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Bis[2,6-bis(4,5-dihydro-1*H*-imidazol-2-yl)-pyridine]manganese(II) bis(perchlorate) acetonitrile solvate

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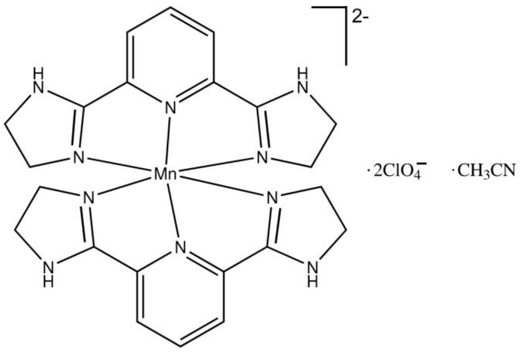
Key indicators: single-crystal X-ray study; T = 273 K; mean σ(C–C) = 0.007 Å; disorder in solvent or counterion; R factor = 0.047; wR factor = 0.121; data-to-parameter ratio = 11.5.

In the cation of the title compound, $[Mn(C_{11}H_{13}N_5)_2](ClO_4)_2\cdot CH_3CN$, the metal atom is located on a twofold rotation axis and is six-coordinated by six N atoms from two different 2,6-bis(4,5-dihydro-1*H*-imidazol-2-yl)pyridine (bip) ligands in a distorted octahedral geometry. The O atoms of the perchlorate anions are disordered with occupancies in the ratio 0.593 (10):0.407 (10). In the crystal, molecules are stabilized by two N—H···O hydrogen bonds, forming zigzag chains along the *a* axis, which are further interconnected by N—H···O hydrogen bonds and π–π interactions [centroid–centroid distance = 3.50 (1) Å] into a three-dimensional network.

Related literature

For the network topologies and potential applications of supramolecular architectures, see: Yaghi *et al.* (1998); Hagrman *et al.* (1999). The protonation and deprotonation of an imidazole ligand is believed to play an important role in the mechanism of the coordination chemistry, see: Bordo *et al.* (2001). Our studies of such complexes involving an imidazole ligand indicate that hydrogen bonding involving this group influences the geometry around the metal atom and the crystallization mechanism, see: Ren *et al.* (2007, 2009); Ren, Ye, He *et al.* (2004); Ren, Ye, Zhu *et al.* (2004). For metal–imidazole bond lengths, see: Stupka *et al.* (2004); Hammes *et al.* (2005); Haga *et al.* (1996); Böca *et al.* (2005). For metal–imidazole bond lengths, see: Ren *et al.* (2009). For the synthesis of 2,6-bis(4,5-dihydro-1*H*-imidazol-2-yl)pyridine, see: Baker *et al.* (1991).



Experimental

Crystal data

$[Mn(C_{11}H_{13}N_5)_2](ClO_4)_2\cdot C_2H_3N$
M_r = 725.42
Monoclinic, $C2/c$
a = 20.521 (5) Å
b = 12.732 (5) Å
c = 14.602 (6) Å
β = 123.893 (10)°
V = 3167.0 (19) Å^3
Z = 4
Mo $K\alpha$ radiation
μ = 0.65 mm^{-1}
T = 273 K
0.28 × 0.21 × 0.14 mm

Data collection

Bruker SMART CCD area-detector diffractometer
Absorption correction: multi-scan (*SADABS*; Bruker, 1998)
T_{min} = 0.837, T_{max} = 0.912
7799 measured reflections
2821 independent reflections
1277 reflections with $I > 2\sigma(I)$
R_{int} = 0.056

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.047
$wR(F^2)$ = 0.121
S = 0.79
2821 reflections
246 parameters
94 restraints
H-atom parameters constrained
$\Delta\rho_{max}$ = 0.33 e Å^{-3}
$\Delta\rho_{min}$ = −0.27 e Å^{-3}

Table 1
Selected geometric parameters (Å, °).

Mn1—N4	2.247 (3)	Mn1—N1	2.287 (3)
Mn1—N2	2.283 (3)		

Table 2
Hydrogen-bond geometry (Å, °).

D—H···A	D—H	H···A	D···A	D—H···A
N5—H5A···O2^{ii}	0.86	2.50	3.237 (12)	144
N5—H5A···O3^{ii}	0.86	2.25	2.942 (12)	137
N5—H5A···O2$'^{ii}$	0.86	2.11	2.965 (8)	176
N3—H3A···O4^{iii}	0.86	2.52	3.26 (2)	144
N3—H3A···O3$'^{iii}$	0.86	2.16	3.015 (8)	178

Symmetry codes: (ii) $x, -y+1, z+\frac{1}{2}$; (iii) $x+\frac{1}{2}, y+\frac{1}{2}, z+1$.

Data collection: *SMART* (Bruker, 1998); cell refinement: *SAINT-Plus* (Bruker, 1998); data reduction: *SAINT-Plus*; program(s) used to solve structure: *SHELXTL* (Sheldrick, 2008); program(s) used to refine structure: *SHELXTL*; molecular graphics: *SHELXTL*; software used to prepare material for publication: *SHELXTL*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: TK2501).

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m1024 Shang *et al.* • $[Mn(C_{11}H_{13}N_5)_2](ClO_4)_2 \cdot C_2H_3N$ *Acta Cryst.* (2009). E**65**, m1023–m1024

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Acta Cryst. (2009). E**65**, m1023–m1024   [doi:10.1107/S1600536809029195]

Bis[2,6-bis(4,5-dihydro-1*H*-imidazol-2-yl)pyridine]manganese(II) bis(perchlorate) acetonitrile solvate

Shao-Ming Shang, Chun-Xia Ren, Xin Wang, Lu-De Lu and Xu-Jie Yang

S1. Comment

The construction supramolecular architectures is currently of great interest owing to their intriguing network topologies and potential functions such as adsorption, ion exchange, shape-selective catalysis, non-linear and magnetic materials (Yaghi *et al.*, 1998; Hagrman *et al.*, 1999). The protonation and deprotonation of an imidazole ligand is believed to play an important role in the mechanism of the coordination chemistry (Bordo, *et al.*, 2001). We described previously a number of such metal complexes, including imidazole ligand, and have concluded that hydrogen bonding involving this group influences the geometry around the metal atom and the crystallization mechanism (Ren, Ye, He *et al.*, 2004; Ren, Ye, Zhu *et al.*, 2004; Ren *et al.*, 2007, 2009). We report here the preparation and crystal structure of a mononuclear coordination complex, $[Mn(bip)_2](ClO_4)_2.CH_3CN$ (I) (bip is 2,6-bis(4,5-dihydro-1*H*-imidazol-2-yl)pyridine).

The crystal structure of (I) crystallizes in the monoclinic space group *C*2/*c*. As shown in Fig. 1, the title compound consists of a $[Mn(bip)_2]^{2+}$ cation, two perchlorate counter ions and one Acetonitride molecular. The manganese(II) atom in the cation is in a distorted tetrahedral geometry, being coordinated with six nitrogen atoms from two neutral tridentate ligands bip. The Mn(1)—N bond lengths of The equatorial 2.292 (4), 2.284 (4), 2.251 (4) Å, which are slightly shorter than the metal-imidazole (Stupka, *et al.*, 2004; Hammes *et al.*, 2005; Haga *et al.*, 1996; Böca *et al.*, 2005) and longer than the metal-imidazole (Ren, *et al.*, 2009). The N—Mn(1)—N bond angle is in the range of 70.17 (15)–147.1 (2) /%. Two bip ligands of adjacent molecules are parallel to each other with a distance of 3.50 Å, showing the presence of *π*-*π* interaction. The molecules further interconnected into three-dimensional network through hydrogen bond between the oxygen atom of perchlorate counter-ion and the uncoordination nitrogen atoms of bip ligands.

S2. Experimental

All the reagents and solvents employed were commercially available and used as received without further purification. The ligand 2,6-bis(4,5-dihydro-1*H*-imidazol-2-yl)pyridine (bip) was synthesized by literature methods (Baker *et al.*, 1991).

A solution of $MnCl_24H_2O$ (0.2 mmol, 40 mg) and $NaClO_4$ (0.4 mmol, 50 mg) in acetonitrile (10 ml), was added dropwise to a stirred solution of bip (0.4 mmol, 86 mg) in methanol(10 ml) at 60 K. Yellow single crystals suitable for X-ray diffraction were obtained by slow diffusion of diethyl ether into the clear filtrate for two days in 60% yield. Elemental analysis, Found: C, 39.68; H, 3.93; N, 21.15%. Calc. for $C_{24}H_{29}C_{l2}MnN_{11}O_8$: C, 39.70; H, 4.00; N, 21.23%. Main IR bands (KBr, cm_{-1}): 3370*m*, 3204 s, 2938*m*, 2887*m*, 1595*m*, 1567 s, 1531 s, 1453 s, 1283 s, 1209w, 1144 s,1116 s, 1089 s, 1028w, 1010*m*, 953w, 830w, 752w, 663w, 636*m*, 628*m*.

S3. Refinement

The H atom attached to N(2) atom was refined isotropically. All the other H atoms were placed in geometrically idealized positions and constrained to ride on their parent atoms with N—H and C—H distances of 0.90 Å and 0.96 Å, respectively, and $U_{\rm iso}$(H) = 1.2 times of those of their parent atoms (Å^2). The O atoms are resolved into two positions by PART instructions. The occupancy for the unprimed O atoms is set at 21 and that of the primed atoms at -21. The clorine-oxygen distances were restrained to 1.44 Å (and the oxygen-oxygen interaction to 2.35 Å). Additionally, the vibration of the oxygen atoms were made isotropic by an ISOR restraint. The O atoms are resolved into two positions and give the site occupany of the major component.

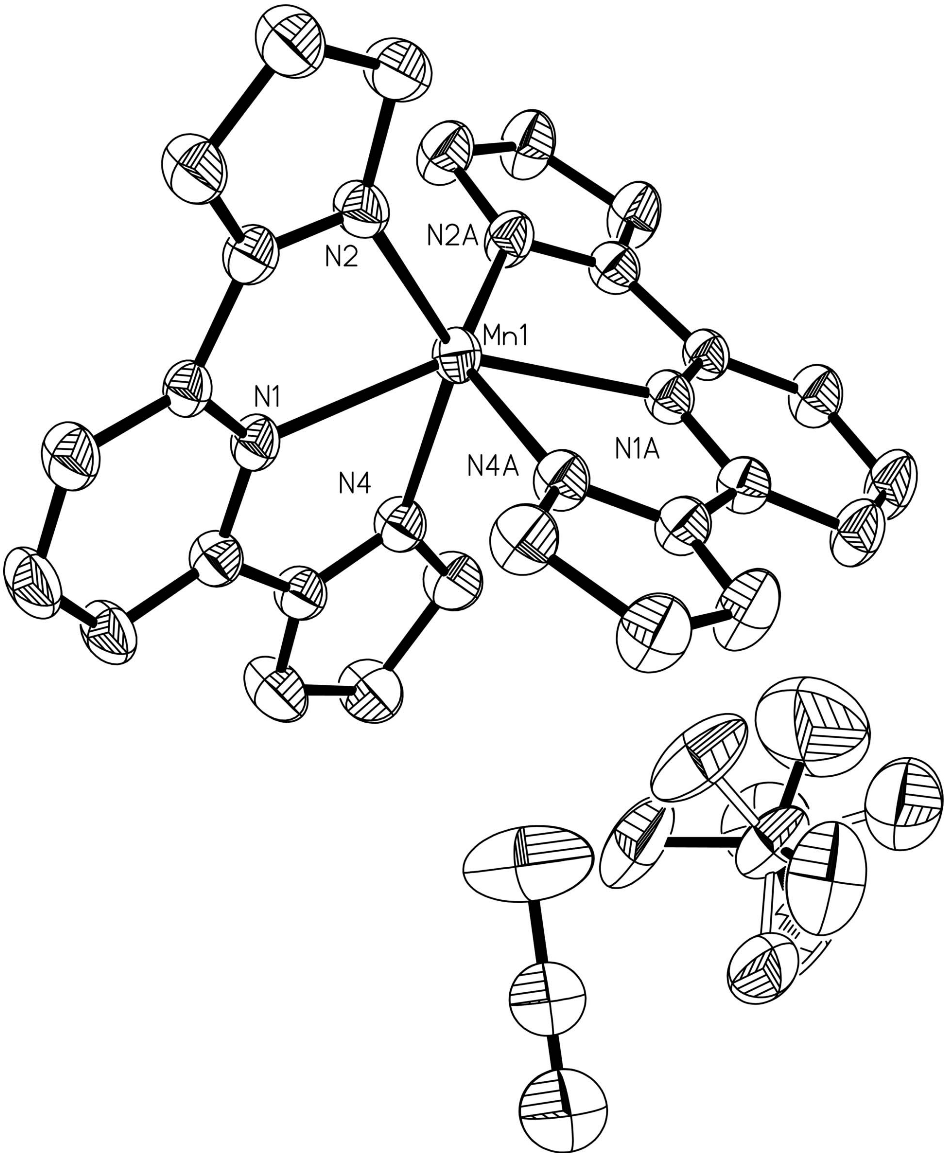

Figure 1

The structure of the complex [Mn(bip)$_2$]$^{2+}$ showing 30% probability displacement ellipsoids and H atoms have been omitted for clarity.

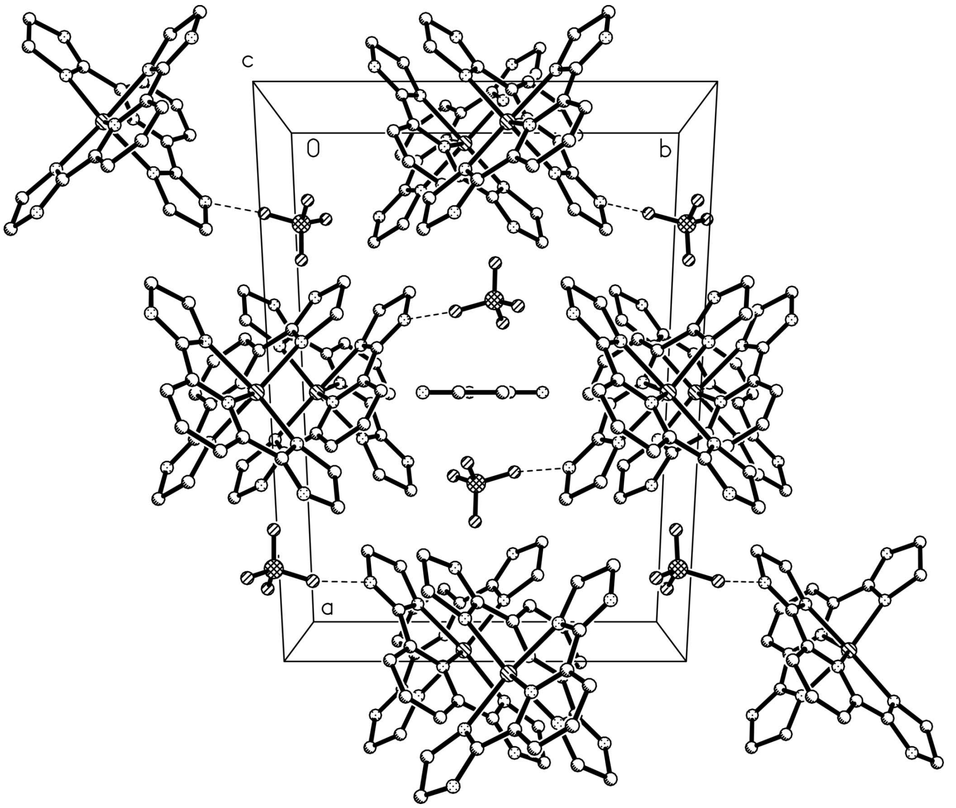

Figure 2

The framework of [Mn(bip)$_2$](ClO$_4$)$_2$CH$_3$CN viewed along the *c* axis. H atoms have been omitted for clarity.

Bis[2,6-bis(4,5-dihydro-1*H*-imidazol-2-yl)pyridine]manganese(II) bis(perchlorate) acetonitrile solvate

Crystal data

[Mn(C$_{11}$H$_{13}$N$_5$)$_2$](ClO$_4$)$_2$·C$_2$H$_3$N
M_r = 725.42
Monoclinic, $C2/c$
a = 20.521 (5) Å
b = 12.732 (5) Å
c = 14.602 (6) Å
β = 123.893 (10)°
V = 3167.0 (19) Å^3
Z = 4

F(000) = 1492
D_x = 1.521 Mg m^{-3}
Mo $K\alpha$ radiation, λ = 0.71073 Å
Cell parameters from 668 reflections
θ = 2.0–25.1°
μ = 0.65 mm^{-1}
T = 273 K
Block, yellow
0.28 × 0.21 × 0.14 mm

Data collection

Bruker SMART CCD area-detector diffractometer
Radiation source: fine-focus sealed tube
Graphite monochromator
φ and ω scans
Absorption correction: multi-scan (*SADABS*; Bruker, 1998)
$T_{\min}$ = 0.837, $T_{\max}$ = 0.912

7799 measured reflections
2821 independent reflections
1277 reflections with $I > 2\sigma(I)$
R_{int} = 0.056
$\theta_{\max}$ = 25.1°, $\theta_{\min}$ = 2.0°
$h = -22 \rightarrow 24$
$k = -11 \rightarrow 15$
$l = -16 \rightarrow 17$

Refinement

Refinement on F^2
Least-squares matrix: full
$R[F^2 > 2\sigma(F^2)] = 0.047$
$wR(F^2) = 0.121$
$S = 0.79$
2821 reflections
246 parameters
94 restraints
Primary atom site location: structure-invariant direct methods
Secondary atom site location: difference Fourier map
Hydrogen site location: inferred from neighbouring sites
H-atom parameters constrained
$w = 1/[\sigma^2(F_o^2) + (0.0647P)^2]$ where $P = (F_o^2 + 2F_c^2)/3$
$(\Delta/\sigma)_{\max} = 0.001$
$\Delta\rho_{\max} = 0.33$ e Å^{-3}
$\Delta\rho_{\min} = -0.27$ e Å^{-3}

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted *R*-factor *wR* and goodness of fit *S* are based on F^2, conventional *R*-factors *R* are based on *F*, with *F* set to zero for negative F^2. The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating *R*-factors(gt) *etc*. and is not relevant to the choice of reflections for refinement. *R*-factors based on F^2 are statistically about twice as large as those based on *F*, and *R*- factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	*y*	*z*	$U_{\rm iso}$*/$U_{\rm eq}$	Occ. (<1)
Mn1	0.5000	0.94417 (7)	0.7500	0.0488 (3)	
N1	0.51856 (17)	0.8931 (3)	0.9132 (2)	0.0468 (8)	
N2	0.59227 (18)	1.0487 (2)	0.8898 (2)	0.0539 (8)	
N3	0.6779 (2)	1.0731 (3)	1.0725 (3)	0.0708 (10)	
H3A	0.6979	1.0635	1.1416	0.085*	
N4	0.40797 (17)	0.8262 (2)	0.7177 (2)	0.0524 (8)	
N5	0.3602 (2)	0.7083 (3)	0.7785 (3)	0.0833 (12)	
H5A	0.3581	0.6730	0.8271	0.100*	
C1	0.6436 (2)	1.1375 (4)	0.9029 (3)	0.0746 (13)	
H1A	0.6130	1.2013	0.8721	0.090*	
H1B	0.6693	1.1225	0.8652	0.090*	
C2	0.7041 (3)	1.1503 (4)	1.0259 (3)	0.0785 (14)	
H2A	0.7566	1.1347	1.0453	0.094*	
H2B	0.7032	1.2208	1.0504	0.094*	
C3	0.6168 (2)	1.0191 (3)	0.9886 (3)	0.0509 (10)	
C4	0.5788 (2)	0.9317 (3)	1.0073 (3)	0.0450 (9)	
C5	0.6011 (2)	0.8891 (3)	1.1076 (3)	0.0630 (12)	
H5	0.6430	0.9169	1.1733	0.076*	
C6	0.5594 (3)	0.8037 (4)	1.1076 (3)	0.0757 (14)	
H6	0.5736	0.7730	1.1741	0.091*	
C7	0.4970 (2)	0.7637 (4)	1.0097 (3)	0.0678 (13)	
H7	0.4689	0.7062	1.0090	0.081*	
C8	0.4778 (2)	0.8114 (3)	0.9136 (3)	0.0506 (10)	
C9	0.4140 (2)	0.7806 (3)	0.8007 (3)	0.0525 (11)	
C10	0.3053 (3)	0.6981 (4)	0.6593 (3)	0.0818 (14)	

H10A	0.2523	0.7166	0.6361	0.098*	
H10B	0.3054	0.6279	0.6335	0.098*	
N6	0.5000	0.3462 (7)	0.7500	0.186 (4)	
C12	0.5000	0.5403 (7)	0.7500	0.178 (5)	
H12B	0.5294	0.5712	0.7238	0.214*	0.50
H12A	0.5189	0.5583	0.8247	0.214*	0.50
H12C	0.4453	0.5583	0.7073	0.214*	0.50
C13	0.5000	0.4339 (7)	0.7500	0.116 (3)	
C11	0.3378 (2)	0.7797 (3)	0.6178 (3)	0.0664 (12)	
H11A	0.3520	0.7454	0.5720	0.080*	
H11B	0.2991	0.8334	0.5742	0.080*	
Cl1	0.32690 (9)	0.52285 (13)	0.39585 (12)	0.1026 (5)	
O1	0.3119 (7)	0.5391 (11)	0.4824 (8)	0.147 (6)	0.407 (10)
O2	0.4047 (4)	0.4898 (11)	0.4424 (8)	0.152 (7)	0.407 (10)
O3	0.2745 (6)	0.4409 (11)	0.3284 (13)	0.209 (9)	0.407 (10)
O4	0.3090 (11)	0.6156 (10)	0.3332 (15)	0.264 (11)	0.407 (10)
O1′	0.3558 (5)	0.6031 (6)	0.4718 (7)	0.151 (5)	0.593 (10)
O2′	0.3455 (6)	0.4194 (5)	0.4361 (7)	0.139 (5)	0.593 (10)
O3′	0.2464 (3)	0.5354 (7)	0.3137 (5)	0.139 (4)	0.593 (10)
O4′	0.3628 (7)	0.5401 (8)	0.3310 (10)	0.217 (7)	0.593 (10)

Atomic displacement parameters (Å^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Mn1	0.0517 (6)	0.0556 (6)	0.0338 (5)	0.000	0.0205 (4)	0.000
N1	0.047 (2)	0.054 (2)	0.0324 (18)	−0.0002 (17)	0.0180 (17)	−0.0028 (15)
N2	0.059 (2)	0.058 (2)	0.0378 (19)	−0.0088 (18)	0.0229 (16)	0.0011 (16)
N3	0.074 (2)	0.084 (3)	0.0357 (19)	−0.030 (2)	0.0196 (19)	−0.0089 (19)
N4	0.051 (2)	0.059 (2)	0.0362 (18)	−0.0072 (17)	0.0177 (17)	−0.0039 (16)
N5	0.091 (3)	0.093 (3)	0.046 (2)	−0.046 (3)	0.027 (2)	−0.007 (2)
C1	0.088 (3)	0.072 (3)	0.054 (3)	−0.023 (3)	0.034 (3)	−0.004 (2)
C2	0.081 (3)	0.091 (4)	0.053 (3)	−0.030 (3)	0.032 (3)	−0.009 (2)
C3	0.049 (3)	0.055 (3)	0.042 (2)	−0.002 (2)	0.022 (2)	−0.004 (2)
C4	0.047 (2)	0.051 (3)	0.034 (2)	0.000 (2)	0.020 (2)	−0.002 (2)
C5	0.065 (3)	0.076 (3)	0.035 (2)	−0.010 (3)	0.019 (2)	−0.003 (2)
C6	0.093 (4)	0.089 (4)	0.039 (3)	−0.014 (3)	0.033 (3)	0.008 (2)
C7	0.080 (3)	0.076 (3)	0.040 (3)	−0.022 (3)	0.029 (2)	−0.001 (2)
C8	0.047 (2)	0.056 (3)	0.043 (2)	−0.005 (2)	0.021 (2)	−0.003 (2)
C9	0.054 (3)	0.055 (3)	0.040 (2)	−0.008 (2)	0.021 (2)	−0.003 (2)
C10	0.081 (3)	0.087 (4)	0.058 (3)	−0.034 (3)	0.027 (3)	−0.015 (3)
N6	0.304 (12)	0.120 (8)	0.196 (9)	0.000	0.179 (9)	0.000
C12	0.163 (10)	0.098 (8)	0.154 (9)	0.000	0.013 (7)	0.000
C13	0.153 (8)	0.094 (8)	0.124 (7)	0.000	0.091 (6)	0.000
C11	0.062 (3)	0.073 (3)	0.044 (2)	−0.018 (2)	0.017 (2)	−0.010 (2)
Cl1	0.0956 (11)	0.0967 (12)	0.0876 (10)	0.0154 (9)	0.0339 (9)	−0.0296 (9)
O1	0.165 (12)	0.155 (13)	0.094 (7)	0.048 (10)	0.056 (8)	−0.039 (7)
O2	0.077 (6)	0.243 (16)	0.111 (10)	0.016 (7)	0.038 (6)	−0.023 (10)
O3	0.105 (9)	0.242 (16)	0.247 (18)	−0.054 (11)	0.079 (11)	−0.189 (15)

O4	0.31 (3)	0.224 (15)	0.27 (2)	0.084 (15)	0.17 (2)	0.105 (15)
O1′	0.134 (7)	0.129 (7)	0.109 (6)	0.003 (5)	0.017 (5)	−0.077 (6)
O2′	0.211 (13)	0.103 (5)	0.137 (7)	0.059 (7)	0.118 (8)	0.025 (5)
O3′	0.110 (5)	0.159 (9)	0.065 (4)	−0.011 (5)	−0.004 (4)	−0.018 (4)
O4′	0.320 (15)	0.167 (10)	0.300 (16)	−0.050 (10)	0.256 (16)	−0.056 (9)

Geometric parameters (Å, º)

Mn1—N4	2.247 (3)	C5—C6	1.384 (5)
Mn1—N4^{i}	2.247 (3)	C5—H5	0.9300
Mn1—N2	2.283 (3)	C6—C7	1.378 (5)
Mn1—N2^{i}	2.283 (3)	C6—H6	0.9300
Mn1—N1	2.287 (3)	C7—C8	1.370 (5)
Mn1—N1^{i}	2.287 (3)	C7—H7	0.9300
N1—C4	1.328 (4)	C8—C9	1.476 (5)
N1—C8	1.337 (4)	C10—C11	1.530 (5)
N2—C3	1.292 (4)	C10—H10A	0.9700
N2—C1	1.484 (5)	C10—H10B	0.9700
N3—C3	1.354 (5)	N6—C13	1.117 (10)
N3—C2	1.458 (5)	C12—C13	1.3547
N3—H3A	0.8600	C12—H12B	0.9600
N4—C9	1.287 (4)	C12—H12A	0.9600
N4—C11	1.486 (4)	C12—H12C	0.9600
N5—C9	1.332 (4)	C11—H11A	0.9700
N5—C10	1.459 (5)	C11—H11B	0.9700
N5—H5A	0.8600	Cl1—O1′	1.376 (5)
C1—C2	1.519 (5)	Cl1—O2′	1.406 (5)
C1—H1A	0.9700	Cl1—O2	1.407 (6)
C1—H1B	0.9700	Cl1—O3′	1.409 (5)
C2—H2A	0.9700	Cl1—O4	1.411 (7)
C2—H2B	0.9700	Cl1—O3	1.426 (7)
C3—C4	1.468 (5)	Cl1—O1	1.474 (7)
C4—C5	1.379 (5)	Cl1—O4′	1.505 (6)

N4—Mn1—N4^{i}	96.14 (16)	C8—C7—H7	121.0
N4—Mn1—N2	139.28 (11)	C6—C7—H7	121.0
N4^{i}—Mn1—N2	91.20 (11)	N1—C8—C7	121.8 (4)
N4—Mn1—N2^{i}	91.20 (11)	N1—C8—C9	111.6 (3)
N4^{i}—Mn1—N2^{i}	139.28 (11)	C7—C8—C9	126.6 (4)
N2—Mn1—N2^{i}	108.70 (16)	N4—C9—N5	116.9 (3)
N4—Mn1—N1	70.18 (11)	N4—C9—C8	119.4 (4)
N4^{i}—Mn1—N1	87.66 (10)	N5—C9—C8	123.7 (4)
N2—Mn1—N1	70.18 (12)	N5—C10—C11	101.3 (3)
N2^{i}—Mn1—N1	132.10 (10)	N5—C10—H10A	111.3
N4—Mn1—N1^{i}	87.66 (10)	C11—C10—H10A	110.3
N4^{i}—Mn1—N1^{i}	70.18 (11)	N5—C10—H10B	112.2
N2—Mn1—N1^{i}	132.10 (11)	C11—C10—H10B	112.1
N2^{i}—Mn1—N1^{i}	70.18 (11)	H10A—C10—H10B	109.4

N1—Mn1—N1^{i}	146.97 (17)	C13—C12—H12B	114.2
C4—N1—C8	120.4 (3)	C13—C12—H12A	103.9
C4—N1—Mn1	119.4 (2)	H12B—C12—H12A	114.2
C8—N1—Mn1	119.0 (2)	C13—C12—H12C	103.9
C3—N2—C1	105.8 (3)	H12B—C12—H12C	114.2
C3—N2—Mn1	116.2 (3)	H12A—C12—H12C	105.4
C1—N2—Mn1	137.3 (2)	N6—C13—C12	180.000 (6)
C3—N3—C2	108.5 (3)	N4—C11—C10	106.3 (3)
C3—N3—H3A	125.7	N4—C11—H11A	110.9
C2—N3—H3A	125.7	C10—C11—H11A	109.5
C9—N4—C11	106.0 (3)	N4—C11—H11B	110.7
C9—N4—Mn1	118.4 (3)	C10—C11—H11B	111.2
C11—N4—Mn1	135.5 (2)	H11A—C11—H11B	108.3
C9—N5—C10	109.6 (3)	O1′—Cl1—O2′	117.7 (5)
C9—N5—H5A	125.2	O1′—Cl1—O2	88.3 (5)
C10—N5—H5A	125.2	O2′—Cl1—O2	63.0 (5)
N2—C1—C2	106.7 (3)	O1′—Cl1—O3′	112.1 (4)
N2—C1—H1A	110.4	O2′—Cl1—O3′	112.2 (5)
C2—C1—H1A	110.4	O2—Cl1—O3′	157.2 (5)
N2—C1—H1B	110.4	O1′—Cl1—O4	74.9 (8)
C2—C1—H1B	110.4	O2′—Cl1—O4	165.2 (8)
H1A—C1—H1B	108.6	O2—Cl1—O4	112.3 (7)
N3—C2—C1	102.0 (3)	O3′—Cl1—O4	66.1 (7)
N3—C2—H2A	111.4	O1′—Cl1—O3	157.0 (5)
C1—C2—H2A	111.4	O2′—Cl1—O3	61.8 (7)
N3—C2—H2B	111.4	O2—Cl1—O3	109.3 (6)
C1—C2—H2B	111.4	O3′—Cl1—O3	54.5 (6)
H2A—C2—H2B	109.2	O4—Cl1—O3	109.9 (7)
N2—C3—N3	116.7 (4)	O1′—Cl1—O1	53.6 (5)
N2—C3—C4	120.8 (3)	O2′—Cl1—O1	84.9 (6)
N3—C3—C4	122.5 (3)	O2—Cl1—O1	110.6 (6)
N1—C4—C5	121.3 (4)	O3′—Cl1—O1	90.5 (5)
N1—C4—C3	111.9 (3)	O4—Cl1—O1	109.6 (6)
C5—C4—C3	126.8 (4)	O3—Cl1—O1	104.9 (6)
C4—C5—C6	118.1 (4)	O1′—Cl1—O4′	104.9 (5)
C4—C5—H5	121.0	O2′—Cl1—O4′	106.6 (4)
C6—C5—H5	121.0	O2—Cl1—O4′	61.7 (5)
C7—C6—C5	120.4 (4)	O3′—Cl1—O4′	101.6 (5)
C7—C6—H6	119.8	O4—Cl1—O4′	60.8 (7)
C5—C6—H6	119.8	O3—Cl1—O4′	96.7 (6)
C8—C7—C6	118.0 (4)	O1—Cl1—O4′	158.3 (6)

Symmetry code: (i) −*x*+1, *y*, −*z*+3/2.

Hydrogen-bond geometry (Å, º)

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
N5—H5*A*···O2^{ii}	0.86	2.50	3.237 (12)	144

N5—H5*A*···O3^{ii}	0.86	2.25	2.942 (12)	137
N5—H5*A*···O2′^{ii}	0.86	2.11	2.965 (8)	176
N3—H3*A*···O4^{iii}	0.86	2.52	3.26 (2)	144
N3—H3*A*···O3′^{iii}	0.86	2.16	3.015 (8)	178

Symmetry codes: (ii) x, $-y+1$, $z+1/2$; (iii) $x+1/2$, $y+1/2$, $z+1$.