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N-(9H-Fluoren-9-ylidene)-4-methyl-aniline

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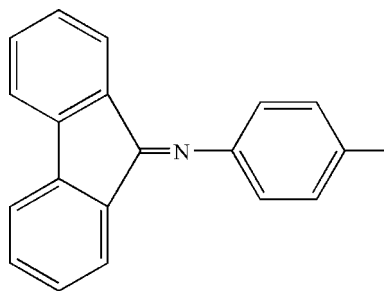
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Key indicators: single-crystal X-ray study; $T = 294$ K; mean $\sigma(\text{C}-\text{C}) = 0.003$ Å; R factor = 0.042; wR factor = 0.115; data-to-parameter ratio = 14.1.

In the title compound, $\text{C}_{20}\text{H}_{15}\text{N}$, the fluorene unit is essentially planar [r.m.s. deviation 0.0334 Å] and the benzene ring bound to the imine N atom bears a methyl group which is nearly coplanar [dihedral angle 0.5 (1)°]. The dihedral angle between the substituted benzene ring and the 9H-fluoren-9-imine unit is 71.1 (3)°. Intermolecular π - π interactions between the benzene rings of adjacent fluorene units [centroid-centroid distance 3.8081 (13) Å] are present in the crystal structure, resulting in a one-dimensional supramolecular architecture.

Related literature

For the properties of Schiff bases, see: Xu *et al.* (2007); Tanaka *et al.* (2006). For the properties of fluorene derivatives, see: Saragi *et al.* (2004). For related structures, see: Glagovich *et al.* (2004); Peters *et al.* (1998); Pierre *et al.* (1997).



Experimental

Crystal data

$\text{C}_{20}\text{H}_{15}\text{N}$	$V = 1460.9 (5) \text{ \AA}^3$
$M_r = 269.33$	$Z = 4$
Monoclinic, $P2_1/n$	Mo $K\alpha$ radiation
$a = 5.6423 (10) \text{ \AA}$	$\mu = 0.07 \text{ mm}^{-1}$
$b = 12.187 (2) \text{ \AA}$	$T = 294 \text{ K}$
$c = 21.310 (4) \text{ \AA}$	$0.35 \times 0.17 \times 0.09 \text{ mm}$
$\beta = 94.441 (2)^\circ$	

Data collection

Bruker SMART APEX CCD area-detector diffractometer	10793 measured reflections
Absorption correction: multi-scan (SADABS; Sheldrick, 1996)	2711 independent reflections
$T_{\min} = 0.976$, $T_{\max} = 0.994$	1779 reflections with $I > 2\sigma(I)$
	$R_{\text{int}} = 0.042$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.042$	192 parameters
$wR(F^2) = 0.115$	H-atom parameters constrained
$S = 1.01$	$\Delta\rho_{\text{max}} = 0.13 \text{ e \AA}^{-3}$
2711 reflections	$\Delta\rho_{\text{min}} = -0.14 \text{ e \AA}^{-3}$

Data collection: SMART (Bruker, 2004); cell refinement: SAINT (Bruker, 2004); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: SHELXTL (Sheldrick, 2008); software used to prepare material for publication: SHELXTL.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: SI2179).

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supplementary materials

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N-(9*H*-Fluoren-9-ylidene)-4-methylaniline

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Comment

Schiff bases have received much attention during the past decades because of their strong coordination capability and diverse biological activities (Xu *et al.*, 2007; Tanaka *et al.*, 2006). In addition, fluorene derivatives have found many applications in chemistry, especially in the optoelectronic area (Saragi *et al.*, 2004). In view of these important properties, the crystal structure of the title compound has been determined.

In the title compound (Fig. 1), the C12—N1—C14 angle of 120.76 (15)° and the N1—C12 bond distance of 1.278 (2) Å are in close agreement with the similar *N*fluorenylideneaniline (Glagovich *et al.*, 2004; Peters *et al.*, 1998; Pierre *et al.*, 1997). The fluorene unit is essentially planar and the benzene ring bound to the imine N atom bears a methyl that is nearly coplanar. The dihedral angle between the substituent benzene ring and the 9*H*-fluoren-9-imine unit is 108.9 (3)°. Intermolecular $\pi\cdots\pi$ interactions between the benzene rings of adjacent fluorene units [centroid-centroid distance is 3.8081 (13) Å, the average perpendicular distance is 3.469 Å, the dihedral angle between the rings is 3.7°, symmetry code = $-1 + x, y, z$] are present in the crystal structure, resulting in a one-dimensional supramolecular architecture (Fig. 2).

Experimental

The title compound was obtained from the condensation reaction of 9-fluorenone and 4-methylaniline as described in literature (Glagovich *et al.*, 2004) and recrystallized from ethanol solution at room temperature to give the desired product as yellow crystals suitable for single-crystal X-ray diffraction.

Refinement

H atoms attached to C atoms of the title compound were placed in geometrically idealized positions and treated as riding with C—H distances constrained to 0.93–0.96 Å, and with $U_{\text{iso}}(\text{H})=1.2U_{\text{eq}}(\text{C})$ (1.5 U_{eq} for methyl H).

Figures

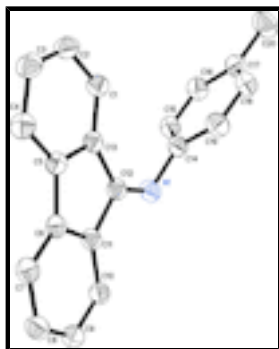


Fig. 1. The molecular structure of the title compound with displacement ellipsoids at the 30% probability level. H atoms are omitted for clarity.



Fig. 2. Partial view of the crystal packing showing the formation of the chain motif of molecules formed by the intermolecular $\pi \cdots \pi$ interactions. H atoms are omitted for clarity.

N-(9*H*-Fluoren-9-ylidene)-4-methylaniline

Crystal data

$C_{20}H_{15}N$	$F_{000} = 568$
$M_r = 269.33$	$D_x = 1.225 \text{ Mg m}^{-3}$
Monoclinic, $P2_1/n$	Mo $K\alpha$ radiation
$a = 5.6423 (10) \text{ \AA}$	$\lambda = 0.71073 \text{ \AA}$
$b = 12.187 (2) \text{ \AA}$	Cell parameters from 1911 reflections
$c = 21.310 (4) \text{ \AA}$	$\theta = 2.5\text{--}22.3^\circ$
$\beta = 94.441 (2)^\circ$	$\mu = 0.07 \text{ mm}^{-1}$
$V = 1460.9 (5) \text{ \AA}^3$	$T = 294 \text{ K}$
$Z = 4$	Block, yellow
	$0.35 \times 0.17 \times 0.09 \text{ mm}$

Data collection

Bruker SMART APEX CCD area-detector diffractometer	2711 independent reflections
Radiation source: fine-focus sealed tube	1779 reflections with $I > 2\sigma(I)$
Monochromator: graphite	$R_{\text{int}} = 0.042$
$T = 294 \text{ K}$	$\theta_{\text{max}} = 25.5^\circ$
phi and ω scans	$\theta_{\text{min}} = 2.5^\circ$
Absorption correction: multi-scan (SADABS; Sheldrick, 1996)	$h = -6 \rightarrow 6$
$T_{\text{min}} = 0.976$, $T_{\text{max}} = 0.994$	$k = -14 \rightarrow 14$
10793 measured reflections	$l = -25 \rightarrow 25$

Refinement

Refinement on F^2	Hydrogen site location: inferred from neighbouring sites
Least-squares matrix: full	H-atom parameters constrained
$R[F^2 > 2\sigma(F^2)] = 0.042$	$w = 1/[\sigma^2(F_o^2) + (0.052P)^2 + 0.1725P]$
$wR(F^2) = 0.115$	where $P = (F_o^2 + 2F_c^2)/3$
$S = 1.01$	$(\Delta/\sigma)_{\text{max}} < 0.001$
2711 reflections	$\Delta\rho_{\text{max}} = 0.13 \text{ e \AA}^{-3}$
192 parameters	$\Delta\rho_{\text{min}} = -0.14 \text{ e \AA}^{-3}$
Primary atom site location: structure-invariant direct methods	Extinction correction: SHELXL, $F_c^* = kFc[1 + 0.001xFc^2\lambda^3/\sin(2\theta)]^{-1/4}$
Secondary atom site location: difference Fourier map	Extinction coefficient: 0.0105 (19)

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes)

are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and

goodness of fit S are based on F^2 , conventional R -factors R are based

on F , with F set to zero for negative F^2 . The threshold expression of

$F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is

not relevant to the choice of reflections for refinement. R -factors based

on F^2 are statistically about twice as large as those based on F , and R -

factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
C1	0.0001 (3)	0.30235 (15)	0.13059 (9)	0.0507 (5)
H1	-0.0938	0.3086	0.0929	0.061*
C2	-0.0545 (4)	0.36053 (16)	0.18339 (10)	0.0582 (5)
H2	-0.1869	0.4062	0.1810	0.070*
C3	0.0836 (4)	0.35188 (16)	0.23918 (10)	0.0604 (5)
H3	0.0452	0.3931	0.2737	0.072*
C4	0.2787 (3)	0.28289 (15)	0.24478 (9)	0.0544 (5)
H4	0.3701	0.2762	0.2828	0.065*
C5	0.3348 (3)	0.22414 (13)	0.19248 (8)	0.0435 (4)
C6	0.5230 (3)	0.14217 (13)	0.18583 (8)	0.0433 (4)
C7	0.7076 (3)	0.10822 (14)	0.22712 (9)	0.0508 (5)
H7	0.7270	0.1372	0.2676	0.061*
C8	0.8643 (3)	0.03010 (15)	0.20743 (9)	0.0549 (5)
H8	0.9901	0.0065	0.2349	0.066*
C9	0.8350 (3)	-0.01288 (16)	0.14735 (9)	0.0557 (5)
H9	0.9409	-0.0657	0.1350	0.067*
C10	0.6513 (3)	0.02124 (14)	0.10530 (9)	0.0506 (5)
H10	0.6328	-0.0077	0.0648	0.061*
C11	0.4954 (3)	0.09953 (13)	0.12487 (8)	0.0432 (4)
C12	0.2980 (3)	0.15766 (13)	0.08907 (8)	0.0440 (4)

supplementary materials

C13	0.1986 (3)	0.23442 (13)	0.13519 (8)	0.0428 (4)
C14	0.0834 (3)	0.21043 (15)	-0.00479 (8)	0.0481 (5)
C15	-0.1301 (3)	0.17101 (16)	-0.03143 (9)	0.0544 (5)
H15	-0.1735	0.0986	-0.0246	0.065*
C16	-0.2793 (4)	0.23849 (16)	-0.06821 (9)	0.0598 (5)
H16	-0.4241	0.2110	-0.0852	0.072*
C17	-0.2197 (4)	0.34635 (17)	-0.08065 (9)	0.0579 (5)
C18	-0.0037 (4)	0.38372 (16)	-0.05457 (9)	0.0612 (5)
H18	0.0417	0.4554	-0.0625	0.073*
C19	0.1475 (4)	0.31784 (16)	-0.01699 (9)	0.0592 (5)
H19	0.2921	0.3454	0.0001	0.071*
C20	-0.3829 (4)	0.4195 (2)	-0.12150 (12)	0.0880 (8)
H20A	-0.3234	0.4263	-0.1623	0.132*
H20B	-0.3899	0.4907	-0.1024	0.132*
H20C	-0.5393	0.3880	-0.1257	0.132*
N1	0.2459 (3)	0.14086 (12)	0.03045 (7)	0.0521 (4)

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
C1	0.0542 (11)	0.0462 (10)	0.0518 (11)	0.0015 (9)	0.0048 (9)	0.0011 (9)
C2	0.0594 (12)	0.0508 (11)	0.0659 (14)	0.0066 (9)	0.0144 (11)	-0.0045 (10)
C3	0.0717 (14)	0.0513 (12)	0.0600 (13)	0.0007 (10)	0.0164 (11)	-0.0119 (10)
C4	0.0629 (12)	0.0494 (11)	0.0507 (11)	-0.0054 (9)	0.0037 (9)	-0.0071 (9)
C5	0.0486 (10)	0.0379 (9)	0.0443 (10)	-0.0061 (8)	0.0062 (8)	-0.0038 (8)
C6	0.0477 (10)	0.0395 (9)	0.0429 (10)	-0.0065 (8)	0.0040 (8)	0.0014 (8)
C7	0.0565 (11)	0.0494 (11)	0.0457 (11)	-0.0039 (9)	-0.0005 (9)	0.0003 (9)
C8	0.0497 (11)	0.0555 (12)	0.0587 (13)	0.0015 (9)	-0.0008 (10)	0.0083 (10)
C9	0.0585 (12)	0.0521 (11)	0.0578 (13)	0.0085 (9)	0.0123 (10)	0.0071 (10)
C10	0.0630 (12)	0.0441 (10)	0.0455 (11)	0.0034 (9)	0.0102 (9)	0.0023 (8)
C11	0.0497 (10)	0.0365 (9)	0.0438 (10)	-0.0017 (8)	0.0053 (8)	0.0022 (8)
C12	0.0509 (11)	0.0381 (9)	0.0434 (11)	-0.0027 (8)	0.0064 (8)	0.0007 (8)
C13	0.0494 (10)	0.0350 (9)	0.0446 (10)	-0.0043 (8)	0.0075 (8)	-0.0004 (8)
C14	0.0591 (12)	0.0486 (10)	0.0366 (10)	0.0046 (9)	0.0029 (9)	-0.0027 (8)
C15	0.0632 (12)	0.0498 (11)	0.0502 (11)	-0.0031 (10)	0.0038 (10)	0.0021 (9)
C16	0.0562 (12)	0.0647 (13)	0.0572 (12)	-0.0059 (10)	-0.0030 (10)	0.0009 (10)
C17	0.0627 (13)	0.0626 (13)	0.0478 (11)	0.0034 (10)	0.0006 (10)	0.0082 (10)
C18	0.0700 (14)	0.0521 (12)	0.0610 (13)	-0.0019 (10)	0.0020 (11)	0.0089 (10)
C19	0.0609 (12)	0.0554 (12)	0.0597 (13)	-0.0027 (10)	-0.0046 (10)	0.0022 (10)
C20	0.0844 (17)	0.0904 (18)	0.0862 (18)	0.0081 (14)	-0.0126 (14)	0.0261 (14)
N1	0.0638 (10)	0.0489 (9)	0.0431 (9)	0.0066 (8)	0.0002 (8)	-0.0012 (7)

Geometric parameters (\AA , $^\circ$)

C1—C2	1.385 (3)	C10—H10	0.9300
C1—C13	1.390 (2)	C11—C12	1.481 (2)
C1—H1	0.9300	C12—N1	1.278 (2)
C2—C3	1.374 (3)	C12—C13	1.497 (2)
C2—H2	0.9300	C14—C15	1.378 (2)

C3—C4	1.383 (3)	C14—C19	1.388 (3)
C3—H3	0.9300	C14—N1	1.420 (2)
C4—C5	1.382 (2)	C15—C16	1.378 (3)
C4—H4	0.9300	C15—H15	0.9300
C5—C13	1.397 (2)	C16—C17	1.387 (3)
C5—C6	1.473 (2)	C16—H16	0.9300
C6—C7	1.374 (2)	C17—C18	1.377 (3)
C6—C11	1.397 (2)	C17—C20	1.509 (3)
C7—C8	1.386 (3)	C18—C19	1.381 (3)
C7—H7	0.9300	C18—H18	0.9300
C8—C9	1.381 (3)	C19—H19	0.9300
C8—H8	0.9300	C20—H20A	0.9600
C9—C10	1.381 (2)	C20—H20B	0.9600
C9—H9	0.9300	C20—H20C	0.9600
C10—C11	1.384 (2)		
C2—C1—C13	118.45 (18)	C6—C11—C12	109.05 (15)
C2—C1—H1	120.8	N1—C12—C11	122.27 (16)
C13—C1—H1	120.8	N1—C12—C13	132.28 (16)
C3—C2—C1	121.13 (18)	C11—C12—C13	105.42 (14)
C3—C2—H2	119.4	C1—C13—C5	120.05 (16)
C1—C2—H2	119.4	C1—C13—C12	131.81 (16)
C2—C3—C4	121.03 (18)	C5—C13—C12	108.00 (15)
C2—C3—H3	119.5	C15—C14—C19	118.94 (17)
C4—C3—H3	119.5	C15—C14—N1	121.18 (17)
C5—C4—C3	118.41 (18)	C19—C14—N1	119.68 (17)
C5—C4—H4	120.8	C16—C15—C14	120.18 (18)
C3—C4—H4	120.8	C16—C15—H15	119.9
C4—C5—C13	120.91 (17)	C14—C15—H15	119.9
C4—C5—C6	129.85 (17)	C15—C16—C17	121.85 (19)
C13—C5—C6	109.18 (15)	C15—C16—H16	119.1
C7—C6—C11	120.53 (16)	C17—C16—H16	119.1
C7—C6—C5	131.23 (16)	C18—C17—C16	117.14 (18)
C11—C6—C5	108.22 (15)	C18—C17—C20	121.3 (2)
C6—C7—C8	118.85 (18)	C16—C17—C20	121.6 (2)
C6—C7—H7	120.6	C17—C18—C19	121.97 (19)
C8—C7—H7	120.6	C17—C18—H18	119.0
C9—C8—C7	120.51 (18)	C19—C18—H18	119.0
C9—C8—H8	119.7	C18—C19—C14	119.90 (18)
C7—C8—H8	119.7	C18—C19—H19	120.1
C10—C9—C8	121.14 (18)	C14—C19—H19	120.1
C10—C9—H9	119.4	C17—C20—H20A	109.5
C8—C9—H9	119.4	C17—C20—H20B	109.5
C9—C10—C11	118.36 (17)	H20A—C20—H20B	109.5
C9—C10—H10	120.8	C17—C20—H20C	109.5
C11—C10—H10	120.8	H20A—C20—H20C	109.5
C10—C11—C6	120.60 (17)	H20B—C20—H20C	109.5
C10—C11—C12	130.17 (16)	C12—N1—C14	120.76 (15)
C13—C1—C2—C3	0.1 (3)	C2—C1—C13—C5	1.4 (3)

supplementary materials

C1—C2—C3—C4	-1.4 (3)	C2—C1—C13—C12	176.50 (17)
C2—C3—C4—C5	1.2 (3)	C4—C5—C13—C1	-1.6 (3)
C3—C4—C5—C13	0.3 (3)	C6—C5—C13—C1	176.08 (15)
C3—C4—C5—C6	-176.88 (17)	C4—C5—C13—C12	-177.78 (15)
C4—C5—C6—C7	-6.8 (3)	C6—C5—C13—C12	-0.07 (18)
C13—C5—C6—C7	175.76 (17)	N1—C12—C13—C1	8.8 (3)
C4—C5—C6—C11	175.26 (18)	C11—C12—C13—C1	-173.40 (17)
C13—C5—C6—C11	-2.18 (18)	N1—C12—C13—C5	-175.62 (18)
C11—C6—C7—C8	-0.7 (3)	C11—C12—C13—C5	2.14 (17)
C5—C6—C7—C8	-178.41 (17)	C19—C14—C15—C16	-1.7 (3)
C6—C7—C8—C9	-0.1 (3)	N1—C14—C15—C16	-176.54 (17)
C7—C8—C9—C10	0.6 (3)	C14—C15—C16—C17	1.2 (3)
C8—C9—C10—C11	-0.3 (3)	C15—C16—C17—C18	0.0 (3)
C9—C10—C11—C6	-0.4 (3)	C15—C16—C17—C20	179.4 (2)
C9—C10—C11—C12	174.14 (17)	C16—C17—C18—C19	-0.7 (3)
C7—C6—C11—C10	1.0 (3)	C20—C17—C18—C19	179.9 (2)
C5—C6—C11—C10	179.15 (15)	C17—C18—C19—C14	0.2 (3)
C7—C6—C11—C12	-174.66 (15)	C15—C14—C19—C18	1.0 (3)
C5—C6—C11—C12	3.53 (18)	N1—C14—C19—C18	175.92 (17)
C10—C11—C12—N1	-0.5 (3)	C11—C12—N1—C14	-169.33 (16)
C6—C11—C12—N1	174.52 (16)	C13—C12—N1—C14	8.1 (3)
C10—C11—C12—C13	-178.57 (17)	C15—C14—N1—C12	-115.92 (19)
C6—C11—C12—C13	-3.52 (18)	C19—C14—N1—C12	69.2 (2)

Fig. 1

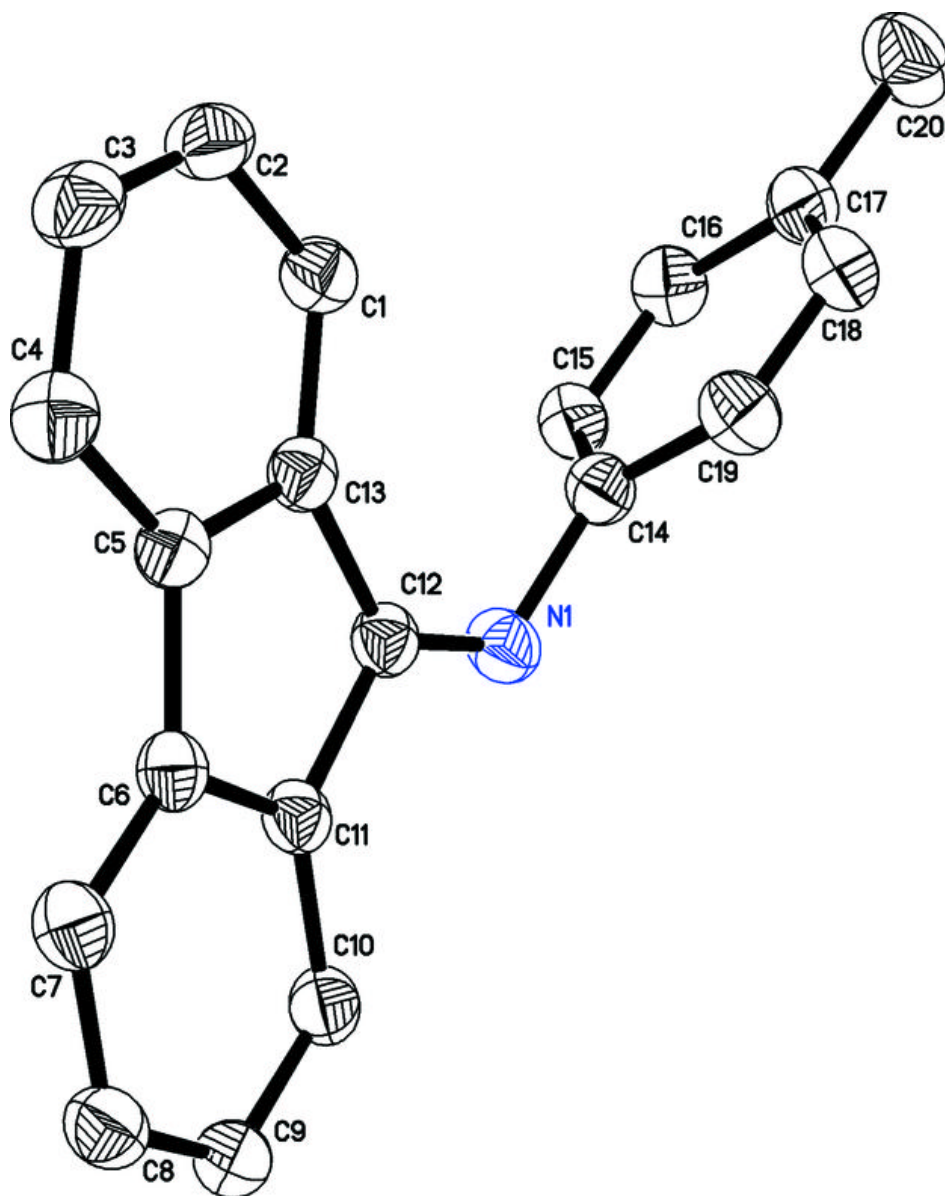


Fig. 2

