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# Tetraaquabis(6-carboxy-1*H*-benzimidazole-5-carboxylato- $\kappa$ N<sup>3</sup>)nickel(II) dimethylformamide disolvate dihydrate

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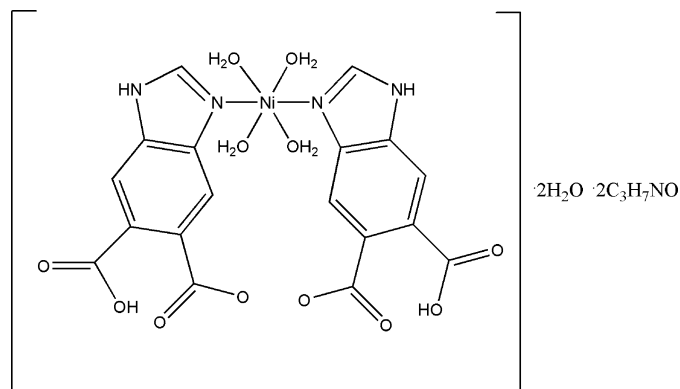
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Key indicators: single-crystal X-ray study;  $T = 293$  K; mean  $\sigma(\text{C}-\text{C}) = 0.004$  Å;  $R$  factor = 0.037;  $wR$  factor = 0.137; data-to-parameter ratio = 12.6.

The title compound,  $[\text{Ni}(\text{C}_9\text{H}_4\text{N}_2\text{O}_4)_2(\text{H}_2\text{O})_4] \cdot 2\text{C}_3\text{H}_7\text{NO} \cdot 2\text{H}_2\text{O}$ , has the Ni<sup>II</sup> center coordinated by four water molecules and two N atoms from two 1*H*-benzimidazole-5,6-dicarboxylate ligands in an octahedral geometry. The molecule interacts with the solvent water and dimethylformamide molecules through N—H...O and O—H...O hydrogen bonds to form a three-dimensional supramolecular network. The metal atom lies on a center of inversion.

## Related literature

For the crystal structures of 1*H*-benzimidazole-5,6-dicarboxylate complexes, see: Gao *et al.* (2008); Lo *et al.* (2007); Song *et al.* (2009).



## Experimental

### Crystal data

$[\text{Ni}(\text{C}_9\text{H}_4\text{N}_2\text{O}_4)_2(\text{H}_2\text{O})_4] \cdot 2\text{C}_3\text{H}_7\text{NO} \cdot 2\text{H}_2\text{O}$   
 $M_r = 723.30$   
 Triclinic,  $P\bar{1}$   
 $a = 8.5327$  (17) Å  
 $b = 9.1387$  (18) Å  
 $c = 11.624$  (2) Å  
 $\alpha = 100.80$  (3)°  
 $\beta = 103.03$  (3)°  
 $\gamma = 114.04$  (3)°  
 $V = 765.7$  (3) Å<sup>3</sup>  
 $Z = 1$   
 Mo  $K\alpha$  radiation  
 $\mu = 0.72$  mm<sup>-1</sup>  
 $T = 293$  K  
 $0.27 \times 0.18 \times 0.17$  mm

### Data collection

Rigaku/MSM Mercury CCD diffractometer  
 Absorption correction: multi-scan (*REQAB*; Jacobson, 1998)  
 $T_{\min} = 0.830$ ,  $T_{\max} = 0.888$   
 6116 measured reflections  
 2737 independent reflections  
 2613 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.020$

### Refinement

$R[F^2 > 2\sigma(F^2)] = 0.037$   
 $wR(F^2) = 0.137$   
 $S = 1.20$   
 2737 reflections  
 217 parameters  
 9 restraints  
 H-atom parameters constrained  
 $\Delta\rho_{\max} = 0.61$  e Å<sup>-3</sup>  
 $\Delta\rho_{\min} = -0.32$  e Å<sup>-3</sup>

Table 1

Hydrogen-bond geometry (Å, °).

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
O1 <i>W</i> —H2 <i>W</i> ...O1 <sup>i</sup>	0.84	1.86	2.693 (3)	173
O1 <i>W</i> —H1 <i>W</i> ...O3 <sup>ii</sup>	0.84	2.00	2.801 (3)	160
O4—H4 <i>A</i> ...O5 <sup>iii</sup>	0.82	1.78	2.585 (3)	167
O2 <i>W</i> —H4 <i>W</i> ...O2 <sup>iv</sup>	0.84	1.79	2.624 (3)	176
O2 <i>W</i> —H3 <i>W</i> ...O1 <i>W</i> <sup>v</sup>	0.84	1.92	2.741 (2)	166
O3 <i>W</i> —H5 <i>W</i> ...O1 <i>W</i> <sup>v</sup>	0.84	2.06	2.810 (3)	148
O3 <i>W</i> —H6 <i>W</i> ...O1 <sup>vi</sup>	0.84	1.81	2.634 (3)	169
N2—H2...O5 <sup>vii</sup>	0.86	1.98	2.779 (3)	155

Symmetry codes: (i)  $-x + 1, -y + 2, -z + 1$ ; (ii)  $x, y, z + 1$ ; (iii)  $-x + 1, -y + 2, -z$ ; (iv)  $x - 1, y - 1, z$ ; (v)  $-x + 1, -y + 1, -z + 1$ ; (vi)  $-x + 2, -y + 2, -z + 1$ ; (vii)  $-x, -y + 1, -z$ .

Data collection: *RAPID-AUTO* (Rigaku, 1998); cell refinement: *RAPID-AUTO*; data reduction: *CrystalStructure* (Rigaku/MSM, 2002); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEPII* (Johnson, 1976); software used to prepare material for publication: *SHELXL97*.

The authors acknowledge Guang Dong Ocean University for supporting this work.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: NG2646).

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**supplementary materials**

*Acta Cryst.* (2009). E65, m1258 [ doi:10.1107/S1600536809038069 ]

**Tetraaquabis(6-carboxy-1*H*-benzimidazole-5-carboxylato- $\kappa$ N<sup>3</sup>)nickel(II) dimethylformamide disolvate dihydrate**

**H. Wang, W.-D. Song, S.-J. Li, P.-W. Qin and S.-W. Hu**

**Comment**

From the structural point of view, 1*H*-benzimidazole-5,6-dicarboxylic acid possesses two nitrogen atoms of imidazole ring and four oxygen atoms of carboxylate groups, and might be used as versatile linker in constructing coordination polymers with abundant hydrogen bonds. And several coordination polymers formed by this ligand have been reported recently: Pentaqua(1*H*-benzimidazole-5,6-dicarboxylato- $\kappa$ N<sup>3</sup>)copper(II) pentahydrate (Gao *et al.*, 2008), Bis(1*H*-benzimidazole-5,6-dicarboxylato)bis[tetraaquadicobalt(II)] pentahydrate (Lo *et al.*, 2007), Pentaqua(1*H*-benzimidazole-5,6-dicarboxylato- $\kappa$ N<sup>3</sup>)cobalt(II)pentahydrate (Song *et al.*, 2009). In the present paper, we synthesized a novel coordination complex  $[\text{Ni}(\text{C}_9\text{H}_4\text{N}_2\text{O}_4)_2(\text{H}_2\text{O})_4] \cdot 2\text{H}_2\text{O} \cdot 2\text{C}_3\text{H}_7\text{NO}$ .

As shown in Figure 1, the Ni<sup>II</sup> atom exhibits an octahedral coordination sphere, defined by two N atoms from two different 1*H*-benzimidazole-5,6-dicarboxylate ligands, and four water molecules. The equatorial plane is defined by O2<sup>w</sup>, O3<sup>w</sup>, O2<sup>w<sup>i</sup></sup> and O3<sup>w<sup>i</sup></sup> atoms, while N1 and N1<sup>i</sup> occupy the axial position (symmetry codes:  $i = 1 - x, 1 - y, 1 - z$ ). Inter/intramolecular O—H $\cdots$ O and N—H $\cdots$ O hydrogen bonds between the carboxylate O atoms of 1*H*-benzimidazole-5,6-dicarboxylate and the coordinated water molecule lead to the structure more stable (Fig 2). The hydrogen bonds are in the normal range (Table 1).

**Experimental**

A C<sub>3</sub>H<sub>7</sub>NO solution (20 mL) containing Ni(NO<sub>3</sub>)<sub>2</sub> (0.1 mmol) and 1*H*-benzimidazole-5,6-dicarboxylic acid (0.2 mmol) was stirred for a few minutes in air, and left to stand at room temperature for about four weeks, then the green crystals were obtained.

**Refinement**

Carbon and nitrogen bound H atoms were placed at calculated positions and were treated as riding on the parent C or N atoms with C—H = 0.93 Å, N—H = 0.86 Å, and with  $U_{\text{iso}}(\text{H}) = 1.2 U_{\text{eq}}(\text{C}, \text{N})$ . The water H-atoms were located in a difference map, and were refined with a distance restraint of O—H = 0.84 Å; their  $U_{\text{iso}}$  values were refined.

## Figures

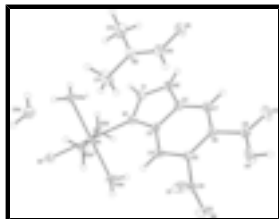


Fig. 1. The structure of the title compound, showing the atomic numbering scheme with 30% probability displacement ellipsoids. [Symmetry codes: (i)  $1 - x, 1 - y, 1 - z.$ ]

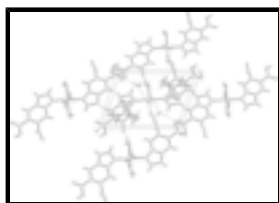


Fig. 2. A view of the three-dimensional network constructed by O—H...O and N—H...O hydrogen bonding interactions.

## Tetraaquabis(6-carboxy-1*H*-benzimidazole-5-carboxylato- $\kappa N^3$ )nickel(II) dimethylformamide disolvate dihydrate

### Crystal data

$[\text{Ni}(\text{C}_9\text{H}_5\text{N}_2\text{O}_4)_2(\text{H}_2\text{O})_4] \cdot 2\text{C}_3\text{H}_7\text{NO} \cdot 2\text{H}_2\text{O}$

$M_r = 723.30$

Triclinic,  $P\bar{1}$

Hall symbol: -P 1

$a = 8.5327$  (17) Å

$b = 9.1387$  (18) Å

$c = 11.624$  (2) Å

$\alpha = 100.80$  (3)°

$\beta = 103.03$  (3)°

$\gamma = 114.04$  (3)°

$V = 765.7$  (3) Å<sup>3</sup>

$Z = 1$

$F_{000} = 378$

$D_x = 1.569$  Mg m<sup>-3</sup>

Mo  $K\alpha$  radiation,  $\lambda = 0.71073$  Å

Cell parameters from 3600 reflections

$\theta = 1.4$ – $28^\circ$

$\mu = 0.72$  mm<sup>-1</sup>

$T = 293$  K

Block, green

$0.27 \times 0.18 \times 0.17$  mm

### Data collection

Rigaku/MSC Mercury CCD diffractometer

Radiation source: fine-focus sealed tube

Monochromator: graphite

$T = 293$  K

$\omega$  scans

Absorption correction: multi-scan (REQAB; Jacobson, 1998)

$T_{\min} = 0.830$ ,  $T_{\max} = 0.888$

6116 measured reflections

2737 independent reflections

2613 reflections with  $I > 2\sigma(I)$

$R_{\text{int}} = 0.020$

$\theta_{\max} = 25.2^\circ$

$\theta_{\min} = 3.3^\circ$

$h = -10 \rightarrow 10$

$k = -10 \rightarrow 9$

$l = -13 \rightarrow 13$

*Refinement*

Refinement on $F^2$	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.037$	H-atom parameters constrained
$wR(F^2) = 0.137$	$w = 1/[\sigma^2(F_o^2) + (0.1051P)^2 + 0.01P]$
$S = 1.20$	where $P = (F_o^2 + 2F_c^2)/3$
2737 reflections	$(\Delta/\sigma)_{\max} < 0.001$
217 parameters	$\Delta\rho_{\max} = 0.61 \text{ e } \text{\AA}^{-3}$
9 restraints	$\Delta\rho_{\min} = -0.32 \text{ e } \text{\AA}^{-3}$
Primary atom site location: structure-invariant direct methods	Extinction correction: none

*Special details*

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
C1	0.2895 (3)	0.4030 (3)	0.2261 (2)	0.0301 (5)
H1	0.2030	0.3062	0.2339	0.036*
N1	0.4387 (2)	0.5153 (2)	0.31852 (17)	0.0272 (4)
Ni1	0.5000	0.5000	0.5000	0.02200 (19)
C2	0.4271 (3)	0.5938 (3)	0.1432 (2)	0.0272 (5)
N2	0.2739 (2)	0.4415 (2)	0.11945 (17)	0.0324 (4)
H2	0.1861	0.3832	0.0499	0.039*
C3	0.5296 (3)	0.6391 (3)	0.2686 (2)	0.0252 (4)
O3W	0.75996 (16)	0.54366 (16)	0.50333 (12)	0.0345 (4)
C4	0.6923 (3)	0.7892 (3)	0.3224 (2)	0.0271 (5)
H4	0.7612	0.8210	0.4056	0.033*
C5	0.7498 (3)	0.8909 (3)	0.2494 (2)	0.0248 (4)
O1	0.9319 (2)	1.1775 (2)	0.37262 (18)	0.0456 (5)
C6	0.6429 (3)	0.8423 (3)	0.1223 (2)	0.0280 (5)
O2	1.0690 (2)	1.0376 (2)	0.3063 (2)	0.0522 (5)
C7	0.4811 (3)	0.6925 (3)	0.0690 (2)	0.0307 (5)
H7	0.4113	0.6596	-0.0141	0.037*

## supplementary materials

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C8	0.6944 (3)	0.9451 (3)	0.0384 (2)	0.0327 (5)
C9	0.9321 (3)	1.0500 (3)	0.31293 (19)	0.0261 (5)
O2W	0.41483 (16)	0.24688 (15)	0.42925 (12)	0.0295 (4)
H6W	0.8553	0.6369	0.5347	0.044*
H5W	0.7568	0.4874	0.4361	0.044*
H3W	0.4755	0.2330	0.3845	0.044*
H4W	0.3028	0.1830	0.3922	0.044*
O3	0.6082 (3)	0.8962 (3)	-0.07127 (17)	0.0591 (6)
O4	0.8343 (3)	1.0927 (3)	0.09199 (18)	0.0633 (7)
H4A	0.8481	1.1446	0.0415	0.095*
O1W	0.3424 (2)	0.7464 (3)	0.69171 (17)	0.0487 (5)
H1W	0.3996	0.7866	0.7687	0.073*
H2W	0.2637	0.7784	0.6716	0.073*
O5	0.0718 (2)	0.7407 (2)	0.05904 (16)	0.0462 (5)
C12	-0.0319 (4)	0.5721 (4)	0.2255 (3)	0.0570 (8)
H12A	-0.1264	0.6015	0.2316	0.086*
H12B	-0.0161	0.5128	0.2835	0.086*
H12C	-0.0653	0.5013	0.1426	0.086*
N3	0.1376 (3)	0.7248 (3)	0.25427 (18)	0.0352 (5)
C11	0.2581 (4)	0.8066 (4)	0.3847 (2)	0.0522 (7)
H11A	0.3625	0.9074	0.3914	0.078*
H11B	0.2971	0.7308	0.4129	0.078*
H11C	0.1933	0.8351	0.4351	0.078*
C10	0.1722 (3)	0.7970 (3)	0.1704 (2)	0.0356 (5)
H10	0.2794	0.8978	0.1949	0.043*

### Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
C1	0.0264 (11)	0.0250 (11)	0.0280 (11)	0.0029 (9)	0.0068 (9)	0.0092 (9)
N1	0.0272 (9)	0.0227 (9)	0.0259 (9)	0.0057 (7)	0.0082 (7)	0.0098 (7)
Ni1	0.0196 (3)	0.0191 (3)	0.0217 (3)	0.00424 (18)	0.00625 (17)	0.00652 (17)
C2	0.0216 (10)	0.0230 (10)	0.0262 (11)	0.0024 (8)	0.0062 (8)	0.0055 (9)
N2	0.0256 (9)	0.0274 (10)	0.0254 (9)	-0.0008 (8)	0.0021 (7)	0.0071 (8)
C3	0.0242 (10)	0.0248 (10)	0.0264 (11)	0.0093 (8)	0.0112 (8)	0.0099 (9)
O3W	0.0231 (8)	0.0295 (8)	0.0380 (9)	0.0049 (6)	0.0105 (6)	0.0003 (7)
C4	0.0231 (10)	0.0266 (11)	0.0223 (10)	0.0059 (8)	0.0037 (8)	0.0060 (8)
C5	0.0223 (10)	0.0237 (11)	0.0245 (10)	0.0075 (8)	0.0074 (8)	0.0078 (8)
O1	0.0282 (9)	0.0305 (9)	0.0559 (11)	0.0043 (7)	0.0102 (7)	-0.0071 (8)
C6	0.0252 (10)	0.0278 (11)	0.0258 (11)	0.0073 (9)	0.0082 (8)	0.0098 (9)
O2	0.0233 (9)	0.0324 (9)	0.0833 (15)	0.0073 (7)	0.0126 (8)	0.0002 (9)
C7	0.0283 (11)	0.0312 (11)	0.0217 (10)	0.0069 (9)	0.0037 (8)	0.0072 (9)
C8	0.0295 (11)	0.0328 (12)	0.0266 (12)	0.0063 (9)	0.0078 (9)	0.0113 (10)
C9	0.0220 (10)	0.0249 (11)	0.0249 (11)	0.0059 (9)	0.0050 (8)	0.0095 (9)
O2W	0.0250 (7)	0.0233 (7)	0.0316 (8)	0.0058 (6)	0.0076 (6)	0.0054 (6)
O3	0.0561 (12)	0.0527 (12)	0.0269 (9)	-0.0082 (9)	0.0002 (8)	0.0191 (9)
O4	0.0599 (12)	0.0464 (11)	0.0327 (10)	-0.0161 (9)	-0.0030 (9)	0.0224 (9)
O1W	0.0396 (10)	0.0590 (12)	0.0390 (10)	0.0258 (9)	0.0087 (8)	-0.0036 (9)

O5	0.0409 (10)	0.0468 (10)	0.0272 (9)	0.0019 (8)	0.0025 (7)	0.0156 (8)
C12	0.0560 (17)	0.0591 (18)	0.0607 (19)	0.0195 (14)	0.0312 (15)	0.0326 (15)
N3	0.0421 (11)	0.0367 (11)	0.0260 (10)	0.0184 (9)	0.0101 (9)	0.0103 (9)
C11	0.078 (2)	0.0531 (17)	0.0285 (13)	0.0379 (16)	0.0077 (13)	0.0140 (12)
C10	0.0363 (12)	0.0301 (12)	0.0302 (12)	0.0083 (10)	0.0065 (10)	0.0104 (10)

*Geometric parameters (Å, °)*

C1—N1	1.317 (3)	C6—C7	1.386 (3)
C1—N2	1.344 (3)	C6—C8	1.492 (3)
C1—H1	0.9300	O2—C9	1.236 (3)
N1—C3	1.397 (3)	C7—H7	0.9300
N1—Ni1	2.1014 (18)	C8—O3	1.209 (3)
Ni1—O2W	2.0501 (14)	C8—O4	1.293 (3)
Ni1—O2W <sup>i</sup>	2.0501 (14)	O2W—H3W	0.8401
Ni1—O3W <sup>i</sup>	2.0773 (13)	O2W—H4W	0.8400
Ni1—O3W	2.0773 (13)	O4—H4A	0.8200
Ni1—N1 <sup>i</sup>	2.1014 (18)	O1W—H1W	0.8408
C2—C7	1.379 (3)	O1W—H2W	0.8405
C2—N2	1.390 (3)	O5—C10	1.252 (3)
C2—C3	1.401 (3)	C12—N3	1.454 (4)
N2—H2	0.8600	C12—H12A	0.9600
C3—C4	1.391 (3)	C12—H12B	0.9600
O3W—H6W	0.8402	C12—H12C	0.9600
O3W—H5W	0.8394	N3—C10	1.298 (3)
C4—C5	1.391 (3)	N3—C11	1.470 (3)
C4—H4	0.9300	C11—H11A	0.9600
C5—C6	1.424 (3)	C11—H11B	0.9600
C5—C9	1.522 (3)	C11—H11C	0.9600
O1—C9	1.240 (3)	C10—H10	0.9300
N1—C1—N2	113.71 (18)	C4—C5—C9	115.96 (18)
N1—C1—H1	123.1	C6—C5—C9	123.63 (19)
N2—C1—H1	123.1	C7—C6—C5	120.7 (2)
C1—N1—C3	104.99 (18)	C7—C6—C8	115.74 (19)
C1—N1—Ni1	123.63 (15)	C5—C6—C8	123.59 (19)
C3—N1—Ni1	131.30 (15)	C2—C7—C6	117.85 (19)
O2W—Ni1—O2W <sup>i</sup>	180.0	C2—C7—H7	121.1
O2W—Ni1—O3W <sup>i</sup>	91.85 (6)	C6—C7—H7	121.1
O2W <sup>i</sup> —Ni1—O3W <sup>i</sup>	88.15 (6)	O3—C8—O4	122.2 (2)
O2W—Ni1—O3W	88.15 (6)	O3—C8—C6	122.5 (2)
O2W <sup>i</sup> —Ni1—O3W	91.85 (6)	O4—C8—C6	115.3 (2)
O3W <sup>i</sup> —Ni1—O3W	180.0	O2—C9—O1	125.6 (2)
O2W—Ni1—N1 <sup>i</sup>	90.06 (7)	O2—C9—C5	116.7 (2)
O2W <sup>i</sup> —Ni1—N1 <sup>i</sup>	89.94 (7)	O1—C9—C5	117.58 (18)
O3W <sup>i</sup> —Ni1—N1 <sup>i</sup>	90.14 (7)	Ni1—O2W—H3W	109.0
O3W—Ni1—N1 <sup>i</sup>	89.86 (7)	Ni1—O2W—H4W	117.7

## supplementary materials

O2W—Ni1—N1	89.94 (7)	H3W—O2W—H4W	110.9
O2W <sup>i</sup> —Ni1—N1	90.06 (7)	C8—O4—H4A	109.5
O3W <sup>i</sup> —Ni1—N1	89.86 (7)	H1W—O1W—H2W	111.4
O3W—Ni1—N1	90.14 (7)	N3—C12—H12A	109.5
N1 <sup>i</sup> —Ni1—N1	180.0	N3—C12—H12B	109.5
C7—C2—N2	131.9 (2)	H12A—C12—H12B	109.5
C7—C2—C3	122.56 (19)	N3—C12—H12C	109.5
N2—C2—C3	105.48 (19)	H12A—C12—H12C	109.5
C1—N2—C2	106.79 (18)	H12B—C12—H12C	109.5
C1—N2—H2	126.6	C10—N3—C12	121.1 (2)
C2—N2—H2	126.6	C10—N3—C11	120.7 (2)
C4—C3—N1	131.2 (2)	C12—N3—C11	117.8 (2)
C4—C3—C2	119.76 (19)	N3—C11—H11A	109.5
N1—C3—C2	109.02 (19)	N3—C11—H11B	109.5
Ni1—O3W—H6W	126.3	H11A—C11—H11B	109.5
Ni1—O3W—H5W	111.1	N3—C11—H11C	109.5
H6W—O3W—H5W	111.6	H11A—C11—H11C	109.5
C5—C4—C3	118.77 (19)	H11B—C11—H11C	109.5
C5—C4—H4	120.6	O5—C10—N3	124.8 (2)
C3—C4—H4	120.6	O5—C10—H10	117.6
C4—C5—C6	120.39 (19)	N3—C10—H10	117.6

Symmetry codes: (i)  $-x+1, -y+1, -z+1$ .

### Hydrogen-bond geometry ( $\text{\AA}, ^\circ$ )

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
O1W—H2W $\cdots$ O1 <sup>ii</sup>	0.84	1.86	2.693 (3)	173
O1W—H1W $\cdots$ O3 <sup>iii</sup>	0.84	2.00	2.801 (3)	160
O4—H4A $\cdots$ O5 <sup>iv</sup>	0.82	1.78	2.585 (3)	167
O2W—H4W $\cdots$ O2 <sup>v</sup>	0.84	1.79	2.624 (3)	176
O2W—H3W $\cdots$ O1W <sup>i</sup>	0.84	1.92	2.741 (2)	166
O3W—H5W $\cdots$ O1W <sup>i</sup>	0.84	2.06	2.810 (3)	148
O3W—H6W $\cdots$ O1 <sup>vi</sup>	0.84	1.81	2.634 (3)	169
N2—H2 $\cdots$ O5 <sup>vii</sup>	0.86	1.98	2.779 (3)	155

Symmetry codes: (ii)  $-x+1, -y+2, -z+1$ ; (iii)  $x, y, z+1$ ; (iv)  $-x+1, -y+2, -z$ ; (v)  $x-1, y-1, z$ ; (i)  $-x+1, -y+1, -z+1$ ; (vi)  $-x+2, -y+2, -z+1$ ; (vii)  $-x, -y+1, -z$ .



Fig. 2

