

Acta Crystallographica Section E

Structure Reports

Online

ISSN 1600-5368

2,3,5,6-Tetramethoxypiperazine-1,4-dicarbaldehyde

Sayed Mojtaba Moosavi and Amir Taheri*

Department of Chemistry, Imam Hossein University, Tehran, Iran
Correspondence e-mail: amir.tahery1@gmail.com

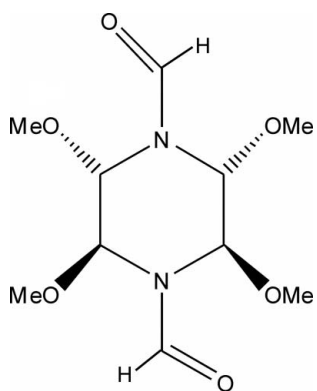
Received 24 August 2009; accepted 1 September 2009

Key indicators: single-crystal X-ray study; $T = 120$ K; mean $\sigma(\text{C}-\text{C}) = 0.004$ Å;
 R factor = 0.059; wR factor = 0.127; data-to-parameter ratio = 17.6.

The asymmetric unit of the title compound, $\text{C}_{10}\text{H}_{18}\text{N}_2\text{O}_6$, contains two halves of two independent centrosymmetric molecules with almost identical conformations. Weak intermolecular $\text{C}-\text{H}\cdots\text{O}$ hydrogen bonds consolidate the crystal packing.

Related literature

For details of the synthesis, see: Ferguson (1968*a,b*). For a closely related compound with acetyl substituents, see: Vedachalam *et al.* (1991). For anomeric interactions, see: Reed & Schleyer (1988). For glycoside structures, see: Schleifer *et al.* (1990).



Experimental

Crystal data

 $\text{C}_{10}\text{H}_{18}\text{N}_2\text{O}_6$ $M_r = 262.26$ Monoclinic, $P2_1/n$
 $a = 14.331$ (4) Å
 $b = 6.6044$ (18) Å
 $c = 14.332$ (4) Å
 $\beta = 114.800$ (3)°
 $V = 1231.4$ (6) Å³ $Z = 4$
Mo $K\alpha$ radiation
 $\mu = 0.12$ mm⁻¹
 $T = 120$ K
 $0.5 \times 0.05 \times 0.05$ mm

Data collection

Bruker SMART 1000 CCD area-
detector diffractometer
Absorption correction: multi-scan
(*SADABS*; Sheldrick, 1998)
 $T_{\min} = 0.980$, $T_{\max} = 0.995$ 12344 measured reflections
2958 independent reflections
2242 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.037$

Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.059$
 $wR(F^2) = 0.127$
 $S = 1.00$
2958 reflections168 parameters
H-atom parameters constrained
 $\Delta\rho_{\text{max}} = 0.56$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.30$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{C5}-\text{H5C}\cdots\text{O6}$	0.96	2.53	3.260 (4)	133
$\text{C3}-\text{H3A}\cdots\text{O1}^i$	0.93	2.46	3.371 (3)	167
$\text{C8}-\text{H8A}\cdots\text{O4}^{ii}$	0.93	2.47	3.364 (4)	163
$\text{C10}-\text{H10B}\cdots\text{O3}^{iii}$	0.96	2.60	3.208 (3)	122

Symmetry codes: (i) $-x + \frac{3}{2}, y + \frac{1}{2}, -z + \frac{3}{2}$; (ii) $-x + \frac{3}{2}, y + \frac{1}{2}, -z + \frac{3}{2}$; (iii) $x + \frac{1}{2}, -y + \frac{1}{2}, z + \frac{1}{2}$.

Data collection: *SMART* (Bruker, 1998); cell refinement: *SAINT-Plus* (Bruker, 1998); data reduction: *SAINT-Plus*; program(s) used to solve structure: *SHELXTL* (Sheldrick, 2008); program(s) used to refine structure: *SHELXTL*; molecular graphics: *SHELXTL*; software used to prepare material for publication: *SHELXTL*.

We thank the Chemistry Group of Imam Hossain University for their cooperation.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: CV2606).

References

- Bruker (1998). *SAINT-Plus* and *SMART*. Bruker AXS Inc., Madison, Wisconsin, USA.
- Ferguson, A. N. (1968*a*). US Patent 3 369 020.
- Ferguson, A. N. (1968*b*). US Patent 3 365 454.
- Reed, A. E. & Schleyer, P. V. R. (1988). *Inorg. Chem.* **27**, 3969–3987.
- Schleifer, L., Senderowitz, H., Aped, P., Tartakovsky, E. & Fuchs, B. (1990). *Carbohydr. Res.* **206**, 21–39.
- Sheldrick, G. M. (1998). *SADABS*. Bruker AXS Inc., Madison, Wisconsin, USA.
- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
- Vedachalam, M., Ramakrishnan, V. T., Boyer, J. H., Dagley, I. J., Nelson, K. A., Adolph, H. G., Gilardi, R., George, C. & Anderson, J. L. F. (1991). *J. Org. Chem.* **56**, 3413–3419.

supplementary materials

Acta Cryst. (2009). E65, o2338 [doi:10.1107/S1600536809035259]

2,3,5,6-Tetramethoxypiperazine-1,4-dicarbaldehyde

S. M. Moosavi and A. Taheri

Comment

The molecule of 1,4-diformyl-2,3,5,6-tetra-methoxypiperazine (Ferguson,1968a, b) (I), contains two formamyl groups in the two contrary-sides of the six-membered ring and four methoxy groups in the same axial situations. There are four asymmetric carbon atoms, S, S and R, R configuration which are mirror images of each other and diastereoisomer. The skeleton is symmetric by a central symmetry. The crystal contains two halves of two independent centrosymmetric molecules with almost identical conformations, which are comparable with similar molecule with acetyl substitutions (Vedachalam *et al.*, 1991).

Despite the presence of six O and two N atoms carrying lone-pair electrons potentially are available for hydrogen-bond formation, there are, in fact, no powerful intramolecular C—H \cdots N or C—H \cdots O hydrogen bonds, rather some weak C—H \cdots O intermolecular interactions are observed (Table 1, Fig. 2).

The Properties of the anomeric effect in the infected system, $n_{\text{O}} \rightarrow \sigma^*_{\text{C}-\text{N}}$ interactions are recognizable (Reed & Schleyer, 1988), The O—C bond is shorter than the C—N bond and the stability of *gauche* (axial) forms is definitely more than anti (equatorial) ones (Figure 1). This electron transfer results in lengthening the exocyclic anomeric C—O bonds, in the endocyclic contraction of the C—N bonds by increasing its double-bond character in the opening of the O—C—N angle as compared with its normal tetrahedral value. It can be considered that aa, ag and gg conformations of the C—O—C—N—C moieties, where a = anti (antiperiplanar) and g = *gauche* (synclinal) which aa and ag refer to equatorial and gg to axial. All of the Methoxy groups are in the gg conformations, comparable with gg in glycoside conformations that are more stable than aa, equatorial conformations (Schleifer *et al.*, 1990). The acetyl groups as substitutions select axial positions than equatorials (Vedachalam *et al.*, 1991).

In (I) *via* resonance of N-lone-pairs with *pi*-electrons of the carbonyl double-bond, N-pyramidalities (about 360°) are larger than natural forms (about 320°) at the same of molecule with acetyl substitutions (Vedachalam *et al.*, 1991).

Experimental

1,4-Diformyltetrachloropiperazine (7.04 g,0.04 mol) (Ferguson, 1968a,b) were mixed with barium carbonate (78 g) and absolute methanol (32 ml, 0.79 mol) were added to the mixture and heated at 313 K for 3 h with stirring. The barium carbonate was filtered off and the methanol solution evaporated. The solid residue was extracted by 20 ml of hot chloroform and the extract evaporated to give white precipitation and washed with a little cold water and dried in oven by 303 K of temperature affording 1.83 g (70% yield) of (I). This solid was crystallized by hot mixture of methanol and benzene (m.p. 469 K). ^1H NMR (CDCl_3): δ_{H} 8.47 (s, 2H, CH), 5.29 (d, 2H, CH), 4.82 (d, 2H, CH), 3.21 (m, 12H,CH). ^{13}C NMR (CDCl_3): δ_{C} 165 (2CH), 83.4 (2CH), 79.1(2CH), 54.5 (6CH), 56.2 (6CH).

Refinement

All H atoms were geometrically positioned (C—H 0.93 - 0.98 Å) and refined as riding, with $U_{\text{iso}}(\text{H}) = 1.2\text{--}1.5 U_{\text{eq}}(\text{C})$.

Figures

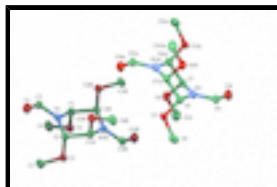


Fig. 1. Two independent molecules of (I) showing the atomic numbering and 50% probability displacement ellipsoids [symmetry codes: (A) $1 - x, 1 - y, 1 - z$; (B) $2 - x, 1 - y, 1 - z$]. H atoms omitted for clarity.

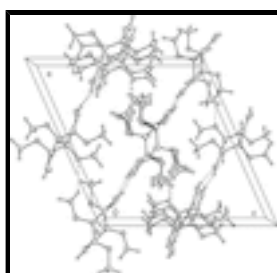


Fig. 2. A portion of the crystal packing viewed along crystallographic axis b . Dashed lines denote hydrogen bonds.

2,3,5,6-Tetramethoxypiperazine-1,4-dicarbaldehyde

Crystal data

$\text{C}_{10}\text{H}_{18}\text{N}_2\text{O}_6$

$M_r = 262.26$

Monoclinic, $P2_1/n$

Hall symbol: $-P\ 2_1/n$

$a = 14.331\ (4)\ \text{\AA}$

$b = 6.6044\ (18)\ \text{\AA}$

$c = 14.332\ (4)\ \text{\AA}$

$\beta = 114.800\ (3)^\circ$

$V = 1231.4\ (6)\ \text{\AA}^3$

$Z = 4$

$F_{000} = 560$

$D_x = 1.415\ \text{Mg m}^{-3}$

Mo $K\alpha$ radiation, $\lambda = 0.71073\ \text{\AA}$

Cell parameters from 1254 reflections

$\theta = 2\text{--}25^\circ$

$\mu = 0.12\ \text{mm}^{-1}$

$T = 120\ \text{K}$

Needle, colourless

$0.5 \times 0.05 \times 0.05\ \text{mm}$

Data collection

Bruker SMART 1000 CCD area-detector diffractometer

Radiation source: fine-focus sealed tube

Monochromator: graphite

$T = 120\ \text{K}$

φ and ω scans

Absorption correction: multi-scan (SADABS; Sheldrick, 1998)

2958 independent reflections

2242 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.037$

$\theta_{\text{max}} = 28.0^\circ$

$\theta_{\text{min}} = 1.6^\circ$

$h = -18 \rightarrow 18$

$T_{\min} = 0.980$, $T_{\max} = 0.995$
12344 measured reflections

$k = -8 \rightarrow 8$
 $l = -18 \rightarrow 18$

Refinement

Refinement on F^2

Secondary atom site location: difference Fourier map

Least-squares matrix: full

Hydrogen site location: inferred from neighbouring sites

$R[F^2 > 2\sigma(F^2)] = 0.059$

H-atom parameters constrained

$wR(F^2) = 0.127$

$w = 1/[\sigma^2(F_o^2) + (0.0108P)^2 + 2.6266P]$
where $P = (F_o^2 + 2F_c^2)/3$

$S = 1.00$

$(\Delta/\sigma)_{\max} = 0.001$

2958 reflections

$\Delta\rho_{\max} = 0.56 \text{ e } \text{Å}^{-3}$

168 parameters

$\Delta\rho_{\min} = -0.29 \text{ e } \text{Å}^{-3}$

Primary atom site location: structure-invariant direct methods

Extinction correction: none

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
O1	0.67186 (15)	0.1494 (3)	0.66884 (16)	0.0354 (5)
O2	0.62334 (13)	0.7282 (3)	0.49476 (14)	0.0279 (4)
O3	0.55403 (13)	0.3233 (3)	0.39458 (14)	0.0291 (4)
O4	1.17420 (15)	0.1502 (3)	0.66699 (16)	0.0354 (5)
O5	0.99381 (14)	0.7197 (3)	0.62522 (15)	0.0311 (5)
O6	0.89855 (13)	0.3088 (3)	0.55478 (15)	0.0322 (5)
N1	0.59356 (15)	0.4210 (3)	0.56729 (16)	0.0237 (5)
N2	1.06918 (15)	0.4199 (4)	0.59207 (17)	0.0245 (5)
C1	0.58449 (18)	0.6410 (4)	0.5610 (2)	0.0239 (5)
H1A	0.6218	0.6977	0.6300	0.029*
C2	0.52907 (18)	0.2971 (4)	0.47925 (19)	0.0245 (5)
H2A	0.5365	0.1541	0.4993	0.029*
C3	0.66071 (19)	0.3326 (4)	0.6564 (2)	0.0260 (6)
H3A	0.7002	0.4165	0.7108	0.031*
C4	0.73438 (19)	0.7406 (5)	0.5418 (2)	0.0332 (7)
H4A	0.7577	0.7963	0.4934	0.050*

supplementary materials

H4B	0.7628	0.6077	0.5618	0.050*
H4C	0.7563	0.8262	0.6013	0.050*
C5	0.6535 (2)	0.2490 (5)	0.4132 (2)	0.0360 (7)
H5A	0.6669	0.2731	0.3539	0.054*
H5B	0.6565	0.1063	0.4268	0.054*
H5C	0.7041	0.3176	0.4715	0.054*
C6	1.06031 (19)	0.6398 (4)	0.5843 (2)	0.0259 (6)
H6A	1.1287	0.6991	0.6214	0.031*
C7	0.98182 (18)	0.2938 (4)	0.5274 (2)	0.0268 (6)
H7A	1.0040	0.1522	0.5327	0.032*
C8	1.16039 (18)	0.3332 (4)	0.6578 (2)	0.0259 (6)
H8A	1.2143	0.4181	0.6972	0.031*
C9	1.0398 (2)	0.7262 (5)	0.7347 (2)	0.0384 (7)
H9A	0.9897	0.7693	0.7587	0.058*
H9B	1.0962	0.8199	0.7579	0.058*
H9C	1.0644	0.5939	0.7614	0.058*
C10	0.9185 (2)	0.2250 (6)	0.6521 (2)	0.0412 (8)
H10A	0.8587	0.2393	0.6657	0.062*
H10B	0.9753	0.2948	0.7041	0.062*
H10C	0.9351	0.0841	0.6526	0.062*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
O1	0.0309 (10)	0.0386 (13)	0.0360 (11)	0.0059 (9)	0.0134 (9)	0.0068 (9)
O2	0.0191 (8)	0.0373 (11)	0.0310 (10)	-0.0045 (8)	0.0141 (7)	0.0020 (8)
O3	0.0198 (8)	0.0421 (12)	0.0274 (9)	0.0021 (8)	0.0119 (7)	-0.0010 (9)
O4	0.0257 (9)	0.0393 (13)	0.0407 (12)	0.0066 (9)	0.0134 (9)	0.0067 (10)
O5	0.0205 (8)	0.0430 (12)	0.0334 (10)	0.0052 (8)	0.0148 (8)	-0.0060 (9)
O6	0.0167 (8)	0.0475 (13)	0.0341 (10)	0.0017 (8)	0.0124 (8)	0.0062 (9)
N1	0.0180 (9)	0.0285 (12)	0.0258 (11)	-0.0011 (8)	0.0103 (8)	0.0003 (9)
N2	0.0152 (9)	0.0295 (12)	0.0284 (11)	0.0010 (8)	0.0088 (8)	-0.0010 (9)
C1	0.0184 (11)	0.0290 (14)	0.0252 (12)	-0.0023 (10)	0.0099 (10)	-0.0007 (11)
C2	0.0177 (11)	0.0320 (15)	0.0256 (12)	-0.0008 (10)	0.0108 (10)	0.0005 (11)
C3	0.0184 (11)	0.0350 (16)	0.0260 (13)	0.0011 (11)	0.0107 (10)	0.0025 (11)
C4	0.0189 (12)	0.0384 (17)	0.0462 (17)	-0.0040 (11)	0.0175 (12)	-0.0005 (14)
C5	0.0220 (13)	0.0527 (19)	0.0379 (16)	0.0055 (13)	0.0172 (12)	-0.0032 (14)
C6	0.0171 (11)	0.0305 (15)	0.0312 (14)	0.0003 (10)	0.0112 (10)	-0.0049 (11)
C7	0.0172 (11)	0.0332 (15)	0.0313 (14)	0.0010 (10)	0.0113 (10)	-0.0033 (11)
C8	0.0153 (11)	0.0367 (16)	0.0275 (13)	0.0023 (10)	0.0107 (10)	0.0042 (12)
C9	0.0395 (16)	0.0463 (19)	0.0335 (15)	0.0073 (14)	0.0193 (13)	-0.0043 (14)
C10	0.0279 (14)	0.056 (2)	0.0452 (18)	0.0047 (14)	0.0203 (13)	0.0170 (16)

Geometric parameters (\AA , $^\circ$)

O1—C3	1.224 (3)	C2—H2A	0.9800
O2—C1	1.409 (3)	C3—H3A	0.9300
O2—C4	1.447 (3)	C4—H4A	0.9600
O3—C2	1.412 (3)	C4—H4B	0.9600

O3—C5	1.423 (3)	C4—H4C	0.9600
O4—C8	1.223 (3)	C5—H5A	0.9600
O5—C6	1.414 (3)	C5—H5B	0.9600
O5—C9	1.425 (3)	C5—H5C	0.9600
O6—C7	1.408 (3)	C6—C7 ⁱⁱ	1.519 (4)
O6—C10	1.414 (3)	C6—H6A	0.9800
N1—C3	1.367 (3)	C7—C6 ⁱⁱ	1.519 (4)
N1—C1	1.458 (4)	C7—H7A	0.9800
N1—C2	1.462 (3)	C8—H8A	0.9300
N2—C8	1.375 (3)	C9—H9A	0.9600
N2—C6	1.458 (4)	C9—H9B	0.9600
N2—C7	1.465 (3)	C9—H9C	0.9600
C1—C2 ⁱ	1.537 (3)	C10—H10A	0.9600
C1—H1A	0.9800	C10—H10B	0.9600
C2—C1 ⁱ	1.537 (3)	C10—H10C	0.9600
C1—O2—C4	112.2 (2)	O3—C5—H5B	109.5
C2—O3—C5	113.0 (2)	H5A—C5—H5B	109.5
C6—O5—C9	112.8 (2)	O3—C5—H5C	109.5
C7—O6—C10	113.8 (2)	H5A—C5—H5C	109.5
C3—N1—C1	119.6 (2)	H5B—C5—H5C	109.5
C3—N1—C2	120.6 (2)	O5—C6—N2	113.1 (2)
C1—N1—C2	119.8 (2)	O5—C6—C7 ⁱⁱ	106.9 (2)
C8—N2—C6	119.6 (2)	N2—C6—C7 ⁱⁱ	110.5 (2)
C8—N2—C7	120.7 (2)	O5—C6—H6A	108.7
C6—N2—C7	119.6 (2)	N2—C6—H6A	108.7
O2—C1—N1	113.5 (2)	C7 ⁱⁱ —C6—H6A	108.7
O2—C1—C2 ⁱ	107.1 (2)	O6—C7—N2	112.5 (2)
N1—C1—C2 ⁱ	109.9 (2)	O6—C7—C6 ⁱⁱ	105.4 (2)
O2—C1—H1A	108.7	N2—C7—C6 ⁱⁱ	111.1 (2)
N1—C1—H1A	108.7	O6—C7—H7A	109.3
C2 ⁱ —C1—H1A	108.7	N2—C7—H7A	109.3
O3—C2—N1	112.1 (2)	C6 ⁱⁱ —C7—H7A	109.3
O3—C2—C1 ⁱ	104.8 (2)	O4—C8—N2	123.4 (3)
N1—C2—C1 ⁱ	111.2 (2)	O4—C8—H8A	118.3
O3—C2—H2A	109.5	N2—C8—H8A	118.3
N1—C2—H2A	109.5	O5—C9—H9A	109.5
C1 ⁱ —C2—H2A	109.5	O5—C9—H9B	109.5
O1—C3—N1	123.8 (3)	H9A—C9—H9B	109.5
O1—C3—H3A	118.1	O5—C9—H9C	109.5
N1—C3—H3A	118.1	H9A—C9—H9C	109.5
O2—C4—H4A	109.5	H9B—C9—H9C	109.5
O2—C4—H4B	109.5	O6—C10—H10A	109.5
H4A—C4—H4B	109.5	O6—C10—H10B	109.5
O2—C4—H4C	109.5	H10A—C10—H10B	109.5
H4A—C4—H4C	109.5	O6—C10—H10C	109.5
H4B—C4—H4C	109.5	H10A—C10—H10C	109.5

supplementary materials

O3—C5—H5A	109.5	H10B—C10—H10C	109.5
C4—O2—C1—N1	-78.1 (3)	C9—O5—C6—N2	77.0 (3)
C4—O2—C1—C2 ⁱ	160.4 (2)	C9—O5—C6—C7 ⁱⁱ	-161.1 (2)
C3—N1—C1—O2	110.9 (3)	C8—N2—C6—O5	-111.3 (2)
C2—N1—C1—O2	-70.5 (3)	C7—N2—C6—O5	70.6 (3)
C3—N1—C1—C2 ⁱ	-129.1 (2)	C8—N2—C6—C7 ⁱⁱ	128.9 (2)
C2—N1—C1—C2 ⁱ	49.4 (3)	C7—N2—C6—C7 ⁱⁱ	-49.1 (3)
C5—O3—C2—N1	66.7 (3)	C10—O6—C7—N2	-67.6 (3)
C5—O3—C2—C1 ⁱ	-172.6 (2)	C10—O6—C7—C6 ⁱⁱ	171.2 (3)
C3—N1—C2—O3	-114.5 (3)	C8—N2—C7—O6	113.5 (3)
C1—N1—C2—O3	67.0 (3)	C6—N2—C7—O6	-68.5 (3)
C3—N1—C2—C1 ⁱ	128.5 (2)	C8—N2—C7—C6 ⁱⁱ	-128.7 (2)
C1—N1—C2—C1 ⁱ	-50.0 (3)	C6—N2—C7—C6 ⁱⁱ	49.4 (3)
C1—N1—C3—O1	178.7 (2)	C6—N2—C8—O4	-178.6 (2)
C2—N1—C3—O1	0.2 (4)	C7—N2—C8—O4	-0.5 (4)

Symmetry codes: (i) $-x+1, -y+1, -z+1$; (ii) $-x+2, -y+1, -z+1$.

Hydrogen-bond geometry ($\text{\AA}, ^\circ$)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
C5—H5C \cdots O6	0.96	2.53	3.260 (4)	133
C3—H3A \cdots O1 ⁱⁱⁱ	0.93	2.46	3.371 (3)	167
C8—H8A \cdots O4 ^{iv}	0.93	2.47	3.364 (4)	163
C10—H10B \cdots O3 ^v	0.96	2.60	3.208 (3)	122

Symmetry codes: (iii) $-x+3/2, y+1/2, -z+3/2$; (iv) $-x+5/2, y+1/2, -z+3/2$; (v) $x+1/2, -y+1/2, z+1/2$.

Fig. 1

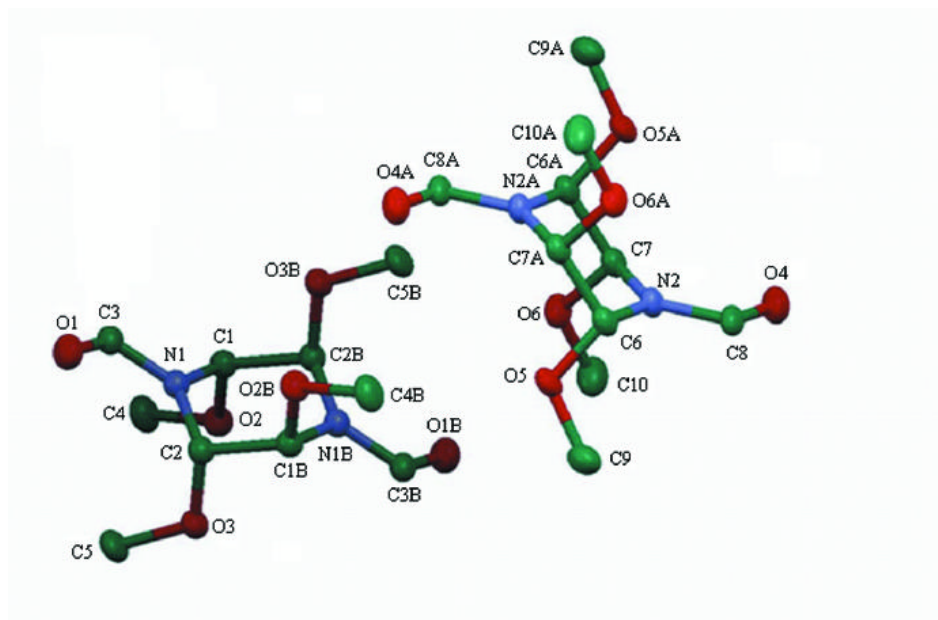


Fig. 2

