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4-Carbamoylpyridinium perchlorate

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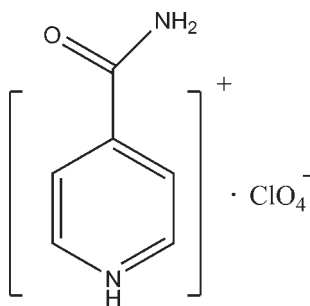
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Key indicators: single-crystal X-ray study; $T = 293$ K; mean $\sigma(\text{C}-\text{C}) = 0.004$ Å; R factor = 0.064; wR factor = 0.176; data-to-parameter ratio = 16.0.

In the cation of the title compound, $\text{C}_6\text{H}_7\text{N}_2\text{O}^+\cdot\text{ClO}_4^-$, the amide group is oriented at a dihedral angle of 10.41 (17)° to the benzene ring. The crystal structure is stabilized by intermolecular $\text{N}-\text{H}\cdots\text{O}$ hydrogen bonding.

Related literature

For general background to structural features and physical properties of simple molecular-ionic crystals containing organic cations and acid radicals (1:1 molar ratio), see: Czupirski *et al.* (2002); Katrusiak & Szafranski (1999, 2006). For the crystal structure of 4-carbamoylpyridinium dihydrogen phosphate, see: Gholivand *et al.* (2007) and for that of 3-(aminocarbonyl)pyridinium perchlorate, see: Athimoolam & Natarajan (2007).



Experimental

Crystal data

 $\text{C}_6\text{H}_7\text{N}_2\text{O}^+\cdot\text{ClO}_4^-$ $M_r = 222.59$

Monoclinic, $P2_1/c$
 $a = 10.935$ (2) Å
 $b = 10.082$ (2) Å
 $c = 8.2021$ (16) Å
 $\beta = 99.37$ (3)°
 $V = 892.2$ (3) Å³

$Z = 4$
Mo $K\alpha$ radiation
 $\mu = 0.43$ mm⁻¹
 $T = 293$ K
 $0.30 \times 0.25 \times 0.22$ mm

Data collection

Rigaku SCXmini diffractometer
Absorption correction: multi-scan
(*CrystalClear*; Rigaku, 2005)
 $T_{\min} = 0.87$, $T_{\max} = 0.90$

8863 measured reflections
2033 independent reflections
1421 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.074$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.064$
 $wR(F^2) = 0.176$
 $S = 1.04$
2033 reflections

127 parameters
H-atom parameters constrained
 $\Delta\rho_{\max} = 0.42$ e Å⁻³
 $\Delta\rho_{\min} = -0.42$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{N1}-\text{H1A}\cdots\text{O3}^{\text{i}}$	0.86	2.10	2.932 (4)	162
$\text{N2}-\text{H2A}\cdots\text{O1}^{\text{ii}}$	0.86	2.32	3.162 (4)	168
$\text{N2}-\text{H2B}\cdots\text{O5}^{\text{iii}}$	0.86	2.17	3.004 (4)	164

Symmetry codes: (i) $-x + 1, -y + 1, -z + 1$; (ii) $-x, -y + 1, -z$; (iii) $-x, y - \frac{1}{2}, -z + \frac{1}{2}$.

Data collection: *CrystalClear* (Rigaku, 2005); cell refinement: *CrystalClear*; data reduction: *CrystalClear*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEP-3 for Windows* (Farrugia, 1997); software used to prepare material for publication: *WinGX* (Farrugia, 1999).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: XU2591).

References

- Athimoolam, S. & Natarajan, S. (2007). *Acta Cryst.* **C63**, o263–o266.
Czupirski, O., Bator, G., Ciunik, Z., Jakubas, R., Medycki, W. & Świergiel, J. (2002). *J. Phys. Condens. Matter*, **14**, 8497–8512.
Farrugia, L. J. (1997). *J. Appl. Cryst.* **30**, 565.
Farrugia, L. J. (1999). *J. Appl. Cryst.* **32**, 837–838.
Gholivand, K., Zare, K., Afshar, F., Shariatinia, Z. & Khavasi, H. R. (2007). *Acta Cryst.* **E63**, o4027.
Katrusiak, A. & Szafranski, M. (1999). *Phys. Rev. Lett.* **82**, 576–579.
Katrusiak, A. & Szafranski, M. (2006). *J. Am. Chem. Soc.* **128**, 15775–15785.
Rigaku (2005). *CrystalClear*. Rigaku Corporation, Tokyo, Japan.
Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.

supplementary materials

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4-Carbamoylpyridinium perchlorate

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Comment

Recently much attention has been devoted to simple molecular–ionic crystals containing organic cations and acid radicals (1:1 molar ratio) due to the tunability of their special structural features and their interesting physical properties (Czupiński *et al.*, 2002; Katrusiak & Szafranski, 1999; Katrusiak & Szafranski, 2006). The crystal structures of 4-carbamoylpyridinium dihydrogen phosphate (Gholivand *et al.*, 2007) and 3-(aminocarbonyl)pyridinium perchlorate (Athimoolam & Natarajan, 2007) have been reported previously. In our laboratory, a compound containing 4-carbamoylpyridinium cation and ClO_4^- anion has been synthesized, its crystal structure is reported herein.

The asymmetric unit of the title compound (Fig. 1) consists of one 4-carbamoylpyridinium cation and one ClO_4^- anion. The crystal structure is stabilized by intermolecular $\text{N—H}\cdots\text{O}$ hydrogen bonds (Table 1).

Experimental

4-Carbamoylpyridine (2.44 g, 20 mmol) and 10% aqueous solution (15 ml) of HClO_4 were dissolved in 30 ml water. The solution was heated at 343 K for 0.5 h, forming a clear solution. The reaction mixture was cooled slowly to room temperature, block crystals of the title compound were formed.

Refinement

All H atoms were placed in calculated positions with $\text{C—H} = 0.93$ and $\text{N—H} = 0.86$ Å, and refined using a riding model with $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C,N})$.

Figures

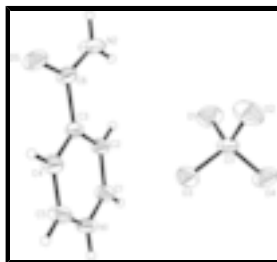


Fig. 1. The asymmetric unit of the title compound with atom labels. Displacement ellipsoids were drawn at the 30% probability level.

4-Carbamoylpyridinium perchlorate

Crystal data

$C_6H_7N_2O^+ \cdot ClO_4^-$

$M_r = 222.59$

Monoclinic, $P2_1/c$

Hall symbol: -P 2ybc

$a = 10.935 (2) \text{ \AA}$

$b = 10.082 (2) \text{ \AA}$

$c = 8.2021 (16) \text{ \AA}$

$\beta = 99.37 (3)^\circ$

$V = 892.2 (3) \text{ \AA}^3$

$Z = 4$

$F_{000} = 456$

$D_x = 1.657 \text{ Mg m}^{-3}$

Mo $K\alpha$ radiation, $\lambda = 0.71073 \text{ \AA}$

Cell parameters from 1421 reflections

$\theta = 3.2\text{--}27.5^\circ$

$\mu = 0.43 \text{ mm}^{-1}$

$T = 293 \text{ K}$

Block, colorless

$0.30 \times 0.25 \times 0.22 \text{ mm}$

Data collection

Rigaku SCXmini
diffractometer

Radiation source: fine-focus sealed tube

Monochromator: graphite

Detector resolution: $13.6612 \text{ pixels mm}^{-1}$

$T = 293 \text{ K}$

ω scans

Absorption correction: multi-scan
(CrystalClear; Rigaku, 2005)

$T_{\min} = 0.87$, $T_{\max} = 0.90$

8863 measured reflections

2033 independent reflections

1421 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.074$

$\theta_{\max} = 27.5^\circ$

$\theta_{\min} = 3.2^\circ$

$h = -14 \rightarrow 14$

$k = -13 \rightarrow 13$

$l = -10 \rightarrow 10$

Refinement

Refinement on F^2

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.064$

$wR(F^2) = 0.176$

$S = 1.04$

2033 reflections

127 parameters

Primary atom site location: structure-invariant direct
methods

Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring
sites

H-atom parameters constrained

$w = 1/[\sigma^2(F_o^2) + (0.0844P)^2 + 0.3072P]$

where $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{\max} < 0.001$

$\Delta\rho_{\max} = 0.42 \text{ e \AA}^{-3}$

$\Delta\rho_{\min} = -0.42 \text{ e \AA}^{-3}$

Extinction correction: none

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
C11	0.35049 (7)	0.39747 (7)	0.09900 (9)	0.0473 (3)
O5	0.0368 (2)	0.8048 (2)	0.2309 (3)	0.0650 (7)
N2	-0.0251 (3)	0.5945 (3)	0.1873 (4)	0.0698 (10)
H2A	-0.0862	0.6169	0.1127	0.084*
H2B	-0.0128	0.5123	0.2130	0.084*
C6	0.0493 (3)	0.6856 (3)	0.2619 (4)	0.0463 (7)
C4	0.2469 (3)	0.7332 (3)	0.4470 (4)	0.0479 (8)
H4A	0.2386	0.8205	0.4104	0.057*
N1	0.3573 (2)	0.5717 (3)	0.6122 (3)	0.0541 (8)
H1A	0.4218	0.5492	0.6813	0.065*
C3	0.1569 (2)	0.6406 (3)	0.3883 (3)	0.0399 (7)
O4	0.4364 (2)	0.2906 (2)	0.1204 (3)	0.0707 (8)
O3	0.3944 (2)	0.4992 (2)	0.2176 (3)	0.0653 (7)
C2	0.1692 (3)	0.5117 (3)	0.4485 (4)	0.0457 (8)
H2C	0.1096	0.4479	0.4117	0.055*
C5	0.3479 (3)	0.6961 (4)	0.5590 (4)	0.0549 (9)
H5A	0.4094	0.7575	0.5973	0.066*
O2	0.3407 (3)	0.4505 (3)	-0.0636 (3)	0.0843 (9)
O1	0.2315 (2)	0.3517 (3)	0.1231 (4)	0.0759 (8)
C1	0.2713 (3)	0.4798 (3)	0.5636 (4)	0.0536 (9)
H1B	0.2803	0.3945	0.6072	0.064*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
C11	0.0405 (5)	0.0430 (5)	0.0519 (5)	0.0012 (3)	-0.0116 (3)	0.0000 (3)
O5	0.0581 (15)	0.0403 (14)	0.0867 (19)	0.0094 (11)	-0.0176 (12)	0.0095 (12)
N2	0.0575 (19)	0.0488 (18)	0.087 (2)	-0.0020 (13)	-0.0377 (17)	0.0116 (15)
C6	0.0397 (17)	0.0405 (18)	0.0548 (19)	0.0032 (13)	-0.0039 (13)	0.0050 (14)
C4	0.0466 (18)	0.0361 (16)	0.057 (2)	-0.0056 (13)	-0.0036 (14)	0.0033 (13)
N1	0.0460 (16)	0.0565 (18)	0.0509 (17)	0.0061 (12)	-0.0187 (13)	-0.0010 (12)
C3	0.0346 (16)	0.0371 (15)	0.0448 (17)	0.0021 (11)	-0.0030 (12)	-0.0012 (12)

supplementary materials

O4	0.0550 (15)	0.0510 (15)	0.097 (2)	0.0147 (11)	-0.0145 (14)	-0.0040 (13)
O3	0.0595 (16)	0.0570 (15)	0.0721 (17)	-0.0015 (11)	-0.0110 (12)	-0.0208 (12)
C2	0.0469 (18)	0.0319 (15)	0.0520 (19)	-0.0018 (12)	-0.0109 (14)	-0.0034 (12)
C5	0.0438 (19)	0.055 (2)	0.060 (2)	-0.0120 (15)	-0.0079 (15)	-0.0007 (16)
O2	0.106 (2)	0.087 (2)	0.0518 (17)	-0.0028 (17)	-0.0107 (15)	0.0126 (14)
O1	0.0434 (15)	0.0839 (19)	0.096 (2)	-0.0120 (13)	-0.0021 (13)	-0.0088 (16)
C1	0.060 (2)	0.0381 (17)	0.055 (2)	0.0063 (15)	-0.0145 (16)	-0.0002 (14)

Geometric parameters (Å, °)

C11—O4	1.421 (2)	C4—C3	1.385 (4)
C11—O2	1.424 (3)	C4—H4A	0.9300
C11—O1	1.425 (3)	N1—C5	1.327 (4)
C11—O3	1.441 (2)	N1—C1	1.334 (4)
O5—C6	1.231 (3)	N1—H1A	0.8600
N2—C6	1.310 (4)	C3—C2	1.389 (4)
N2—H2A	0.8600	C2—C1	1.377 (4)
N2—H2B	0.8600	C2—H2C	0.9300
C6—C3	1.507 (4)	C5—H5A	0.9300
C4—C5	1.369 (5)	C1—H1B	0.9300
O4—C11—O2	110.33 (19)	C5—N1—C1	123.0 (3)
O4—C11—O1	109.76 (17)	C5—N1—H1A	118.5
O2—C11—O1	108.68 (18)	C1—N1—H1A	118.5
O4—C11—O3	108.36 (15)	C4—C3—C2	119.0 (3)
O2—C11—O3	109.30 (18)	C4—C3—C6	117.9 (3)
O1—C11—O3	110.41 (17)	C2—C3—C6	123.1 (3)
C6—N2—H2A	120.0	C1—C2—C3	118.9 (3)
C6—N2—H2B	120.0	C1—C2—H2C	120.6
H2A—N2—H2B	120.0	C3—C2—H2C	120.6
O5—C6—N2	123.2 (3)	N1—C5—C4	119.3 (3)
O5—C6—C3	118.9 (3)	N1—C5—H5A	120.3
N2—C6—C3	117.8 (3)	C4—C5—H5A	120.3
C5—C4—C3	120.0 (3)	N1—C1—C2	119.7 (3)
C5—C4—H4A	120.0	N1—C1—H1B	120.1
C3—C4—H4A	120.0	C2—C1—H1B	120.1
C5—C4—C3—C2	1.9 (5)	C4—C3—C2—C1	-0.6 (5)
C5—C4—C3—C6	-178.4 (3)	C6—C3—C2—C1	179.7 (3)
O5—C6—C3—C4	-9.0 (5)	C1—N1—C5—C4	-1.0 (5)
N2—C6—C3—C4	169.5 (3)	C3—C4—C5—N1	-1.2 (5)
O5—C6—C3—C2	170.7 (3)	C5—N1—C1—C2	2.4 (5)
N2—C6—C3—C2	-10.9 (5)	C3—C2—C1—N1	-1.5 (5)

Hydrogen-bond geometry (Å, °)

<i>D</i> —H \cdots <i>A</i>	<i>D</i> —H	H \cdots <i>A</i>	<i>D</i> \cdots <i>A</i>	<i>D</i> —H \cdots <i>A</i>
N1—H1A \cdots O3 ⁱ	0.86	2.10	2.932 (4)	162
N2—H2A \cdots O1 ⁱⁱ	0.86	2.32	3.162 (4)	168
N2—H2B \cdots O5 ⁱⁱⁱ	0.86	2.17	3.004 (4)	164

Symmetry codes: (i) $-x+1, -y+1, -z+1$; (ii) $-x, -y+1, -z$; (iii) $-x, y-1/2, -z+1/2$.

Fig. 1

