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## 2-Bromo-4-chloro-6-[(*E*)-(2-chlorophenyl)iminomethyl]phenol

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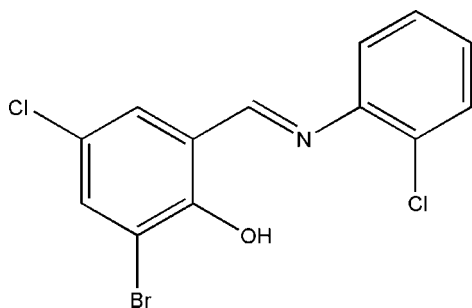
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 Key indicators: single-crystal X-ray study;  $T = 298$  K; mean  $\sigma(\text{C}-\text{C}) = 0.005$  Å;  $R$  factor = 0.032;  $wR$  factor = 0.077; data-to-parameter ratio = 14.0.

The title compound,  $\text{C}_{13}\text{H}_8\text{BrCl}_2\text{NO}$ , was obtained by reaction of 3-bromo-5-chlorosalicylaldehyde and 2-chlorobenzeneamine in methanol. The molecule displays an *E* configuration with respect to the imine  $\text{C}=\text{N}$  double bond. The dihedral angle between the two benzene rings is  $4.57$  ( $11$ )°. The molecular conformation is stabilized by an intramolecular  $\text{O}-\text{H}\cdots\text{N}$  hydrogen bond. In the crystal structure, molecules are linked by intermolecular  $\text{C}-\text{H}\cdots\text{O}$  hydrogen-bonding interactions into zigzag chains running parallel to the *b* axis. Intermolecular  $\text{Br}\cdots\text{Cl}$  [ $3.5289$  ( $11$ ) Å] and  $\text{Cl}\cdots\text{Cl}$  [ $3.5042$  ( $12$ ) Å] interactions are present.

### Related literature

For the biological activities of Schiff base complexes, see: Cukurovali *et al.* (2002); Tarafder *et al.* (2002); Ali *et al.* (2002). For halogen-halogen interactions, see: Saruma *et al.* (1986); Moorthy *et al.* (2002).



### Experimental

#### Crystal data

 $\text{C}_{13}\text{H}_8\text{BrCl}_2\text{NO}$ 
 $M_r = 345.01$ 

 Monoclinic,  $P2_1/c$   
 $a = 8.4299$  (10) Å  
 $b = 14.0115$  (16) Å  
 $c = 11.4194$  (14) Å  
 $\beta = 104.5120$  (10)°  
 $V = 1305.8$  (3) Å<sup>3</sup>
 $Z = 4$   
 Mo  $K\alpha$  radiation  
 $\mu = 3.54$  mm<sup>-1</sup>  
 $T = 298$  K  
 $0.45 \times 0.38 \times 0.36$  mm

#### Data collection

 Siemens SMART CCD area-detector diffractometer  
 Absorption correction: multi-scan (SADABS; Siemens, 1996)  
 $T_{\min} = 0.230$ ,  $T_{\max} = 0.279$ 

 6450 measured reflections  
 2295 independent reflections  
 1726 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.038$ 

#### Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.032$   
 $wR(F^2) = 0.077$   
 $S = 1.03$   
 2295 reflections

 164 parameters  
 H-atom parameters constrained  
 $\Delta\rho_{\max} = 0.39$  e Å<sup>-3</sup>  
 $\Delta\rho_{\min} = -0.58$  e Å<sup>-3</sup>
**Table 1**

Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{O1}-\text{H1}\cdots\text{N1}$	0.82	1.86	2.586 (3)	147
$\text{C11}-\text{H11}\cdots\text{O1}^i$	0.93	2.56	3.324 (5)	139

 Symmetry code: (i)  $-x, y - \frac{1}{2}, -z + \frac{1}{2}$ 

Data collection: SMART (Siemens, 1996); cell refinement: SAINT (Siemens, 1996); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: SHELXTL (Sheldrick, 2008); software used to prepare material for publication: SHELXTL.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: RZ2298).

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**supplementary materials**

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### Comment

Schiff base complexes are of great interests for inorganic and bioinorganic chemistry. To the best of our knowledge, in the past two decades Schiff base ligands have demonstrated significant biological activities and new examples have been tested for their antitumor, antimicrobial and antiviral activities (Tarafder *et al.*, 2002; Cukurovali *et al.*, 2002; Ali *et al.*, 2002). As an extension of the work on the structural characterization of Schiff base compounds, the crystal structure of the title compound is reported here.

The molecular structure and crystal packing of the title compound are illustrated in Fig. 1 and 2, respectively. The C1=N1 bond distance (1.279 (4) Å) is shorter than expected. The molecule is not strictly planar, the maximum deviations from the planarity are 0.199 (5) and 0.162 (5) for atoms C11 and C12. The dihedral angle formed by the benzene rings is 4.57 (11)°. The molecular conformation is stabilized by an intramolecular O—H...N hydrogen bond (Table 1). In the crystal packing, the molecules are linked via intermolecular C—H...O hydrogen bonds into zig-zag chains running parallel to the *b* axis. In addition, intermolecular Br...Cl and Cl...Cl interactions are observed (Fig. 2) falling in the typical range of halogen-halogen interactions (Saruma & Desiraju, 1986, Moorthy *et al.*, 2002): Br1...Cl1<sup>i</sup> = 3.5289 (11) Å; Cl1...Cl2<sup>ii</sup> = 3.5042 (12) Å; symmetry codes: (i) *x*, 3/2-*y*, -1/2+*z*; (ii) 1+*x*, *y*, 1+*z*.

### Experimental

3-Bromo-5-chlorosalicylaldehyde (0.1 mmol, 23.6 mg) and 2-chlorobenzeneamine (0.1 mmol, 12.8 mg) were dissolved in methanol (10 ml). The mixture was stirred at room temperature for 1 h and then filtered. After allowing the filtrate to stand in air for 7 d, yellow block-shaped crystals of the title compound were formed by slow evaporation of the solvent. The crystals were collected, washed with methanol and dried in a vacuum desiccator using anhydrous CaCl<sub>2</sub> (yield 52%).

### Refinement

All H atoms were placed in geometrically idealized positions and constrained to ride on their parent atoms, with C—H = 0.93 Å, O—H = 0.82 Å and  $U_{\text{iso}}(\text{H}) = 1.2 U_{\text{eq}}(\text{C}, \text{O})$ .

### Figures

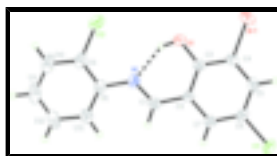


Fig. 1. The structure of the title compound with 30% probability ellipsoids. H atoms are shown as spheres of arbitrary radii. The dashed line represents a hydrogen bond.

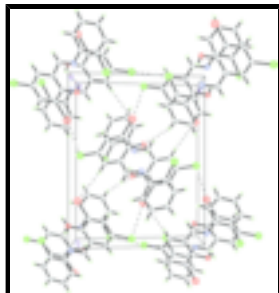


Fig. 2. The crystal packing of the title compound viewed along the *a* axis. Halogen-halogen interactions are shown as dashed lines.

## 2-Bromo-4-chloro-6-[(*E*)-(2-chlorophenyl)iminomethyl]phenol

### Crystal data

$C_{13}H_8BrCl_2NO$	$F_{000} = 680$
$M_r = 345.01$	$D_x = 1.755 \text{ Mg m}^{-3}$
Monoclinic, $P2_1/c$	Mo $K\alpha$ radiation
Hall symbol: -P 2ybc	$\lambda = 0.71073 \text{ \AA}$
$a = 8.4299 (10) \text{ \AA}$	Cell parameters from 2353 reflections
$b = 14.0115 (16) \text{ \AA}$	$\theta = 2.4\text{--}25.2^\circ$
$c = 11.4194 (14) \text{ \AA}$	$\mu = 3.54 \text{ mm}^{-1}$
$\beta = 104.5120 (10)^\circ$	$T = 298 \text{ K}$
$V = 1305.8 (3) \text{ \AA}^3$	Block, yellow
$Z = 4$	$0.45 \times 0.38 \times 0.36 \text{ mm}$

### Data collection

Siemens SMART CCD area-detector diffractometer	2295 independent reflections
Radiation source: fine-focus sealed tube	1726 reflections with $I > 2\sigma(I)$
Monochromator: graphite	$R_{\text{int}} = 0.038$
$T = 298 \text{ K}$	$\theta_{\text{max}} = 25.0^\circ$
$\varphi$ and $\omega$ scans	$\theta_{\text{min}} = 2.4^\circ$
Absorption correction: multi-scan (SADABS; Siemens, 1996)	$h = -10 \rightarrow 9$
$T_{\text{min}} = 0.230$ , $T_{\text{max}} = 0.279$	$k = -14 \rightarrow 16$
6450 measured reflections	$l = -13 \rightarrow 13$

### Refinement

Refinement on $F^2$	Hydrogen site location: inferred from neighbouring sites
Least-squares matrix: full	H-atom parameters constrained
$R[F^2 > 2\sigma(F^2)] = 0.032$	$w = 1/[\sigma^2(F_o^2) + (0.0256P)^2 + 0.9771P]$
$wR(F^2) = 0.077$	where $P = (F_o^2 + 2F_c^2)/3$
$S = 1.03$	$(\Delta/\sigma)_{\text{max}} = 0.001$
	$\Delta\rho_{\text{max}} = 0.39 \text{ e \AA}^{-3}$

2295 reflections

$$\Delta\rho_{\min} = -0.57 \text{ e } \text{\AA}^{-3}$$

164 parameters

Extinction correction: SHELXL97 (Sheldrick, 2008),

$$F_c^* = kFc[1+0.001xFc^2\lambda^3/\sin(2\theta)]^{-1/4}$$

Primary atom site location: structure-invariant direct methods

Extinction coefficient: 0.0127 (8)

Secondary atom site location: difference Fourier map

### Special details

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

### Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
Br1	0.54350 (5)	0.75675 (2)	0.54203 (4)	0.06096 (18)
Cl1	0.70198 (11)	0.53128 (7)	0.95100 (8)	0.0596 (3)
Cl2	-0.01427 (12)	0.49417 (8)	0.22925 (8)	0.0767 (3)
N1	0.1581 (3)	0.43674 (18)	0.4723 (2)	0.0434 (7)
O1	0.3040 (3)	0.59685 (15)	0.45663 (19)	0.0500 (6)
H1	0.2386	0.5532	0.4344	0.075*
C1	0.2481 (4)	0.4266 (2)	0.5798 (3)	0.0436 (8)
H1A	0.2359	0.3725	0.6238	0.052*
C2	0.3687 (4)	0.4978 (2)	0.6346 (3)	0.0381 (7)
C3	0.3916 (4)	0.5800 (2)	0.5697 (3)	0.0387 (7)
C4	0.5108 (4)	0.6456 (2)	0.6259 (3)	0.0420 (8)
C5	0.6058 (4)	0.6305 (2)	0.7418 (3)	0.0461 (8)
H5	0.6860	0.6744	0.7777	0.055*
C6	0.5809 (4)	0.5499 (2)	0.8042 (3)	0.0453 (8)
C7	0.4644 (4)	0.4838 (2)	0.7520 (3)	0.0436 (8)
H7	0.4492	0.4296	0.7950	0.052*
C8	0.0439 (4)	0.3675 (2)	0.4143 (3)	0.0457 (8)
C9	-0.0435 (4)	0.3866 (3)	0.2963 (3)	0.0528 (9)
C10	-0.1548 (5)	0.3222 (4)	0.2313 (4)	0.0741 (13)
H10	-0.2122	0.3362	0.1524	0.089*
C11	-0.1804 (5)	0.2369 (3)	0.2839 (5)	0.0814 (14)
H11	-0.2550	0.1931	0.2403	0.098*
C12	-0.0967 (5)	0.2165 (3)	0.3997 (5)	0.0751 (12)
H12	-0.1145	0.1589	0.4349	0.090*
C13	0.0149 (5)	0.2818 (3)	0.4650 (4)	0.0622 (10)
H13	0.0711	0.2677	0.5440	0.075*

## supplementary materials

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### Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
Br1	0.0721 (3)	0.0385 (2)	0.0755 (3)	-0.00233 (18)	0.0246 (2)	0.00929 (18)
Cl1	0.0696 (6)	0.0611 (6)	0.0377 (5)	0.0104 (5)	-0.0059 (4)	-0.0064 (4)
Cl2	0.0649 (6)	0.1111 (9)	0.0489 (6)	-0.0086 (6)	0.0043 (5)	0.0204 (6)
N1	0.0443 (15)	0.0433 (15)	0.0400 (16)	-0.0012 (12)	0.0058 (13)	-0.0021 (13)
O1	0.0568 (14)	0.0463 (13)	0.0419 (14)	-0.0019 (11)	0.0033 (11)	0.0102 (11)
C1	0.053 (2)	0.0336 (17)	0.045 (2)	0.0022 (15)	0.0139 (17)	0.0005 (15)
C2	0.0456 (18)	0.0347 (16)	0.0339 (16)	0.0025 (14)	0.0100 (14)	-0.0008 (14)
C3	0.0413 (18)	0.0372 (17)	0.0376 (18)	0.0075 (14)	0.0100 (15)	0.0012 (14)
C4	0.0472 (19)	0.0302 (16)	0.051 (2)	0.0046 (14)	0.0159 (16)	0.0037 (14)
C5	0.0457 (19)	0.0396 (18)	0.051 (2)	0.0001 (15)	0.0078 (17)	-0.0093 (16)
C6	0.052 (2)	0.0444 (19)	0.0365 (18)	0.0114 (16)	0.0050 (15)	-0.0022 (15)
C7	0.056 (2)	0.0365 (17)	0.0375 (18)	0.0046 (16)	0.0096 (16)	0.0017 (15)
C8	0.0420 (18)	0.048 (2)	0.047 (2)	0.0008 (15)	0.0108 (16)	-0.0116 (16)
C9	0.042 (2)	0.073 (2)	0.043 (2)	0.0002 (18)	0.0108 (16)	-0.0120 (18)
C10	0.057 (2)	0.111 (4)	0.052 (2)	-0.013 (2)	0.011 (2)	-0.029 (3)
C11	0.066 (3)	0.088 (3)	0.091 (4)	-0.021 (2)	0.020 (3)	-0.052 (3)
C12	0.073 (3)	0.055 (2)	0.099 (4)	-0.018 (2)	0.023 (3)	-0.022 (3)
C13	0.065 (2)	0.051 (2)	0.065 (3)	-0.0089 (19)	0.006 (2)	-0.0048 (19)

### Geometric parameters ( $\text{\AA}$ , $^\circ$ )

Br1—C4	1.885 (3)	C5—H5	0.9300
Cl1—C6	1.750 (3)	C6—C7	1.373 (5)
Cl2—C9	1.736 (4)	C7—H7	0.9300
N1—C1	1.279 (4)	C8—C13	1.381 (5)
N1—C8	1.409 (4)	C8—C9	1.390 (5)
O1—C3	1.338 (3)	C9—C10	1.377 (5)
O1—H1	0.8200	C10—C11	1.379 (6)
C1—C2	1.449 (4)	C10—H10	0.9300
C1—H1A	0.9300	C11—C12	1.364 (7)
C2—C7	1.394 (4)	C11—H11	0.9300
C2—C3	1.409 (4)	C12—C13	1.388 (5)
C3—C4	1.394 (4)	C12—H12	0.9300
C4—C5	1.381 (4)	C13—H13	0.9300
C5—C6	1.378 (4)		
C1—N1—C8	123.1 (3)	C6—C7—H7	120.0
C3—O1—H1	109.5	C2—C7—H7	120.0
N1—C1—C2	121.4 (3)	C13—C8—C9	117.8 (3)
N1—C1—H1A	119.3	C13—C8—N1	125.0 (3)
C2—C1—H1A	119.3	C9—C8—N1	117.2 (3)
C7—C2—C3	119.9 (3)	C10—C9—C8	121.4 (4)
C7—C2—C1	119.5 (3)	C10—C9—C12	118.9 (3)
C3—C2—C1	120.6 (3)	C8—C9—C12	119.7 (3)
O1—C3—C4	119.2 (3)	C9—C10—C11	119.5 (4)

O1—C3—C2	122.4 (3)	C9—C10—H10	120.3
C4—C3—C2	118.3 (3)	C11—C10—H10	120.3
C5—C4—C3	121.2 (3)	C12—C11—C10	120.3 (4)
C5—C4—Br1	119.4 (2)	C12—C11—H11	119.8
C3—C4—Br1	119.4 (2)	C10—C11—H11	119.8
C6—C5—C4	119.6 (3)	C11—C12—C13	119.9 (4)
C6—C5—H5	120.2	C11—C12—H12	120.1
C4—C5—H5	120.2	C13—C12—H12	120.1
C7—C6—C5	120.9 (3)	C8—C13—C12	121.1 (4)
C7—C6—C11	119.8 (3)	C8—C13—H13	119.5
C5—C6—C11	119.3 (3)	C12—C13—H13	119.5
C6—C7—C2	120.1 (3)		
C8—N1—C1—C2	177.5 (3)	C11—C6—C7—C2	-179.4 (2)
N1—C1—C2—C7	-179.9 (3)	C3—C2—C7—C6	-0.2 (5)
N1—C1—C2—C3	-0.8 (5)	C1—C2—C7—C6	178.9 (3)
C7—C2—C3—O1	179.6 (3)	C1—N1—C8—C13	0.1 (5)
C1—C2—C3—O1	0.6 (4)	C1—N1—C8—C9	-178.6 (3)
C7—C2—C3—C4	0.0 (4)	C13—C8—C9—C10	-0.4 (5)
C1—C2—C3—C4	-179.1 (3)	N1—C8—C9—C10	178.4 (3)
O1—C3—C4—C5	-179.1 (3)	C13—C8—C9—C12	178.9 (3)
C2—C3—C4—C5	0.6 (4)	N1—C8—C9—C12	-2.3 (4)
O1—C3—C4—Br1	0.8 (4)	C8—C9—C10—C11	0.0 (6)
C2—C3—C4—Br1	-179.6 (2)	C12—C9—C10—C11	-179.3 (3)
C3—C4—C5—C6	-1.0 (5)	C9—C10—C11—C12	0.3 (6)
Br1—C4—C5—C6	179.2 (2)	C10—C11—C12—C13	-0.1 (7)
C4—C5—C6—C7	0.8 (5)	C9—C8—C13—C12	0.6 (6)
C4—C5—C6—C11	180.0 (2)	N1—C8—C13—C12	-178.1 (3)
C5—C6—C7—C2	-0.2 (5)	C11—C12—C13—C8	-0.4 (6)

Hydrogen-bond geometry ( $\text{\AA}$ ,  $^\circ$ )

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
O1—H1 $\cdots$ N1	0.82	1.86	2.586 (3)	147
C11—H11 $\cdots$ O1 <sup>i</sup>	0.93	2.56	3.324 (5)	139

Symmetry codes: (i)  $-x, y-1/2, -z+1/2$ .

Fig. 1

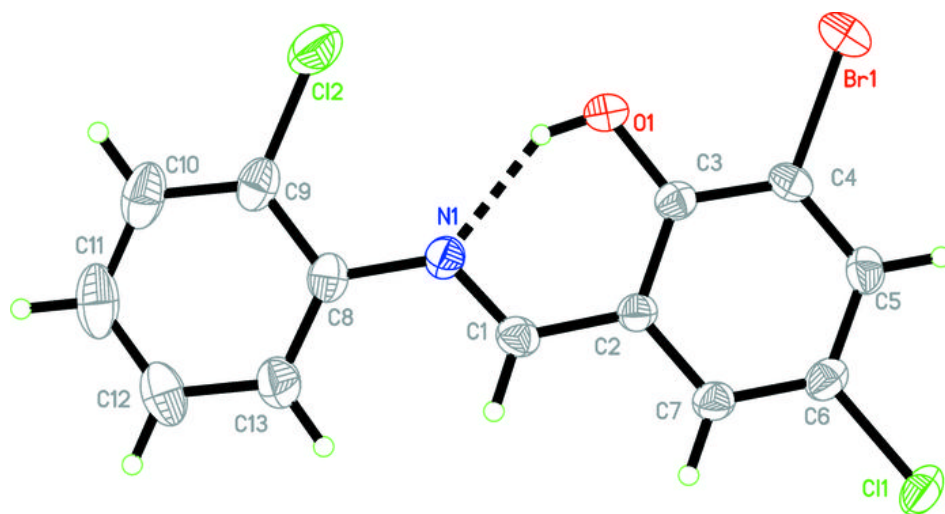


Fig. 2

