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Tetra- μ -benzoato- κ^8 O:O'-bis[(benzoic acid- κ O)nickel(II)]

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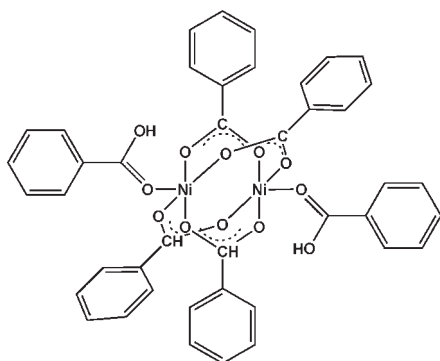
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Key indicators: single-crystal X-ray study; $T = 273$ K; mean $\sigma(\text{C}-\text{C}) = 0.003$ Å; R factor = 0.027; wR factor = 0.055; data-to-parameter ratio = 18.3.

The title compound, $[\text{Ni}_2(\text{C}_7\text{H}_5\text{O}_2)_4(\text{C}_7\text{H}_6\text{O}_2)_2]$, is composed of two Ni^{II} ions, four bridging benzoate anions and two η^1 -benzoic acid molecules. The $[\text{Ni}_2(\text{PhCOO})_4]$ unit adopts a typical paddle-wheel conformation. The center between the two Ni^{II} atoms represents a crystallographic center of inversion. In addition, each Ni^{II} ion also coordinates to one O atom from a benzoic acid molecule. The crystal packing is realised by intermolecular hydrogen-bonding interactions and π - π stacking interactions, with a centroid-centroid distance of 3.921 (1) Å.

Related literature

For related benzoate complexes, see: Cotton *et al.* (2005, 1987, 1988); Bellitto *et al.* (1985); Figuerola *et al.* (2007); Gavrilenko *et al.* (2008); Shi *et al.* (2004); Zheng *et al.* (2004); Zhong *et al.* (2007, 2008).



Experimental

Crystal data

 $[\text{Ni}_2(\text{C}_7\text{H}_5\text{O}_2)_4(\text{C}_7\text{H}_6\text{O}_2)_2]$
 $M_r = 846.10$

 Monoclinic, $P2_1/n$
 $a = 10.7685$ (8) Å

 $b = 11.7173$ (7) Å
 $c = 15.258$ (1) Å
 $\beta = 91.354$ (3)°
 $V = 1924.7$ (2) Å³
 $Z = 2$

 Mo $K\alpha$ radiation
 $\mu = 1.04$ mm⁻¹
 $T = 273$ K
 $0.28 \times 0.26 \times 0.20$ mm

Data collection

 Bruker SMART CCD area-detector diffractometer
 Absorption correction: multi-scan (SADABS; Sheldrick, 2008)
 $T_{\text{min}} = 0.759$, $T_{\text{max}} = 0.819$
 17534 measured reflections
 4639 independent reflections
 3789 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.022$

Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.027$
 $wR(F^2) = 0.055$
 $S = 0.99$
 4639 reflections

 254 parameters
 H-atom parameters constrained
 $\Delta\rho_{\text{max}} = 0.22$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.29$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{O6}-\text{H6}\cdots\text{O3}^{\text{i}}$	0.82	1.81	2.626 (2)	170

Symmetry code: (i) $-x + 1, -y, -z + 2$.

Data collection: SMART (Bruker, 2004); cell refinement: SAINT (Bruker, 2004); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: SHELXTL (Sheldrick, 2008); software used to prepare material for publication: SHELXTL.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: IM2153).

References

- Bellitto, C., Dessy, G. & Fares, V. (1985). *Inorg. Chem.* **24**, 2815–2820.
 Bruker (2004). SMART and SAINT. Bruker AXS Inc., Madison, Wisconsin, USA.
 Cotton, F.-A., Matusz, M. & Poli, R. (1987). *Inorg. Chem.* **26**, 1472–1474.
 Cotton, F.-A., Matusz, M., Poli, R. & Feng, X. (1988). *J. Am. Chem. Soc.* **110**, 1144–1154.
 Cotton, F. A., Murillo, C. A. & Walton, R. A. (2005). *Multiple Bonds between Metal Atoms*. New York: Springer Science and Business Media Inc.
 Figuerola, A., Tangoulis, V., Ribas, J., Hartl, H., Brudgam, I., Maestro, M. & Diaz, C. (2007). *Inorg. Chem.* **46**, 11017–11024.
 Gavrilenko, K.-S., Cadore, O., Bernot, K., Rosa, P., Sessoli, R., Golhen, S., Pavlishchuk, V.-V. & Ouahab, L. (2008). *Chem. Eur. J.* **14**, 2034–2043.
 Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
 Shi, X., Zhu, G.-S., Fang, Q.-R., Wu, G., Tian, G., Wang, R.-W., Zhang, D.-L., Xue, M. & Qiu, S.-L. (2004). *Eur. J. Inorg. Chem.* pp. 185–191.
 Zheng, S.-L., Yang, J.-H., Yu, X.-L., Chen, X.-M. & Wong, W.-T. (2004). *Inorg. Chem.* **43**, 830–838.
 Zhong, D.-C., Guo, G.-Q., Deng, J.-H., Zhu, R.-H. & Zhou, Y.-B. (2007). *Acta Cryst.* **E63**, m3091–m3092.
 Zhong, D.-C., Guo, G.-Q., Zuo, X.-H., Deng, J.-H., Yuan, L. & Zhu, R.-H. (2008). *Acta Cryst.* **E64**, m88–m89.

supplementary materials

Acta Cryst. (2009). E65, m1484 [doi:10.1107/S1600536809044766]

Tetra- μ -benzoato- κ^8 O:O'-bis[(benzoic acid- κ O)nickel(II)]

J.-H. Deng, Y.-P. Yi, Z.-X. Xiong, L. Yuan and G.-Q. Mei

Comment

Benzoic acid is the most simple aromatic carboxyl compound with well-known antibacterial activity. Over the past years, many metal complexes based on benzoic acid or benzoate ligands have been synthesized and characterized (Figuerola *et al.*, 2007; Gavrilenko *et al.*, 2008; Shi *et al.*, 2004; Zheng *et al.*, 2004). We also reported the Co(II) and Cd(II) complexes with benzoate and 2-aminopyridine ligands (Zhong *et al.*, 2007; Zhong *et al.*, 2008). As a continuation of this work, we herein report the synthesis and crystal structure of a nickel (II) complex exhibiting benzoate as well as benzoic acid ligands.

The title compound (I) is a typical paddle-wheel complex that have previously been observed (Bellitto *et al.*, 1985; Cotton *et al.*, 1987; Cotton *et al.*, 1988). Two Ni^{II} ions are bridged by four benzoate anions ligand using a μ -COO⁻ coordination mode. Each Ni^{II} also coordinates to one oxygen atom from one benzoic acid molecule in the axial position (Fig. 1). The center between the two nickel atoms represents a crystallographic center of inversion. The Ni—O bond lengths, ranging from 1.945 (1) Å to 2.193 (1) Å, are in the normal value range. The almost equivalent bond distances of O1—C8 and O2—C8 (1.260 (2) Å and 1.258 (2) Å) in one benzoate ligand and O3—C1 and O4—C1 (1.270 (2) Å and 1.255 (2) Å) in the other reflect the expected delocalization in the carboxylate unit. On the other hand bond distances of O5—C15 and O6—C15 (1.213 (2) Å and 1.317 (2) Å) in the axial benzoic acid ligands prove that it is not deprotonated and accordingly there is one single and one double bond. Albeit the short Ni \cdots Ni distance of 2.6062 (4) Å there is no bonding interaction between both metal centers due to the d⁸ electron configuration of Ni²⁺ leading to an overall bond order of zero (Cotton *et al.*, 2005)

The carboxyl group of the benzoic acid ligand (O6) acts as an intramolecular hydrogen bond donor site towards another oxygen atom of one of the bridging benzoate anions (O3). In addition, intermolecular contacts in terms of π - π stacking interactions with centroid centroid distances of 3.921 (1) Å are observed between the phenyl groups.

Experimental

All reagents are commercially available and were used without further purification. A mixture of Ni(NO₃)₂ × 4 H₂O (0.5 mmol), sodium benzoate (1 mmol) and 2,2'-bipyridine (0.5 mmol) was dissolved in 10 ml water/methanol (1/1). After stirring for 30 min, the mixture was placed in a 20 ml Teflon-lined reactor and heated to 110 °C in an oven for 7 days. The resulting solution was filtered and the filtrate was allowed to stay at room temperature. Well shaped blue crystals suitable for X-rays diffraction were obtained after one week. Yield: 65.8% based on sodium benzoate.

Refinement

All H atoms were placed geometrically and were refined using a riding model with C—H and O—H distances of 0.93 Å and 0.82 Å and $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$, $U_{\text{iso}}(\text{H}) = 1.5U_{\text{eq}}(\text{O})$.

Figures

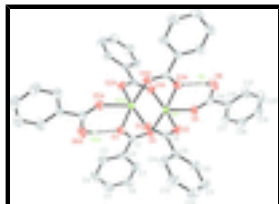


Fig. 1. The structure of **(I)**, showing 30% probability displacement ellipsoids and the atom-labeling scheme. [Symmetry code: (i) $-x+1, -y, -z+2$.]

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Crystal data

$[\text{Ni}_2(\text{C}_7\text{H}_5\text{O}_2)_4(\text{C}_7\text{H}_6\text{O}_2)_2]$

$M_r = 846.10$

Monoclinic, $P2_1/n$

Hall symbol: $-P\ 2_1n$

$a = 10.7685$ (8) Å

$b = 11.7173$ (7) Å

$c = 15.258$ (1) Å

$\beta = 91.354$ (3)°

$V = 1924.7$ (2) Å³

$Z = 2$

$F_{000} = 872$

$D_x = 1.460$ Mg m⁻³

Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å

Cell parameters from 7941 reflections

$\theta = 2.3$ – 28.0 °

$\mu = 1.04$ mm⁻¹

$T = 273$ K

Block, blue

$0.28 \times 0.26 \times 0.20$ mm

Data collection

Bruker SMART CCD area-detector diffractometer

Radiation source: fine-focus sealed tube

Monochromator: graphite

$T = 273$ K

φ and ω scans

Absorption correction: multi-scan (SADABS; Sheldrick, 2008)

$T_{\min} = 0.759$, $T_{\max} = 0.819$

17534 measured reflections

4639 independent reflections

3789 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.022$

$\theta_{\text{max}} = 28.1$ °

$\theta_{\text{min}} = 2.2$ °

$h = -14 \rightarrow 12$

$k = -15 \rightarrow 14$

$l = -18 \rightarrow 20$

Refinement

Refinement on F^2

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.027$

$wR(F^2) = 0.055$

$S = 0.99$

Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring sites

H-atom parameters constrained

$w = 1/[\sigma^2(F_o^2) + 1.1987P]$

where $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{\text{max}} = 0.002$

4639 reflections $\Delta\rho_{\max} = 0.22 \text{ e } \text{\AA}^{-3}$
 254 parameters $\Delta\rho_{\min} = -0.29 \text{ e } \text{\AA}^{-3}$
 Primary atom site location: structure-invariant direct methods Extinction correction: none

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
Ni1	0.517820 (18)	0.106516 (17)	0.979210 (13)	0.03259 (6)
O1	0.42315 (11)	-0.11994 (11)	0.89928 (7)	0.0471 (3)
O2	0.45485 (12)	0.06289 (11)	0.86345 (8)	0.0499 (3)
O3	0.31459 (11)	-0.05001 (10)	1.05426 (8)	0.0473 (3)
O4	0.34725 (11)	0.13053 (10)	1.01561 (8)	0.0483 (3)
O5	0.56860 (12)	0.28015 (11)	0.94023 (9)	0.0528 (3)
O6	0.75910 (14)	0.25119 (13)	0.89184 (13)	0.0800 (5)
H6	0.7451	0.1858	0.9080	0.120*
C1	0.27981 (15)	0.05301 (15)	1.04542 (10)	0.0397 (4)
C2	0.15133 (15)	0.08413 (16)	1.07061 (11)	0.0428 (4)
C3	0.07758 (19)	0.0110 (2)	1.11745 (14)	0.0656 (6)
H3	0.1080	-0.0598	1.1353	0.079*
C4	-0.0418 (2)	0.0433 (2)	1.13776 (17)	0.0817 (7)
H4	-0.0910	-0.0053	1.1704	0.098*
C5	-0.0880 (2)	0.1463 (2)	1.11020 (17)	0.0768 (7)
H5	-0.1690	0.1667	1.1230	0.092*
C6	-0.0157 (2)	0.2189 (2)	1.06420 (18)	0.0779 (7)
H6A	-0.0472	0.2889	1.0456	0.093*
C7	0.10426 (18)	0.18862 (18)	1.04522 (15)	0.0623 (6)
H7	0.1540	0.2393	1.0149	0.075*
C8	0.41868 (15)	-0.03704 (15)	0.84671 (11)	0.0402 (4)
C9	0.36303 (15)	-0.05843 (16)	0.75772 (11)	0.0429 (4)
C10	0.34030 (19)	0.03208 (18)	0.70105 (12)	0.0567 (5)
H10	0.3640	0.1056	0.7170	0.068*
C11	0.2822 (2)	0.0125 (2)	0.62069 (14)	0.0708 (6)
H11	0.2670	0.0731	0.5826	0.085*
C12	0.2468 (2)	-0.0961 (2)	0.59682 (14)	0.0715 (6)
H12	0.2077	-0.1087	0.5427	0.086*

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C13	0.2689 (2)	-0.1858 (2)	0.65259 (14)	0.0710 (6)
H13	0.2442	-0.2591	0.6365	0.085*
C14	0.32799 (19)	-0.16738 (18)	0.73287 (12)	0.0564 (5)
H14	0.3442	-0.2286	0.7702	0.068*
C15	0.66085 (18)	0.31531 (16)	0.90494 (12)	0.0464 (4)
C16	0.67474 (18)	0.43389 (16)	0.87387 (12)	0.0470 (4)
C17	0.5774 (2)	0.50892 (17)	0.88474 (13)	0.0563 (5)
H17	0.5048	0.4842	0.9105	0.068*
C18	0.5882 (2)	0.62111 (19)	0.85726 (14)	0.0684 (6)
H18	0.5231	0.6720	0.8651	0.082*
C19	0.6951 (3)	0.6569 (2)	0.81847 (15)	0.0758 (7)
H19	0.7021	0.7323	0.8000	0.091*
C20	0.7914 (3)	0.5829 (2)	0.80675 (17)	0.0809 (7)
H20	0.8631	0.6077	0.7798	0.097*
C21	0.7825 (2)	0.47092 (19)	0.83488 (15)	0.0670 (6)
H21	0.8484	0.4208	0.8277	0.080*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Ni1	0.03193 (10)	0.03109 (10)	0.03470 (10)	0.00072 (9)	-0.00051 (7)	0.00127 (9)
O1	0.0553 (7)	0.0457 (7)	0.0400 (6)	-0.0009 (6)	-0.0056 (5)	-0.0001 (6)
O2	0.0593 (8)	0.0479 (7)	0.0420 (7)	-0.0071 (6)	-0.0089 (6)	0.0007 (6)
O3	0.0406 (7)	0.0409 (7)	0.0607 (8)	0.0068 (6)	0.0076 (6)	0.0041 (6)
O4	0.0381 (6)	0.0430 (7)	0.0640 (8)	0.0051 (6)	0.0062 (6)	0.0046 (6)
O5	0.0531 (8)	0.0408 (7)	0.0648 (8)	0.0010 (6)	0.0111 (6)	0.0076 (6)
O6	0.0642 (10)	0.0483 (9)	0.1289 (15)	0.0090 (8)	0.0353 (9)	0.0193 (9)
C1	0.0377 (9)	0.0440 (10)	0.0373 (9)	0.0044 (8)	-0.0017 (7)	-0.0025 (7)
C2	0.0347 (9)	0.0493 (11)	0.0443 (9)	0.0041 (8)	0.0010 (7)	-0.0063 (8)
C3	0.0504 (12)	0.0710 (15)	0.0759 (15)	0.0086 (11)	0.0157 (10)	0.0131 (12)
C4	0.0524 (14)	0.094 (2)	0.1000 (19)	0.0019 (13)	0.0287 (13)	0.0100 (16)
C5	0.0427 (12)	0.0887 (19)	0.0995 (19)	0.0124 (12)	0.0143 (12)	-0.0141 (15)
C6	0.0511 (13)	0.0690 (16)	0.114 (2)	0.0214 (12)	0.0130 (13)	0.0013 (14)
C7	0.0466 (11)	0.0550 (13)	0.0860 (15)	0.0113 (10)	0.0142 (10)	0.0041 (11)
C8	0.0348 (9)	0.0466 (10)	0.0393 (9)	0.0026 (8)	0.0013 (7)	-0.0023 (8)
C9	0.0380 (9)	0.0526 (11)	0.0381 (9)	-0.0003 (8)	-0.0004 (7)	-0.0031 (8)
C10	0.0632 (13)	0.0571 (13)	0.0493 (11)	-0.0062 (10)	-0.0107 (9)	0.0030 (9)
C11	0.0832 (16)	0.0773 (16)	0.0508 (12)	0.0016 (13)	-0.0206 (11)	0.0083 (11)
C12	0.0783 (16)	0.0869 (18)	0.0483 (12)	0.0023 (14)	-0.0214 (11)	-0.0103 (12)
C13	0.0864 (17)	0.0681 (15)	0.0578 (13)	-0.0087 (13)	-0.0151 (12)	-0.0166 (11)
C14	0.0677 (13)	0.0549 (12)	0.0463 (10)	-0.0010 (10)	-0.0069 (9)	-0.0046 (9)
C15	0.0518 (11)	0.0414 (10)	0.0460 (10)	0.0004 (9)	0.0013 (8)	-0.0009 (8)
C16	0.0575 (11)	0.0395 (10)	0.0441 (10)	-0.0042 (9)	-0.0002 (8)	-0.0006 (8)
C17	0.0650 (13)	0.0499 (12)	0.0539 (12)	0.0032 (10)	-0.0006 (10)	0.0029 (9)
C18	0.0915 (17)	0.0472 (13)	0.0662 (14)	0.0111 (12)	-0.0076 (12)	0.0016 (10)
C19	0.112 (2)	0.0447 (13)	0.0706 (15)	-0.0124 (14)	-0.0019 (14)	0.0078 (11)
C20	0.0904 (19)	0.0598 (16)	0.0932 (18)	-0.0205 (14)	0.0171 (15)	0.0114 (13)
C21	0.0697 (15)	0.0512 (13)	0.0808 (15)	-0.0051 (11)	0.0141 (12)	0.0040 (11)

Geometric parameters (Å, °)

Ni1—O2	1.9452 (12)	C7—H7	0.9300
Ni1—O4	1.9519 (12)	C8—C9	1.492 (2)
Ni1—O1 ⁱ	1.9520 (11)	C9—C14	1.382 (3)
Ni1—O3 ⁱ	1.9999 (12)	C9—C10	1.386 (3)
Ni1—O5	2.1926 (13)	C10—C11	1.382 (3)
Ni1—Ni1 ⁱ	2.6062 (4)	C10—H10	0.9300
O1—C8	1.260 (2)	C11—C12	1.375 (3)
O1—Ni1 ⁱ	1.9520 (11)	C11—H11	0.9300
O2—C8	1.258 (2)	C12—C13	1.370 (3)
O3—C1	1.270 (2)	C12—H12	0.9300
O3—Ni1 ⁱ	1.9999 (12)	C13—C14	1.384 (3)
O4—C1	1.255 (2)	C13—H13	0.9300
O5—C15	1.213 (2)	C14—H14	0.9300
O6—C15	1.317 (2)	C15—C16	1.477 (3)
O6—H6	0.8200	C16—C17	1.381 (3)
C1—C2	1.490 (2)	C16—C21	1.386 (3)
C2—C3	1.379 (3)	C17—C18	1.385 (3)
C2—C7	1.377 (3)	C17—H17	0.9300
C3—C4	1.382 (3)	C18—C19	1.373 (3)
C3—H3	0.9300	C18—H18	0.9300
C4—C5	1.368 (3)	C19—C20	1.367 (3)
C4—H4	0.9300	C19—H19	0.9300
C5—C6	1.359 (3)	C20—C21	1.384 (3)
C5—H5	0.9300	C20—H20	0.9300
C6—C7	1.377 (3)	C21—H21	0.9300
C6—H6A	0.9300		
O2—Ni1—O4	89.20 (5)	O2—C8—O1	125.54 (16)
O2—Ni1—O1 ⁱ	169.19 (5)	O2—C8—C9	117.17 (16)
O4—Ni1—O1 ⁱ	90.32 (5)	O1—C8—C9	117.27 (16)
O2—Ni1—O3 ⁱ	88.77 (5)	C14—C9—C10	119.50 (17)
O4—Ni1—O3 ⁱ	168.91 (5)	C14—C9—C8	120.40 (17)
O1 ⁱ —Ni1—O3 ⁱ	89.64 (5)	C10—C9—C8	120.01 (17)
O2—Ni1—O5	94.67 (5)	C11—C10—C9	119.7 (2)
O4—Ni1—O5	100.71 (5)	C11—C10—H10	120.1
O1 ⁱ —Ni1—O5	96.03 (5)	C9—C10—H10	120.1
O3 ⁱ —Ni1—O5	90.32 (5)	C12—C11—C10	120.4 (2)
O2—Ni1—Ni1 ⁱ	85.36 (4)	C12—C11—H11	119.8
O4—Ni1—Ni1 ⁱ	85.63 (4)	C10—C11—H11	119.8
O1 ⁱ —Ni1—Ni1 ⁱ	83.83 (4)	C13—C12—C11	120.16 (19)
O3 ⁱ —Ni1—Ni1 ⁱ	83.34 (4)	C13—C12—H12	119.9
O5—Ni1—Ni1 ⁱ	173.66 (4)	C11—C12—H12	119.9
C8—O1—Ni1 ⁱ	123.31 (11)	C12—C13—C14	120.0 (2)

supplementary materials

C8—O2—Ni1	121.90 (11)	C12—C13—H13	120.0
C1—O3—Ni1 ⁱ	123.58 (11)	C14—C13—H13	120.0
C1—O4—Ni1	123.73 (11)	C13—C14—C9	120.3 (2)
C15—O5—Ni1	130.36 (12)	C13—C14—H14	119.9
C15—O6—H6	109.5	C9—C14—H14	119.9
O4—C1—O3	123.69 (15)	O5—C15—O6	122.89 (17)
O4—C1—C2	117.72 (15)	O5—C15—C16	123.51 (18)
O3—C1—C2	118.58 (16)	O6—C15—C16	113.59 (17)
C3—C2—C7	119.05 (17)	C17—C16—C21	119.83 (19)
C3—C2—C1	122.08 (17)	C17—C16—C15	118.53 (18)
C7—C2—C1	118.87 (17)	C21—C16—C15	121.64 (18)
C2—C3—C4	119.7 (2)	C16—C17—C18	119.9 (2)
C2—C3—H3	120.1	C16—C17—H17	120.0
C4—C3—H3	120.1	C18—C17—H17	120.0
C5—C4—C3	120.4 (2)	C19—C18—C17	119.8 (2)
C5—C4—H4	119.8	C19—C18—H18	120.1
C3—C4—H4	119.8	C17—C18—H18	120.1
C6—C5—C4	120.1 (2)	C18—C19—C20	120.6 (2)
C6—C5—H5	119.9	C18—C19—H19	119.7
C4—C5—H5	119.9	C20—C19—H19	119.7
C5—C6—C7	119.9 (2)	C19—C20—C21	120.2 (2)
C5—C6—H6A	120.0	C19—C20—H20	119.9
C7—C6—H6A	120.0	C21—C20—H20	119.9
C6—C7—C2	120.7 (2)	C20—C21—C16	119.7 (2)
C6—C7—H7	119.6	C20—C21—H21	120.2
C2—C7—H7	119.6	C16—C21—H21	120.2

Symmetry codes: (i) $-x+1, -y, -z+2$.

Hydrogen-bond geometry ($\text{\AA}, ^\circ$)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
O6—H6 ⁱ ···O3 ⁱ	0.82	1.81	2.626 (2)	170

Symmetry codes: (i) $-x+1, -y, -z+2$.

Fig. 1

