

Acta Crystallographica Section E

Structure Reports

Online

ISSN 1600-5368

Diiodido{4-nitro-2-[2-(piperidin-1-yl)-ethyliminomethyl]phenolato}zinc(II)

Xue-Wen Zhu,* Xu-Zhao Yang, Chun-Xia Zhang,
Gang-Sen Li and Zhi-Gang Yin

Key Laboratory of Surface and Interface Science of Henan, School of Material & Chemical Engineering, Zhengzhou University of Light Industry, Zhengzhou 450002, People's Republic of China

Correspondence e-mail: xuewen-zhu@163.com

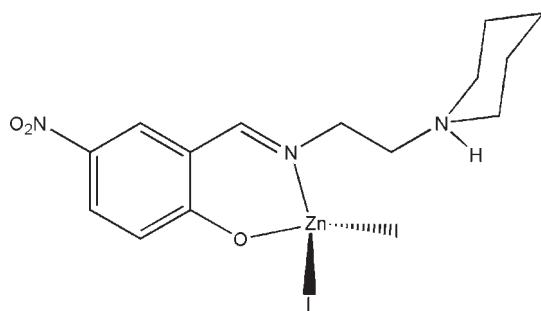
Received 19 September 2009; accepted 5 October 2009

Key indicators: single-crystal X-ray study; $T = 298$ K; mean $\sigma(\text{C}-\text{C}) = 0.006$ Å; R factor = 0.036; wR factor = 0.080; data-to-parameter ratio = 19.1.

In the title complex, $[\text{ZnI}_2(\text{C}_{14}\text{H}_{19}\text{N}_3\text{O}_3)]$, the Zn^{II} atom is four-coordinated by the imine N and phenolate O atoms of the Schiff base ligand, and by two iodide ions in a distorted tetrahedral coordination. In the crystal structure, molecules are linked through intermolecular $\text{N}-\text{H}\cdots\text{O}$ hydrogen bonds, forming dimers.

Related literature

For background to the chemistry of Schiff base complexes, see: Ali *et al.* (2008); Biswas *et al.* (2008); Chen *et al.* (2008); Darensbourg & Frantz (2007); Habibi *et al.* (2007); Kawamoto *et al.* (2008); Lipscomb & Sträter (1996); Tomat *et al.* (2007); Wu *et al.* (2008); Yuan *et al.* (2007). For related structures see: Zhu (2008); Zhu & Yang (2008*a,b,c*); Qiu (2006*a,b*); Wei *et al.* (2007); Zhu *et al.* (2007).



Experimental

Crystal data

$[\text{ZnI}_2(\text{C}_{14}\text{H}_{19}\text{N}_3\text{O}_3)]$
 $M_r = 596.49$
Triclinic, $P\bar{1}$
 $a = 8.7467$ (2) Å
 $b = 10.7114$ (3) Å
 $c = 10.9541$ (2) Å

$\alpha = 89.553$ (2)°
 $\beta = 89.334$ (2)°
 $\gamma = 68.984$ (2)°
 $V = 957.94$ (4) Å³
 $Z = 2$
Mo $K\alpha$ radiation

$\mu = 4.52$ mm⁻¹
 $T = 298$ K

0.20 × 0.20 × 0.18 mm

Data collection

Bruker APEXII CCD area-detector diffractometer
Absorption correction: multi-scan (SADABS; Sheldrick, 2004)
 $T_{\text{min}} = 0.465$, $T_{\text{max}} = 0.497$

5795 measured reflections
4034 independent reflections
2994 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.017$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.036$
 $wR(F^2) = 0.080$
 $S = 1.01$
4034 reflections
211 parameters
1 restraint

H atoms treated by a mixture of independent and constrained refinement
 $\Delta\rho_{\text{max}} = 0.50$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.62$ e Å⁻³

Table 1

Selected geometric parameters (Å, °).

Zn1—O1	1.948 (3)	Zn1—I2	2.5666 (6)
Zn1—N1	2.037 (3)	Zn1—I1	2.5690 (6)
O1—Zn1—N1	94.96 (13)	O1—Zn1—I1	121.56 (10)
O1—Zn1—I2	107.27 (10)	N1—Zn1—I1	102.97 (10)
N1—Zn1—I2	113.81 (10)	I2—Zn1—I1	114.61 (2)

Table 2

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
N2—H2 \cdots O2 ⁱ	0.89 (4)	2.00 (3)	2.777 (5)	144 (5)

Symmetry code: (i) $-x, -y + 1, -z + 1$.

Data collection: APEX2 (Bruker, 2004); cell refinement: SAINT (Bruker, 2004); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: SHELXTL (Sheldrick, 2008); software used to prepare material for publication: SHELXL97.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: OM2281).

References

- Ali, H. M., Mohamed Mustafa, M. I., Rizal, M. R. & Ng, S. W. (2008). *Acta Cryst. E* **64**, m718–m719.
Biswas, C., Drew, M. G. B. & Ghosh, A. (2008). *Inorg. Chem.* **47**, 4513–4519.
Bruker (2004). APEX2 and SAINT. Bruker AXS Inc., Madison, Wisconsin, USA.
Chen, Z., Morimoto, H., Matsunaga, S. & Shibasaki, M. (2008). *J. Am. Chem. Soc.* **130**, 2170–2171.
Darensbourg, D. J. & Frantz, E. B. (2007). *Inorg. Chem.* **46**, 5967–5978.
Habibi, M. H., Askari, E., Chantrapromma, S. & Fun, H.-K. (2007). *Acta Cryst. E* **63**, m2905–m2906.
Kawamoto, T., Nishiwaki, M., Tsunekawa, Y., Nozaki, K. & Konno, T. (2008). *Inorg. Chem.* **47**, 3095–3104.
Lipscomb, W. N. & Sträter, N. (1996). *Chem. Rev.* **96**, 2375–2434.
Qiu, X.-Y. (2006*a*). *Acta Cryst. E* **62**, m717–m718.
Qiu, X.-Y. (2006*b*). *Acta Cryst. E* **62**, m2173–m2174.
Sheldrick, G. M. (2004). SADABS. University of Göttingen, Germany.

- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
- Tomat, E., Cuesta, L., Lynch, V. M. & Sessler, J. L. (2007). *Inorg. Chem.* **46**, 6224–6226.
- Wei, Y.-J., Wang, F.-W. & Zhu, Q.-Y. (2007). *Acta Cryst.* **E63**, m654–m655.
- Wu, J.-C., Liu, S.-X., Keene, T. D., Neels, A., Mereacre, V., Powell, A. K. & Decurtins, S. (2008). *Inorg. Chem.* **47**, 3452–3459.
- Yuan, M., Zhao, F., Zhang, W., Wang, Z.-M. & Gao, S. (2007). *Inorg. Chem.* **46**, 11235–11242.
- Zhu, X.-W. (2008). *Acta Cryst.* **E64**, m1456–m1457.
- Zhu, Q.-Y., Wei, Y.-J. & Wang, F.-W. (2007). *Acta Cryst.* **E63**, m1431–m1432.
- Zhu, X.-W. & Yang, X.-Z. (2008a). *Acta Cryst.* **E64**, m1090–m1091.
- Zhu, X.-W. & Yang, X.-Z. (2008b). *Acta Cryst.* **E64**, m1092–m1093.
- Zhu, X.-W. & Yang, X.-Z. (2008c). *Acta Cryst.* **E64**, m1094–m1095.

supplementary materials

Acta Cryst. (2009). E65, m1332-m1333 [doi:10.1107/S1600536809040495]

Diiodido{4-nitro-2-[2-(piperidin-1-yl)ethyliminomethyl]phenolato}zinc(II)

X.-W. Zhu, X.-Z. Yang, C.-X. Zhang, G.-S. Li and Z.-G. Yin

Comment

Schiff bases are interesting ligands in coordination chemistry (Biswas *et al.*, 2008; Wu *et al.*, 2008; Kawamoto *et al.*, 2008; Ali *et al.*, 2008; Habibi *et al.*, 2007), and their complexes have been investigated in many fields (Chen *et al.*, 2008; Yuan *et al.*, 2007; Tomat *et al.*, 2007; Darensbourg & Frantz, 2007). Zinc(II) is an important element in biological systems and functions as the active site of hydrolytic enzymes, such as carboxypeptidase and carbonic anhydrase where it is in a hard-donor coordination environment of nitrogen and oxygen ligands (Lipscomb & Sträter, 1996). Recently, we have reported a few Schiff base zinc complexes (Zhu, 2008; Zhu & Yang, 2008a,b,c). In this paper, the title new zinc(II) complex, Fig. 1, is reported.

In the title complex, the Zn^{II} atom is four-coordinated by the imine N and phenolate O atoms of the Schiff base ligand, and by two iodide ions in a tetrahedral coordination. The coordinate bond lengths (Table 1) are typical and comparable to the corresponding values observed in the Schiff base zinc complexes we reported previously and other similar Schiff base zinc complexes (Zhu *et al.*, 2007; Wei *et al.*, 2007; Qiu, 2006a,b).

In the crystal structure, molecules are linked through intermolecular N—H···O hydrogen bonds, forming dimers (Table 2, Fig. 2).

Experimental

The Schiff base compound was prepared by the condensation of equimolar amounts of 5-nitrosalicylaldehyde with 2-piperidin-1-ylethylamine in a methanol solution. The complex was prepared by the following method. To an anhydrous methanol solution (5 ml) of ZnI₂ (31.9 mg, 0.1 mmol) was added a methanol solution (10 ml) of the Schiff base compound (27.7 mg, 0.1 mmol) with stirring. The mixture was stirred for 30 min at room temperature and filtered. Upon keeping the filtrate in air for a few days, colorless block-shaped crystals were formed.

Refinement

H2 was located from a difference Fourier map and refined isotropically, with N—H distance restrained to 0.90 (1) Å. Other H atoms were placed in geometrically idealized positions and constrained to ride on their parent atoms, with C—H distances in the range 0.93–0.97 Å, and with $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$.

Figures

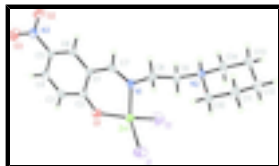


Fig. 1. The molecular structure of the title complex, with ellipsoids drawn at the 30% probability level.

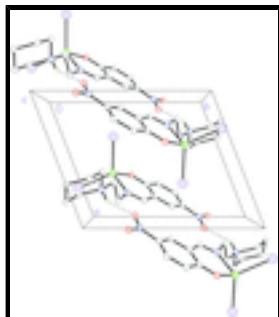


Fig. 2. The crystal packing of the title complex, viewed along the *c* axis. Intermolecular hydrogen bonds are shown as dashed lines.

Diiodido{4-nitro-2-[2-(piperidin-1-yl)ethyliminomethyl]phenolato}zinc(II)

Crystal data

[ZnI₂(C₁₄H₁₉N₃O₃)]

M_r = 596.49

Triclinic, *P* $\bar{1}$

a = 8.7467 (2) Å

b = 10.7114 (3) Å

c = 10.9541 (2) Å

α = 89.553 (2)°

β = 89.334 (2)°

γ = 68.984 (2)°

V = 957.94 (4) Å³

Z = 2

*F*₀₀₀ = 568

D_x = 2.068 Mg m⁻³

Mo *K*α radiation, λ = 0.71073 Å

Cell parameters from 1800 reflections

θ = 2.6–25.5°

μ = 4.52 mm⁻¹

T = 298 K

Block, colorless

0.20 × 0.20 × 0.18 mm

Data collection

Bruker APEXII CCD area-detector diffractometer

4034 independent reflections

Radiation source: fine-focus sealed tube

2994 reflections with *I* > 2σ(*I*)

Monochromator: graphite

*R*_{int} = 0.017

T = 298 K

θ_{\max} = 27.0°

ω scans

θ_{\min} = 1.9°

Absorption correction: multi-scan (SADABS; Sheldrick, 2004)

h = -11→10

*T*_{min} = 0.465, *T*_{max} = 0.497

k = -11→13

5795 measured reflections

l = -13→13

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.036$	H atoms treated by a mixture of independent and constrained refinement
$wR(F^2) = 0.080$	$w = 1/[\sigma^2(F_o^2) + (0.0311P)^2 + 0.4408P]$
$S = 1.01$	where $P = (F_o^2 + 2F_c^2)/3$
4034 reflections	$(\Delta/\sigma)_{\max} = 0.001$
211 parameters	$\Delta\rho_{\max} = 0.50 \text{ e } \text{\AA}^{-3}$
1 restraint	$\Delta\rho_{\min} = -0.62 \text{ e } \text{\AA}^{-3}$
Primary atom site location: structure-invariant direct methods	Extinction correction: none

Special details

Geometry. All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R-factor wR and goodness of fit S are based on F^2 , conventional R-factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > 2\sigma(F^2)$ is used only for calculating R-factors(gt) etc. and is not relevant to the choice of reflections for refinement. R-factors based on F^2 are statistically about twice as large as those based on F , and R- factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
Zn1	0.38637 (6)	0.75022 (5)	0.70365 (5)	0.04264 (14)
I1	0.26005 (4)	1.00680 (3)	0.70856 (3)	0.06259 (13)
I2	0.68572 (4)	0.65555 (4)	0.63039 (4)	0.07424 (14)
N1	0.2296 (4)	0.7001 (4)	0.5954 (3)	0.0390 (8)
N2	0.2722 (4)	0.8331 (3)	0.2827 (3)	0.0363 (8)
N3	0.0106 (4)	0.3085 (3)	0.8808 (3)	0.0435 (9)
O1	0.3662 (4)	0.6417 (3)	0.8421 (3)	0.0533 (8)
O2	-0.0766 (4)	0.3055 (3)	0.7938 (3)	0.0579 (9)
O3	0.0124 (4)	0.2483 (3)	0.9762 (3)	0.0587 (9)
C1	0.1913 (5)	0.5441 (4)	0.7469 (3)	0.0345 (9)
C2	0.2852 (5)	0.5627 (4)	0.8453 (4)	0.0355 (9)
C3	0.2900 (5)	0.4870 (4)	0.9535 (4)	0.0415 (10)
H3	0.3534	0.4954	1.0183	0.050*
C4	0.2054 (5)	0.4026 (4)	0.9657 (4)	0.0393 (10)
H4	0.2104	0.3549	1.0377	0.047*
C5	0.1107 (5)	0.3891 (4)	0.8682 (3)	0.0344 (9)

supplementary materials

C6	0.1057 (5)	0.4572 (4)	0.7614 (4)	0.0378 (9)
H6	0.0438	0.4454	0.6970	0.045*
C7	0.1648 (5)	0.6156 (4)	0.6312 (4)	0.0378 (10)
H7	0.0934	0.5985	0.5774	0.045*
C8	0.1724 (5)	0.7702 (5)	0.4791 (4)	0.0482 (11)
H8A	0.1146	0.8647	0.4939	0.058*
H8B	0.0975	0.7346	0.4405	0.058*
C9	0.3181 (5)	0.7512 (5)	0.3965 (4)	0.0459 (11)
H9A	0.3989	0.7758	0.4401	0.055*
H9B	0.3676	0.6575	0.3748	0.055*
C10	0.2279 (6)	0.9798 (4)	0.3042 (4)	0.0508 (12)
H10A	0.3208	0.9959	0.3385	0.061*
H10B	0.1380	1.0103	0.3626	0.061*
C11	0.1785 (6)	1.0575 (5)	0.1866 (4)	0.0569 (13)
H11A	0.1527	1.1519	0.2022	0.068*
H11B	0.0809	1.0462	0.1553	0.068*
C12	0.3150 (6)	1.0100 (5)	0.0919 (4)	0.0537 (12)
H12A	0.2793	1.0582	0.0159	0.064*
H12B	0.4100	1.0277	0.1200	0.064*
C13	0.3597 (7)	0.8626 (5)	0.0713 (4)	0.0609 (14)
H13A	0.4504	0.8317	0.0137	0.073*
H13B	0.2672	0.8464	0.0362	0.073*
C14	0.4070 (6)	0.7853 (5)	0.1887 (4)	0.0609 (14)
H14A	0.4314	0.6911	0.1730	0.073*
H14B	0.5052	0.7954	0.2201	0.073*
H2	0.183 (4)	0.825 (5)	0.250 (4)	0.080*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Zn1	0.0420 (3)	0.0466 (3)	0.0436 (3)	-0.0209 (2)	-0.0078 (2)	0.0040 (2)
I1	0.0622 (2)	0.0466 (2)	0.0738 (3)	-0.01273 (16)	-0.01062 (18)	-0.00726 (17)
I2	0.03816 (19)	0.0782 (3)	0.0974 (3)	-0.00982 (18)	-0.00365 (18)	-0.0088 (2)
N1	0.0397 (19)	0.047 (2)	0.0332 (19)	-0.0191 (17)	-0.0049 (16)	0.0070 (16)
N2	0.040 (2)	0.039 (2)	0.0326 (19)	-0.0174 (16)	-0.0002 (15)	0.0030 (15)
N3	0.042 (2)	0.037 (2)	0.048 (2)	-0.0101 (17)	-0.0009 (18)	0.0072 (17)
O1	0.065 (2)	0.070 (2)	0.0409 (17)	-0.0426 (19)	-0.0175 (15)	0.0099 (16)
O2	0.068 (2)	0.062 (2)	0.060 (2)	-0.0424 (19)	-0.0181 (18)	0.0107 (17)
O3	0.063 (2)	0.055 (2)	0.059 (2)	-0.0233 (17)	-0.0031 (17)	0.0232 (17)
C1	0.033 (2)	0.042 (2)	0.026 (2)	-0.0120 (18)	-0.0035 (17)	0.0023 (17)
C2	0.033 (2)	0.040 (2)	0.034 (2)	-0.0144 (19)	-0.0044 (17)	0.0017 (18)
C3	0.039 (2)	0.051 (3)	0.032 (2)	-0.013 (2)	-0.0081 (18)	0.0002 (19)
C4	0.038 (2)	0.040 (2)	0.032 (2)	-0.0050 (19)	0.0018 (18)	0.0087 (18)
C5	0.037 (2)	0.032 (2)	0.033 (2)	-0.0111 (18)	-0.0016 (18)	0.0033 (17)
C6	0.039 (2)	0.040 (2)	0.035 (2)	-0.0142 (19)	-0.0066 (18)	0.0030 (18)
C7	0.038 (2)	0.048 (3)	0.029 (2)	-0.017 (2)	-0.0046 (18)	0.0038 (19)
C8	0.048 (3)	0.064 (3)	0.036 (2)	-0.025 (2)	-0.006 (2)	0.012 (2)
C9	0.046 (3)	0.047 (3)	0.042 (3)	-0.013 (2)	0.000 (2)	0.012 (2)

C10	0.068 (3)	0.041 (3)	0.043 (3)	-0.020 (2)	0.006 (2)	0.003 (2)
C11	0.063 (3)	0.043 (3)	0.056 (3)	-0.008 (2)	0.011 (3)	0.012 (2)
C12	0.053 (3)	0.062 (3)	0.048 (3)	-0.024 (3)	0.003 (2)	0.014 (2)
C13	0.071 (3)	0.066 (4)	0.044 (3)	-0.022 (3)	0.014 (3)	0.005 (2)
C14	0.067 (3)	0.049 (3)	0.055 (3)	-0.007 (3)	0.025 (3)	0.004 (2)

Geometric parameters (Å, °)

Zn1—O1	1.948 (3)	C5—C6	1.367 (5)
Zn1—N1	2.037 (3)	C6—H6	0.9300
Zn1—I2	2.5666 (6)	C7—H7	0.9300
Zn1—I1	2.5690 (6)	C8—C9	1.508 (6)
N1—C7	1.284 (5)	C8—H8A	0.9700
N1—C8	1.473 (5)	C8—H8B	0.9700
N2—C9	1.492 (5)	C9—H9A	0.9700
N2—C10	1.496 (5)	C9—H9B	0.9700
N2—C14	1.503 (5)	C10—C11	1.509 (6)
N2—H2	0.89 (4)	C10—H10A	0.9700
N3—O3	1.220 (4)	C10—H10B	0.9700
N3—O2	1.235 (5)	C11—C12	1.518 (6)
N3—C5	1.439 (5)	C11—H11A	0.9700
O1—C2	1.285 (5)	C11—H11B	0.9700
C1—C6	1.396 (6)	C12—C13	1.502 (7)
C1—C2	1.420 (5)	C12—H12A	0.9700
C1—C7	1.452 (5)	C12—H12B	0.9700
C2—C3	1.423 (5)	C13—C14	1.503 (7)
C3—C4	1.363 (6)	C13—H13A	0.9700
C3—H3	0.9300	C13—H13B	0.9700
C4—C5	1.398 (6)	C14—H14A	0.9700
C4—H4	0.9300	C14—H14B	0.9700
O1—Zn1—N1	94.96 (13)	N1—C8—H8A	109.9
O1—Zn1—I2	107.27 (10)	C9—C8—H8A	109.9
N1—Zn1—I2	113.81 (10)	N1—C8—H8B	109.9
O1—Zn1—I1	121.56 (10)	C9—C8—H8B	109.9
N1—Zn1—I1	102.97 (10)	H8A—C8—H8B	108.3
I2—Zn1—I1	114.61 (2)	N2—C9—C8	112.2 (3)
C7—N1—C8	117.3 (4)	N2—C9—H9A	109.2
C7—N1—Zn1	121.3 (3)	C8—C9—H9A	109.2
C8—N1—Zn1	121.0 (3)	N2—C9—H9B	109.2
C9—N2—C10	113.3 (3)	C8—C9—H9B	109.2
C9—N2—C14	110.7 (3)	H9A—C9—H9B	107.9
C10—N2—C14	110.1 (3)	N2—C10—C11	110.8 (4)
C9—N2—H2	111 (3)	N2—C10—H10A	109.5
C10—N2—H2	105 (3)	C11—C10—H10A	109.5
C14—N2—H2	107 (3)	N2—C10—H10B	109.5
O3—N3—O2	122.8 (4)	C11—C10—H10B	109.5
O3—N3—C5	119.6 (4)	H10A—C10—H10B	108.1
O2—N3—C5	117.6 (3)	C10—C11—C12	110.9 (4)
C2—O1—Zn1	126.5 (3)	C10—C11—H11A	109.5

supplementary materials

C6—C1—C2	119.3 (4)	C12—C11—H11A	109.5
C6—C1—C7	114.7 (4)	C10—C11—H11B	109.5
C2—C1—C7	125.8 (4)	C12—C11—H11B	109.5
O1—C2—C1	124.3 (4)	H11A—C11—H11B	108.0
O1—C2—C3	118.5 (4)	C13—C12—C11	109.5 (4)
C1—C2—C3	117.2 (4)	C13—C12—H12A	109.8
C4—C3—C2	122.4 (4)	C11—C12—H12A	109.8
C4—C3—H3	118.8	C13—C12—H12B	109.8
C2—C3—H3	118.8	C11—C12—H12B	109.8
C3—C4—C5	119.0 (4)	H12A—C12—H12B	108.2
C3—C4—H4	120.5	C12—C13—C14	111.2 (4)
C5—C4—H4	120.5	C12—C13—H13A	109.4
C6—C5—C4	120.6 (4)	C14—C13—H13A	109.4
C6—C5—N3	118.8 (4)	C12—C13—H13B	109.4
C4—C5—N3	120.5 (4)	C14—C13—H13B	109.4
C5—C6—C1	121.4 (4)	H13A—C13—H13B	108.0
C5—C6—H6	119.3	N2—C14—C13	111.4 (4)
C1—C6—H6	119.3	N2—C14—H14A	109.3
N1—C7—C1	126.9 (4)	C13—C14—H14A	109.3
N1—C7—H7	116.5	N2—C14—H14B	109.3
C1—C7—H7	116.5	C13—C14—H14B	109.3
N1—C8—C9	109.1 (4)	H14A—C14—H14B	108.0

Hydrogen-bond geometry (\AA , $^\circ$)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
N2—H2 \cdots O2 ⁱ	0.89 (4)	2.00 (3)	2.777 (5)	144 (5)

Symmetry codes: (i) $-x, -y+1, -z+1$.

Fig. 1

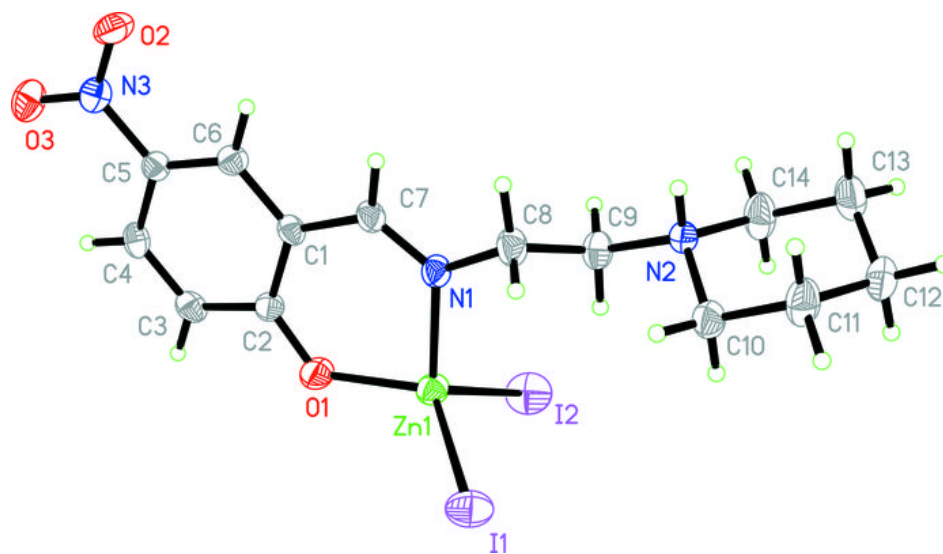


Fig. 2

