-pyridine-2,5-dicarboxylato] dihydrate]

Sheng Li, Fu-Li Zhang, Shou-Bin Wang and Hui-Ling Bai

S1. Comment

The asymmetric unit of the title compound is shown in Fig. 1. Atom Tb1 displays octa-coordination through two water molecules, four carboxylate O atoms and two pyridyl N atoms from two 2,5-pydc and two 2,5-Hpydc ligands (2,5-pydc = pyridine-2,5-dicarboxylate). The 2,5-pydc and 2,5-Hpydc ligands bridge between Tb^{III} atoms to generate helical coordination polymers along [001] (Fig. 2). An extensive network of O—H···O hydrogen bonds is formed between the coordination polymers and the lattice water molecules (Table 1 and Fig. 3).

S2. Experimental

A mixture of terbium oxide (0.5 mmol), pyridine-2,5-dicarboxylic acid (0.5 mmol), in H_2O (8 ml) and ethanol (8 ml) was sealed in a 25 ml Teflon-lined stainless steel autoclave and kept at 413 K for three days. Colourless crystals were obtained after cooling to room temperature with a yield of 27%. Elemental analysis calculated for $C_{14}H_{15}N_2TbO_{12}$: C 30.68, H 2.74, N 5.11%; Found: C 30.62, H 2.72, N 5.06%.

S3. Refinement

H atoms bound to C atoms were placed in calculated positions with C—H = 0.93 Å and refined as riding with U_{iso}(H) = 1.2U_{eq}(C). H atoms of the water molecules were placed so as to form a reasonable H-bond network and refined as riding with U_{iso}(H) = 1.5U_{eq}(O). The Flack parameter was refined as a full least-squares variable, and the refined value of 0.54 (2) suggests inversion twinning.

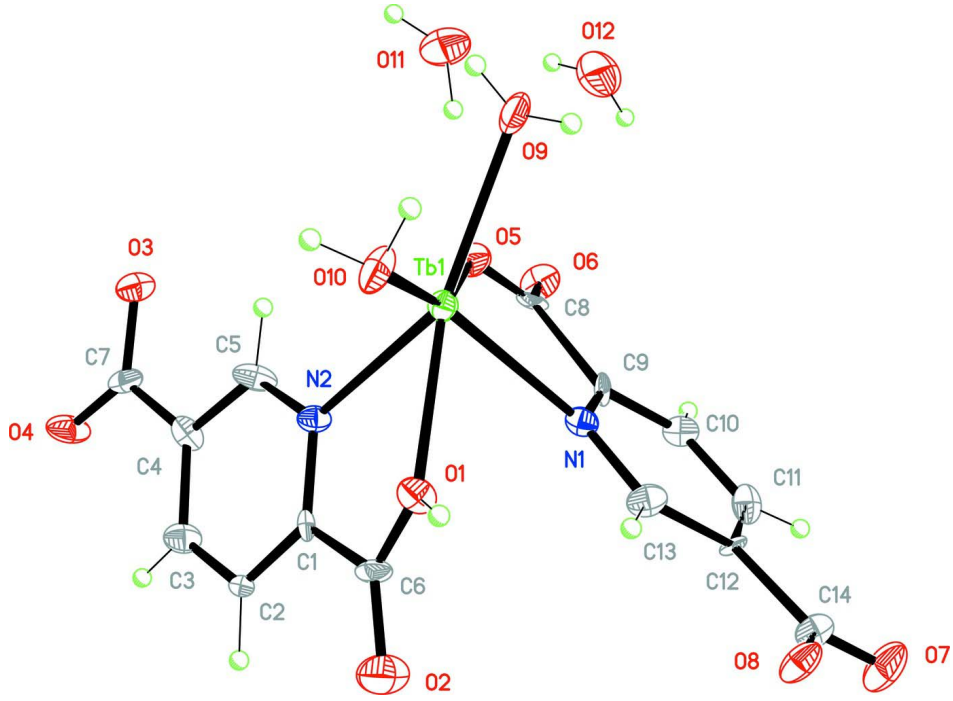

Figure 1

Asymmetric unit of the title compound, showing 50% probability displacement ellipsoids for non-H atoms.

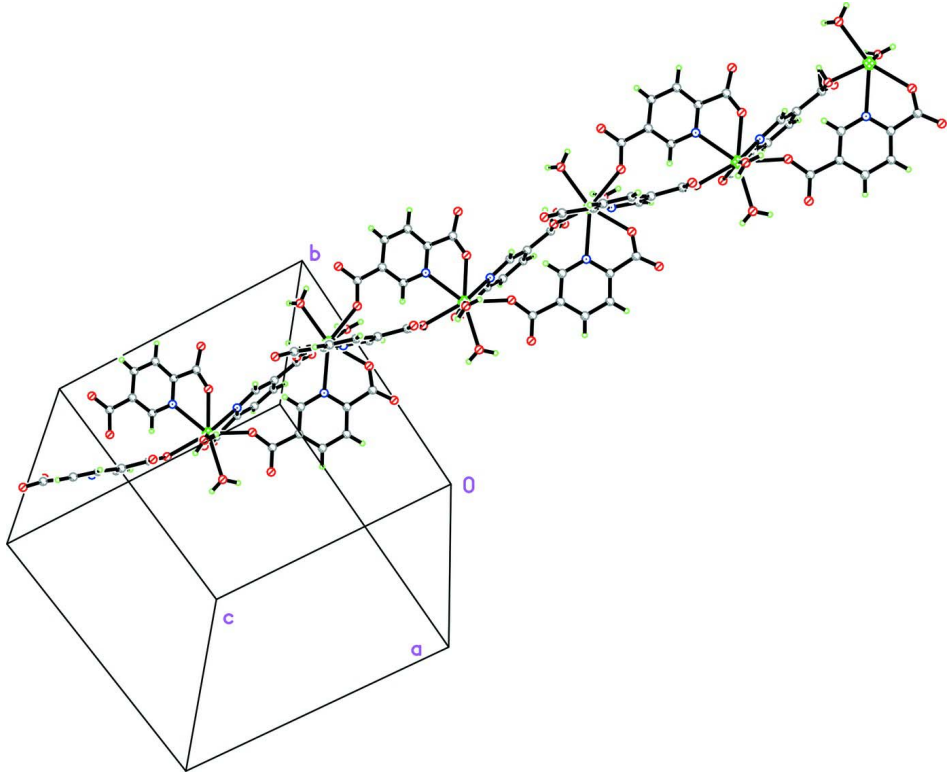

Figure 2

One-dimensional coordination polymer running along [001].

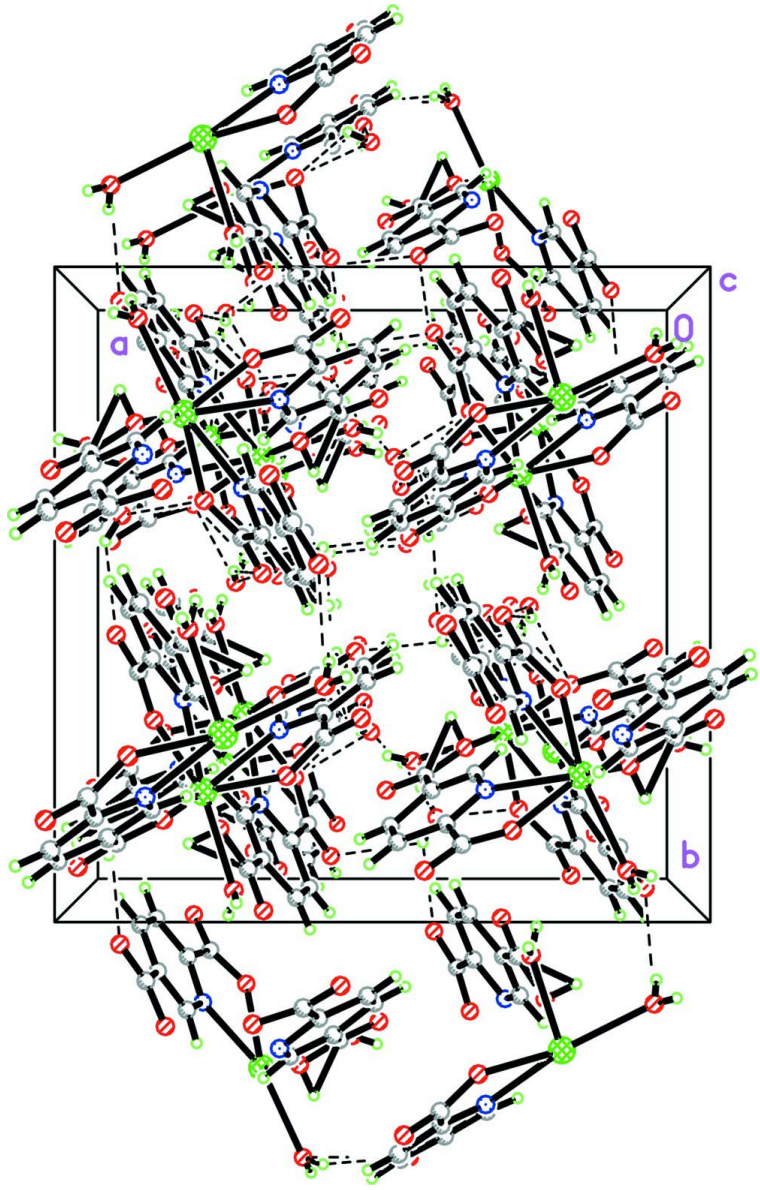

Figure 3

Projection along [001], showing the tetragonal arrangement of coordination polymers. O—H···O hydrogen bonds are shown as dashed lines.

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$[Tb(C_7H_3NO_4)(C_7H_4NO_4)(H_2O)_2]\cdot 2H_2O$
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Tetragonal, $I\bar{4}$
Hall symbol: I -4
$a = 15.107\ (2)$ Å
$c = 14.8587\ (15)$ Å
$V = 3391.1\ (7)$ Å^3
$Z = 8$
$F(000) = 2192$
$D_x = 2.202$ Mg m^{-3}
Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å
Cell parameters from 3001 reflections
$\theta = 1.9$–$25.3°$
$\mu = 4.25$ mm^{-1}
$T = 298$ K
Block, colourless
0.12 × 0.11 × 0.08 mm

Data collection

Bruker APEXII CCD diffractometer
Radiation source: fine-focus sealed tube
Graphite monochromator
φ and ω scans
Absorption correction: multi-scan (*SADABS*; Bruker, 2001)
T_{min} = 0.617, T_{max} = 0.713
6901 measured reflections
3001 independent reflections
2886 reflections with $I > 2\sigma(I)$
R_{int} = 0.072
θ_{max} = 25.3°, θ_{min} = 1.9°
h = −12→18
k = −18→17
l = −13→17

Refinement

Refinement on F^2
Least-squares matrix: full
$R[F^2 > 2\sigma(F^2)]$ = 0.053
$wR(F^2)$ = 0.131
S = 1.03
3001 reflections
263 parameters
0 restraints
Primary atom site location: structure-invariant direct methods
Secondary atom site location: difference Fourier map
Hydrogen site location: inferred from neighbouring sites
H-atom parameters constrained
$w = 1/[\sigma^2(F_o^2) + (0.1007P)^2 + 0.8682P]$ where $P = (F_o^2 + 2F_c^2)/3$
$(\Delta/\sigma)_{max} < 0.001$
$\Delta\rho_{max}$ = 3.51 e Å^{-3}
$\Delta\rho_{min}$ = −1.17 e Å^{-3}
Absolute structure: Flack (1983), 1387 Freidel pairs
Absolute structure parameter: 0.54 (2)

Special details

Geometry. All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R-factor wR and goodness of fit S are based on F^2, conventional R-factors R are based on F, with F set to zero for negative F^2. The threshold expression of $F^2 > 2$sigma(F^2) is used only for calculating R-factors(gt) etc. and is not relevant to the choice of reflections for refinement. R-factors based on F^2 are statistically about twice as large as those based on F, and R- factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	*y*	*z*	U_{iso}*/U_{eq}
Tb1	0.30210 (3)	0.22535 (3)	0.22736 (3)	0.01388 (17)
C1	0.1833 (6)	0.4100 (6)	0.1946 (6)	0.0116 (19)
C2	0.1334 (7)	0.4753 (6)	0.1553 (6)	0.0148 (19)
H2A	0.1128	0.5222	0.1900	0.018*
C3	0.1132 (7)	0.4720 (7)	0.0639 (7)	0.019 (2)
H3A	0.0796	0.5166	0.0376	0.023*
C4	0.1433 (6)	0.4021 (6)	0.0129 (7)	0.015 (2)
C5	0.1879 (7)	0.3364 (7)	0.0575 (8)	0.021 (2)
H5A	0.2046	0.2864	0.0251	0.025*
C6	0.2161 (6)	0.4123 (6)	0.2921 (6)	0.015 (2)
C7	0.1302 (7)	0.3959 (7)	−0.0870 (7)	0.017 (2)
C8	0.1196 (7)	0.1238 (6)	0.1716 (7)	0.014 (2)
C9	0.0975 (6)	0.1488 (6)	0.2710 (7)	0.0129 (18)
C10	0.0233 (6)	0.1192 (7)	0.3157 (7)	0.017 (2)
H10A	−0.0195	0.0865	0.2854	0.021*

C11	0.0131 (7)	0.1383 (7)	0.4052 (7)	0.018 (2)
H11A	−0.0351	0.1158	0.4366	0.022*
C12	0.0768 (6)	0.1929 (6)	0.4506 (6)	0.0115 (18)
C13	0.1449 (7)	0.2207 (7)	0.3980 (7)	0.019 (2)
H13A	0.1865	0.2577	0.4249	0.023*
C14	0.0684 (7)	0.2114 (7)	0.5505 (7)	0.018 (2)
N1	0.1583 (5)	0.2001 (5)	0.3112 (6)	0.0142 (16)
N2	0.2094 (5)	0.3400 (5)	0.1465 (6)	0.0142 (17)
O1	0.2739 (5)	0.3551 (4)	0.3099 (5)	0.0178 (15)
H1	0.3058	0.3643	0.3561	0.027*
O2	0.1862 (6)	0.4677 (6)	0.3436 (5)	0.035 (2)
O3	0.1625 (5)	0.3283 (5)	−0.1271 (5)	0.0212 (16)
O4	0.0921 (6)	0.4561 (5)	−0.1264 (5)	0.0262 (18)
O5	0.1946 (5)	0.1457 (5)	0.1459 (5)	0.0208 (16)
O6	0.0625 (5)	0.0865 (5)	0.1277 (5)	0.0240 (17)
O7	0.0056 (6)	0.1862 (6)	0.5922 (6)	0.033 (2)
O8	0.1326 (5)	0.2550 (5)	0.5846 (5)	0.0214 (16)
O9	0.3704 (5)	0.0819 (5)	0.1997 (5)	0.0240 (17)
H91	0.3822	0.0700	0.2543	0.036*
H92	0.4222	0.0849	0.1778	0.036*
O10	0.4504 (5)	0.2736 (5)	0.2620 (5)	0.0236 (16)
H101	0.4641	0.2868	0.2082	0.035*
H102	0.4875	0.2327	0.2729	0.035*
O11	0.2668 (7)	0.0205 (6)	0.0225 (6)	0.045 (2)
H111	0.2452	0.0578	0.0593	0.068*
H112	0.2812	0.0241	−0.0327	0.068*
O12	0.1257 (7)	−0.0901 (6)	0.0634 (7)	0.046 (2)
H121	0.0724	−0.0890	0.0818	0.068*
H122	0.1145	−0.0821	0.0079	0.068*

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Tb1	0.0145 (3)	0.0144 (3)	0.0128 (2)	0.00004 (17)	−0.00014 (19)	0.00015 (19)
C1	0.013 (5)	0.011 (4)	0.011 (4)	−0.002 (4)	0.001 (4)	0.006 (4)
C2	0.022 (5)	0.008 (4)	0.014 (4)	0.002 (4)	0.007 (4)	−0.001 (4)
C3	0.015 (5)	0.020 (5)	0.023 (5)	0.007 (4)	−0.002 (4)	−0.001 (4)
C4	0.017 (5)	0.013 (5)	0.016 (5)	−0.003 (4)	−0.011 (4)	0.000 (4)
C5	0.021 (5)	0.016 (5)	0.026 (5)	0.005 (4)	−0.007 (4)	−0.009 (4)
C6	0.016 (5)	0.012 (4)	0.017 (6)	−0.001 (4)	0.002 (4)	−0.008 (4)
C7	0.012 (5)	0.020 (5)	0.019 (5)	0.000 (4)	0.004 (4)	−0.007 (4)
C8	0.023 (5)	0.006 (4)	0.012 (4)	0.004 (4)	−0.003 (4)	−0.006 (4)
C9	0.018 (4)	0.013 (4)	0.008 (4)	−0.002 (4)	−0.002 (4)	0.006 (4)
C10	0.014 (5)	0.018 (5)	0.020 (5)	−0.001 (4)	−0.003 (4)	−0.004 (4)
C11	0.015 (5)	0.023 (5)	0.018 (5)	0.000 (4)	−0.007 (4)	0.007 (4)
C12	0.008 (4)	0.015 (4)	0.012 (4)	0.004 (4)	0.008 (4)	0.000 (4)
C13	0.018 (5)	0.022 (5)	0.017 (5)	0.001 (4)	−0.002 (4)	−0.006 (4)
C14	0.014 (5)	0.025 (6)	0.014 (5)	0.001 (5)	−0.003 (4)	−0.004 (4)

N1	0.013 (4)	0.012 (4)	0.018 (4)	−0.001 (3)	−0.001 (4)	−0.002 (4)
N2	0.013 (4)	0.015 (4)	0.015 (4)	0.003 (3)	−0.002 (3)	−0.002 (3)
O1	0.021 (3)	0.016 (3)	0.016 (3)	−0.001 (3)	−0.007 (3)	−0.003 (3)
O2	0.050 (5)	0.035 (5)	0.019 (4)	0.019 (4)	−0.006 (4)	−0.002 (4)
O3	0.020 (4)	0.026 (4)	0.017 (4)	0.007 (3)	−0.002 (3)	−0.004 (3)
O4	0.032 (4)	0.024 (4)	0.022 (4)	0.017 (3)	−0.006 (3)	−0.002 (3)
O5	0.017 (4)	0.021 (4)	0.024 (4)	−0.003 (3)	0.003 (3)	−0.004 (3)
O6	0.019 (4)	0.029 (4)	0.024 (4)	−0.005 (3)	0.001 (3)	−0.006 (3)
O7	0.030 (4)	0.045 (5)	0.024 (4)	−0.016 (4)	0.011 (4)	−0.004 (4)
O8	0.019 (4)	0.032 (4)	0.013 (3)	−0.010 (3)	0.002 (3)	−0.007 (3)
O9	0.025 (4)	0.031 (4)	0.016 (3)	0.000 (3)	0.008 (3)	0.007 (3)
O10	0.021 (4)	0.035 (4)	0.014 (4)	−0.001 (3)	0.005 (3)	0.006 (3)
O11	0.064 (6)	0.039 (5)	0.033 (5)	0.012 (5)	0.014 (5)	−0.010 (4)
O12	0.053 (6)	0.045 (6)	0.039 (5)	0.018 (5)	−0.004 (5)	0.011 (5)

Geometric parameters (Å, º)

Tb1—O1	2.351 (7)	C8—C9	1.560 (14)
Tb1—O5	2.356 (7)	C9—N1	1.341 (12)
Tb1—O8^{i}	2.358 (7)	C9—C10	1.379 (14)
Tb1—O3^{ii}	2.371 (7)	C10—C11	1.368 (16)
Tb1—O10	2.412 (7)	C10—H10A	0.930
Tb1—O9	2.435 (8)	C11—C12	1.435 (14)
Tb1—N2	2.531 (8)	C11—H11A	0.930
Tb1—N1	2.534 (8)	C12—C13	1.359 (14)
C1—N2	1.335 (13)	C12—C14	1.516 (13)
C1—C2	1.373 (13)	C13—N1	1.343 (15)
C1—C6	1.531 (13)	C13—H13A	0.930
C2—C3	1.393 (15)	C14—O7	1.195 (14)
C2—H2A	0.930	C14—O8	1.277 (12)
C3—C4	1.378 (15)	O1—H1	0.850
C3—H3A	0.930	O3—Tb1^{i}	2.371 (7)
C4—C5	1.370 (15)	O8—Tb1^{ii}	2.358 (7)
C4—C7	1.500 (15)	O9—H91	0.850
C5—N2	1.363 (15)	O9—H92	0.850
C5—H5A	0.930	O10—H101	0.850
C6—O2	1.221 (13)	O10—H102	0.850
C6—O1	1.256 (12)	O11—H111	0.850
C7—O4	1.226 (13)	O11—H112	0.850
C7—O3	1.279 (13)	O12—H121	0.850
C8—O6	1.219 (12)	O12—H122	0.850
C8—O5	1.242 (13)		
O1—Tb1—O5	124.7 (3)	O1—C6—C1	114.1 (8)
O1—Tb1—O8^{i}	116.1 (3)	O4—C7—O3	123.3 (10)
O5—Tb1—O8^{i}	83.7 (2)	O4—C7—C4	119.3 (9)
O1—Tb1—O3^{ii}	81.4 (3)	O3—C7—C4	117.4 (9)
O5—Tb1—O3^{ii}	116.7 (3)	O6—C8—O5	127.2 (10)

O8^{i}—Tb1—O3^{ii}	140.3 (2)	O6—C8—C9	117.9 (9)
O1—Tb1—O10	78.8 (2)	O5—C8—C9	114.9 (8)
O5—Tb1—O10	154.8 (3)	N1—C9—C10	121.9 (10)
O8^{i}—Tb1—O10	76.4 (2)	N1—C9—C8	114.6 (9)
O3^{ii}—Tb1—O10	72.5 (2)	C10—C9—C8	123.5 (9)
O1—Tb1—O9	155.1 (2)	C11—C10—C9	119.4 (10)
O5—Tb1—O9	75.6 (3)	C11—C10—H10A	120.3
O8^{i}—Tb1—O9	77.5 (3)	C9—C10—H10A	120.3
O3^{ii}—Tb1—O9	75.8 (2)	C10—C11—C12	120.2 (10)
O10—Tb1—O9	84.9 (3)	C10—C11—H11A	119.9
O1—Tb1—N2	64.9 (3)	C12—C11—H11A	119.9
O5—Tb1—N2	74.0 (3)	C13—C12—C11	114.5 (9)
O8^{i}—Tb1—N2	73.6 (3)	C13—C12—C14	124.7 (10)
O3^{ii}—Tb1—N2	142.3 (3)	C11—C12—C14	120.7 (9)
O10—Tb1—N2	114.1 (3)	N1—C13—C12	126.4 (10)
O9—Tb1—N2	139.8 (3)	N1—C13—H13A	116.8
O1—Tb1—N1	73.4 (3)	C12—C13—H13A	116.8
O5—Tb1—N1	65.5 (3)	O7—C14—O8	124.1 (10)
O8^{i}—Tb1—N1	145.1 (3)	O7—C14—C12	121.0 (10)
O3^{ii}—Tb1—N1	72.1 (3)	O8—C14—C12	114.8 (9)
O10—Tb1—N1	137.5 (3)	C9—N1—C13	117.4 (9)
O9—Tb1—N1	108.2 (3)	C9—N1—Tb1	117.1 (7)
N2—Tb1—N1	82.1 (3)	C13—N1—Tb1	124.5 (7)
N2—C1—C2	120.3 (9)	C1—N2—C5	118.7 (9)
N2—C1—C6	115.4 (8)	C1—N2—Tb1	116.8 (6)
C2—C1—C6	124.3 (9)	C5—N2—Tb1	124.4 (7)
C1—C2—C3	120.6 (9)	C6—O1—Tb1	126.1 (6)
C1—C2—H2A	119.7	C6—O1—H1	116.9
C3—C2—H2A	119.7	Tb1—O1—H1	117.0
C4—C3—C2	119.5 (9)	C7—O3—Tb1^{i}	141.6 (7)
C4—C3—H3A	120.2	C8—O5—Tb1	127.4 (6)
C2—C3—H3A	120.2	C14—O8—Tb1^{ii}	137.7 (6)
C5—C4—C3	116.8 (9)	Tb1—O9—H91	96.6
C5—C4—C7	119.9 (9)	Tb1—O9—H92	114.0
C3—C4—C7	123.3 (9)	H91—O9—H92	100.6
N2—C5—C4	123.9 (9)	Tb1—O10—H101	95.5
N2—C5—H5A	118.0	Tb1—O10—H102	115.7
C4—C5—H5A	118.0	H101—O10—H102	100.9
O2—C6—O1	126.6 (9)	H111—O11—H112	132.5
O2—C6—C1	119.3 (9)	H121—O12—H122	96.9

N2—C1—C2—C3	−3.4 (15)	O3^{ii}—Tb1—N1—C13	41.3 (8)
C6—C1—C2—C3	174.1 (9)	O10—Tb1—N1—C13	6.3 (10)
C1—C2—C3—C4	0.4 (15)	O9—Tb1—N1—C13	109.1 (8)
C2—C3—C4—C5	3.7 (15)	N2—Tb1—N1—C13	−111.0 (8)
C2—C3—C4—C7	−175.7 (10)	C2—C1—N2—C5	2.2 (14)
C3—C4—C5—N2	−5.2 (16)	C6—C1—N2—C5	−175.6 (9)
C7—C4—C5—N2	174.2 (10)	C2—C1—N2—Tb1	179.6 (7)

N2—C1—C6—O2	−170.7 (9)	C6—C1—N2—Tb1	1.9 (10)
C2—C1—C6—O2	11.7 (15)	C4—C5—N2—C1	2.3 (16)
N2—C1—C6—O1	10.2 (12)	C4—C5—N2—Tb1	−174.9 (8)
C2—C1—C6—O1	−167.4 (9)	O1—Tb1—N2—C1	−7.5 (6)
C5—C4—C7—O4	−178.2 (10)	O5—Tb1—N2—C1	134.4 (7)
C3—C4—C7—O4	1.1 (16)	O8^{i}—Tb1—N2—C1	−137.6 (7)
C5—C4—C7—O3	−0.4 (15)	O3^{ii}—Tb1—N2—C1	21.2 (9)
C3—C4—C7—O3	178.9 (9)	O10—Tb1—N2—C1	−71.1 (7)
O6—C8—C9—N1	−171.3 (9)	O9—Tb1—N2—C1	176.6 (6)
O5—C8—C9—N1	7.9 (12)	N1—Tb1—N2—C1	67.8 (7)
O6—C8—C9—C10	10.5 (13)	O1—Tb1—N2—C5	169.7 (9)
O5—C8—C9—C10	−170.3 (9)	O5—Tb1—N2—C5	−48.3 (8)
N1—C9—C10—C11	−3.6 (15)	O8^{i}—Tb1—N2—C5	39.7 (8)
C8—C9—C10—C11	174.4 (9)	O3^{ii}—Tb1—N2—C5	−161.5 (7)
C9—C10—C11—C12	3.5 (15)	O10—Tb1—N2—C5	106.2 (8)
C10—C11—C12—C13	−0.9 (14)	O9—Tb1—N2—C5	−6.1 (10)
C10—C11—C12—C14	−177.4 (9)	N1—Tb1—N2—C5	−115.0 (8)
C11—C12—C13—N1	−1.9 (15)	O2—C6—O1—Tb1	161.3 (8)
C14—C12—C13—N1	174.5 (10)	C1—C6—O1—Tb1	−19.6 (12)
C13—C12—C14—O7	178.6 (11)	O5—Tb1—O1—C6	−30.8 (9)
C11—C12—C14—O7	−5.3 (15)	O8^{i}—Tb1—O1—C6	70.2 (8)
C13—C12—C14—O8	−1.4 (14)	O3^{ii}—Tb1—O1—C6	−147.4 (8)
C11—C12—C14—O8	174.7 (9)	O10—Tb1—O1—C6	138.9 (8)
C10—C9—N1—C13	1.0 (14)	O9—Tb1—O1—C6	−171.0 (7)
C8—C9—N1—C13	−177.2 (8)	N2—Tb1—O1—C6	15.4 (8)
C10—C9—N1—Tb1	169.6 (7)	N1—Tb1—O1—C6	−73.5 (8)
C8—C9—N1—Tb1	−8.6 (10)	O4—C7—O3—Tb1^{i}	12.3 (17)
C12—C13—N1—C9	1.8 (16)	C4—C7—O3—Tb1^{i}	−165.4 (7)
C12—C13—N1—Tb1	−165.8 (8)	O6—C8—O5—Tb1	176.0 (8)
O1—Tb1—N1—C9	147.5 (7)	C9—C8—O5—Tb1	−3.1 (12)
O5—Tb1—N1—C9	5.4 (6)	O1—Tb1—O5—C8	−46.5 (9)
O8^{i}—Tb1—N1—C9	35.5 (9)	O8^{i}—Tb1—O5—C8	−164.1 (8)
O3^{ii}—Tb1—N1—C9	−126.4 (7)	O3^{ii}—Tb1—O5—C8	51.8 (8)
O10—Tb1—N1—C9	−161.3 (6)	O10—Tb1—O5—C8	157.8 (7)
O9—Tb1—N1—C9	−58.6 (7)	O9—Tb1—O5—C8	117.3 (8)
N2—Tb1—N1—C9	81.4 (7)	N2—Tb1—O5—C8	−89.3 (8)
O1—Tb1—N1—C13	−44.8 (8)	N1—Tb1—O5—C8	−0.9 (8)
O5—Tb1—N1—C13	173.0 (9)	O7—C14—O8—Tb1^{ii}	23.9 (17)
O8^{i}—Tb1—N1—C13	−156.8 (7)	C12—C14—O8—Tb1^{ii}	−156.1 (7)

Symmetry codes: (i) −*x*+1/2, −*y*+1/2, *z*−1/2; (ii) −*x*+1/2, −*y*+1/2, *z*+1/2.

Hydrogen-bond geometry (Å, º)

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
O1—H1···O12^{iii}	0.85	1.98	2.801 (12)	162
O9—H91···O4^{ii}	0.85	1.86	2.706 (10)	180
O9—H92···O4^{iv}	0.85	1.99	2.842 (11)	180

O10—H101···O7^{i}	0.85	1.83	2.679 (11)	179
O10—H102···O9^{iii}	0.85	2.15	3.000 (11)	180
O11—H111···O5	0.85	2.00	2.851 (12)	180
O11—H112···O2^{i}	0.85	1.91	2.758 (12)	180
O12—H121···O6^{v}	0.85	2.15	3.000 (12)	180
O12—H122···O6^{vi}	0.85	2.08	2.930 (13)	180

Symmetry codes: (i) $-x+1/2, -y+1/2, z-1/2$; (ii) $-x+1/2, -y+1/2, z+1/2$; (iii) $y+1/2, -x+1/2, -z+1/2$; (iv) $-y+1, x, -z$; (v) $-x, -y, z$; (vi) $y, -x, -z$.