

Acta Crystallographica Section E

Structure Reports

Online

ISSN 1600-5368

Tetraaquabis(3,5-di-4-pyridyl-1,2,4-triazolato- κ N)cobalt(II) dihydrate

Lin Yi Dong

School of Pharmacy, Tianjin Medical University, Tianjin 300070, People's Republic of China

Correspondence e-mail: pass2009_good@126.com

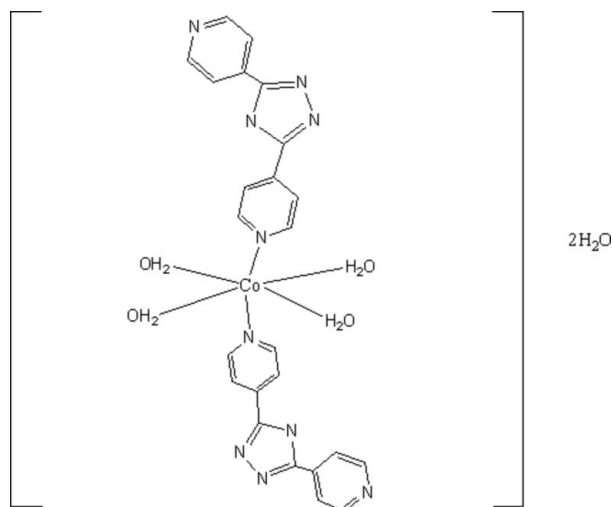
Received 26 March 2009; accepted 31 March 2009

 Key indicators: single-crystal X-ray study; $T = 293$ K; mean $\sigma(\text{C}-\text{C}) = 0.005$ Å; R factor = 0.054; wR factor = 0.108; data-to-parameter ratio = 10.0.

The Co^{II} atom in the title compound, $[\text{Co}(\text{C}_{12}\text{H}_8\text{N}_5)_2(\text{H}_2\text{O})_4] \cdot 2\text{H}_2\text{O}$, lies on a center of inversion and is bonded to two N -heterocycles and to four water molecules in a slightly distorted octahedral coordination. The coordinated and lattice water molecules interact with the N -heterocycles through $\text{O}-\text{H} \cdots \text{N}$ hydrogen bonds, generating a three-dimensional supramolecular architecture.

Related literature

For magnetic studies of transition metal complexes with 1,2,4-triazole derivatives, see: Haasnoot (2000). For the potential applications of complexes containing substituted 1,2,4-triazole ligands with spin-crossover properties in molecular-based memory devices, displays and optical switches, see: Kahn & Martinez (1998). For 3,5-di(4-pyridine)-1,2,4-triazole, see: Zhang *et al.* (2006); Sreenivasulu & Vittal (2004). For the structure of water, see: Tajkhorshid *et al.* (2002). For the synthesis, see: Basu & Dutta (1964). For a trinuclear water cluster, see: König (1944).



Experimental

Crystal data

 $[\text{Co}(\text{C}_{12}\text{H}_8\text{N}_5)_2(\text{H}_2\text{O})_4] \cdot 2\text{H}_2\text{O}$
 $M_r = 611.49$

 Monoclinic, $P2_1/c$
 $a = 7.3660$ (15) Å

 $b = 15.654$ (3) Å

 $c = 11.857$ (2) Å

 $\beta = 107.34$ (3)°

 $V = 1305.1$ (5) Å³
 $Z = 2$

 Mo $K\alpha$ radiation

 $\mu = 0.72$ mm⁻¹
 $T = 293$ K

 $0.40 \times 0.20 \times 0.12$ mm

Data collection

Bruker SMART CCD area-detector diffractometer

Absorption correction: multi-scan

(SADABS; Sheldrick, 1996)

 $T_{\text{min}} = 0.842$, $T_{\text{max}} = 0.917$

11054 measured reflections

2423 independent reflections

 2009 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.065$

Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.054$
 $wR(F^2) = 0.108$
 $S = 1.07$

2420 reflections

243 parameters

H atoms treated by a mixture of independent and constrained refinement

 $\Delta\rho_{\text{max}} = 0.29$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.41$ e Å⁻³

Table 1

Selected geometric parameters (Å, °).

Co1—O1	2.100 (2)	Co1—N1	2.134 (3)
Co1—O2	2.126 (2)		

Table 2

Hydrogen-bond geometry (Å, °).

$D-\text{H} \cdots A$	$D-\text{H}$	$\text{H} \cdots A$	$D \cdots A$	$D-\text{H} \cdots A$
O1—H1A \cdots N3 ⁱⁱ	0.851 (10)	2.42 (3)	3.070 (4)	134 (3)
O1—H1A \cdots N4 ⁱⁱ	0.851 (10)	1.966 (12)	2.803 (4)	167 (4)
O1—H1B \cdots O3 ⁱⁱⁱ	0.852 (10)	1.99 (2)	2.791 (4)	155 (3)
O2—H2A \cdots O3	0.851 (10)	1.973 (14)	2.801 (4)	164 (3)
O2—H2B \cdots N3 ^{iv}	0.850 (10)	1.973 (19)	2.792 (4)	161 (5)
O2—H2B \cdots N4 ^{iv}	0.850 (10)	2.60 (4)	3.220 (4)	130 (4)
O3—H3A \cdots N2 ^v	0.852 (10)	2.077 (12)	2.926 (4)	174 (4)
O3—H3B \cdots N5 ^{vi}	0.849 (10)	1.950 (14)	2.786 (4)	168 (5)

 Symmetry codes: (ii) $x, -y + \frac{3}{2}, z - \frac{1}{2}$; (iii) $-x, -y + 1, -z$; (iv) $-x + 1, y - \frac{1}{2}, -z + \frac{1}{2}$; (v) $x, -y + \frac{3}{2}, z + \frac{1}{2}$; (vi) $-x, -y + 2, -z$.

Data collection: SMART (Bruker, 1997); cell refinement: SAINT (Bruker, 1997); data reduction: SAINT; program(s) used to solve structure: SHELXL97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: DIAMOND (Brandenburg, 1999); software used to prepare material for publication: publCIF (Westrip, 2009).

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: NG2566).

References

- Basu, U. P. & Dutta, S. (1964). *J. Org. Chem.* **30**, 3562–3564.
- Brandenburg, K. (1999). *DIAMOND*. Crystal Impact GbR, Bonn, Germany.
- Bruker (1997). *SAINTE* and *SMART*. Bruker AXS Inc., Madison, Wisconsin, USA.
- Haasnoot, J. G. (2000). *Coord. Chem. Rev.* **200–202**, 131–185.
- Kahn, O. & Martinez, C. J. (1998). *Science*, **279**, 44–48.
- König, H. A. Z. (1944). *Z. Kristallogr.* **105**, 279–286.
- Sheldrick, G. M. (1996). *SADABS*. University of Göttingen, Germany.
- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
- Sreenivasulu, B. & Vittal, J. J. (2004). *Angew. Chem. Int. Ed.* **43**, 5769–5772.
- Tajkhorshid, E., Nollert, P., Jensen, M., Miercke, L. J. W., Connell, J. O., Stroud, R. M. & Schulten, K. (2002). *Science*, **296**, 525–530.
- Westrip (2009). *publCIF*. In preparation.
- Zhang, J. P., Lin, Y. Y., Huang, X. C. & Chen, X. M. (2006). *Cryst. Growth Des.* **6**, 519–523.

supplementary materials

Acta Cryst. (2009). E65, m487-m488 [doi:10.1107/S1600536809011982]

Tetraaquabis(3,5-di-4-pyridyl-1,2,4-triazolato- κN)cobalt(II) dihydrate

L. Y. Dong

Comment

Transition metal complexes with 1,2,4-triazole derivatives as ligands are of great interest as they are the subject of magnetic studies (Haasnoot, 2000). Some complexes containing substituted 1,2,4-triazole ligands have spin-crossover properties, which could be used in molecular-based memory devices, displays and optical switches (Kahn & Martinez, 1998). The ligand 3,5-di(4-pyridine)-1,2,4-triazole (*L*) is of special interest as it contains multi-dentate donor atoms and shows diverse coordination modes. Especially only a few examples about the coordination chemistry of *L* are reported. Some unusual coordination modes of *L* also have been reported forming interesting supramolecular isomerism systems (Zhang *et al.*, 2006). On the other hand, water is quite important for our common life (Tajkhorshid *et al.*, 2002). It has been the focus of intense research interests for their unusual properties in biological system and also plays an important role in biological self-assembly processes (Sreenivasulu & Vittal, 2004).

In this work, we synthesized a new compound $[\text{Co}(L)_2(\text{H}_2\text{O})_4](\text{H}_2\text{O})_2$ (**1**) (*L* = 3,5-di(4-pyridine)-1,2,4-triazole). **1** is composed of one cobalt(II) cation, two *L* ligand, four coordinated and two lattice water molecules. The cobalt(II) cation is six-coordinated in the octahedral geometry. The equatorial site of Cobalt cation is occupied by four aqua molecules while the axial site is occupied by two nitrogen atoms of two mono-dentate *L* ligands. The mono-dentate coordination mode of *L* is different from previously reported di-, tri- or tetra-dentate coordination modes of *L*.

O1, O2 from coordination water molecules and O3 from lattice water molecules generate strong intermolecular hydrogen bondings and construct trinuclear water clusters, in which O3 acts as the hydrogen acceptors and O1, O2 act as hydrogen bonding donors. The hydrogen-bonding distances are 2.791 (4) Å (O1—H1B \cdots O3) and 2.801 (4) Å (O2—H2A \cdots O3), respectively. The average O \cdots O distance is 2.796 (4) Å, which is similar to that (2.75 Å) in the structure of ice (König, 1944).

strong N—H \cdots O hydrogen bonds generated from water molecules and nitrogen atoms of pyridine or triazole groups are also observed resulting in the three-dimensional supramolecular network (Table 2). π - π stacking interactions between two neighboring triazole groups further consolidating the architecture centroid-centroid distance 3.677 (4) Å]

Perspective drawing with the atomic numbering scheme is illustrated in figure 1. Selected geometric parameters (Å, °) for **1** are listed in table 1. Selected hydrogen-bonding geometric parameters (Å, °) for **1** are listed in table 2. The trinuclear water clusters, corresponding N—H \cdots O hydrogen bonds and π - π stacking are shown in figure 2. The three-dimensional supramolecular packing architecture of (**1**) is shown in figure 3.

Experimental

The ligand was prepared according to the previous literature (Basu & Dutta, (1964)). $[\text{Co}(L)_2(\text{H}_2\text{O})_4](\text{H}_2\text{O})$ (**1**) (*L* = 3,5-di(4-pyridine)-1,2,4-triazole) was prepared under the hydrothermal conditions. $[\text{Co}(\text{ClO}_4)_2] \cdot 6\text{H}_2\text{O}$ (0.2 mmol), *L* (0.2 mmol) and 18 ml water was added to a 25 ml reaction vessel. the reaction vessel was then sealed and subsequently placed in an oven for 140 h at 160°C. well shaped red block crystals were obtained and washed with ethanol.

Refinement

The carbon-bound H atoms were positioned geometrically and were allowed to ride on their parent C atoms. The water H atoms were located from a difference density map and were refined with distance restraints of O—H = 0.85±0.01 Å.

Figures

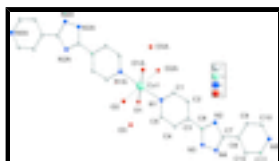


Fig. 1. The molecular structure and atom-labeling scheme of (I).

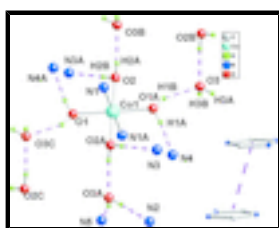


Fig. 2. The trinuclear water clusters stabilizing the packing structure of 1.

Tetraquabis(3,5-di-4-pyridyl-1,2,4-triazolato-κN)cobalt(II) dihydrate

Crystal data

[Co(C₁₂H₈N₅)₂(H₂O)₄]·2(H₂O)

M_r = 611.49

Monoclinic, *P*2₁/*c*

Hall symbol: -*P* 2ybc

a = 7.3660 (15) Å

b = 15.654 (3) Å

c = 11.857 (2) Å

β = 107.34 (3)°

V = 1305.1 (5) Å³

Z = 2

F(000) = 634

D_x = 1.556 Mg m⁻³

Mo *K*α radiation, λ = 0.71073 Å

Cell parameters from 556 reflections

θ = 1.5–25.5°

μ = 0.72 mm⁻¹

T = 293 K

Block, red

0.40 × 0.20 × 0.12 mm

Data collection

Bruker SMART CCD area-detector diffractometer

Radiation source: fine-focus sealed tube graphite

φ and ω scans

Absorption correction: multi-scan (*SADABS*; Sheldrick, 1996)

T_{min} = 0.842, *T_{max}* = 0.917

11054 measured reflections

2423 independent reflections

2009 reflections with *I* > 2σ(*I*)

R_{int} = 0.065

θ_{max} = 25.5°, θ_{min} = 3.2°

h = -8→8

k = -18→18

l = -14→14

Refinement

Refinement on F^2	0 restraints
Least-squares matrix: full	H atoms treated by a mixture of independent and constrained refinement
$R[F^2 > 2\sigma(F^2)] = 0.054$	$w = 1/[\sigma^2(F_o^2) + (0.0298P)^2 + 0.2298P]$
$wR(F^2) = 0.108$	where $P = (F_o^2 + 2F_c^2)/3$
$S = 1.07$	$(\Delta/\sigma)_{\max} < 0.001$
2420 reflections	$\Delta\rho_{\max} = 0.29 \text{ e } \text{\AA}^{-3}$
243 parameters	$\Delta\rho_{\min} = -0.41 \text{ e } \text{\AA}^{-3}$

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
C1	0.3827 (5)	0.6826 (2)	-0.0776 (3)	0.0306 (8)
C2	0.3536 (5)	0.7690 (2)	-0.0713 (3)	0.0299 (8)
C3	0.3919 (4)	0.80864 (19)	0.0377 (3)	0.0218 (7)
C4	0.4634 (5)	0.7565 (2)	0.1360 (3)	0.0290 (8)
C5	0.4886 (5)	0.6711 (2)	0.1224 (3)	0.0294 (8)
C6	0.3518 (5)	0.8987 (2)	0.0514 (3)	0.0229 (7)
C7	0.2496 (5)	1.02433 (19)	0.0220 (3)	0.0241 (7)
C8	0.1690 (5)	1.1049 (2)	-0.0324 (3)	0.0244 (7)
C9	0.1117 (5)	1.1170 (2)	-0.1528 (3)	0.0330 (9)
C10	0.0440 (6)	1.1952 (2)	-0.1990 (3)	0.0368 (9)
C11	0.0861 (6)	1.2510 (3)	-0.0181 (4)	0.0450 (11)
C12	0.1540 (6)	1.1751 (2)	0.0357 (4)	0.0403 (10)
Co1	0.5000	0.5000	0.0000	0.02203 (19)
H1	0.363 (5)	0.657 (2)	-0.147 (3)	0.033 (10)*
H2	0.313 (5)	0.799 (3)	-0.138 (3)	0.046 (12)*
H4	0.495 (5)	0.781 (2)	0.209 (3)	0.034 (10)*
H5	0.532 (5)	0.637 (2)	0.187 (3)	0.036 (10)*
H9	0.118 (5)	1.075 (2)	-0.199 (3)	0.028 (10)*
H10	0.013 (6)	1.203 (3)	-0.278 (4)	0.052 (13)*
H11	0.078 (6)	1.299 (3)	0.026 (4)	0.057 (13)*

supplementary materials

H12	0.188 (6)	1.170 (3)	0.116 (4)	0.051 (12)*
H1A	0.261 (5)	0.484 (3)	-0.2116 (18)	0.069 (16)*
H2A	0.276 (4)	0.5057 (16)	0.135 (3)	0.031 (10)*
H3A	0.163 (5)	0.571 (2)	0.2816 (10)	0.041 (12)*
H1B	0.143 (3)	0.475 (3)	-0.140 (3)	0.077 (17)*
H2B	0.416 (6)	0.446 (3)	0.1861 (19)	0.084 (18)*
H3B	0.081 (6)	0.6269 (11)	0.194 (4)	0.069 (16)*
N1	0.4483 (4)	0.63277 (16)	0.0175 (2)	0.0262 (6)
N2	0.2624 (4)	0.95193 (16)	-0.0378 (2)	0.0229 (6)
N3	0.3929 (4)	0.93481 (17)	0.1575 (2)	0.0295 (7)
N4	0.3253 (4)	1.01645 (17)	0.1389 (2)	0.0308 (7)
N5	0.0298 (4)	1.26290 (18)	-0.1346 (3)	0.0371 (8)
O1	0.2504 (3)	0.49368 (17)	-0.1431 (2)	0.0334 (6)
O2	0.3487 (4)	0.46767 (16)	0.1212 (2)	0.0302 (6)
O3	0.1157 (4)	0.57516 (16)	0.2069 (2)	0.0329 (6)

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
C1	0.047 (2)	0.0220 (18)	0.0201 (18)	0.0038 (16)	0.0054 (16)	-0.0033 (15)
C2	0.044 (2)	0.0184 (17)	0.0245 (19)	0.0057 (15)	0.0057 (16)	0.0020 (15)
C3	0.0241 (18)	0.0184 (16)	0.0231 (17)	0.0017 (13)	0.0074 (14)	-0.0003 (13)
C4	0.043 (2)	0.0239 (18)	0.0188 (18)	0.0052 (16)	0.0075 (16)	-0.0032 (15)
C5	0.043 (2)	0.0200 (18)	0.0232 (19)	0.0050 (15)	0.0065 (16)	0.0043 (14)
C6	0.0277 (18)	0.0182 (16)	0.0224 (17)	-0.0006 (14)	0.0066 (14)	0.0000 (13)
C7	0.0297 (18)	0.0194 (17)	0.0227 (17)	-0.0001 (13)	0.0072 (14)	0.0002 (13)
C8	0.0247 (18)	0.0211 (17)	0.0269 (18)	-0.0023 (14)	0.0067 (14)	-0.0009 (14)
C9	0.044 (2)	0.0231 (19)	0.028 (2)	0.0046 (16)	0.0043 (17)	-0.0031 (16)
C10	0.044 (2)	0.033 (2)	0.027 (2)	0.0034 (17)	0.0023 (18)	0.0073 (17)
C11	0.067 (3)	0.024 (2)	0.045 (3)	0.013 (2)	0.017 (2)	-0.0019 (18)
C12	0.065 (3)	0.027 (2)	0.028 (2)	0.0129 (18)	0.011 (2)	0.0002 (16)
Co1	0.0282 (4)	0.0162 (3)	0.0206 (3)	0.0025 (3)	0.0055 (2)	0.0004 (3)
N1	0.0358 (17)	0.0170 (14)	0.0251 (15)	0.0032 (12)	0.0081 (12)	-0.0001 (11)
N2	0.0283 (15)	0.0154 (14)	0.0241 (15)	0.0011 (11)	0.0063 (12)	-0.0011 (11)
N3	0.0398 (18)	0.0213 (15)	0.0255 (16)	0.0099 (13)	0.0065 (13)	0.0017 (12)
N4	0.0461 (18)	0.0213 (16)	0.0221 (15)	0.0069 (13)	0.0057 (13)	-0.0013 (11)
N5	0.044 (2)	0.0223 (16)	0.044 (2)	0.0079 (14)	0.0115 (16)	0.0055 (14)
O1	0.0312 (14)	0.0408 (15)	0.0264 (14)	-0.0029 (13)	0.0058 (10)	-0.0009 (12)
O2	0.0346 (15)	0.0295 (13)	0.0296 (14)	0.0062 (11)	0.0142 (12)	0.0021 (11)
O3	0.0436 (16)	0.0239 (14)	0.0286 (15)	0.0017 (12)	0.0067 (12)	0.0018 (11)

Geometric parameters (\AA , $^\circ$)

C1—N1	1.337 (4)	C10—N5	1.328 (5)
C1—C2	1.376 (5)	C10—H10	0.90 (4)
C1—H1	0.89 (4)	C11—N5	1.331 (5)
C2—C3	1.385 (5)	C11—C12	1.370 (5)
C2—H2	0.89 (4)	C11—H11	0.93 (4)
C3—C4	1.392 (5)	C12—H12	0.92 (4)

C3—C6	1.459 (4)	Co1—O1 ⁱ	2.100 (2)
C4—C5	1.365 (5)	Co1—O1	2.100 (2)
C4—H4	0.91 (4)	Co1—O2 ⁱ	2.126 (2)
C5—N1	1.332 (4)	Co1—O2	2.126 (2)
C5—H5	0.91 (4)	Co1—N1 ⁱ	2.134 (3)
C6—N3	1.329 (4)	Co1—N1	2.134 (3)
C6—N2	1.354 (4)	N3—N4	1.365 (4)
C7—N4	1.336 (4)	O1—H1A	0.851 (10)
C7—N2	1.355 (4)	O1—H1B	0.852 (10)
C7—C8	1.459 (4)	O2—H2A	0.851 (10)
C8—C9	1.377 (5)	O2—H2B	0.850 (10)
C8—C12	1.387 (5)	O3—H3A	0.852 (10)
C9—C10	1.373 (5)	O3—H3B	0.849 (10)
C9—H9	0.87 (4)		
N1—C1—C2	123.4 (3)	C11—C12—C8	119.8 (4)
N1—C1—H1	116 (2)	C11—C12—H12	121 (3)
C2—C1—H1	121 (2)	C8—C12—H12	120 (3)
C1—C2—C3	120.0 (3)	O1 ⁱ —Co1—O1	180.0
C1—C2—H2	119 (3)	O1 ⁱ —Co1—O2 ⁱ	91.47 (10)
C3—C2—H2	121 (3)	O1—Co1—O2 ⁱ	88.53 (10)
C2—C3—C4	116.1 (3)	O1 ⁱ —Co1—O2	88.53 (10)
C2—C3—C6	123.0 (3)	O1—Co1—O2	91.47 (10)
C4—C3—C6	120.8 (3)	O2 ⁱ —Co1—O2	180.0
C5—C4—C3	120.4 (3)	O1 ⁱ —Co1—N1 ⁱ	89.19 (10)
C5—C4—H4	122 (2)	O1—Co1—N1 ⁱ	90.81 (10)
C3—C4—H4	118 (2)	O2 ⁱ —Co1—N1 ⁱ	91.23 (10)
N1—C5—C4	123.5 (3)	O2—Co1—N1 ⁱ	88.77 (10)
N1—C5—H5	116 (2)	O1 ⁱ —Co1—N1	90.81 (10)
C4—C5—H5	120 (2)	O1—Co1—N1	89.19 (10)
N3—C6—N2	113.4 (3)	O2 ⁱ —Co1—N1	88.77 (10)
N3—C6—C3	121.3 (3)	O2—Co1—N1	91.23 (10)
N2—C6—C3	125.1 (3)	N1 ⁱ —Co1—N1	180.0
N4—C7—N2	113.1 (3)	C5—N1—C1	116.7 (3)
N4—C7—C8	121.8 (3)	C5—N1—Co1	122.2 (2)
N2—C7—C8	125.0 (3)	C1—N1—Co1	121.0 (2)
C9—C8—C12	116.1 (3)	C6—N2—C7	101.5 (3)
C9—C8—C7	122.5 (3)	C6—N3—N4	106.0 (2)
C12—C8—C7	121.3 (3)	C7—N4—N3	105.8 (2)
C10—C9—C8	120.0 (3)	C10—N5—C11	115.6 (3)
C10—C9—H9	120 (2)	Co1—O1—H1A	118 (3)
C8—C9—H9	120 (2)	Co1—O1—H1B	126 (3)
N5—C10—C9	124.3 (4)	H1A—O1—H1B	109.1 (17)
N5—C10—H10	117 (3)	Co1—O2—H2A	117 (2)
C9—C10—H10	119 (3)	Co1—O2—H2B	115 (3)
N5—C11—C12	124.2 (4)	H2A—O2—H2B	109.4 (15)
N5—C11—H11	115 (3)	H3A—O3—H3B	106 (4)

supplementary materials

C12—C11—H11	121 (3)		
N1—C1—C2—C3	-0.3 (6)	C2—C1—N1—Co1	-177.5 (3)
C1—C2—C3—C4	1.4 (5)	O1 ⁱ —Co1—N1—C5	-36.4 (3)
C1—C2—C3—C6	-175.5 (3)	O1—Co1—N1—C5	143.6 (3)
C2—C3—C4—C5	-1.3 (5)	O2 ⁱ —Co1—N1—C5	-127.9 (3)
C6—C3—C4—C5	175.6 (3)	O2—Co1—N1—C5	52.1 (3)
C3—C4—C5—N1	0.1 (6)	N1 ⁱ —Co1—N1—C5	-122 (27)
C2—C3—C6—N3	-179.1 (3)	O1 ⁱ —Co1—N1—C1	140.0 (3)
C4—C3—C6—N3	4.2 (5)	O1—Co1—N1—C1	-40.0 (3)
C2—C3—C6—N2	4.7 (5)	O2 ⁱ —Co1—N1—C1	48.5 (3)
C4—C3—C6—N2	-172.0 (3)	O2—Co1—N1—C1	-131.5 (3)
N4—C7—C8—C9	171.9 (3)	N1 ⁱ —Co1—N1—C1	55 (27)
N2—C7—C8—C9	-5.8 (5)	N3—C6—N2—C7	-0.6 (4)
N4—C7—C8—C12	-5.0 (5)	C3—C6—N2—C7	175.9 (3)
N2—C7—C8—C12	177.3 (3)	N4—C7—N2—C6	0.0 (4)
C12—C8—C9—C10	-0.2 (6)	C8—C7—N2—C6	177.9 (3)
C7—C8—C9—C10	-177.4 (3)	N2—C6—N3—N4	0.8 (4)
C8—C9—C10—N5	0.5 (6)	C3—C6—N3—N4	-175.7 (3)
N5—C11—C12—C8	0.6 (7)	N2—C7—N4—N3	0.5 (4)
C9—C8—C12—C11	-0.3 (6)	C8—C7—N4—N3	-177.5 (3)
C7—C8—C12—C11	176.9 (4)	C6—N3—N4—C7	-0.8 (4)
C4—C5—N1—C1	1.1 (5)	C9—C10—N5—C11	-0.2 (6)
C4—C5—N1—Co1	177.6 (3)	C12—C11—N5—C10	-0.3 (6)
C2—C1—N1—C5	-1.0 (5)		

Symmetry codes: (i) $-x+1, -y+1, -z$.

Hydrogen-bond geometry ($\text{\AA}, ^\circ$)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
O1—H1A \cdots N3 ⁱⁱ	0.85 (1)	2.42 (3)	3.070 (4)	134 (3)
O1—H1A \cdots N4 ⁱⁱ	0.85 (1)	1.97 (1)	2.803 (4)	167 (4)
O1—H1B \cdots O3 ⁱⁱⁱ	0.85 (1)	1.99 (2)	2.791 (4)	155 (3)
O2—H2A \cdots O3	0.85 (1)	1.97 (1)	2.801 (4)	164 (3)
O2—H2B \cdots N3 ^{iv}	0.85 (1)	1.97 (2)	2.792 (4)	161 (5)
O2—H2B \cdots N4 ^{iv}	0.85 (1)	2.60 (4)	3.220 (4)	130 (4)
O3—H3A \cdots N2 ^v	0.85 (1)	2.08 (1)	2.926 (4)	174 (4)
O3—H3B \cdots N5 ^{vi}	0.85 (1)	1.95 (1)	2.786 (4)	168 (5)

Symmetry codes: (ii) $x, -y+3/2, z-1/2$; (iii) $-x, -y+1, -z$; (iv) $-x+1, y-1/2, -z+1/2$; (v) $x, -y+3/2, z+1/2$; (vi) $-x, -y+2, -z$.

Fig. 1

