

Acta Crystallographica Section E

## Structure Reports

Online

ISSN 1600-5368

# Bis(2,2'-bi-1*H*-imidazole- $\kappa^2N^3,N^{3'}$ )(thiocyanato- $\kappa N$ )copper(II) chloride

Qi Ma,<sup>a</sup> Shuang-Ming Meng,<sup>a</sup> Feng Feng,<sup>a</sup> Li-Ping Lu<sup>b</sup> and Miao-Li Zhu<sup>b\*</sup>

<sup>a</sup>College of Chemistry and Chemical Engineering, Shanxi Datong University, Datong, Shanxi 037009, People's Republic of China, and <sup>b</sup>Institute of Molecular Science, Key Laboratory of Chemical Biology and Molecular Engineering of the Education Ministry, Shanxi University, Taiyuan, Shanxi 030006, People's Republic of China  
Correspondence e-mail: miaoli@sxu.edu.cn

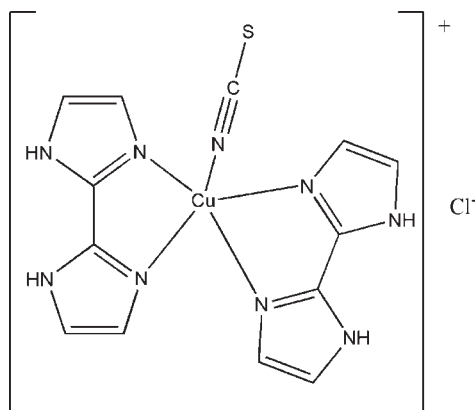
Received 26 October 2009; accepted 27 October 2009

Key indicators: single-crystal X-ray study;  $T = 298$  K; mean  $\sigma(\text{C}-\text{C}) = 0.009$  Å;  $R$  factor = 0.047;  $wR$  factor = 0.096; data-to-parameter ratio = 10.7.

In the title salt,  $[\text{Cu}(\text{NCS})(\text{C}_6\text{H}_6\text{N}_4)_2]\text{Cl}$ , the  $\text{Cu}^{\text{II}}$  atom adopts a five-coordinated square-pyramidal geometry consisting of an N atom from a thiocyanate anion and four N atoms from two chealting biimidazole ligands. The thiocyanate ligand occupies the axial position and is, like the  $\text{Cu}^{\text{II}}$  centre, located on a mirror plane. The cation and anion are linked into a linear chain by  $\text{N}-\text{H}\cdots\text{S}$  and  $\text{N}-\text{H}\cdots\text{Cl}$  hydrogen bonds.

## Related literature

For the neutral molecule 2,2'-biimidazole ( $\text{H}_2\text{biim}$ ) and its monoanionic derivative ( $\text{Hbiim}^-$ ), see: Tadokoro & Nakasuji (2000). Thiocyanate is a versatile bridging ligand, see: Ribas *et al.* (1998). For  $\text{Cu}-\text{N}$  bond lengths in biimidazole-Cu complexes, see: Govor *et al.* (2008);



## Experimental

### Crystal data

$[\text{Cu}(\text{NCS})(\text{C}_6\text{H}_6\text{N}_4)_2]\text{Cl}$   
 $M_r = 425.37$   
Orthorhombic,  $Cmc2_1$   
 $a = 12.900$  (5) Å  
 $b = 9.442$  (4) Å  
 $c = 12.888$  (5) Å

$V = 1569.8$  (11) Å<sup>3</sup>  
 $Z = 4$   
Mo  $K\alpha$  radiation  
 $\mu = 1.71$  mm<sup>-1</sup>  
 $T = 298$  K  
 $0.10 \times 0.10 \times 0.10$  mm

### Data collection

Bruker SMART CCD area-detector diffractometer  
Absorption correction: multi-scan (*SADABS*; Sheldrick, 1996)  
 $T_{\text{min}} = 0.610$ ,  $T_{\text{max}} = 0.847$

3518 measured reflections  
1291 independent reflections  
1254 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.046$

### Refinement

$R[F^2 > 2\sigma(F^2)] = 0.047$   
 $wR(F^2) = 0.096$   
 $S = 1.20$   
1291 reflections  
121 parameters  
1 restraint

H-atom parameters constrained  
 $\Delta\rho_{\text{max}} = 0.45$  e Å<sup>-3</sup>  
 $\Delta\rho_{\text{min}} = -0.45$  e Å<sup>-3</sup>  
Absolute structure: Flack (1983), 539 Friedel pairs  
Flack parameter: 0.06 (3)

**Table 1**

Selected bond lengths (Å).

Cu1—N1	2.014 (6)	Cu1—N5	2.344 (10)
Cu1—N3	2.031 (5)		

**Table 2**

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
N2—H2A $\cdots$ Cl <sup>ii</sup>	0.86	2.27	3.092 (5)	160
N4—H4 $\cdots$ Cl <sup>i</sup>	0.86	2.52	3.305 (6)	153
N2—H2A $\cdots$ S1 <sup>iii</sup>	0.86	3.36	3.782 (6)	113
N4—H4 $\cdots$ S1 <sup>iv</sup>	0.86	3.38	3.818 (6)	114

Symmetry codes: (ii)  $x + 1, y, z$ ; (iii)  $-x + \frac{3}{2}, -y + \frac{3}{2}, z + \frac{1}{2}$ ; (iv)  $x - \frac{1}{2}, y - \frac{1}{2}, z$ .

Data collection: *SMART* (Bruker, 2000); cell refinement: *SAINT* (Bruker, 2000); data reduction: *SAINT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL/PC* (Sheldrick, 2008); software used to prepare material for publication: *publCIF* (Westrip, 2009).

The authors acknowledge financial support from the National Natural Science Foundation of China (grant No. 20471033), the Natural Science Foundation of Shanxi Province (grant No. 20051013) and the Overseas Returned Scholar Foundation of Shanxi Province in 2008.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: NG2677).

**References**

- Bruker (2000). *SMART* and *SAINT*. Bruker AXS Inc., Madison, Wisconsin, USA.
- Flack, H. D. (1983). *Acta Cryst.* **A39**, 876–881.
- Govor, E. V., Lysenko, A. B., Domasevitch, K. V., Rusanov, E. B. & Chernega, A. N. (2008). *Acta Cryst.* **C64**, m117–m120.
- Ribas, J., Diaz, C., Costa, R., Tercero, J., Solans, X., Font-Bardia, M. & Stoeckli-Evans, H. (1998). *Inorg. Chem.* **37**, 233–239.
- Sheldrick, G. M. (1996). *SADABS*. University of Göttingen, Germany.
- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
- Tadokoro, M. & Nakasuji, K. (2000). *Coord. Chem. Rev.* **198**, 205–218.
- Westrip, S. P. (2009). *publCIF*. In preparation.

**supplementary materials**

*Acta Cryst.* (2009). E65, m1488-m1489 [ doi:10.1107/S1600536809044717 ]

## Bis(2,2'-bi-1*H*-imidazole- $\kappa^2N^3,N^{3'}$ )(thiocyanato- $\kappa N$ )copper(II) chloride

Q. Ma, S.-M. Meng, F. Feng, L.-P. Lu and M.-L. Zhu

### Comment

The neutral molecule 2,2'-biimidazole ( $H_2biim$ ) and its monoanionic derivative ( $Hbiim^-$ ) is a particular organic target for construction of hybrid materials. Its molecular moieties possess a double property, namely they can be coordinated to metal centres and can act as a donor in hydrogen bonding interactions (Tadokoro *et al.*, 2000). The thiocyanate is a versatile bridging ligand (Ribas *et al.*, 1998), we think that the self-assembly of these ligand with metal ions should yield structure fascinating compounds. Thus, the title compound (I) was synthesized and its crystal structure is reported here.

The X-ray crystallographic analysis shows that the crystal structure of (I) consists of  $[Cu(NCS)(biim)_2]^+$  cation and  $Cl^-$  anions (Fig 1). Cu(II) ion adopts a five coordinated square pyramidal geometry consisting of a nitrogen atom (N5) from thiocyanato anion and four nitrogen atoms (N1, N1<sup>i</sup>, N3 and N3<sup>i</sup>) from two coordinating biimidazole ligands. Four nitrogen atoms of two chelating  $H_2biim$  ligands form the basal plane of the pyramid and the apical position is occupied by the thiocyanate ligand which is coordinated in the axial position through the nitrogen atom. Bond distances of the Cu—N1 and Cu—N3(biim) [2.014 (6) and 2.031 (5) Å] (Table 1) are shorter than the apical Cu1—N5(SCN<sup>-</sup>) distance [2.344 (10) Å]. These distances lie in the range reported for biimidazole-Cu complexes (Govor *et al.*, 2008). The chelating  $H_2biim$  ligands are almost planar and dihedral angle of two biimidazole plane is 6.32°. Meanwhile, In the crystal, molecules are linked by hydrogen bond interaction (N—H...Cl) forming the three-dimensional architecture (Fig 2).

### Experimental

All chemicals were of reagent grade, were commercially available and were used without further purification. CuCl (0.099 g, 0.10 mmol) dissolved in 10.0 ml of ethanol solution was added to 10.0 ml of ethanol solution of  $H_2biim$  (0.0134 g, 0.10 mmol) with stirring. Half an hour later, KSCN (0.0195 g, 0.20 mmol) was slowly added above the mixture. Black green crystal of  $[Cu(biim)_2(SCN)]Cl$  were separated from the mother liquor by slow evaporation at room temperature after two weeks.

### Refinement

H atoms attached to C and N(biimidazole) atoms of (I) were placed in geometrically idealized positions with  $Csp^2-H = 0.93\text{Å}$  and  $N-H = 0.86\text{Å}$  and constrained to ride on their parent atoms.

## Figures

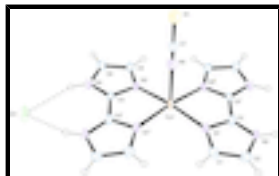


Fig. 1. A view of the structure of (I) with displacement ellipsoids drawn at the 30% probability level. Symmetrical code (i)  $2 - x, y, z$

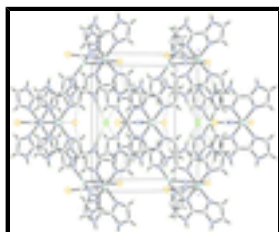


Fig. 2. The packing view in (I). Cu (Light cyan); Cl (green); S (yellow); N (blue); C (gray)

## Bis(2,2'-bi-1H-imidazole- $\kappa^2N^3, N^{3'}$ )(thiocyanato- $\kappa N$ )copper(II) chloride

### Crystal data

[Cu(NCS)(C<sub>6</sub>H<sub>6</sub>N<sub>4</sub>)<sub>2</sub>]Cl

$M_r = 425.37$

Orthorhombic, *Cmc*2<sub>1</sub>

Hall symbol: C 2c -2

$a = 12.900 (5) \text{ \AA}$

$b = 9.442 (4) \text{ \AA}$

$c = 12.888 (5) \text{ \AA}$

$V = 1569.8 (11) \text{ \AA}^3$

$Z = 4$

$F_{000} = 860$

$D_x = 1.800 \text{ Mg m}^{-3}$

Mo  $K\alpha$  radiation,  $\lambda = 0.71073 \text{ \AA}$

Cell parameters from 4272 reflections

$\theta = 2.1\text{--}26.6^\circ$

$\mu = 1.71 \text{ mm}^{-1}$

$T = 298 \text{ K}$

Block, green

$0.10 \times 0.10 \times 0.10 \text{ mm}$

### Data collection

Bruker SMART CCD area-detector diffractometer

Radiation source: fine-focus sealed tube

Monochromator: graphite

$T = 298 \text{ K}$

$\omega$  scans

Absorption correction: multi-scan (SADABS; Sheldrick, 1996)

$T_{\min} = 0.610$ ,  $T_{\max} = 0.847$

3518 measured reflections

1291 independent reflections

1254 reflections with  $I > 2\sigma(I)$

$R_{\text{int}} = 0.046$

$\theta_{\max} = 25.0^\circ$

$\theta_{\min} = 3.1^\circ$

$h = -14 \rightarrow 15$

$k = -11 \rightarrow 10$

$l = -13 \rightarrow 15$

### Refinement

Refinement on  $F^2$

Least-squares matrix: full

Hydrogen site location: inferred from neighbouring sites

H-atom parameters constrained

$R[F^2 > 2\sigma(F^2)] = 0.047$	$w = 1/[\sigma^2(F_o^2) + (0.0005P)^2 + 5.3473P]$
$wR(F^2) = 0.096$	where $P = (F_o^2 + 2F_c^2)/3$
$S = 1.20$	$(\Delta/\sigma)_{\max} < 0.001$
1291 reflections	$\Delta\rho_{\max} = 0.45 \text{ e } \text{\AA}^{-3}$
121 parameters	$\Delta\rho_{\min} = -0.45 \text{ e } \text{\AA}^{-3}$
1 restraint	Extinction correction: none
Primary atom site location: structure-invariant direct methods	Absolute structure: Flack (1983), 539 Friedel pairs
Secondary atom site location: difference Fourier map	Flack parameter: 0.06 (3)

### Special details

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

### Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )

	$x$	$y$	$z$	$U_{\text{iso}}^*/U_{\text{eq}}$
Cu1	0.5000	0.39548 (11)	0.38462 (9)	0.0307 (3)
S1	0.5000	0.7886 (3)	0.1136 (2)	0.0476 (7)
N1	0.6167 (3)	0.4872 (6)	0.4638 (4)	0.0317 (12)
N2	0.7857 (4)	0.4990 (6)	0.4874 (4)	0.0316 (13)
H2A	0.8509	0.4821	0.4808	0.038*
N3	0.3808 (4)	0.2803 (5)	0.3256 (4)	0.0294 (12)
N4	0.2133 (4)	0.2519 (6)	0.3172 (4)	0.0333 (13)
H4	0.1484	0.2647	0.3294	0.040*
N5	0.5000	0.5634 (10)	0.2506 (7)	0.044 (2)
C1	0.6381 (5)	0.5905 (6)	0.5367 (5)	0.0321 (15)
H1	0.5893	0.6466	0.5703	0.039*
C2	0.7418 (5)	0.5963 (8)	0.5509 (6)	0.0369 (18)
H2	0.7767	0.6563	0.5962	0.044*
C3	0.7087 (4)	0.4344 (7)	0.4372 (5)	0.0302 (15)
C4	0.2915 (4)	0.3246 (6)	0.3605 (4)	0.0248 (14)
C5	0.3575 (5)	0.1711 (8)	0.2579 (5)	0.0359 (16)
H5	0.4062	0.1168	0.2225	0.043*
C6	0.2537 (5)	0.1550 (10)	0.2509 (6)	0.0345 (18)
H6	0.2176	0.0911	0.2094	0.041*
C7	0.5000	0.6548 (10)	0.1940 (8)	0.034 (2)
Cl1	0.0000	0.3818 (3)	0.4227 (2)	0.0501 (8)

## supplementary materials

---

### Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
Cu1	0.0153 (5)	0.0344 (6)	0.0425 (6)	0.000	0.000	-0.0068 (6)
S1	0.0303 (13)	0.0513 (17)	0.0612 (19)	0.000	0.000	0.0124 (14)
N1	0.019 (2)	0.039 (3)	0.037 (3)	0.004 (2)	0.004 (2)	0.007 (3)
N2	0.015 (2)	0.033 (3)	0.047 (4)	0.001 (2)	-0.002 (2)	-0.004 (3)
N3	0.019 (2)	0.033 (3)	0.036 (3)	-0.005 (2)	-0.002 (2)	-0.005 (2)
N4	0.021 (3)	0.037 (3)	0.042 (3)	-0.001 (2)	-0.005 (2)	0.002 (3)
N5	0.031 (4)	0.055 (6)	0.046 (6)	0.000	0.000	-0.006 (4)
C1	0.032 (3)	0.025 (4)	0.040 (4)	-0.003 (3)	0.003 (3)	-0.004 (3)
C2	0.037 (4)	0.041 (5)	0.033 (4)	-0.005 (3)	-0.001 (3)	-0.002 (4)
C3	0.021 (3)	0.035 (4)	0.035 (4)	0.003 (3)	0.002 (3)	0.003 (3)
C4	0.018 (3)	0.032 (3)	0.024 (4)	-0.003 (2)	-0.002 (2)	0.010 (3)
C5	0.022 (3)	0.043 (4)	0.042 (4)	0.001 (3)	0.005 (3)	-0.001 (3)
C6	0.026 (4)	0.039 (4)	0.038 (4)	-0.010 (3)	0.001 (3)	-0.004 (3)
C7	0.021 (4)	0.026 (5)	0.054 (7)	0.000	0.000	0.006 (5)
Cl1	0.0206 (11)	0.0479 (15)	0.082 (2)	0.000	0.000	-0.0179 (14)

### Geometric parameters ( $\text{\AA}$ , $^\circ$ )

Cu1—N1	2.014 (6)	N4—C4	1.342 (7)
Cu1—N1 <sup>i</sup>	2.014 (6)	N4—C6	1.356 (10)
Cu1—N3	2.031 (5)	N4—H4	0.8600
Cu1—N3 <sup>i</sup>	2.031 (5)	N5—C7	1.130 (13)
Cu1—N5	2.344 (10)	C1—C2	1.351 (9)
S1—C7	1.634 (11)	C1—H1	0.9300
N1—C3	1.332 (7)	C2—H2	0.9300
N1—C1	1.382 (8)	C3—C4 <sup>i</sup>	1.432 (8)
N2—C3	1.333 (8)	C4—C3 <sup>i</sup>	1.432 (8)
N2—C2	1.354 (9)	C5—C6	1.352 (9)
N2—H2A	0.8600	C5—H5	0.9300
N3—C4	1.305 (7)	C6—H6	0.9300
N3—C5	1.383 (8)		
N1—Cu1—N1 <sup>i</sup>	96.7 (3)	C6—N4—H4	125.7
N1—Cu1—N3	170.3 (2)	C7—N5—Cu1	172.8 (8)
N1 <sup>i</sup> —Cu1—N3	81.62 (18)	C2—C1—N1	108.6 (6)
N1—Cu1—N3 <sup>i</sup>	81.62 (18)	C2—C1—H1	125.7
N1 <sup>i</sup> —Cu1—N3 <sup>i</sup>	170.3 (2)	N1—C1—H1	125.7
N3—Cu1—N3 <sup>i</sup>	98.4 (3)	C1—C2—N2	107.8 (7)
N1—Cu1—N5	94.7 (2)	C1—C2—H2	126.1
N1 <sup>i</sup> —Cu1—N5	94.7 (2)	N2—C2—H2	126.1
N3—Cu1—N5	95.0 (2)	N1—C3—N2	111.6 (6)
N3 <sup>i</sup> —Cu1—N5	95.0 (2)	N1—C3—C4 <sup>i</sup>	116.5 (6)
C3—N1—C1	105.1 (5)	N2—C3—C4 <sup>i</sup>	131.9 (6)

C3—N1—Cu1	112.0 (5)	N3—C4—N4	110.9 (5)
C1—N1—Cu1	142.9 (4)	N3—C4—C3 <sup>i</sup>	118.2 (5)
C3—N2—C2	106.9 (5)	N4—C4—C3 <sup>i</sup>	131.0 (6)
C3—N2—H2A	126.5	C6—C5—N3	110.0 (7)
C2—N2—H2A	126.5	C6—C5—H5	125.0
C4—N3—C5	105.3 (5)	N3—C5—H5	125.0
C4—N3—Cu1	111.5 (4)	C5—C6—N4	105.2 (7)
C5—N3—Cu1	143.1 (4)	C5—C6—H6	127.4
C4—N4—C6	108.6 (5)	N4—C6—H6	127.4
C4—N4—H4	125.7	N5—C7—S1	179.1 (10)
N1 <sup>i</sup> —Cu1—N1—C3	172.7 (3)	C1—N1—C3—N2	-0.8 (7)
N3 <sup>i</sup> —Cu1—N1—C3	2.3 (4)	Cu1—N1—C3—N2	177.9 (4)
N5—Cu1—N1—C3	-92.0 (5)	C1—N1—C3—C4 <sup>i</sup>	-179.4 (5)
N1 <sup>i</sup> —Cu1—N1—C1	-9.4 (9)	Cu1—N1—C3—C4 <sup>i</sup>	-0.6 (7)
N3 <sup>i</sup> —Cu1—N1—C1	-179.7 (7)	C2—N2—C3—N1	1.1 (8)
N5—Cu1—N1—C1	85.9 (7)	C2—N2—C3—C4 <sup>i</sup>	179.4 (6)
N1 <sup>i</sup> —Cu1—N3—C4	3.8 (4)	C5—N3—C4—N4	-1.3 (7)
N3 <sup>i</sup> —Cu1—N3—C4	173.9 (3)	Cu1—N3—C4—N4	177.1 (4)
N5—Cu1—N3—C4	-90.3 (4)	C5—N3—C4—C3 <sup>i</sup>	177.0 (5)
N1 <sup>i</sup> —Cu1—N3—C5	-178.8 (8)	Cu1—N3—C4—C3 <sup>i</sup>	-4.6 (6)
N3 <sup>i</sup> —Cu1—N3—C5	-8.6 (9)	C6—N4—C4—N3	0.2 (8)
N5—Cu1—N3—C5	87.1 (7)	C6—N4—C4—C3 <sup>i</sup>	-177.8 (7)
C3—N1—C1—C2	0.2 (7)	C4—N3—C5—C6	2.0 (8)
Cu1—N1—C1—C2	-177.8 (6)	Cu1—N3—C5—C6	-175.5 (6)
N1—C1—C2—N2	0.5 (8)	N3—C5—C6—N4	-1.9 (9)
C3—N2—C2—C1	-1.0 (8)	C4—N4—C6—C5	1.1 (9)

Symmetry codes: (i)  $-x+1, y, z$ .

*Hydrogen-bond geometry* ( $\text{\AA}$ ,  $^\circ$ )

<i>D</i> —H $\cdots$ <i>A</i>	<i>D</i> —H	H $\cdots$ <i>A</i>	<i>D</i> $\cdots$ <i>A</i>	<i>D</i> —H $\cdots$ <i>A</i>
N2—H2A $\cdots$ C11 <sup>ii</sup>	0.86	2.27	3.092 (5)	160
N4—H4 $\cdots$ C11	0.86	2.52	3.305 (6)	153
N2—H2A $\cdots$ S1 <sup>iii</sup>	0.86	3.36	3.782 (6)	113
N4—H4 $\cdots$ S1 <sup>iv</sup>	0.86	3.38	3.818 (6)	114

Symmetry codes: (ii)  $x+1, y, z$ ; (iii)  $-x+3/2, -y+3/2, z+1/2$ ; (iv)  $x-1/2, y-1/2, z$ .

Fig. 1

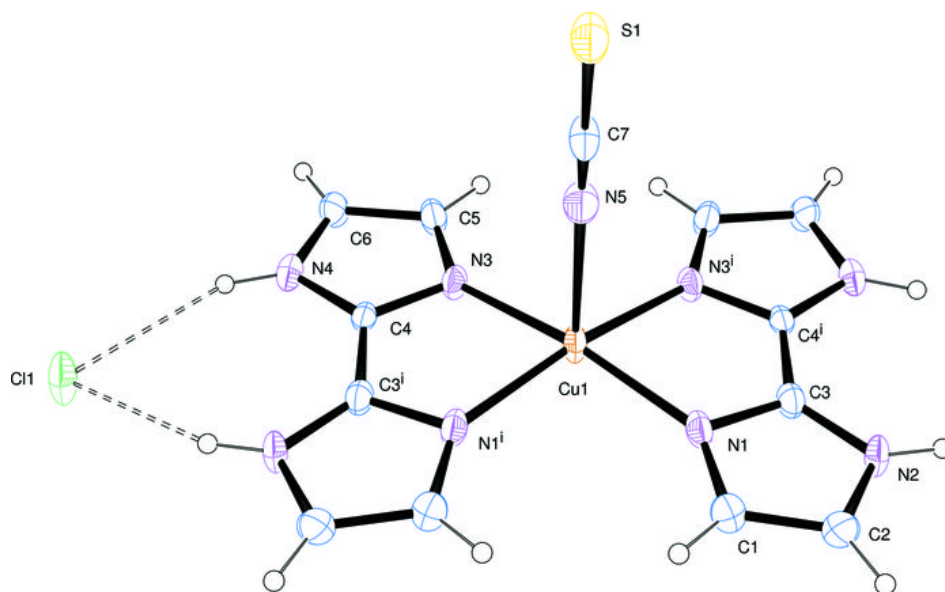


Fig. 2

