

Acta Crystallographica Section E

## Structure Reports

Online

ISSN 1600-5368

4,4'-Bipyridinium bis(oxalato- $\kappa^2O^1, O^2$ )-cuprate(II): an ion-pair complexLai-Jun Zhang,<sup>a,b\*</sup> Xing-Can Shen<sup>c</sup> and Hong Liang<sup>c</sup>

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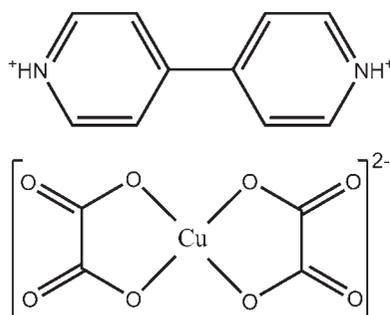
Received 14 September 2009; accepted 26 September 2009

Key indicators: single-crystal X-ray study;  $T = 293$  K; mean  $\sigma(C-C) = 0.004$  Å;  $R$  factor = 0.027;  $wR$  factor = 0.073; data-to-parameter ratio = 10.2.

The title compound,  $(C_{10}H_{10}N_2)[Cu(C_2O_4)_2]$  or  $(4,4'-H_2bpy)^+ [Cu(ox)_2]^{2-}$  (bpy is 4,4'-bipyridine and ox is oxalate), is an ion-pair complex comprising a protonated 4,4'-bipyridinium dication and a square-planar dioxalato-copper(II) dianion. In the centrosymmetric dianion, the  $Cu^{II}$  centre is coordinated by four O atoms from the two discrete oxalate ligands [ $Cu-O = 1.9245$  (19) and  $1.9252$  (17) Å], while the planar dications are also centrosymmetric. Inter-species  $N-H \cdots O$  hydrogen bonds link the cations and anions into one-dimensional chains and, together with weak intra-ion  $C-H \cdots O$  interactions, give a two-dimensional sheet structure.

## Related literature

For related background, see: Ren *et al.* (2007). For related structures, see, for example: Bukowska-Strzyzewska & Tosik (1979); Crawford *et al.* (2004); Diallo *et al.* (2008); Dou *et al.* (2007); Madhu & Das (2004); Näther *et al.* (2001); Tosik *et al.* (1990); Willett *et al.* (2006).



## Experimental

## Crystal data

$(C_{10}H_{10}N_2)[Cu(C_2O_4)_2]$   
 $M_r = 397.79$   
 Triclinic,  $P\bar{1}$   
 $a = 3.6900$  (7) Å  
 $b = 9.950$  (2) Å  
 $c = 10.230$  (2) Å  
 $\alpha = 113.77$  (3)°  
 $\beta = 98.43$  (3)°

$\gamma = 97.89$  (3)°  
 $V = 331.93$  (15) Å<sup>3</sup>  
 $Z = 1$   
 Mo  $K\alpha$  radiation  
 $\mu = 1.70$  mm<sup>-1</sup>  
 $T = 293$  K  
 $0.41 \times 0.27 \times 0.22$  mm

## Data collection

Bruker APEXII CCD area-detector diffractometer  
 Absorption correction: multi-scan (SADABS; Bruker, 2005)  
 $T_{min} = 0.527$ ,  $T_{max} = 0.705$

1761 measured reflections  
 1168 independent reflections  
 1126 reflections with  $I > 2\sigma(I)$   
 $R_{int} = 0.013$

## Refinement

$R[F^2 > 2\sigma(F^2)] = 0.027$   
 $wR(F^2) = 0.073$   
 $S = 1.09$   
 1168 reflections

115 parameters  
 H-atom parameters constrained  
 $\Delta\rho_{max} = 0.30$  e Å<sup>-3</sup>  
 $\Delta\rho_{min} = -0.29$  e Å<sup>-3</sup>

Table 1

Hydrogen-bond geometry (Å, °).

$D-H \cdots A$	$D-H$	$H \cdots A$	$D \cdots A$	$D-H \cdots A$
$N1-H1A \cdots O4^i$	0.86	2.02	2.755 (3)	143
$N1-H1A \cdots O1^i$	0.86	2.21	2.880 (3)	135
$C1-H1 \cdots O4^{ii}$	0.93	2.49	3.381 (3)	160
$C2-H2 \cdots O3^{ii}$	0.93	2.57	3.195 (3)	125
$C4-H4 \cdots O2$	0.93	2.42	3.272 (3)	153
$C5-H5 \cdots O1$	0.93	2.46	3.215 (3)	138

Symmetry codes: (i)  $-x + 2, -y + 1, -z + 1$ ; (ii)  $x + 1, y, z - 1$ .

Data collection: APEX2 (Bruker, 2005); cell refinement: SAINT-Plus (Bruker, 2005); data reduction: SAINT-Plus; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: SHELXTL (Sheldrick, 2008); software used to prepare material for publication: SHELXTL.

We are grateful for the support of the National Natural Science Foundation of China (No. 20701010), the Natural Science Foundation of Guangxi Province (No. 0728094) and the Department of Education of Jiangxi Province [grant No. GanJiaoJiZi (2007)348].

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: ZS2010).

## References

- Bruker (2005). APEX2, SAINT-Plus and SADABS. Bruker AXS Inc., Madison, Wisconsin, USA.  
 Bukowska-Strzyzewska, M. & Tosik, A. (1979). *Pol. J. Chem.* **53**, 2423–2428.  
 Crawford, P. C., Gillon, A. L., Green, J., Orpen, A. G., Podesta, T. J. & Pritchard, S. V. (2004). *CrystEngComm*, **6**, 419–428.  
 Diallo, M., Tamboura, F. B., Gaye, M., Barry, A. H. & Bah, Y. (2008). *Acta Cryst.* **E64**, m1124–m1125.  
 Dou, Q.-Q., He, Y.-K., Zhang, L.-T. & Han, Z.-B. (2007). *Acta Cryst.* **E63**, m2908–m2909.  
 Madhu, V. & Das, S. K. (2004). *Polyhedron*, **23**, 1235–1242.  
 Näther, C., Jess, I. & Bolte, M. (2001). *Acta Cryst.* **E57**, m78–m79.

Ren, H., Xu, J.-N., Yu, L.-X., Ye, J.-W., Bi, M.-H., Zhang, P. & Song, T.-Y. (2007). *Chem. J. Chin. Univ.* **28**, 1014–1017.  
Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.

Tosik, A. M., Bukowska-Strzyewska, M. & Mrozinski, J. (1990). *J. Coord. Chem.* **21**, 253–261.  
Willett, R. D., Butcher, R. E., Landee, C. P. & Twamley, B. (2006). *Polyhedron*, **25**, 2093–2100.

## supporting information

*Acta Cryst.* (2009). E65, m1276–m1277 [https://doi.org/10.1107/S1600536809039117]

**4,4'-Bipyridinium bis(oxalato- $\kappa^2 O^1, O^2$ )cuprate(II): an ion-pair complex****Lai-Jun Zhang, Xing-Can Shen and Hong Liang****S1. Comment**

Currently, *in situ* chemical reactions are of considerable interest, providing a powerful route for the preparation of novel crystals with unexpected structures, e.g. Ren *et al.* (2007) has reported the hydrothermal synthesis of a new coordination polymer  $[\text{Cu}(\text{ox})(4,4'\text{-bpy})]_n$  (bpy = 4,4'-bipyridine; ox = oxalate<sup>2-</sup>) resulting from the decomposition of furan-2-carboxylic acid to give the ox<sup>2-</sup> ligand in the reaction of this acid with  $\text{Cu}(\text{NO}_3)_2 \cdot 3\text{H}_2\text{O}$  and 4,4'-bpy. In the work reported here, a novel ion-pair complex,  $(4,4'\text{-H}_2\text{bpy})[\text{Cu}(\text{ox})_2]$  (I) was obtained by using sodium 2-hydroxyphosphonocarboxylate as the ox<sup>2-</sup> source and  $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$  as the copper source. Copper(II) is a relatively strong oxidant with the ability to oxidise the sodium 2-hydroxyphosphonocarboxylic giving ox<sup>2-</sup> which is generated *in situ*, resulting in formation of the title compound, the structure of which is reported here.

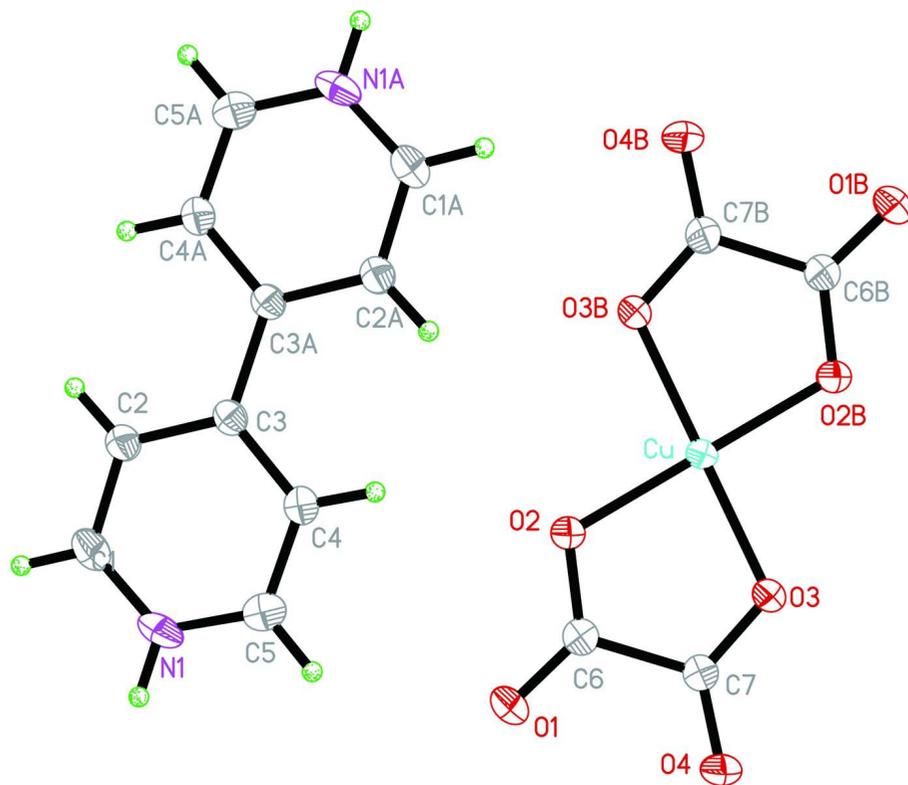
The structure of the (I) shows an ion-pair complex comprising a protonated 4,4'-bipyridine dication,  $(4,4'\text{-H}_2\text{bpy})^{2+}$  (Bukowska-Strzyzewska *et al.*, 1979; Crawford *et al.*, 2004; Dou *et al.*, 2007; Madhu *et al.*, 2004; Näther *et al.*, 2001; Tosik *et al.*, 1990; Willett *et al.*, 2006), and a  $[\text{Cu}(\text{C}_2\text{O}_4)_2]^{2-}$  anion (Fig. 1). The discrete bis-(oxalato)copper(II) dianions have square planar  $\text{CuO}_4$  stereochemistry, comprising four O donor atoms from two oxalate ligands [Cu–O, 1.9245 (19), 1.9252 (17) Å] and lie across inversion centres in the unit cell. In the axial sites, the Cu–O<sub>oxalate</sub> contacts are 2.920 (2) Å. With the 4,4'-bipyridine dications, the pyridine rings are coplanar and also have crystallographic inversion symmetry. Interamolecular N—H $\cdots$ O hydrogen bonds (Table 1) link the cations and anions into one-dimensional chains and together with weak intramolecular C—H $\cdots$ O interactions give two-dimensional sheet structures which layer down the *a* direction in the unit cell (Fig. 2).

**S2. Experimental**

Sodium 2-hydroxyphosphonocarboxylate (1 mmol) was dissolved in 10 ml of 50% ethanol-water after which 1 mmol of copper(II) chloride dihydrate and 1 mmol of 4,4'-bipyridine were added in sequence. After stirring for 10 min, the resulting mixture was transferred into a Teflon-lined stainless steel vessel (15 ml) which was sealed and heated at 110 °C for four days, then allowed to cool to room temperature. The blue-green single crystal blocks of the title compound (I), together with yellow-red prismatic crystals of a second unidentified component were obtained. The yield of (I) was *ca.* 25% (based on copper). IR (cm<sup>-1</sup>, KBr): 3439, 3110, 3060, 2921, 1665, 1581, 1487, 1409, 1273, 1205 1106, 1000, 892, 827, 798.

**S3. Refinement**

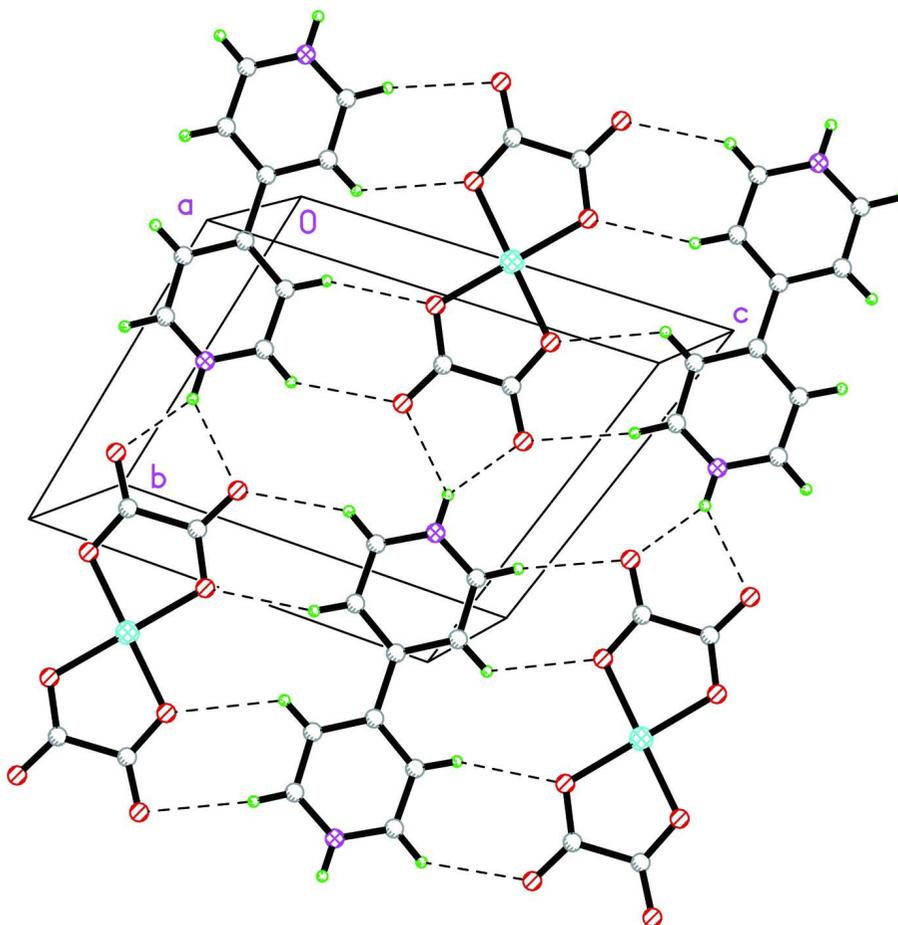
All H atoms were geometrically placed and refined using a riding model approximation, with C—H = 0.93 Å or N—H = 0.86 Å and  $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C}, \text{N})$ .



**Figure 1**

Molecular configuration and atom numbering scheme for the discrete dication and dianion species in (I). Displacement ellipsoids are drawn at the 50% probability level. Both cation and anion have crystallographic inversion symmetry.

Symmetry codes: (A)  $-x + 1, -y, -z$ ; (B)  $-x, -y, -z + 1$ .



**Figure 2**

The two-dimensional hydrogen-bonded sheet structure of (I) showing intra- and intermolecular hydrogen bonds as dashed lines.

#### 4,4'-Bipyridinium bis(oxalato- $\kappa^2O^1, O^2$ )cuprate(II)

##### Crystal data

(C<sub>10</sub>H<sub>10</sub>N<sub>2</sub>)[Cu(C<sub>2</sub>O<sub>4</sub>)<sub>2</sub>]

$M_r = 397.79$

Triclinic,  $P\bar{1}$

Hall symbol: -P 1

$a = 3.6900$  (7) Å

$b = 9.950$  (2) Å

$c = 10.230$  (2) Å

$\alpha = 113.77$  (3)°

$\beta = 98.43$  (3)°

$\gamma = 97.89$  (3)°

$V = 331.93$  (15) Å<sup>3</sup>

$Z = 1$

$F(000) = 201$

$D_x = 1.990$  Mg m<sup>-3</sup>

Mo  $K\alpha$  radiation,  $\lambda = 0.71073$  Å

Cell parameters from 1168 reflections

$\theta = 2.2$ – $25.1$ °

$\mu = 1.70$  mm<sup>-1</sup>

$T = 293$  K

Block, green-blue

$0.41 \times 0.27 \times 0.22$  mm

##### Data collection

Bruker APEXII CCD area-detector  
diffractometer

Radiation source: fine-focus sealed tube  
Graphite monochromator

$\varphi$  and  $\omega$  scans

Absorption correction: multi-scan  
(*SADABS*; Bruker, 2005)

$T_{\min} = 0.527$ ,  $T_{\max} = 0.705$

1761 measured reflections  
 1168 independent reflections  
 1126 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.013$

$\theta_{\text{max}} = 25.1^\circ$ ,  $\theta_{\text{min}} = 2.2^\circ$   
 $h = -4 \rightarrow 4$   
 $k = -9 \rightarrow 11$   
 $l = -11 \rightarrow 12$

*Refinement*

Refinement on  $F^2$   
 Least-squares matrix: full  
 $R[F^2 > 2\sigma(F^2)] = 0.027$   
 $wR(F^2) = 0.073$   
 $S = 1.09$   
 1168 reflections  
 115 parameters  
 0 restraints  
 Primary atom site location: structure-invariant  
 direct methods

Secondary atom site location: difference Fourier  
 map  
 Hydrogen site location: inferred from  
 neighbouring sites  
 H-atom parameters constrained  
 $w = 1/[\sigma^2(F_o^2) + (0.0434P)^2 + 0.1814P]$   
 where  $P = (F_o^2 + 2F_c^2)/3$   
 $(\Delta/\sigma)_{\text{max}} < 0.001$   
 $\Delta\rho_{\text{max}} = 0.30 \text{ e } \text{\AA}^{-3}$   
 $\Delta\rho_{\text{min}} = -0.29 \text{ e } \text{\AA}^{-3}$

*Special details*

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
Cu1	0.0000	0.0000	0.5000	0.02273 (16)
C1	0.9693 (6)	0.2924 (3)	0.0264 (3)	0.0263 (5)
H1	1.0658	0.3392	-0.0273	0.032*
C2	0.7599 (6)	0.1481 (3)	-0.0443 (3)	0.0228 (5)
H2	0.7140	0.0971	-0.1459	0.027*
C3	0.6154 (6)	0.0779 (2)	0.0367 (2)	0.0182 (4)
C4	0.6955 (7)	0.1599 (3)	0.1887 (3)	0.0265 (5)
H4	0.6054	0.1165	0.2461	0.032*
C5	0.9057 (7)	0.3037 (3)	0.2534 (3)	0.0312 (6)
H5	0.9580	0.3582	0.3547	0.037*
C6	0.4733 (6)	0.2586 (3)	0.5616 (3)	0.0232 (5)
C7	0.3414 (6)	0.2770 (3)	0.7045 (2)	0.0224 (5)
N1	1.0353 (5)	0.3657 (2)	0.1713 (2)	0.0273 (4)
H1A	1.1664	0.4564	0.2135	0.033*
O1	0.6801 (5)	0.3636 (2)	0.5609 (2)	0.0342 (4)
O2	0.3514 (5)	0.12699 (19)	0.45407 (17)	0.0286 (4)
O3	0.1187 (5)	0.16185 (19)	0.69418 (18)	0.0265 (4)
O4	0.4569 (5)	0.39568 (19)	0.81441 (18)	0.0313 (4)

Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
Cu1	0.0285 (2)	0.0171 (2)	0.0166 (2)	-0.00496 (16)	0.00626 (16)	0.00411 (17)
C1	0.0250 (12)	0.0251 (12)	0.0318 (13)	0.0015 (10)	0.0060 (10)	0.0167 (11)
C2	0.0249 (11)	0.0219 (12)	0.0211 (11)	0.0009 (9)	0.0046 (9)	0.0102 (10)
C3	0.0155 (10)	0.0184 (11)	0.0216 (11)	0.0041 (9)	0.0043 (8)	0.0095 (9)
C4	0.0319 (13)	0.0245 (12)	0.0218 (12)	-0.0003 (10)	0.0082 (10)	0.0101 (10)
C5	0.0360 (14)	0.0262 (13)	0.0237 (13)	-0.0002 (11)	0.0047 (10)	0.0059 (11)
C6	0.0244 (11)	0.0224 (12)	0.0214 (12)	0.0018 (9)	0.0051 (9)	0.0092 (10)
C7	0.0235 (11)	0.0221 (12)	0.0208 (12)	0.0022 (9)	0.0046 (9)	0.0096 (10)
N1	0.0242 (10)	0.0160 (10)	0.0371 (12)	-0.0022 (8)	0.0040 (9)	0.0099 (9)
O1	0.0425 (10)	0.0237 (9)	0.0309 (10)	-0.0083 (8)	0.0111 (8)	0.0101 (8)
O2	0.0381 (9)	0.0205 (8)	0.0196 (8)	-0.0057 (7)	0.0102 (7)	0.0037 (7)
O3	0.0332 (9)	0.0208 (8)	0.0198 (8)	-0.0059 (7)	0.0079 (7)	0.0062 (7)
O4	0.0389 (10)	0.0203 (9)	0.0227 (9)	-0.0070 (7)	0.0063 (7)	0.0018 (7)

Geometric parameters ( $\text{\AA}$ ,  $^\circ$ )

Cu1—O3	1.9245 (19)	C4—C5	1.367 (4)
Cu1—O3 <sup>i</sup>	1.9245 (19)	C4—H4	0.9300
Cu1—O2 <sup>i</sup>	1.9252 (17)	C5—N1	1.331 (3)
Cu1—O2	1.9252 (17)	C5—H5	0.9300
C1—N1	1.327 (3)	C6—O1	1.210 (3)
C1—C2	1.369 (3)	C6—O2	1.288 (3)
C1—H1	0.9300	C6—C7	1.557 (3)
C2—C3	1.397 (3)	C7—O4	1.221 (3)
C2—H2	0.9300	C7—O3	1.272 (3)
C3—C4	1.395 (3)	N1—H1A	0.8600
C3—C3 <sup>ii</sup>	1.485 (4)		
O3—Cu1—O3 <sup>i</sup>	180.0	C5—C4—H4	119.9
O3—Cu1—O2 <sup>i</sup>	93.91 (7)	C3—C4—H4	119.9
O3 <sup>i</sup> —Cu1—O2 <sup>i</sup>	86.09 (7)	N1—C5—C4	119.9 (2)
O3—Cu1—O2	86.09 (7)	N1—C5—H5	120.1
O3 <sup>i</sup> —Cu1—O2	93.91 (7)	C4—C5—H5	120.1
O2 <sup>i</sup> —Cu1—O2	180.00 (8)	O1—C6—O2	126.2 (2)
N1—C1—C2	120.4 (2)	O1—C6—C7	119.2 (2)
N1—C1—H1	119.8	O2—C6—C7	114.60 (19)
C2—C1—H1	119.8	O4—C7—O3	125.9 (2)
C1—C2—C3	119.7 (2)	O4—C7—C6	119.3 (2)
C1—C2—H2	120.1	O3—C7—C6	114.8 (2)
C3—C2—H2	120.1	C1—N1—C5	122.3 (2)
C4—C3—C2	117.5 (2)	C1—N1—H1A	118.9
C4—C3—C3 <sup>ii</sup>	121.4 (2)	C5—N1—H1A	118.9

C2—C3—C3 <sup>ii</sup>	121.1 (3)	C6—O2—Cu1	111.74 (14)
C5—C4—C3	120.3 (2)	C7—O3—Cu1	112.38 (14)

Symmetry codes: (i)  $-x, -y, -z+1$ ; (ii)  $-x+1, -y, -z$ .

*Hydrogen-bond geometry (Å, °)*

<i>D—H...A</i>	<i>D—H</i>	<i>H...A</i>	<i>D...A</i>	<i>D—H...A</i>
N1—H1A...O4 <sup>iii</sup>	0.86	2.02	2.755 (3)	143
N1—H1A...O1 <sup>iii</sup>	0.86	2.21	2.880 (3)	135
C1—H1...O4 <sup>iv</sup>	0.93	2.49	3.381 (3)	160
C2—H2...O3 <sup>iv</sup>	0.93	2.57	3.195 (3)	125
C4—H4...O2	0.93	2.42	3.272 (3)	153
C5—H5...O1	0.93	2.46	3.215 (3)	138

Symmetry codes: (iii)  $-x+2, -y+1, -z+1$ ; (iv)  $x+1, y, z-1$ .