Acta Crystallographica Section E
Structure Reports Online
ISSN 1600-5368

5,6-Dihydroxy-1,10-phenanthrolin-1-ium chloride dihydrate

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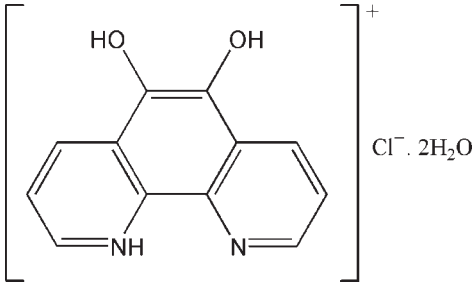
Received 4 August 2009; accepted 24 August 2009

Key indicators: single-crystal X-ray study; T = 296 K; mean σ(C–C) = 0.002 Å; R factor = 0.036; wR factor = 0.111; data-to-parameter ratio = 16.0.

The title compound, $C_{12}H_9N_2O_2^+\cdot Cl^-\cdot 2H_2O$, exhibits a layered structure which is stabilized by intermolecular O—H···O, O—H···Cl^- and N^+—H···Cl^- hydrogen bonds, and π–π interactions (centroid–centroid distances = 3.654 and 3.583 Å). The distances between the molecules are 3.371 and 3.294 Å.

Related literature

For a related structure, see: Borel & Bond (2008).



Experimental

Crystal data

$C_{12}H_9N_2O_2^+\cdot Cl^-\cdot 2H_2O$
M_r = 284.69
Triclinic, $P\overline{1}$
a = 7.7627 (1) Å
b = 8.6974 (1) Å
c = 9.6432 (1) Å
α = 86.116 (1)°
β = 86.859 (1)°
γ = 74.580 (1)°
V = 625.73 (1) Å^3
Z = 2
Mo $K\alpha$ radiation
μ = 0.32 mm^{-1}
T = 296 K
0.25 × 0.12 × 0.03 mm

Data collection

Bruker P4 diffractometer
Absorption correction: none
9780 measured reflections
2872 independent reflections
2336 reflections with $I > 2\sigma(I)$
R_{int} = 0.023

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.036
$wR(F^2)$ = 0.111
S = 1.06
2872 reflections
180 parameters
H atoms treated by a mixture of independent and constrained refinement
$\Delta\rho_{max}$ = 0.29 e Å^{-3}
$\Delta\rho_{min}$ = −0.22 e Å^{-3}

Table 1
Hydrogen-bond geometry (Å, °).

D—H···A	D—H	H···A	D···A	D—H···A
O1—H01···O4	0.82	1.86	2.6782 (18)	179
O2—H02···O3	0.82	1.89	2.6669 (18)	157
O3—H03B···Cl1	0.81 (3)	2.40 (3)	3.2133 (15)	174 (3)
O3—H03A···Cl1^{i}	0.78 (3)	2.45 (3)	3.2323 (15)	176 (3)
O4—H04B···Cl1	0.91 (3)	2.33 (3)	3.2185 (14)	165 (3)
O4—H04A···Cl1^{ii}	0.94 (3)	2.30 (3)	3.2216 (14)	168 (3)
N2—H9···Cl1^{iii}	0.86	2.37	3.1635 (13)	153

Symmetry codes: (i) $-x+1, -y+1, -z$; (ii) $-x, -y+1, -z$; (iii) $x, y-1, z+1$.

Data collection: *XSCANS* (Bruker, 1999); cell refinement: *XSCANS*); data reduction: *SHELXTL* (Sheldrick, 2008); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL*; software used to prepare material for publication: *SHELXTL*.

We are grateful for financial support from the National Science Foundation of Fujian Province of China (No. E0610017, 2003 F006).

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HG2548).

References

Borel, C. & Bond, A. D. (2008). *Acta Cryst.* E**64**, o34.
Bruker (1999). *XSCANS*. Bruker AXS Inc., Madison, Wisconsin, USA.
Sheldrick, G. M. (2008). *Acta Cryst.* A**64**, 112–122.

supplementary materials

Acta Cryst. (2009). E**65**, o2367 [**doi:10.1107/S1600536809033777**]

5,6-Dihydroxy-1,10-phenanthrolin-1-ium chloride dihydrate

X.-Y. Lin, S.-J. Tang and W.-S. Wu

Comment

The title compound, $[C_{12}H_9N_2O_2]Cl.2H_2O$, was obtained unintentionally as the product of an attempted synthesis of a condensation product between 1,10-phenanthroline-5,6-dione and picoloylhydrazide. Compared with a similar compound (5,6-dioxo-1,10-phenanthrolin-1-ium chloride) reported (Borel & Bond, 2008), the Cl—H distance is slightly longer (2.372 vs 2.274Å).

The structure of the title compound is shown in Fig. 1. It exhibits a layered structure which is stabilized by inter-molecular O—H···O, O—H···Cl^-, N^+—H···Cl^- hydrogen bonds, detailed in Fig. 2 and Table 1, as well as π-π interactions and C—H···O, C—H···Cl^- interactions. With Cl^- as the connecting point, it occurs two different shape parallelograms made up of O and Cl^-. The dihedral angle between the two planes, which possess different shapes, is 78.67°. The distances between the layers, which belong to offset face to face, are 3.371 Å and 3.294 Å, reflecting π-π interactions.

Experimental

1,10-phenanthroline-5,6-dione (300 mg, 1.53 mmol) was dissolved in a mixed solution of 10 ml CH_2Cl_2 and 30 ml EtOH when heating with stirring. When all of the compound dissolved, picoloylhydrazide (200 mg, 1.46 mmol) was added and refluxed 8hrs. Then HCl(aq) was added until the pH was 6. Red crystals of the title compound were obtained by slow evaporation of solvent at room temperature. Analysis: Found C 50.45, H 4.82, N 9.71%, calc. for $C_{12}H_{13}ClN_2O_4$, C 50.63, H 4.60, N 9.84%.

Refinement

The positions of the O1- , O2- and N2-bound H atoms were placed at fixed positions and refined accord to the riding model. O3- and O4-bound H atoms were located in a difference Fourier map and refined freely. The C-bound H atoms were included in the riding model approximation with C—H = 0.93 Å and U_{iso} of each H atom = $1.2U_{eq}$(C).

Figures

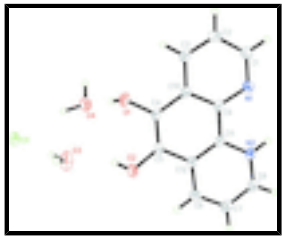


Fig. 1. View of the title compound showing the atom-labelling scheme. Displacement ellipsoids are drawn at the 30% probability level. H atoms are represented by circles of arbitrary radius.

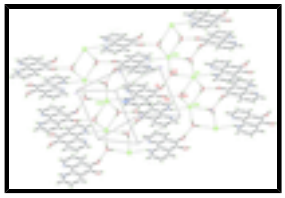


Fig. 2. Crystal Packing diagram of the title compound, showing the H-bonded interactions (dashed lines). Cl11, Cl12, Cl13 represent $Cl1^{i}$, $Cl1^{ii}$, $Cl1^{iii}$, respectively.

5,6-Dihydroxy-1,10-phenanthrolin-1-ium chloride dihydrate

Crystal data

$C_{12}H_9N_2O_2^+ \cdot Cl^- \cdot 2H_2O$	$Z = 2$
$M_r = 284.69$	$F_{000} = 296$
Triclinic, $P\overline{1}$	$D_x = 1.511$ Mg m^{-3}
Hall symbol: -P 1	Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å
$a = 7.7627$ (1) Å	Cell parameters from 2922 reflections
$b = 8.6974$ (1) Å	$\theta = 2.1$–27.7°
$c = 9.6432$ (1) Å	$\mu = 0.32$ mm^{-1}
$\alpha = 86.116$ (1)°	$T = 296$ K
$\beta = 86.859$ (1)°	Block, red
$\gamma = 74.580$ (1)°	0.25 × 0.12 × 0.03 mm
$V = 625.728$ (13) Å^3	

Data collection

Bruker P4 diffractometer	2872 independent reflections
Radiation source: fine-focus sealed tube	2336 reflections with $I > 2\sigma(I)$
Monochromator: graphite	$R_{int} = 0.023$
Detector resolution: 0 pixels mm^{-1}	$\theta_{max} = 27.7^\circ$
$T = 296$ K	$\theta_{min} = 2.1^\circ$
ω scans	$h = -9 \rightarrow 9$
Absorption correction: none	$k = -11 \rightarrow 11$
9780 measured reflections	$l = -12 \rightarrow 12$

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.036$	H atoms treated by a mixture of independent and constrained refinement
$wR(F^2) = 0.111$	$w = 1/[\sigma^2(F_o^2) + (0.0638P)^2 + 0.0835P]$ where $P = (F_o^2 + 2F_c^2)/3$
$S = 1.06$	$(\Delta/\sigma)_{max} < 0.001$
2872 reflections	$\Delta\rho_{max} = 0.29$ e Å^{-3}
180 parameters	$\Delta\rho_{min} = -0.22$ e Å^{-3}

Primary atom site location: structure-invariant direct methods

Extinction correction: none

Special details

Geometry. All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted *R*-factor *wR* and goodness of fit *S* are based on F^2, conventional *R*-factors *R* are based on *F*, with *F* set to zero for negative F^2. The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating *R*-factors(gt) *etc*. and is not relevant to the choice of reflections for refinement. *R*-factors based on F^2 are statistically about twice as large as those based on *F*, and *R*- factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	*y*	*z*	U_{iso}*/U_{eq}
Cl1	0.19595 (6)	0.69904 (5)	0.01405 (4)	0.04841 (15)
O1	0.30334 (16)	0.13935 (13)	0.25607 (10)	0.0444 (3)
H01	0.2307	0.2235	0.2339	0.067*
O2	0.37121 (17)	0.34384 (13)	0.45497 (11)	0.0463 (3)
H02	0.3972	0.3484	0.3715	0.069*
O3	0.4713 (2)	0.4352 (2)	0.20054 (15)	0.0614 (4)
H03B	0.404 (3)	0.497 (3)	0.148 (3)	0.078 (8)*
H03A	0.555 (4)	0.400 (3)	0.152 (3)	0.086 (9)*
O4	0.06665 (19)	0.41360 (17)	0.18085 (14)	0.0554 (3)
H04B	0.089 (4)	0.492 (4)	0.121 (3)	0.105 (9)*
H04A	−0.024 (4)	0.389 (4)	0.133 (3)	0.101 (9)*
N1	0.18014 (17)	−0.20833 (15)	0.63900 (12)	0.0357 (3)
N2	0.24864 (16)	−0.00702 (15)	0.82063 (12)	0.0351 (3)
H9	0.2179	−0.0925	0.8478	0.042*
C1	0.1482 (2)	−0.30386 (18)	0.54877 (16)	0.0392 (3)
H1	0.1168	−0.3958	0.5828	0.047*
C2	0.1591 (2)	−0.27406 (19)	0.40435 (16)	0.0394 (3)
H2	0.1357	−0.3454	0.3451	0.047*
C3	0.2039 (2)	−0.14094 (18)	0.35128 (15)	0.0358 (3)
H3	0.2101	−0.1197	0.2556	0.043*
C4	0.28852 (19)	0.10836 (17)	0.39662 (13)	0.0321 (3)
C5	0.32397 (19)	0.20722 (16)	0.48912 (14)	0.0318 (3)
C6	0.3461 (2)	0.26396 (19)	0.73659 (16)	0.0389 (3)
H6	0.3798	0.3567	0.7090	0.047*
C7	0.3317 (2)	0.2208 (2)	0.87618 (16)	0.0456 (4)
H7	0.3559	0.2838	0.9428	0.055*
C8	0.2811 (2)	0.0834 (2)	0.91623 (15)	0.0436 (4)
H8	0.2696	0.0541	1.0102	0.052*
C9	0.22636 (18)	−0.07646 (16)	0.58613 (13)	0.0298 (3)
C10	0.24091 (17)	−0.03478 (16)	0.44305 (13)	0.0298 (3)
C11	0.26175 (17)	0.02913 (16)	0.68159 (13)	0.0299 (3)

C12	0.31023 (18)	0.16851 (16)	0.63634 (14)	0.0309 (3)

Atomic displacement parameters (Å^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Cl1	0.0659 (3)	0.0498 (3)	0.0337 (2)	−0.0246 (2)	−0.00378 (17)	0.00808 (16)
O1	0.0661 (8)	0.0399 (6)	0.0248 (5)	−0.0116 (5)	0.0009 (5)	0.0034 (4)
O2	0.0697 (8)	0.0360 (6)	0.0367 (6)	−0.0223 (5)	0.0041 (5)	0.0027 (5)
O3	0.0576 (9)	0.0664 (9)	0.0517 (8)	−0.0087 (7)	0.0078 (7)	0.0172 (7)
O4	0.0623 (8)	0.0580 (8)	0.0512 (7)	−0.0280 (6)	−0.0123 (6)	0.0164 (6)
N1	0.0414 (7)	0.0340 (6)	0.0329 (6)	−0.0132 (5)	0.0009 (5)	0.0003 (5)
N2	0.0447 (7)	0.0361 (7)	0.0263 (6)	−0.0146 (5)	0.0005 (5)	0.0019 (5)
C1	0.0455 (9)	0.0327 (8)	0.0410 (8)	−0.0138 (6)	0.0024 (6)	−0.0018 (6)
C2	0.0442 (9)	0.0366 (8)	0.0392 (8)	−0.0115 (6)	−0.0011 (6)	−0.0112 (6)
C3	0.0396 (8)	0.0382 (8)	0.0283 (7)	−0.0074 (6)	−0.0011 (5)	−0.0036 (6)
C4	0.0369 (7)	0.0333 (7)	0.0235 (6)	−0.0061 (5)	0.0002 (5)	0.0021 (5)
C5	0.0347 (7)	0.0279 (7)	0.0318 (7)	−0.0078 (5)	0.0000 (5)	0.0037 (5)
C6	0.0454 (9)	0.0364 (8)	0.0380 (8)	−0.0160 (6)	−0.0007 (6)	−0.0045 (6)
C7	0.0586 (10)	0.0495 (10)	0.0343 (8)	−0.0218 (8)	−0.0033 (7)	−0.0102 (7)
C8	0.0556 (10)	0.0525 (10)	0.0251 (7)	−0.0183 (7)	−0.0004 (6)	−0.0029 (6)
C9	0.0306 (7)	0.0306 (7)	0.0273 (6)	−0.0070 (5)	−0.0004 (5)	0.0002 (5)
C10	0.0297 (7)	0.0313 (7)	0.0271 (6)	−0.0056 (5)	−0.0008 (5)	−0.0019 (5)
C11	0.0297 (7)	0.0335 (7)	0.0256 (6)	−0.0072 (5)	−0.0012 (5)	0.0003 (5)
C12	0.0305 (7)	0.0323 (7)	0.0296 (7)	−0.0081 (5)	−0.0005 (5)	−0.0006 (5)

Geometric parameters (Å, °)

O1—C4	1.3685 (16)	C2—H2	0.9300
O1—H01	0.8200	C3—C10	1.414 (2)
O2—C5	1.3491 (18)	C3—H3	0.9300
O2—H02	0.8200	C4—C5	1.367 (2)
O3—H03B	0.81 (3)	C4—C10	1.428 (2)
O3—H03A	0.78 (3)	C5—C12	1.4415 (18)
O4—H04B	0.91 (3)	C6—C7	1.380 (2)
O4—H04A	0.94 (3)	C6—C12	1.401 (2)
N1—C1	1.3192 (19)	C6—H6	0.9300
N1—C9	1.3504 (18)	C7—C8	1.379 (2)
N2—C8	1.327 (2)	C7—H7	0.9300
N2—C11	1.3608 (17)	C8—H8	0.9300
N2—H9	0.8600	C9—C10	1.4088 (18)
C1—C2	1.401 (2)	C9—C11	1.4293 (19)
C1—H1	0.9300	C11—C12	1.3980 (19)
C2—C3	1.355 (2)		
C4—O1—H01	109.5	C4—C5—C12	119.73 (13)
C5—O2—H02	109.5	C7—C6—C12	120.17 (14)
H03B—O3—H03A	103 (2)	C7—C6—H6	119.9
H04B—O4—H04A	99 (2)	C12—C6—H6	119.9
C1—N1—C9	116.77 (12)	C8—C7—C6	119.53 (14)

C8—N2—C11	123.10 (13)	C8—C7—H7	120.2
C8—N2—H9	118.4	C6—C7—H7	120.2
C11—N2—H9	118.4	N2—C8—C7	119.93 (14)
N1—C1—C2	123.39 (14)	N2—C8—H8	120.0
N1—C1—H1	118.3	C7—C8—H8	120.0
C2—C1—H1	118.3	N1—C9—C10	124.54 (12)
C3—C2—C1	119.85 (14)	N1—C9—C11	117.94 (12)
C3—C2—H2	120.1	C10—C9—C11	117.52 (12)
C1—C2—H2	120.1	C9—C10—C3	116.20 (13)
C2—C3—C10	119.24 (13)	C9—C10—C4	120.66 (12)
C2—C3—H3	120.4	C3—C10—C4	123.13 (12)
C10—C3—H3	120.4	N2—C11—C12	118.90 (12)
C5—C4—O1	121.50 (13)	N2—C11—C9	119.19 (12)
C5—C4—C10	121.14 (12)	C12—C11—C9	121.91 (12)
O1—C4—C10	117.32 (12)	C11—C12—C6	118.36 (13)
O2—C5—C4	125.29 (12)	C11—C12—C5	119.02 (12)
O2—C5—C12	114.97 (12)	C6—C12—C5	122.62 (13)
C9—N1—C1—C2	−0.4 (2)	O1—C4—C10—C9	178.79 (12)
N1—C1—C2—C3	−0.2 (2)	C5—C4—C10—C3	−179.72 (13)
C1—C2—C3—C10	0.7 (2)	O1—C4—C10—C3	−2.0 (2)
O1—C4—C5—O2	1.8 (2)	C8—N2—C11—C12	−0.5 (2)
C10—C4—C5—O2	179.45 (13)	C8—N2—C11—C9	179.24 (13)
O1—C4—C5—C12	−178.75 (12)	N1—C9—C11—N2	0.33 (19)
C10—C4—C5—C12	−1.1 (2)	C10—C9—C11—N2	179.86 (12)
C12—C6—C7—C8	−0.3 (3)	N1—C9—C11—C12	−179.92 (12)
C11—N2—C8—C7	−0.4 (2)	C10—C9—C11—C12	−0.4 (2)
C6—C7—C8—N2	0.8 (3)	N2—C11—C12—C6	1.0 (2)
C1—N1—C9—C10	0.5 (2)	C9—C11—C12—C6	−178.75 (13)
C1—N1—C9—C11	−179.99 (13)	N2—C11—C12—C5	−179.91 (12)
N1—C9—C10—C3	−0.1 (2)	C9—C11—C12—C5	0.3 (2)
C11—C9—C10—C3	−179.56 (12)	C7—C6—C12—C11	−0.6 (2)
N1—C9—C10—C4	179.20 (12)	C7—C6—C12—C5	−179.68 (14)
C11—C9—C10—C4	−0.29 (19)	O2—C5—C12—C11	179.92 (12)
C2—C3—C10—C9	−0.5 (2)	C4—C5—C12—C11	0.4 (2)
C2—C3—C10—C4	−179.81 (14)	O2—C5—C12—C6	−1.0 (2)
C5—C4—C10—C9	1.1 (2)	C4—C5—C12—C6	179.47 (13)

Hydrogen-bond geometry (Å, °)

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
O1—H01···O4	0.82	1.86	2.6782 (18)	179
O2—H02···O3	0.82	1.89	2.6669 (18)	157
O3—H03B···Cl1	0.81 (3)	2.40 (3)	3.2133 (15)	174 (3)
O3—H03A···Cl1^{i}	0.78 (3)	2.45 (3)	3.2323 (15)	176 (3)
O4—H04B···Cl1	0.91 (3)	2.33 (3)	3.2185 (14)	165 (3)
O4—H04A···Cl1^{ii}	0.94 (3)	2.30 (3)	3.2216 (14)	168 (3)
N2—H9···Cl1^{iii}	0.86	2.37	3.1635 (13)	153

Symmetry codes: (i) −*x*+1, −*y*+1, −*z*; (ii) −*x*, −*y*+1, −*z*; (iii) *x*, *y*−1, *z*+1.

Fig. 1

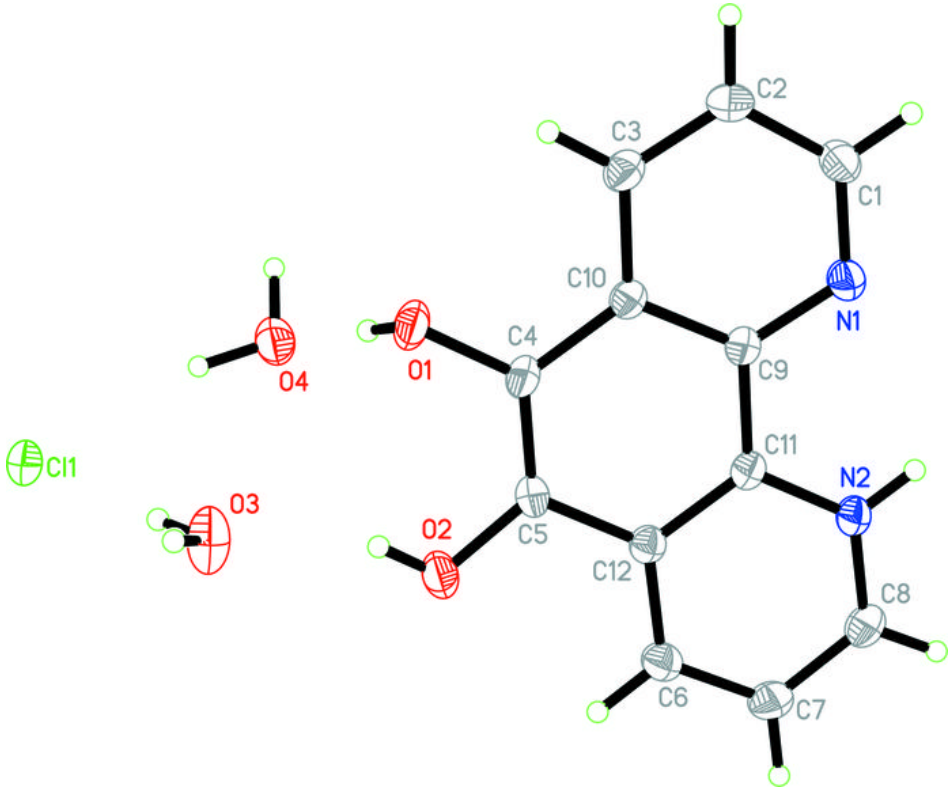


Fig. 2