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## Structure Reports

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**(Z)-2,2,2-Trichloro-*N*<sup>2</sup>-cyanoacetamide**

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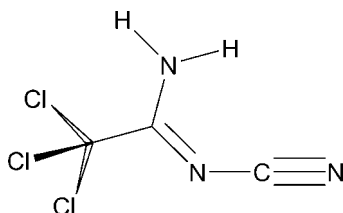
Received 1 August 2009; accepted 5 August 2009

Key indicators: single-crystal X-ray study;  $T = 173$  K; mean  $\sigma(\text{C}-\text{C}) = 0.002$  Å;  $R$  factor = 0.019;  $wR$  factor = 0.050; data-to-parameter ratio = 18.7.

The title compound,  $\text{C}_3\text{H}_2\text{Cl}_3\text{N}_3$ , crystallizes as the *Z* isomer with respect to the  $\text{C}=\text{N}$  bond. The  $-\text{C}(\text{NH}_2)=\text{NCN}$  functional group is effectively planar (r.m.s. deviation = 0.016 Å), with only the three Cl atoms out of the molecular plane. A strong network of  $\text{N}-\text{H}\cdots\text{N}$  hydrogen bonds forms dimers which are associated into ribbons in the crystal structure. Hydrogen bonding is suspected to be the cause of the near-equivalence of the formal  $\text{C}-\text{N}$  and  $\text{C}=\text{N}$  bonds ( $\Delta_{\text{CN}} = 0.008$  Å)

## Related literature

For literature related to characterization, see: Huffman & Schaefer (1963). For comparable structures of *N'*-cyanoamidines; see Allen (2002). For the crystal structures of *N*<sup>2</sup>-cyano-3-[2-diaminomethyleneamino]-4-thiazolylmethylthio]-propionamidinemonohydrate, (II) and 3-[2-[amino(methylamino)methyleneamino]-4-thiazolylmethylthio]-*N*<sup>2</sup>-cyano-propionamide, (III), see Ishida *et al.* (1989). For the crystal structure of (*E*)-1,2-bis(1-amino-1-(cyanoimino)-2-methylprop-2-yl)diazene-1,2-dioxide, (IV), see: Tretyakov *et al.* (2006). For the sole other acyclic trichloromethyl amidine with a reported crystal structure, *N*-(4-amino-3-furanzanyl)-2,2,2-trichloro-*N*-methoxyacetamide, (V), see: George & Gilardi (1986). For background to the  $\Delta_{\text{CN}}$  parameter, see: Boéré, *et al.* (1998).



## Experimental

## Crystal data

$\text{C}_3\text{H}_2\text{Cl}_3\text{N}_3$	$V = 673.89$ (8) Å <sup>3</sup>
$M_r = 186.43$	$Z = 4$
Monoclinic, $P2_1/n$	Mo $K\alpha$ radiation
$a = 5.5388$ (4) Å	$\mu = 1.26$ mm <sup>-1</sup>
$b = 6.6127$ (4) Å	$T = 173$ K
$c = 18.4727$ (12) Å	$0.41 \times 0.27 \times 0.21$ mm
$\beta = 95.122$ (1)°	

## Data collection

Bruker APEXII CCD area-detector diffractometer	7459 measured reflections
Absorption correction: multi-scan (SADABS; Bruker, 2006)	1552 independent reflections
$T_{\text{min}} = 0.616$ , $T_{\text{max}} = 0.770$	1479 reflections with $I > 2\sigma(I)$
	$R_{\text{int}} = 0.017$

## Refinement

$R[F^2 > 2\sigma(F^2)] = 0.019$	83 parameters
$wR(F^2) = 0.050$	H-atom parameters constrained
$S = 1.06$	$\Delta\rho_{\text{max}} = 0.41$ e Å <sup>-3</sup>
1552 reflections	$\Delta\rho_{\text{min}} = -0.27$ e Å <sup>-3</sup>

Table 1

Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{N1}-\text{H1A}\cdots\text{N3}^i$	0.88	2.10	2.9583 (15)	164
$\text{N1}-\text{H1B}\cdots\text{N3}^{ii}$	0.88	2.40	3.1893 (15)	150

Symmetry codes: (i)  $-x + 2, -y, -z$ ; (ii)  $x, y + 1, z$ .

Table 2

Comparative distances (Å) and angles (°) in amidines (I)–(V).

Value	(I)	(II)	(III)	(IV)	(V)
$\text{C2}-\text{N1}$	1.3115 (15)	1.308 (4)	1.308 (3)	1.307 (2)	1.387 (4)
$\text{C2}-\text{N2}$	1.3032 (15)	1.320 (4)	1.317 (3)	1.306 (2)	1.2737 (4)
$\Delta_{\text{CN}}$	0.008	-0.012	-0.009	0.001	0.114
$\text{C2}-\text{C1}$	1.5396 (15)	1.520 (4)	1.513 (3)	1.522 (3)	1.525 (5)
$\text{C3}-\text{N3}$	1.1533 (17)	1.164 (4)	1.1567 (3)	1.153 (3)	
$\text{C3}-\text{N2}$	1.3226 (16)	1.320 (4)	1.333 (3)	1.322 (3)	
$\text{N2}-\text{C2}-\text{N1}$	127.94 (11)	118.0 (2)	125.9 (2)	126.0 (2)	127.7 (3)
$\text{N2}-\text{C2}-\text{C1}$	114.43 (10)	124.1 (2)	116.8 (2)	114.9 (1)	117.2 (3)
$\text{N1}-\text{C2}-\text{C1}$	117.52 (10)	117.9 (2)	117.3 (2)	118.7 (1)	115.0 (3)
$\text{C2}-\text{N2}-\text{C3}$	121.04 (10)	119.1 (2)	118.7 (1)		
$\text{N3}-\text{C3}-\text{N2}$	172.16 (13)	173.2 (2)	173.9 (3)	173.2 (2)	

Compound (I) corresponds to the title compound and (II)–(V) are defined in the *Related Literature* section. Atom numbering corresponds to that in Fig. 1 (in Supplementary materials).

Data collection: APEX2 (Bruker, 2006); cell refinement: SAINT-Plus (Bruker, 2006); data reduction: SAINT-Plus; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXTL (Sheldrick, 2008); molecular graphics: Mercury (Macrae *et al.*, 2006); software used to prepare material for publication: publCIF (Westrip, 2009).

The Natural Sciences and Engineering Research Council of Canada (NSERC) is gratefully acknowledged for a Discovery Grant. The diffractometer was purchased with the help of NSERC and the University of Lethbridge. Tracy Burton (formerly of this Department) is acknowledged for the synthesis of the title compound.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: WN2341).

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**supplementary materials**

*Acta Cryst.* (2009). E65, o2134-o2135 [ doi:10.1107/S1600536809031079 ]

## (Z)-2,2,2-Trichloro-*N*<sup>2</sup>-cyanoacetamidine

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### Comment

The structure of the title compound, (I), is shown in Fig. 1. Molecular dimensions are available in the archived CIF. Structure (I) crystallizes as the *Z* isomer with respect to the imino bond (Fig. 1). The structure is essentially planar except for the CCl<sub>3</sub> group (r.m.s. mean deviation for the —C(NH<sub>2</sub>)=NCN group is 0.016 Å), while Cl2 is almost perpendicular to this plane; thus Cl1 deviates by 0.65 and Cl3 by 0.84 Å from the plane. The parameter  $\Delta_{\text{CN}} = d(\text{C—N}) - d(\text{C=N})$  has been found to range between 0 and 0.178 Å for many amidines for which the structures are known (Boeré *et al.*, 1998). For (I),  $\Delta_{\text{CN}} = 0.008$  Å, which is very small for a monomeric amidine with such unsymmetrical substitution. The N2–C3–N3 angle is almost linear, at 172.16 (13)°. There is a network of N—H···N hydrogen bonds (Table 1) linking centrosymmetric pairs of molecules into planar ribbons along the *b* axis (Fig. 2). Short contacts of 3.203 (1) Å between Cl2 (the upward- and downward-facing chlorine atoms) and N2 (imino nitrogen) link these layers into a 3-D network in the crystal structure. Finally, there are 3.4132 (5) Å short contacts between Cl1 and Cl2 bridging molecules. It is likely that this strong intermolecular hydrogen bonding is responsible for the small value of  $\Delta_{\text{CN}}$ .

Of nine *N*-cyanoamidines in the literature, six are *E* (refcodes HANBAA, ILIPAU, JATLIZ, TAHHOA, TESQAK, WAXXUO; Allen, 2002) and two are *Z* (refcodes JATMAS, NERKAX; Allen, 2002) with respect to the imino bond; for the last, VOVPUR (Allen, 2002), the isomer is not specified. The most relevant for comparison with (I) are *N*<sup>2</sup>-cyano-3-[2-diaminomethyleneamino]-4-thiazolylmethylthio]propionamidinemonohydrate, (II), 3-{2-[amino(methylamino)methyleneamino]-4-thiazolylmethylthio}-*N*<sup>2</sup>-cyanopropionamidine, (III) (Ishida *et al.*, 1989) and (*E*)-1,2-bis(1-amino-1-(cyanoimino)-2-methylprop-2-yl)diazene-1,2-dioxide, (IV) (Tretyakov *et al.*, 2006), which all bear the NH<sub>2</sub> group in addition to the nitrile on *N*. Each of these structures shares the high degree of planarity of the —C(NH<sub>2</sub>)=NCN group (r.m.s. deviations for (II) - (IV) are 0.008, 0.025 and 0.069 Å, respectively.) Of these three examples, (II) is *E* while (III) and (IV) are both *Z*; note that (II) and (III) differ only in methylation at a very remote amino group. There is only one acyclic trichloromethyl amidine with a crystal structure reported in the literature, viz. *N*-(4-amino-3-furanzanyl)-2,2,2-trichloro-*N*-methoxyacetamidine, (V) (George & Gilardi, 1986) and this is the *Z* isomer. The structure of (IV), which is arguably the most similar structure, electronically and chemically, to (I) also shows a very similar pattern of hydrogen bonding where centrosymmetric dimers are linked in ribbons within the crystal structure by additional hydrogen bonds.

Key geometrical parameters for structures (I) - (V) are compared in Table 2, which includes values for  $\Delta_{\text{CN}}$ , all of which fall within the known range. However, (II) and (III) are highly unusual in having the wrong sign for this parameter. That is, the imino bond is actually longer than the amino. We are not aware of other instances of this occurrence; the locations of the NH<sub>2</sub> hydrogen atoms in both structures were corroborated by expected hydrogen bonding. It is likely that this powerful hydrogen bonding is responsible for the inversion in expected bond distances, perhaps augmented by the strong electron-withdrawing cyano substituent on *N*'.

## Experimental

General Procedures: Reagent grade methanol was dried by distillation with Mg and catalytic I<sub>2</sub>. Sodium methoxide was transferred to the flask within a glove box under nitrogen.

Preparation of methyl trichloroacetimidate: 50 ml of dried methanol and 21.66 g (150 mmol) of trichloroacetonitrile were added to 0.50 g (10 mmol) of sodium methoxide. After stirring for 48 h at room temperature, the solution was saturated with CO<sub>2</sub>(s) to eliminate remaining sodium methoxide. Methanol was then distilled off at 335–7 K, whereafter the liquid methyl trichloroacetimidate was distilled at 415 K at a reduced pressure. Yield 19.59 g (110 mmol, 74%).

Preparation of 2,2,2-trichloro-*N*'-cyano-acetamidine: 0.42 g (10 mmol) of cyanamide was dissolved in 5 ml of anhydrous methanol. With stirring, 1.76 g (10 mmol) of methyl trichloroacetimidate was added dropwise. An ice bath may be required to maintain temperature during addition of methyl trichloroacetimidate. The solution was stirred for 3 h at RT. Methanol was removed by rotary evaporation followed by high vacuum. The solid residue was dissolved in a minimum volume (3.5 ml) of hot CH<sub>3</sub>CN, cooled to room temperature, and placed within the 238 K freezer. The colourless crystals produced were filtered and vacuum dried yielding 0.121 g (0.649 mmol, 6.51% yield), mp 433–7 K (Huffman & Schaefer, 1963).

## Refinement

Both H atoms were located in a difference Fourier map. They were refined using a riding model and U<sub>iso</sub>(H) was set equal to 1.2U<sub>eq</sub>(N1). The highest residual peak has a fraction of the electron density of a single H atom and is located 0.76 Å from Cl1.

## Figures

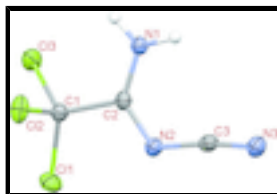


Fig. 1. A view of (I), plotted with displacement ellipsoids drawn at the 50% probability level. H atoms are shown as spheres of arbitrary radius.

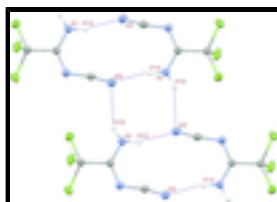


Fig. 2. The network of hydrogen bonds (dashed lines) linking centrosymmetric pairs of molecules into planar ribbons along the *b* axis. Symmetry equivalents are  $-x, 1 + y, -z$ ;  $2 - x, -y, -x$  and  $2 - x, 1 - y, -x$ . These ribbons lie parallel to the (207) Miller planes.

## (*Z*)-2,2,2-Trichloro-*N*'-cyanoacetamidine

### Crystal data

C<sub>3</sub>H<sub>2</sub>Cl<sub>3</sub>N<sub>3</sub>

*M<sub>r</sub>* = 186.43

Monoclinic, *P*2<sub>1</sub>/*n*

*F*<sub>000</sub> = 368

*D<sub>x</sub>* = 1.838 Mg m<sup>-3</sup>

Melting point: 441 K

Hall symbol: -P 2yn  
 $a = 5.5388$  (4) Å  
 $b = 6.6127$  (4) Å  
 $c = 18.4727$  (12) Å  
 $\beta = 95.122$  (1)°  
 $V = 673.89$  (8) Å<sup>3</sup>  
 $Z = 4$

Mo  $K\alpha$  radiation,  $\lambda = 0.71073$  Å  
 Cell parameters from 4497 reflections  
 $\theta = 2.2$ – $27.6$ °  
 $\mu = 1.26$  mm<sup>-1</sup>  
 $T = 173$  K  
 Block, colourless  
 $0.41 \times 0.27 \times 0.21$  mm

### Data collection

Bruker APEXII CCD area-detector diffractometer  
 Radiation source: Molybdenum  
 Monochromator: graphite  
 $T = 173$  K  
 $P = 101$  kPa  
 $\varphi$  and  $\omega$  scans  
 Absorption correction: multi-scan (SADABS; Bruker, 2006)  
 $T_{\min} = 0.616$ ,  $T_{\max} = 0.770$   
 7459 measured reflections

1552 independent reflections  
 1479 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.017$   
 $\theta_{\text{max}} = 27.5$ °  
 $\theta_{\text{min}} = 2.2$ °  
 $h = -7 \rightarrow 7$   
 $k = -8 \rightarrow 8$   
 $l = -24 \rightarrow 23$

### Refinement

Refinement on  $F^2$   
 Least-squares matrix: full  
 $R[F^2 > 2\sigma(F^2)] = 0.019$   
 $wR(F^2) = 0.050$   
 $S = 1.06$   
 1552 reflections  
 83 parameters  
 Primary atom site location: structure-invariant direct methods  
 Secondary atom site location: difference Fourier map

Hydrogen site location: difference Fourier map  
 H-atom parameters constrained  
 $w = 1/[\sigma^2(F_o^2) + (0.0247P)^2 + 0.2765P]$   
 where  $P = (F_o^2 + 2F_c^2)/3$   
 $(\Delta/\sigma)_{\text{max}} = 0.001$   
 $\Delta\rho_{\text{max}} = 0.41$  e Å<sup>-3</sup>  
 $\Delta\rho_{\text{min}} = -0.27$  e Å<sup>-3</sup>  
 Extinction correction: SHELXTL (Sheldrick, 2008),  
 $F_c^* = kF_c[1 + 0.001 \times F_c^2 \lambda^3 / \sin(2\theta)]^{-1/4}$   
 Extinction coefficient: 0.0231 (18)

### Special details

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) etc. and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

## supplementary materials

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### Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
C11	0.23790 (5)	0.15285 (5)	0.181208 (18)	0.02983 (10)
C12	0.66794 (6)	0.37010 (5)	0.233709 (17)	0.02976 (10)
C13	0.33903 (6)	0.53919 (5)	0.119806 (17)	0.02989 (10)
C2	0.62444 (19)	0.20753 (17)	0.10178 (6)	0.0179 (2)
C3	0.7632 (2)	-0.09854 (17)	0.06561 (7)	0.0224 (2)
C1	0.4692 (2)	0.31232 (17)	0.15584 (6)	0.0196 (2)
N2	0.62859 (18)	0.01089 (15)	0.10646 (5)	0.0229 (2)
N1	0.74493 (18)	0.32334 (15)	0.05999 (6)	0.0230 (2)
H1A	0.8430	0.2689	0.0305	0.028*
H1B	0.7277	0.4555	0.0614	0.028*
N3	0.8680 (2)	-0.21297 (17)	0.03320 (6)	0.0298 (2)

### Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
C11	0.02508 (16)	0.02924 (17)	0.03733 (18)	-0.00196 (11)	0.01472 (13)	0.00338 (12)
C12	0.03001 (17)	0.03173 (17)	0.02701 (16)	0.00592 (12)	-0.00034 (12)	-0.01080 (12)
C13	0.03524 (18)	0.02289 (16)	0.03307 (17)	0.01263 (12)	0.01157 (13)	0.00600 (11)
C2	0.0173 (5)	0.0184 (5)	0.0181 (5)	0.0013 (4)	0.0022 (4)	-0.0011 (4)
C3	0.0260 (6)	0.0158 (5)	0.0259 (6)	-0.0021 (4)	0.0060 (5)	0.0019 (4)
C1	0.0199 (5)	0.0175 (5)	0.0221 (5)	0.0025 (4)	0.0054 (4)	0.0012 (4)
N2	0.0269 (5)	0.0166 (5)	0.0264 (5)	0.0015 (4)	0.0096 (4)	0.0000 (4)
N1	0.0270 (5)	0.0168 (5)	0.0270 (5)	0.0010 (4)	0.0121 (4)	0.0001 (4)
N3	0.0364 (6)	0.0194 (5)	0.0355 (6)	0.0012 (4)	0.0136 (5)	-0.0025 (4)

### Geometric parameters ( $\text{\AA}$ , $^\circ$ )

C11—C1	1.7549 (12)	C2—C1	1.5396 (15)
C12—C1	1.7733 (12)	C3—N3	1.1533 (17)
C13—C1	1.7686 (12)	C3—N2	1.3226 (16)
C2—N2	1.3032 (15)	N1—H1A	0.8800
C2—N1	1.3115 (15)	N1—H1B	0.8800
N2—C2—N1	127.94 (11)	C2—C1—C12	106.31 (7)
N2—C2—C1	114.43 (10)	C11—C1—C12	109.15 (6)
N1—C2—C1	117.52 (10)	C13—C1—C12	108.98 (6)
N3—C3—N2	172.16 (13)	C2—N2—C3	121.04 (10)
C2—C1—C11	111.47 (8)	C2—N1—H1A	120.0
C2—C1—C13	111.73 (8)	C2—N1—H1B	120.0
C11—C1—C13	109.12 (6)	H1A—N1—H1B	120.0
N2—C2—C1—C11	-26.28 (12)	N2—C2—C1—C12	92.56 (10)
N1—C2—C1—C11	157.21 (9)	N1—C2—C1—C12	-83.95 (11)
N2—C2—C1—C13	-148.66 (9)	N1—C2—N2—C3	-0.9 (2)
N1—C2—C1—C13	34.83 (12)	C1—C2—N2—C3	-176.95 (10)

*Hydrogen-bond geometry* ( $\text{\AA}$ ,  $^\circ$ )

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$N1-H1A\cdots N3^i$	0.88	2.10	2.9583 (15)	164
$N1-H1B\cdots N3^{ii}$	0.88	2.40	3.1893 (15)	150

Symmetry codes: (i)  $-x+2, -y, -z$ ; (ii)  $x, y+1, z$ .

**Table 2**

*Comparative distances* ( $\text{\AA}$ ) *and angles* ( $^\circ$ ) *in amidines (I)–(V)*

Value	(I)	(II)	(III)	(IV)	(V)
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$\Delta_{CN}$	0.008	-0.012	-0.009	0.001	0.114
C2-C1	1.5396 (15)	1.520 (4)	1.513 (3)	1.522 (3)	1.525 (5)
C3-N3	1.1533 (17)	1.164 (4)	1.1567 (3)	1.153 (3)	
C3-N2	1.3226 (16)	1.320 (4)	1.333 (3)	1.322 (3)	
N2-C2-N1	127.94 (11)	118.0 (2)	125.9 (2)	126.0 (2)	127.7 (3)
N2-C2-C1	114.43 (10)	124.1 (2)	116.8 (2)	114.9 (1)	117.2 (3)
N1-C2-C1	117.52 (10)	117.9 (2)	117.3 (2)	118.7 (1)	115.0 (3)
C2-N2-C3	121.04 (10)	119.1 (2)	118.7 (1)		
N3-C3-N2	172.16 (13)	173.2 (2)	173.9 (3)	173.2 (2)	

Atom numbering corresponds to that in Fig. 1.

Fig. 1

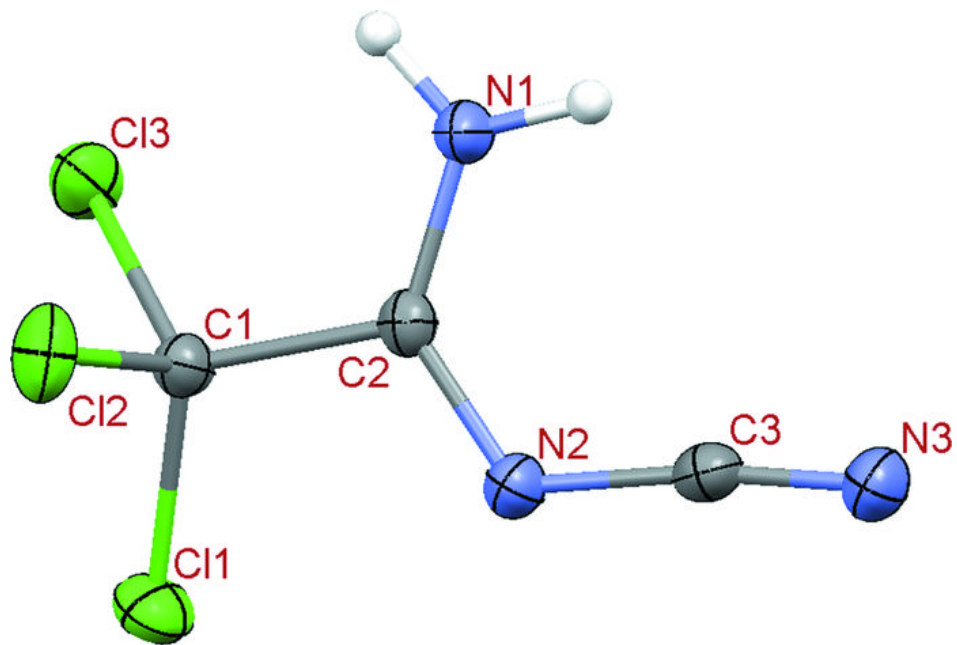


Fig. 2

