

Synchrotron study of poly[[di- μ -aqua(μ -2,2'-bipyridyl-5,5'-dicarboxylato)di-potassium] dihydrate]

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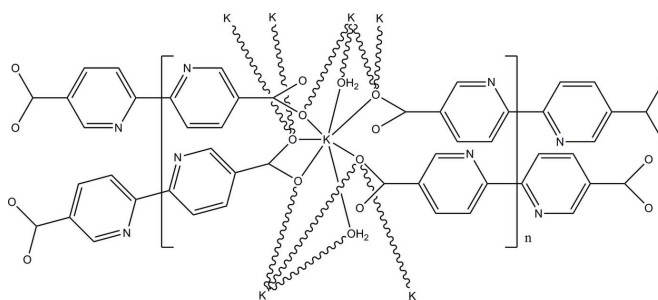
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Key indicators: single-crystal synchrotron study; $T = 150$ K; mean $\sigma(\text{C}-\text{C}) = 0.004$ Å; H-atom completeness 72%; R factor = 0.045; wR factor = 0.128; data-to-parameter ratio = 13.9.

The title compound, $\{[\text{K}_2(\text{C}_{12}\text{H}_6\text{N}_2\text{O}_4)(\text{H}_2\text{O})_2] \cdot 2\text{H}_2\text{O}\}_n$, forms a three-dimensional coordination polymer in the solid state. The asymmetric unit consists of one K^+ ion, half of a 2,2'-bipyridyl-5,5'-dicarboxylate ligand, one coordinated water molecule and one solvent water molecule. The K^+ ion is 7-coordinated by the oxygen atoms of two water molecules and by five oxygen atoms of four carboxylate groups, one of which is chelating. The extended structure can be described as a binodal network in which each K^+ is a six-connected node, bonding to four carboxylate groups and two bridging water molecules, and the 2,2'-bipyridyl-5,5'-dicarboxylate linkers are eight-connected nodes, with each carboxylate group bridging four metal centers. Overall, this arrangement generates a complex network with point symbol $\{3^4.4^{12}.5^{12}\}\{3^4.4^4.5^4.6^3\}_2$. Both of the bridging water molecules participate as donors in hydrogen-bonding interactions; one to solvent water molecules and a second to an oxygen atom of a carboxylate group.

Related literature

For topological analysis, see: Blatov (2007). For background to metal-organic frameworks (MOFs), see: MacDougall *et al.* (2005). The 2,2'-bipyridyl-5,5'-dicarboxylate ligand has been used as a linear linker for a variety of MOFs, see: Finn & Zubieta, (2002); Schoknecht & Kempe (2004); Szeto *et al.*, (2008). It is a particularly interesting linker due to the fact that the pyridyl N atoms have the ability to act as Lewis bases for binding metal centers (Szeto *et al.*, 2008). There has been only one structural example of this ligand bound to an alkali metal reported, *viz.* Rb (Hafizovic *et al.*, 2007). For synthetic details, see: Anderson *et al.* (1985).



Experimental

Crystal data

$[\text{K}_2(\text{C}_{12}\text{H}_6\text{N}_2\text{O}_4)(\text{H}_2\text{O})_2] \cdot 2\text{H}_2\text{O}$
 $M_r = 392.46$
 Monoclinic, $P2_1/c$
 $a = 3.6769$ (6) Å
 $b = 8.2042$ (14) Å
 $c = 26.292$ (4) Å
 $\beta = 92.924$ (2)°

$V = 792.1$ (2) Å³

$Z = 2$

Synchrotron radiation

$\lambda = 0.77490$ Å

$\mu = 0.76$ mm⁻¹

$T = 150$ K

$0.04 \times 0.03 \times 0.01$ mm

Data collection

Bruker APEXII diffractometer
 Absorption correction: multi-scan
 (SADABS; Sheldrick, 2008a)
 $T_{\min} = 0.654$, $T_{\max} = 0.746$

8561 measured reflections

1622 independent reflections

1301 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.045$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.045$

$wR(F^2) = 0.128$

$S = 1.08$

1622 reflections

117 parameters

H atoms treated by a mixture of independent and constrained refinement

$\Delta\rho_{\text{max}} = 1.03$ e Å⁻³

$\Delta\rho_{\text{min}} = -0.51$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D-H \cdots A$	$D-H$	$H \cdots A$	$D \cdots A$	$D-H \cdots A$
$\text{O3}-\text{H3A} \cdots \text{O1}^{\text{i}}$	0.84 (5)	1.93 (5)	2.770 (4)	177 (5)
$\text{O3}-\text{H3B} \cdots \text{O4}^{\text{ii}}$	0.84 (6)	2.15 (6)	2.976 (5)	169 (5)

Symmetry codes: (i) $-x + 1, y - \frac{1}{2}, -z + \frac{1}{2}$; (ii) $-x + 1, y + \frac{1}{2}, -z + \frac{1}{2}$.

Data collection: APEX2 (Bruker, 2008); cell refinement: SAINT (Bruker, 2008); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008b); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008b); molecular graphics: SHELXTL (Sheldrick, 2008b); software used to prepare material for publication: SHELXTL.

Samples for synchrotron crystallographic analysis were submitted through the SCrALS (Service Crystallography at Advanced Light Source) program. Crystallographic data were collected at Beamline 11.3.1 at the Advanced Light Source (ALS), Lawrence Berkeley National Laboratory. The ALS is supported by the US Department of Energy, Office of Energy Sciences, under contract DE-AC02-05CH11231.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: WM2256).

References

- Anderson, S., Constable, E. C., Seddon, K. R., Turp, J. E., Baggott, J. E. & Pilling, M. J. (1985). *J. Chem. Soc. Dalton Trans.* pp. 2247-2261.
- Blatov, V. A. (2007). <http://www.topos.ssu.samara.ru/>.
- Bruker (2008). *APEX2* and *SAINT*. Bruker AXS Inc., Madison, Wisconsin, USA.
- Finn, R. C. & Zubieta, J. (2002). *Solid State Sci.* **4**, 83–86.
- Hafizovic, J., Olsbye, U. & Lillerud, K. P. (2007). *Acta Cryst.* **E63**, m962–m964.
- MacDougall, D. J., Morris, J. J., Noll, B. C. & Henderson, K. W. (2005). *Chem. Commun.* pp. 456–458.
- Schoknecht, B. & Kempe, R. (2004). *Z. Anorg. Allg. Chem.* **630**, 1377–1379.
- Sheldrick, G. M. (2008a). *SADABS*. University of Göttingen, Germany.
- Sheldrick, G. M. (2008b). *Acta Cryst.* **A64**, 112–122.
- Szeto, K. C., Kongshaug, K. O., Jakobsen, S., Tilset, M. & Lillerud, K. P. (2008). *Dalton Trans.* pp. 2054–2060.

supplementary materials

Acta Cryst. (2009). E65, m1265-m1266 [doi:10.1107/S1600536809038756]

Synchrotron study of poly[[di- μ -aqua(μ -2,2'-bipyridyl-5,5'-dicarboxylato)dipotassium] dihydrate]

J. A. Bertke, A. G. Oliver and K. W. Henderson

Comment

We are interested in using pre-assembled and structurally well defined aggregates as secondary building units (SBUs) to build coordination polymers or metal-organic frameworks (MOFs) (MacDougall *et al.*, 2005). The title compound was synthesized as a precursor for use as a linker in the construction of MOFs. The 2,2'-bipyridyl-5,5'-dicarboxylate ligand has been used as a linear linker for a variety of MOFs (Finn & Zubieta, 2002; Schoknecht & Kempe, 2004); Szeto *et al.*, 2008). It is a particularly interesting linker due to the fact that the pyridyl N atoms have the ability to act as Lewis bases for binding metal centers (Szeto *et al.*, 2008).

The asymmetric unit consists of one K^+ ion, half of a 2,2'-bipyridyl-5,5'-dicarboxylate ligand, one coordinated water molecule and one solvent water molecule (Fig. 1). The extended crystal structure of the title compound is composed of two-dimensional sheets consisting of potassium-carboxylate and potassium-water interactions. These sheets extend parallel to (001). The three-dimensional assembly is completed by connection of the sheets through the 2,2'-bipyridyl linkers (Fig. 2). Small channels run along the *a* axis and contain the disordered solvent water molecules. The geometry around the seven-coordinate potassium metal center can be described as a distorted monocapped trigonal-prism. The faces deviate from the ideal triangular, with angles of 49.69° (O1—O3—O2), 61.19° (O2—O1—O3) and 69.10° (O1—O2—O3). There is also a short contact distance of 2.826 (6) Å between O4 and N1, which is most likely due to a hydrogen bonding interaction. However, the hydrogen atoms on the solvent water molecule could not be located in the difference map.

For the topological analysis of the structure, the program *TOPOS* was used (Blatov, 2007). There has been only one structural example of the 2,2'-bipyridyl-5,5'-dicarboxylate ligand bound to an alkali metal reported, *viz.* Rb (Hafizovic *et al.*, 2007).

Experimental

The title complex was prepared using a modification of the method of Seddon and Pilling (Anderson *et al.*, 1985) for the preparation of 2,2'-bipyridyl-4,4'-dicarboxylic acid. 5,5'-dimethyl-2,2'-bipyridine and $K[MnO_4]$ were purchased from Aldrich and used without further purification. 1H NMR data were collected on a Varian Unity Plus 300 spectrometer at 298 K. $K[MnO_4]$ (10 g, 63 mmol) was added to a solution of 5,5'-dimethyl-2,2'-bipyridine (1.6 g, 8.7 mmol) in 100 mL H_2O . The mixture was heated to reflux for 12 h. Upon cooling, the mixture was filtered and the black precipitate was washed with water (2x 30 mL). The filtrate and the washings were combined and extracted with diethyl ether to remove unreacted starting material. The aqueous fraction was collected and the solvent was removed under vacuum, leaving a sticky white solid. The solid was heated under vacuum at 373 K overnight to give a fine white powder. Single crystals were grown from an aqueous solution upon slow evaporation of the solvent.

Refinement

Hydrogen atoms were placed at calculated positions and allowed to ride on the position of the parent atom, with the exception of those on the coordinated O3 water molecule which were located in the difference fourier map. H atoms of the uncoordinated water molecule O4 could not be located unambiguously and were eventually excluded from the refinement.

Figures

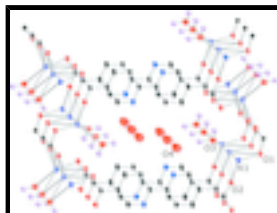


Fig. 1. A portion of the polymeric structure of the title compound. All hydrogen atoms except those located on O3 have been omitted. Thermal ellipsoids are shown at the 50% probability level.

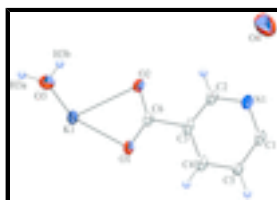


Fig. 2. Thermal ellipsoid plot of the asymmetric unit of the title compound. Thermal ellipsoids are shown at the 50% probability level.

poly[[di- μ -aqua(μ -2,2'-bipyridyl-5,5'-dicarboxylato)dipotassium] dihydrate]

Crystal data

$[\text{K}_2(\text{C}_{12}\text{H}_6\text{N}_2\text{O}_4)(\text{H}_2\text{O})_2] \cdot 2\text{H}_2\text{O}$

$M_r = 392.46$

Monoclinic, $P2_1/c$

Hall symbol: -P 2yc

$a = 3.6769$ (6) Å

$b = 8.2042$ (14) Å

$c = 26.292$ (4) Å

$\beta = 92.924$ (2)°

$V = 792.1$ (2) Å³

$Z = 2$

$F_{000} = 404$

$D_x = 1.646$ Mg m⁻³

Synchrotron radiation, $\lambda = 0.77490$ Å

Cell parameters from 1984 reflections

$\theta = 2.8$ – 27.6 °

$\mu = 0.76$ mm⁻¹

$T = 150$ K

Plate, colorless

$0.04 \times 0.03 \times 0.01$ mm

Data collection

Bruker APEXII
diffractometer

Radiation source: synchrotron

Monochromator: channel-cut Si-<111> crystal

Detector resolution: 83.33 pixels mm⁻¹

$T = 150$ K

ω and φ scans

1622 independent reflections

1301 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.045$

$\theta_{\text{max}} = 29.0$ °

$\theta_{\text{min}} = 2.8$ °

$h = -4 \rightarrow 4$

Absorption correction: multi-scan
(SADABS; Sheldrick, 2008a) $k = -10 \rightarrow 10$
 $T_{\min} = 0.654$, $T_{\max} = 0.746$ $l = -32 \rightarrow 32$
 8561 measured reflections

Refinement

Refinement on F^2 Secondary atom site location: difference Fourier map
 Least-squares matrix: full Hydrogen site location: inferred from neighbouring sites
 $R[F^2 > 2\sigma(F^2)] = 0.045$ H atoms treated by a mixture of independent and constrained refinement
 $wR(F^2) = 0.128$ $w = 1/[\sigma^2(F_o^2) + (0.0609P)^2 + 1.2501P]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $S = 1.08$ $(\Delta/\sigma)_{\max} < 0.001$
 1622 reflections $\Delta\rho_{\max} = 1.03 \text{ e } \text{\AA}^{-3}$
 117 parameters $\Delta\rho_{\min} = -0.51 \text{ e } \text{\AA}^{-3}$
 Primary atom site location: structure-invariant direct methods Extinction correction: none

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
K1	0.06233 (18)	0.72796 (8)	0.22289 (2)	0.0250 (2)
O1	0.5909 (6)	0.7368 (3)	0.30066 (8)	0.0257 (5)
O2	0.4216 (6)	0.4762 (3)	0.29578 (8)	0.0262 (5)
O3	0.5493 (7)	0.5466 (3)	0.16339 (11)	0.0359 (6)
H3A	0.499 (13)	0.454 (6)	0.1748 (18)	0.052 (14)*
H3B	0.528 (16)	0.539 (7)	0.132 (2)	0.080 (19)*
O4	0.537 (2)	0.0661 (6)	0.44962 (17)	0.154 (3)
N1	0.7659 (7)	0.3947 (3)	0.44667 (10)	0.0272 (6)
C1	0.9350 (8)	0.5168 (4)	0.47310 (10)	0.0216 (6)
C2	0.6527 (8)	0.4234 (4)	0.39788 (11)	0.0251 (7)
H2A	0.5352	0.3375	0.3792	0.030*
C3	0.6981 (8)	0.5704 (4)	0.37364 (11)	0.0225 (6)
C4	0.8744 (8)	0.6938 (4)	0.40162 (11)	0.0252 (7)

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H4A	0.9138	0.7968	0.3863	0.030*
C5	0.9916 (8)	0.6669 (4)	0.45143 (11)	0.0209 (6)
H5A	1.1106	0.7510	0.4707	0.025*
C6	0.5582 (8)	0.5960 (4)	0.31898 (11)	0.0200 (6)

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
K1	0.0231 (3)	0.0293 (4)	0.0222 (4)	-0.0001 (3)	-0.0021 (2)	0.0022 (3)
O1	0.0312 (12)	0.0272 (12)	0.0182 (10)	0.0001 (9)	-0.0035 (9)	0.0015 (8)
O2	0.0310 (12)	0.0258 (12)	0.0209 (10)	-0.0008 (9)	-0.0053 (9)	-0.0028 (9)
O3	0.0432 (15)	0.0268 (14)	0.0371 (16)	-0.0043 (11)	-0.0035 (12)	0.0001 (11)
O4	0.345 (9)	0.055 (3)	0.059 (3)	0.017 (4)	-0.009 (4)	-0.001 (2)
N1	0.0293 (14)	0.0311 (15)	0.0209 (13)	-0.0025 (11)	-0.0025 (11)	-0.0005 (11)
C1	0.0182 (13)	0.0305 (16)	0.0159 (14)	0.0023 (12)	-0.0013 (11)	-0.0020 (12)
C2	0.0250 (16)	0.0316 (17)	0.0183 (14)	0.0008 (13)	-0.0024 (12)	-0.0005 (12)
C3	0.0176 (14)	0.0330 (16)	0.0168 (14)	0.0038 (12)	0.0003 (11)	-0.0008 (12)
C4	0.0225 (15)	0.0325 (17)	0.0205 (15)	0.0018 (12)	-0.0003 (11)	-0.0004 (12)
C5	0.0200 (14)	0.0256 (15)	0.0167 (14)	-0.0018 (11)	-0.0040 (11)	-0.0009 (11)
C6	0.0181 (14)	0.0261 (16)	0.0158 (14)	0.0029 (11)	0.0002 (11)	-0.0014 (11)

Geometric parameters (\AA , $^\circ$)

K1—O2 ⁱ	2.732 (2)	N1—C2	1.349 (4)
K1—O1 ⁱⁱ	2.748 (2)	N1—C1	1.353 (4)
K1—O1	2.749 (2)	C1—C5	1.376 (4)
K1—O3 ⁱⁱⁱ	2.814 (3)	C1—C1 ^{iv}	1.496 (6)
K1—O2 ⁱⁱⁱ	2.843 (2)	C2—C3	1.378 (5)
K1—O3	2.855 (3)	C2—H2A	0.9500
K1—O2	3.070 (2)	C3—C4	1.392 (4)
O1—C6	1.259 (4)	C3—C6	1.516 (4)
O2—C6	1.249 (4)	C4—C5	1.376 (4)
O3—H3A	0.84 (5)	C4—H4A	0.9500
O3—H3B	0.84 (6)	C5—H5A	0.9500
O2 ⁱ —K1—O1 ⁱⁱ	71.60 (7)	K1 ^{vi} —O2—K1	97.69 (6)
O2 ⁱ —K1—O1	122.90 (7)	K1 ^{vii} —O2—K1	130.08 (8)
O1 ⁱⁱ —K1—O1	83.96 (6)	K1 ^v —O3—K1	80.87 (7)
O2 ⁱ —K1—O3 ⁱⁱⁱ	83.24 (8)	K1 ^v —O3—H3A	116 (3)
O1 ⁱⁱ —K1—O3 ⁱⁱⁱ	89.84 (8)	K1—O3—H3A	97 (3)
O1—K1—O3 ⁱⁱⁱ	148.85 (8)	K1 ^v —O3—H3B	129 (4)
O2 ⁱ —K1—O2 ⁱⁱⁱ	82.49 (6)	K1—O3—H3B	124 (4)
O1 ⁱⁱ —K1—O2 ⁱⁱⁱ	124.15 (7)	H3A—O3—H3B	106 (5)
O1—K1—O2 ⁱⁱⁱ	69.93 (6)	C2—N1—C1	118.1 (3)
O3 ⁱⁱⁱ —K1—O2 ⁱⁱⁱ	135.82 (8)	N1—C1—C5	121.7 (3)
O2 ⁱ —K1—O3	135.30 (8)	N1—C1—C1 ^{iv}	117.7 (3)

O1 ⁱⁱ —K1—O3	149.13 (8)	C5—C1—C1 ^{iv}	120.6 (3)
O1—K1—O3	88.96 (7)	N1—C2—C3	123.6 (3)
O3 ⁱⁱ —K1—O3	80.87 (7)	N1—C2—H2A	118.2
O2 ⁱⁱⁱ —K1—O3	80.56 (7)	C3—C2—H2A	118.2
O2 ⁱ —K1—O2	150.35 (3)	C2—C3—C4	117.2 (3)
O1 ⁱⁱ —K1—O2	79.68 (6)	C2—C3—C6	121.1 (3)
O1—K1—O2	44.61 (6)	C4—C3—C6	121.7 (3)
O3 ⁱⁱ —K1—O2	104.26 (7)	C5—C4—C3	120.0 (3)
O2 ⁱⁱⁱ —K1—O2	108.66 (3)	C5—C4—H4A	120.0
O3—K1—O2	74.31 (7)	C3—C4—H4A	120.0
C6—O1—K1 ^v	109.68 (18)	C4—C5—C1	119.4 (3)
C6—O1—K1	100.52 (17)	C4—C5—H5A	120.3
K1 ^v —O1—K1	83.96 (6)	C1—C5—H5A	120.3
C6—O2—K1 ^{vi}	156.4 (2)	O2—C6—O1	125.4 (3)
C6—O2—K1 ^{vii}	113.13 (18)	O2—C6—C3	117.5 (3)
K1 ^{vi} —O2—K1 ^{vii}	82.49 (6)	O1—C6—C3	117.1 (3)
C6—O2—K1	85.69 (17)		
O2 ⁱ —K1—O1—C6	134.52 (17)	K1 ^{vi} —O2—C6—K1	-99.4 (5)
O1 ⁱⁱ —K1—O1—C6	71.04 (18)	K1 ^{vii} —O2—C6—K1	132.06 (13)
O3 ⁱⁱ —K1—O1—C6	-8.6 (3)	K1 ^{vi} —O2—C6—K1 ^v	-163.2 (4)
O2 ⁱⁱⁱ —K1—O1—C6	-159.19 (19)	K1 ^{vii} —O2—C6—K1 ^v	68.29 (14)
O3—K1—O1—C6	-78.86 (18)	K1—O2—C6—K1 ^v	-63.77 (7)
O2—K1—O1—C6	-10.37 (16)	K1 ^{vi} —O2—C6—K1 ^{vii}	128.6 (5)
C6 ⁱⁱ —K1—O1—C6	50.56 (15)	K1—O2—C6—K1 ^{vii}	-132.06 (13)
C6 ⁱⁱⁱ —K1—O1—C6	-169.28 (16)	K1 ^v —O1—C6—O2	-64.9 (3)
K1 ⁱⁱ —K1—O1—C6	71.04 (18)	K1—O1—C6—O2	22.4 (3)
K1 ^v —K1—O1—C6	-108.96 (18)	K1 ^v —O1—C6—C3	114.8 (2)
O2 ⁱ —K1—O1—K1 ^v	-116.52 (7)	K1—O1—C6—C3	-157.9 (2)
O1 ⁱⁱ —K1—O1—K1 ^v	180.0	K1 ^v —O1—C6—K1	-87.29 (11)
O3 ⁱⁱ —K1—O1—K1 ^v	100.40 (15)	K1—O1—C6—K1 ^v	87.29 (11)
O2 ⁱⁱⁱ —K1—O1—K1 ^v	-50.23 (6)	K1 ^v —O1—C6—K1 ^{vii}	-4.1 (3)
O3—K1—O1—K1 ^v	30.09 (7)	K1—O1—C6—K1 ^{vii}	83.2 (2)
O2—K1—O1—K1 ^v	98.58 (9)	C2—C3—C6—O2	-4.5 (4)
C6—K1—O1—K1 ^v	108.96 (18)	C4—C3—C6—O2	176.1 (3)
C6 ⁱⁱ —K1—O1—K1 ^v	159.52 (7)	C2—C3—C6—O1	175.8 (3)
C6 ⁱⁱⁱ —K1—O1—K1 ^v	-60.32 (6)	C4—C3—C6—O1	-3.7 (4)
K1 ⁱⁱ —K1—O1—K1 ^v	180.0	C2—C3—C6—K1	108.6 (6)
O2 ⁱ —K1—O2—C6	-67.1 (2)	C4—C3—C6—K1	-70.9 (7)
O1 ⁱⁱ —K1—O2—C6	-81.54 (17)	C2—C3—C6—K1 ^v	-125.9 (3)
O1—K1—O2—C6	10.31 (16)	C4—C3—C6—K1 ^v	54.7 (4)
O3 ⁱⁱ —K1—O2—C6	-168.72 (17)	C2—C3—C6—K1 ^{vii}	-47.7 (3)

supplementary materials

O2 ⁱⁱⁱ —K1—O2—C6	41.19 (16)	C4—C3—C6—K1 ^{vii}	132.8 (3)
O3—K1—O2—C6	115.25 (18)	O2 ⁱ —K1—C6—O2	138.37 (12)
C6 ⁱⁱ —K1—O2—C6	-94.85 (18)	O1 ⁱⁱ —K1—C6—O2	93.78 (17)
C6 ⁱⁱⁱ —K1—O2—C6	37.2 (2)	O1—K1—C6—O2	-160.9 (3)
K1 ⁱⁱ —K1—O2—C6	-116.44 (16)	O3 ⁱⁱ —K1—C6—O2	13.6 (2)
K1 ^v —K1—O2—C6	63.56 (16)	O2 ⁱⁱⁱ —K1—C6—O2	-141.38 (15)
O2 ⁱ —K1—O2—K1 ^{vi}	89.41 (15)	O3—K1—C6—O2	-60.97 (17)
O1 ⁱⁱ —K1—O2—K1 ^{vi}	75.00 (7)	C6 ⁱⁱ —K1—C6—O2	74.81 (17)
O1—K1—O2—K1 ^{vi}	166.84 (12)	C6 ⁱⁱⁱ —K1—C6—O2	-149.60 (19)
O3 ⁱⁱ —K1—O2—K1 ^{vi}	-12.18 (9)	K1 ⁱⁱ —K1—C6—O2	74.81 (17)
O2 ⁱⁱⁱ —K1—O2—K1 ^{vi}	-162.27 (8)	K1 ^v —K1—C6—O2	-105.19 (17)
O3—K1—O2—K1 ^{vi}	-88.21 (8)	O2 ⁱ —K1—C6—O1	-60.7 (2)
C6—K1—O2—K1 ^{vi}	156.5 (2)	O1 ⁱⁱ —K1—C6—O1	-105.34 (19)
C6 ⁱⁱ —K1—O2—K1 ^{vi}	61.68 (7)	O3 ⁱⁱ —K1—C6—O1	174.51 (17)
C6 ⁱⁱⁱ —K1—O2—K1 ^{vi}	-166.30 (7)	O2 ⁱⁱⁱ —K1—C6—O1	19.50 (18)
K1 ⁱⁱ —K1—O2—K1 ^{vi}	40.09 (7)	O3—K1—C6—O1	99.91 (18)
K1 ^v —K1—O2—K1 ^{vi}	-139.91 (7)	O2—K1—C6—O1	160.9 (3)
O2 ⁱ —K1—O2—K1 ^{vii}	176.04 (11)	C6 ⁱⁱ —K1—C6—O1	-124.31 (16)
O1 ⁱⁱ —K1—O2—K1 ^{vii}	161.63 (11)	C6 ⁱⁱⁱ —K1—C6—O1	11.28 (17)
O1—K1—O2—K1 ^{vii}	-106.52 (13)	K1 ⁱⁱ —K1—C6—O1	-124.31 (16)
O3 ⁱⁱ —K1—O2—K1 ^{vii}	74.45 (11)	K1 ^v —K1—C6—O1	55.69 (16)
O2 ⁱⁱⁱ —K1—O2—K1 ^{vii}	-75.64 (13)	O2 ⁱ —K1—C6—C3	17.7 (6)
O3—K1—O2—K1 ^{vii}	-1.58 (10)	O1 ⁱⁱ —K1—C6—C3	-26.9 (6)
C6—K1—O2—K1 ^{vii}	-116.8 (2)	O1—K1—C6—C3	78.4 (6)
C6 ⁱⁱ —K1—O2—K1 ^{vii}	148.31 (12)	O3 ⁱⁱ —K1—C6—C3	-107.0 (6)
C6 ⁱⁱⁱ —K1—O2—K1 ^{vii}	-79.67 (12)	O2 ⁱⁱⁱ —K1—C6—C3	97.9 (6)
K1 ⁱⁱ —K1—O2—K1 ^{vii}	126.73 (8)	O3—K1—C6—C3	178.4 (6)
K1 ^v —K1—O2—K1 ^{vii}	-53.28 (8)	O2—K1—C6—C3	-120.7 (7)
O2 ⁱ —K1—O3—K1 ^v	109.38 (9)	C6 ⁱⁱ —K1—C6—C3	-45.9 (6)
O1 ⁱⁱ —K1—O3—K1 ^v	-105.92 (13)	C6 ⁱⁱⁱ —K1—C6—C3	89.7 (6)
O1—K1—O3—K1 ^v	-29.55 (7)	K1 ⁱⁱ —K1—C6—C3	-45.9 (6)
O3 ⁱⁱ —K1—O3—K1 ^v	180.0	K1 ^v —K1—C6—C3	134.1 (6)
O2 ⁱⁱⁱ —K1—O3—K1 ^v	40.27 (6)	O2 ⁱ —K1—C6—K1 ^v	-116.44 (9)
O2—K1—O3—K1 ^v	-72.29 (7)	O1 ⁱⁱ —K1—C6—K1 ^v	-161.03 (6)
C6—K1—O3—K1 ^v	-51.77 (7)	O1—K1—C6—K1 ^v	-55.69 (16)
C6 ⁱⁱ —K1—O3—K1 ^v	-107.38 (9)	O3 ⁱⁱ —K1—C6—K1 ^v	118.82 (8)
C6 ⁱⁱⁱ —K1—O3—K1 ^v	56.64 (7)	O2 ⁱⁱⁱ —K1—C6—K1 ^v	-36.19 (6)
K1 ⁱⁱ —K1—O3—K1 ^v	180.0	O3—K1—C6—K1 ^v	44.22 (6)
C2—N1—C1—C5	0.1 (4)	O2—K1—C6—K1 ^v	105.19 (17)
C2—N1—C1—C1 ^{iv}	-179.0 (3)	C6 ⁱⁱ —K1—C6—K1 ^v	180.0
C1—N1—C2—C3	-0.4 (5)	C6 ⁱⁱⁱ —K1—C6—K1 ^v	-44.41 (7)

N1—C2—C3—C4	0.7 (5)	K1 ⁱⁱ —K1—C6—K1 ^v	180.0
N1—C2—C3—C6	-178.8 (3)	O2 ⁱ —K1—C6—K1 ^{vii}	173.13 (8)
C2—C3—C4—C5	-0.7 (4)	O1 ⁱⁱ —K1—C6—K1 ^{vii}	128.54 (9)
C6—C3—C4—C5	178.8 (3)	O1—K1—C6—K1 ^{vii}	-126.1 (2)
C3—C4—C5—C1	0.4 (4)	O3 ⁱⁱ —K1—C6—K1 ^{vii}	48.39 (12)
N1—C1—C5—C4	-0.1 (5)	O2 ⁱⁱⁱ —K1—C6—K1 ^{vii}	-106.62 (8)
C1 ^{iv} —C1—C5—C4	179.0 (3)	O3—K1—C6—K1 ^{vii}	-26.21 (8)
K1 ^{vi} —O2—C6—O1	-119.0 (4)	O2—K1—C6—K1 ^{vii}	34.76 (13)
K1 ^{vii} —O2—C6—O1	112.4 (3)	C6 ⁱⁱ —K1—C6—K1 ^{vii}	109.57 (6)
K1—O2—C6—O1	-19.7 (3)	C6 ⁱⁱⁱ —K1—C6—K1 ^{vii}	-114.84 (11)
K1 ^{vi} —O2—C6—C3	61.2 (6)	K1 ⁱⁱ —K1—C6—K1 ^{vii}	109.57 (6)
K1 ^{vii} —O2—C6—C3	-67.3 (3)	K1 ^v —K1—C6—K1 ^{vii}	-70.43 (6)
K1—O2—C6—C3	160.6 (2)		

Symmetry codes: (i) $-x, y+1/2, -z+1/2$; (ii) $x-1, y, z$; (iii) $-x+1, y+1/2, -z+1/2$; (iv) $-x+2, -y+1, -z+1$; (v) $x+1, y, z$; (vi) $-x, y-1/2, -z+1/2$; (vii) $-x+1, y-1/2, -z+1/2$.

Hydrogen-bond geometry ($\text{\AA}, ^\circ$)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
O3—H3A \cdots O1 ^{vii}	0.84 (5)	1.93 (5)	2.770 (4)	177 (5)
O3—H3B \cdots O4 ⁱⁱⁱ	0.84 (6)	2.15 (6)	2.976 (5)	169 (5)

Symmetry codes: (vii) $-x+1, y-1/2, -z+1/2$; (iii) $-x+1, y+1/2, -z+1/2$.

Fig. 1

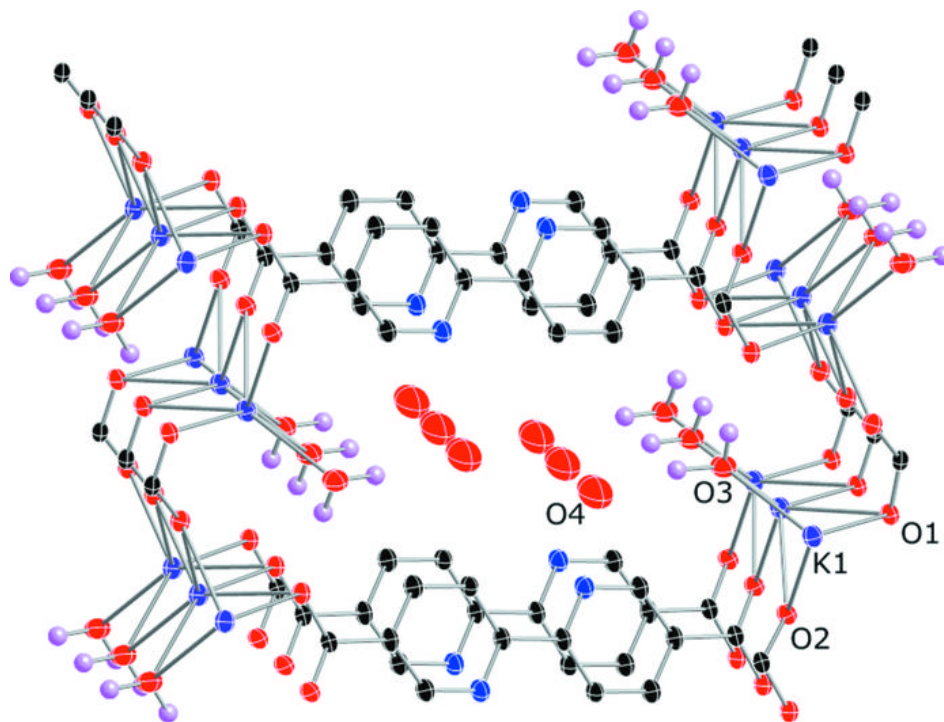


Fig. 2

