

Acta Crystallographica Section E

Structure Reports

Online

ISSN 1600-5368

(Z)-(1,3-Thiazinan-2-ylideneamino)-formonitrileYu-wen Peng^{a*} and Lin-hai Wu^b

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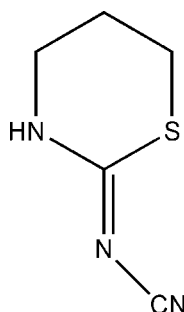
Received 9 March 2009; accepted 11 March 2009

Key indicators: single-crystal X-ray study; $T = 153$ K; mean $\sigma(\text{C}-\text{C}) = 0.006$ Å; R factor = 0.066; wR factor = 0.192; data-to-parameter ratio = 13.3.

In the title molecule, $\text{C}_5\text{H}_7\text{N}_3\text{S}$, the thiazine ring shows a conformation close to a half-boat. The Cremer & Pople puckering parameters of the thiazine ring are $q_2 = 0.4645$ (2) Å, $\theta = 132.4$ (3) and $\varphi = 285.52$ (2)°. The packing is stabilized by intermolecular $\text{N}-\text{H}\cdots\text{N}$ and $\text{C}-\text{H}\cdots\text{S}$ interactions.

Related literature

For the crystal structures of thiazine compounds, see: Kálmán, *et al.* (1977). For the biological activities of thiazine-containing compounds, see: Soloway *et al.* (1978); Tomizawa *et al.* (1995). For bond-length data, see: Allen *et al.* (1987). For puckering parameters, see: Cremer & Pople (1975).

**Experimental***Crystal data*

$\text{C}_5\text{H}_7\text{N}_3\text{S}$
 $M_r = 141.21$
 Monoclinic, $P2_1/c$
 $a = 7.0931$ (14) Å
 $b = 12.689$ (3) Å
 $c = 9.232$ (3) Å
 $\beta = 128.617$ (19)°
 $V = 649.2$ (3) Å³
 $Z = 4$
 Mo $K\alpha$ radiation
 $\mu = 0.40$ mm⁻¹
 $T = 153$ K
 $0.42 \times 0.11 \times 0.06$ mm

Data collection

Rigaku R-AXIS RAPID IP area-detector diffractometer
 Absorption correction: multi-scan (ABSCOR; Higashi 1995)
 $T_{\min} = 0.849$, $T_{\max} = 0.976$
 4731 measured reflections
 1101 independent reflections
 970 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.032$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.066$
 $wR(F^2) = 0.192$
 $S = 1.09$
 1101 reflections
 83 parameters
 H-atom parameters constrained
 $\Delta\rho_{\max} = 1.14$ e Å⁻³
 $\Delta\rho_{\min} = -0.32$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{N1}-\text{H1A}\cdots\text{N3}^{\text{i}}$	0.86	2.12	2.926 (4)	156
$\text{C3}-\text{H3B}\cdots\text{S1}^{\text{ii}}$	0.99	2.74	3.468 (3)	131

Symmetry codes: (i) $-x + 1, y + \frac{1}{2}, -z + \frac{1}{2}$; (ii) $-x + 2, y + \frac{1}{2}, -z + \frac{3}{2}$.

Data collection: *RAPID-AUTO* (Rigaku, 2004); cell refinement: *RAPID-AUTO*; data reduction: *RAPID-AUTO*; program(s) used to solve structure: *SHELXTL* (Sheldrick, 2008); program(s) used to refine structure: *SHELXTL*; molecular graphics: *SHELXTL*; software used to prepare material for publication: *SHELXTL*.

We gratefully acknowledge Dr Jing-Kun Xu for valuable advice given by him in the preparation of the reported compound.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HG2486).

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supporting information

Acta Cryst. (2009). E65, o784 [doi:10.1107/S1600536809008940]

(Z)-(1,3-Thiazinan-2-ylideneamino)formonitrile**Yu-wen Peng and Lin-hai Wu****S1. Comment**

The frequent occurrence of pesticide residues accidents and the introduction of Green Trade Barriers to protect human health threaten export and national agriculture market and highlights the importance of food safety. Therefore, the development of pesticides with new chemical structures and high insecticidal activities with low residues is highly desirable. Consequently, spurred by the need for new pesticidal agents and the fact that many new effective pesticides possess heterocyclic rings in their structure, such as the thiazine ring (Soloway *et al.*, 1978; Tomizawa *et al.*, 1995), over the last few years, we have synthesized some novel thiazine derivatives. Here, we report the crystal structure of (Z)-(1,3-thiazinan-2-ylideneamino)formonitrile.

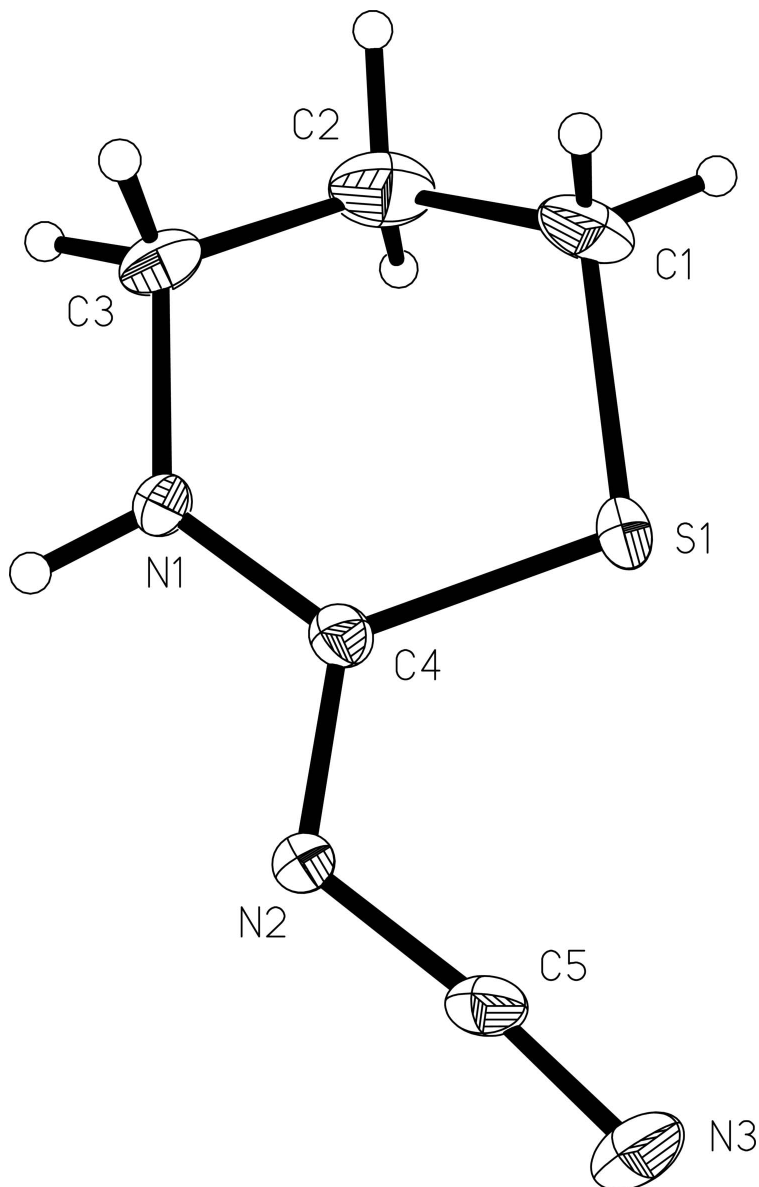
In (Z)-(1,3-thiazinan-2-ylideneamino)formonitrile (Fig. 1), all bond lengths are normal (Allen *et al.*, 1987) and in a good agreement with those reported previously (Kálmán, *et al.*, 1977). It is known that the imino tautomers can exist as two geometrical isomers, *syn* (Z) and *anti* (E), but in this crystal, only Z isomers have been observed. The thiazine ring shows a conformation near to half boat with the C2 atom deviating 0.618 (2) Å above the plane formed by S1, C1, N1, C3 and C4 [maximum least squares plane deviation for S1 0.053 (3) Å]. This geometry is proved by the puckering parameters $q_2 = 0.4645$ (2) Å, $\theta = 132.4$ (3)° and $\varphi = 285.52$ (2)° (Cremer & Pople, 1975). There are some weak N—H⋯N and C—H⋯S intermolecular interactions (see Table 1) which stabilize the title structure.

S2. Experimental

A mixture of dimethyl cyanocarbonimidodithioate 10 mmol (1.46 g) and 3-aminopropane-1-thiol (1.00 g, 11 mmol) was refluxed in absolute EtOH (20 ml) for 3 h. On cooling, the product crystallizes and was filtered and then recrystallized from absolute ethanol. Yield 1.20 g (85%). Single crystals suitable for X-ray measurements were obtained by recrystallization from ethanol at room temperature.

S3. Refinement

H atoms were positioned geometrically and refined using a riding model, with C—H = 0.99 Å, N—H = 0.86 Å with $U_{\text{iso}}(\text{H}) = 1.2$ times $U_{\text{eq}}(\text{C}, \text{N})$.

**Figure 1**

The molecular structure of (I), with atom labels and 40% probability displacement ellipsoids for non-H atoms.

(Z)-(1,3-Thiazinan-2-ylideneamino)formonitrile

Crystal data

$C_5H_7N_3S$

$M_r = 141.21$

Monoclinic, $P2_1/c$

Hall symbol: $-P\ 2_1/c$

$a = 7.0931(14)\ \text{\AA}$

$b = 12.689(3)\ \text{\AA}$

$c = 9.232(3)\ \text{\AA}$

$\beta = 128.617(19)^\circ$

$V = 649.2(3)\ \text{\AA}^3$

$Z = 4$

$F(000) = 296$

$D_x = 1.445\ \text{Mg m}^{-3}$

Mo $K\alpha$ radiation, $\lambda = 0.71073\ \text{\AA}$

Cell parameters from 2501 reflections

$\theta = 2.3\text{--}25.1^\circ$

$\mu = 0.40\ \text{mm}^{-1}$

$T = 153\ \text{K}$

Needle, colorless

$0.42 \times 0.11 \times 0.06\ \text{mm}$

Data collection

Rigaku R-AXIS RAPID IP area-detector
diffractometer
Radiation source: Rotating Anode
Graphite monochromator
 ω Oscillation scans
Absorption correction: multi-scan
(*ABSCOR*; Higashi 1995)
 $T_{\min} = 0.849$, $T_{\max} = 0.976$

4731 measured reflections
1101 independent reflections
970 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.032$
 $\theta_{\max} = 25.0^\circ$, $\theta_{\min} = 3.2^\circ$
 $h = -8 \rightarrow 8$
 $k = -14 \rightarrow 15$
 $l = -10 \rightarrow 10$

Refinement

Refinement on F^2
Least-squares matrix: full
 $R[F^2 > 2\sigma(F^2)] = 0.066$
 $wR(F^2) = 0.192$
 $S = 1.09$
1101 reflections
83 parameters
0 restraints
Primary atom site location: structure-invariant
direct methods
Secondary atom site location: difference Fourier
map

Hydrogen site location: inferred from
neighbouring sites
H-atom parameters constrained
 $w = 1/[\sigma^2(F_o^2) + (0.1515P)^2 + 0.068P]$
where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\max} < 0.001$
 $\Delta\rho_{\max} = 1.14 \text{ e } \text{\AA}^{-3}$
 $\Delta\rho_{\min} = -0.32 \text{ e } \text{\AA}^{-3}$
Extinction correction: *SHELXTL* (Sheldrick,
2008), $F_c^* = kFc[1 + 0.001x \text{Fc}^2 \lambda^3 / \sin(2\theta)]^{-1/4}$
Extinction coefficient: 0.040 (14)

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
S1	0.97192 (13)	0.28689 (6)	0.64894 (10)	0.0263 (5)
N1	0.7395 (4)	0.47140 (17)	0.5152 (3)	0.0213 (6)
H1A	0.6246	0.5077	0.4223	0.026*
N2	0.5695 (4)	0.33420 (19)	0.3151 (3)	0.0231 (7)
N3	0.5490 (5)	0.1469 (2)	0.2340 (4)	0.0324 (8)
C1	1.1467 (7)	0.3667 (3)	0.8561 (5)	0.0406 (10)
H1B	1.0764	0.3613	0.9206	0.049*
H1C	1.3136	0.3390	0.9404	0.049*
C2	1.1539 (7)	0.4785 (3)	0.8161 (5)	0.0403 (9)
H2B	1.2330	0.4840	0.7582	0.048*
H2C	1.2538	0.5179	0.9345	0.048*
C3	0.9066 (6)	0.5297 (3)	0.6889 (5)	0.0316 (8)
H3A	0.8374	0.5342	0.7544	0.038*
H3B	0.9245	0.6023	0.6600	0.038*
C4	0.7430 (5)	0.3711 (2)	0.4833 (4)	0.0189 (7)

C5	0.5676 (5)	0.2335 (2)	0.2802 (4)	0.0228 (7)
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Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
S1	0.0269 (6)	0.0189 (6)	0.0199 (6)	0.0071 (3)	0.0081 (5)	0.0043 (2)
N1	0.0262 (13)	0.0119 (11)	0.0189 (13)	0.0008 (10)	0.0108 (12)	-0.0002 (9)
N2	0.0243 (13)	0.0159 (13)	0.0174 (13)	0.0017 (10)	0.0073 (11)	-0.0020 (9)
N3	0.0362 (17)	0.0226 (15)	0.0376 (18)	-0.0079 (12)	0.0226 (15)	-0.0118 (12)
C1	0.0317 (19)	0.052 (2)	0.0163 (17)	0.0024 (15)	0.0041 (15)	-0.0040 (14)
C2	0.040 (2)	0.044 (2)	0.032 (2)	-0.0136 (16)	0.0198 (18)	-0.0108 (15)
C3	0.0365 (19)	0.0224 (16)	0.0284 (18)	-0.0043 (14)	0.0165 (16)	-0.0111 (13)
C4	0.0206 (14)	0.0179 (14)	0.0165 (15)	0.0000 (11)	0.0107 (13)	0.0014 (11)
C5	0.0180 (14)	0.0265 (16)	0.0181 (15)	-0.0046 (12)	0.0085 (13)	-0.0056 (12)

Geometric parameters (Å, °)

S1—C4	1.737 (3)	C1—C2	1.474 (6)
S1—C1	1.806 (4)	C1—H1B	0.9900
N1—C4	1.310 (4)	C1—H1C	0.9900
N1—C3	1.467 (4)	C2—C3	1.519 (5)
N1—H1A	0.8600	C2—H2B	0.9900
N2—C5	1.315 (4)	C2—H2C	0.9900
N2—C4	1.327 (4)	C3—H3A	0.9900
N3—C5	1.156 (4)	C3—H3B	0.9900
C4—S1—C1	103.52 (16)	C1—C2—H2C	108.8
C4—N1—C3	128.3 (3)	C3—C2—H2C	108.8
C4—N1—H1A	115.8	H2B—C2—H2C	107.7
C3—N1—H1A	115.9	N1—C3—C2	112.6 (3)
C5—N2—C4	119.3 (2)	N1—C3—H3A	109.1
C2—C1—S1	112.5 (3)	C2—C3—H3A	109.1
C2—C1—H1B	109.1	N1—C3—H3B	109.1
S1—C1—H1B	109.1	C2—C3—H3B	109.1
C2—C1—H1C	109.1	H3A—C3—H3B	107.8
S1—C1—H1C	109.1	N1—C4—N2	118.2 (3)
H1B—C1—H1C	107.8	N1—C4—S1	122.6 (2)
C1—C2—C3	113.7 (3)	N2—C4—S1	119.2 (2)
C1—C2—H2B	108.8	N3—C5—N2	174.2 (3)
C3—C2—H2B	108.8		
C4—S1—C1—C2	-34.3 (3)	C5—N2—C4—N1	-179.3 (3)
S1—C1—C2—C3	59.4 (4)	C5—N2—C4—S1	3.5 (4)
C4—N1—C3—C2	25.2 (4)	C1—S1—C4—N1	7.7 (3)
C1—C2—C3—N1	-54.0 (4)	C1—S1—C4—N2	-175.3 (2)
C3—N1—C4—N2	178.6 (3)	C4—N2—C5—N3	-177 (3)
C3—N1—C4—S1	-4.3 (4)		

Hydrogen-bond geometry (Å, °)

<i>D</i> —H \cdots <i>A</i>	<i>D</i> —H	H \cdots <i>A</i>	<i>D</i> \cdots <i>A</i>	<i>D</i> —H \cdots <i>A</i>
N1—H1A \cdots N3 ⁱ	0.86	2.12	2.926 (4)	156
C3—H3B \cdots S1 ⁱⁱ	0.99	2.74	3.468 (3)	131

Symmetry codes: (i) $-x+1, y+1/2, -z+1/2$; (ii) $-x+2, y+1/2, -z+3/2$.