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## Structure Reports

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 Poly[[ $\mu_2$ -acetato-aquadi- $\mu_3$ -isonicotinato-holmium(III)silver(I)] perchlorate]

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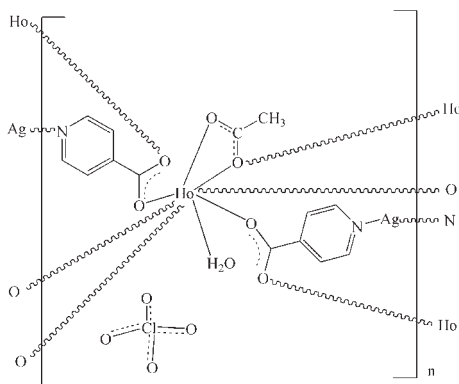
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 Key indicators: single-crystal X-ray study;  $T = 296$  K; mean  $\sigma(\text{C}-\text{C}) = 0.007$  Å; disorder in main residue;  $R$  factor = 0.023;  $wR$  factor = 0.054; data-to-parameter ratio = 10.7.

In the title three-dimensional heterometallic complex,  $\{[\text{AgHo}(\text{C}_6\text{H}_4\text{NO}_2)_2(\text{C}_2\text{H}_3\text{O}_2)(\text{H}_2\text{O})]\text{ClO}_4\}_n$ , the  $\text{Ho}^{\text{III}}$  ion is eight-coordinated by four O atoms from four different isonicotinate ligands, three O atoms from two different acetate ligands and one O atom of a water molecule. The two-coordinate  $\text{Ag}^{\text{I}}$  ion is bonded to two N atoms from two different isonicotinate anions. These metal coordination units are connected by bridging isonicotinate and acetate ligands, generating a three-dimensional network. The coordinated water molecules link the carboxylate group of the acetate ligand and the nitrate ligand by  $\text{O}-\text{H}\cdots\text{O}$  hydrogen bonding. The crystal structure is further stabilized by hydrogen bonds. The perchlorate ion is disordered over two sites with site-occupancy factors 0.539 (12) and 0.461 (12), while the methyl group of the acetate ligand is disordered over two sites with site-occupancy factors 0.51 (4) and 0.49 (4).

## Related literature

For the applications of lanthanide-transition metal heterometallic complexes in ion exchange, magnetism, bimetallic catalysis and as luminescent probes, see: Cheng *et al.* (2006); Kuang *et al.* (2007); Peng *et al.* (2008); Zhu *et al.* (2009).



## Experimental

## Crystal data

$[\text{AgHo}(\text{C}_6\text{H}_4\text{NO}_2)_2(\text{C}_2\text{H}_3\text{O}_2)(\text{H}_2\text{O})]\text{ClO}_4$   
 $M_r = 693.51$   
 Monoclinic,  $P2_1/c$   
 $a = 16.2158$  (10) Å  
 $b = 14.9024$  (9) Å  
 $c = 7.9068$  (5) Å

$\beta = 91.826$  (1)°  
 $V = 1909.7$  (2) Å<sup>3</sup>  
 $Z = 4$   
 Mo  $K\alpha$  radiation  
 $\mu = 5.34$  mm<sup>-1</sup>  
 $T = 296$  K  
 $0.20 \times 0.18 \times 0.15$  mm

## Data collection

Bruker APEXII area-detector diffractometer  
 Absorption correction: multi-scan (SADABS; Sheldrick, 1996)  
 $T_{\text{min}} = 0.357$ ,  $T_{\text{max}} = 0.449$

9735 measured reflections  
 3439 independent reflections  
 3085 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.027$

## Refinement

$R[F^2 > 2\sigma(F^2)] = 0.023$   
 $wR(F^2) = 0.054$   
 $S = 1.04$   
 3439 reflections  
 320 parameters  
 158 restraints

H atoms treated by a mixture of independent and constrained refinement  
 $\Delta\rho_{\text{max}} = 0.64$  e Å<sup>-3</sup>  
 $\Delta\rho_{\text{min}} = -0.71$  e Å<sup>-3</sup>

Table 1

Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{O1W}-\text{H1W}\cdots\text{O4}^{\text{i}}$	0.82 (4)	2.19 (3)	2.898 (4)	145 (4)
$\text{O1W}-\text{H2W}\cdots\text{O6}^{\text{ii}}$	0.81 (4)	1.99 (4)	2.787 (4)	168 (5)

Symmetry codes: (i)  $x, -y + \frac{1}{2}, z + \frac{1}{2}$ ; (ii)  $x, -y + \frac{1}{2}, z - \frac{1}{2}$ .

Data collection: APEX2 (Bruker, 2004); cell refinement: SAINT (Bruker, 2004); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: XP in SHELXTL (Sheldrick, 2008); software used to prepare material for publication: SHELXL97.

The author acknowledges South China Normal University for supporting this work.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: PV2224).

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**supplementary materials**

*Acta Cryst.* (2009). E65, m1619 [ doi:10.1107/S1600536809046601 ]

## Poly[[ $\mu_2$ -acetato-aquadi- $\mu_3$ -isonicotinato-holmium(III)silver(I)] perchlorate]

S. Feng

### Comment

In the past few years, lanthanide-transition metal heterometallic complexes with bridging multifunctional organic ligands are of increasing interest, not only because of their impressive topological structures, but also due to their versatile applications in ion exchange, magnetism, bimetallic catalysis and luminescent probe (Cheng *et al.*, 2006; Kuang *et al.*, 2007; Peng *et al.*, 2008; Zhu *et al.*, 2009). As an extension of this research, the structure of the title compound, a new heterometallic coordination polymer, (I), has been determined which is presented in this article.

The asymmetric unit of the title compound (Fig. 1), contains one of each of the  $\text{Ho}^{\text{III}}$  and  $\text{Ag}^{\text{I}}$  ions, two halves of acetate ligand, two isonicotinate ligands, and one coordinated water molecule. The  $\text{Ho}^{\text{III}}$  ion is eight-coordinated by four O atoms from four different isonicotinate ligands, and three O atoms from two different acetate ligands, and one O atom of water molecule; the Ho center can be described as adopting a bicapped trigonal prism coordination geometry. The two-coordinate  $\text{Ag}^{\text{I}}$  ion is bonded to two N atoms from two different isonicotinate anions. Thus the  $\text{Ag}^{\text{I}}$  ion is in a somewhat linear conformation with  $\text{N1—Ag1—N2}$  angle  $165.55(17)^\circ$ . These metal coordination units are connected by bridging isonicotinate and acetate ligands, generating a three-dimensional network (Fig. 2). The coordinated water molecules link the carboxylate group and acetate ligand by  $\text{O—H}\cdots\text{O}$  hydrogen bonding (Table 1). The crystal structure is further stabilized by hydrogen bonds.

### Experimental

A mixture of  $\text{AgNO}_3$  (0.057 g, 0.33 mmol),  $\text{Ho}_2\text{O}_3$  (0.116 g, 0.33 mmol), isonicotinic acid (0.164 g, 1.33 mmol),  $\text{CH}_3\text{COONa}$  (0.057 g, 0.7 mmol),  $\text{H}_2\text{O}$  (7 ml), and  $\text{HClO}_4$  (0.257 mmol) (pH 2) was sealed in a 20 ml Teflon-lined reaction vessel at 443 K for 6 days then slowly cooled to room temperature. The product was collected by filtration, washed with water and air-dried. Colorless block crystals suitable for X-ray analysis were obtained.

### Refinement

H atoms bonded to C atoms were positioned geometrically and refined as riding, with  $\text{C—H} = 0.93$  or  $0.96 \text{ \AA}$  and  $U_{\text{iso}}(\text{H}) = 1.2$  or  $1.5 U_{\text{eq}}(\text{C})$ . H atoms of water molecules were found from difference Fourier maps and refined isotropically with a restraint of  $\text{O—H} = 0.81 - 0.82 \text{ \AA}$ . The perchlorate ion was disordered over two sites with site occupancy factors 0.539 (12) and 0.461 (12), whereas the methyl group of the acetate ligand was disordered over two sites with site occupancy factors 0.51 (4) and 0.49 (4).

Figures

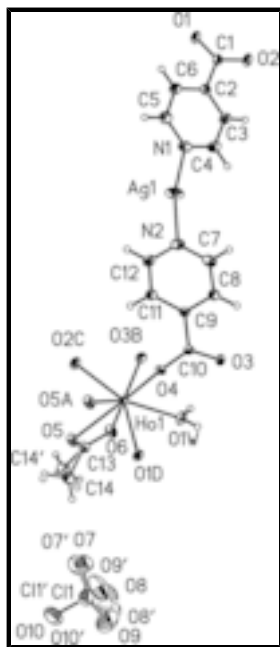


Fig. 1. The molecular structure showing the atomic-numbering scheme and displacement ellipsoids drawn at the 30% probability level. Symmetry codes: (A)  $1 - x, 1 - y, -z$ ; (B)  $x, 0.5 - y, 1/2 + z$ ; (C)  $-x, 1/2 + y, 0.5 - z$ ; (D)  $1 + x, 0.5 - y, -1/2 + z$ .

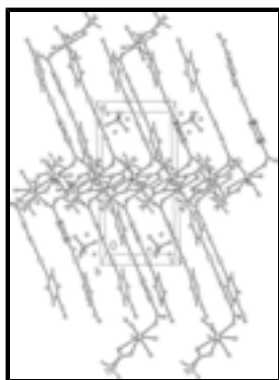


Fig. 2. A view of the three-dimensional structure of the title compound.

**Poly[[ $\mu_2$ -acetato-aquadi- $\mu_3$ -isonicotinato-holmium(III)silver(I)] perchlorate]**

*Crystal data*

[AgHo(C<sub>6</sub>H<sub>4</sub>NO<sub>2</sub>)<sub>2</sub>(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)(H<sub>2</sub>O)]ClO<sub>4</sub>

$M_r = 693.51$

Monoclinic,  $P2_1/c$

Hall symbol: -P 2ybc

$a = 16.2158 (10) \text{ \AA}$

$b = 14.9024 (9) \text{ \AA}$

$c = 7.9068 (5) \text{ \AA}$

$\beta = 91.8260 (10)^\circ$

$V = 1909.7 (2) \text{ \AA}^3$

$Z = 4$

$F_{000} = 1320$

$D_x = 2.412 \text{ Mg m}^{-3}$

Mo  $K\alpha$  radiation,  $\lambda = 0.71073 \text{ \AA}$

Cell parameters from 5514 reflections

$\theta = 2.5\text{--}27.8^\circ$

$\mu = 5.34 \text{ mm}^{-1}$

$T = 296 \text{ K}$

Block, colorless

$0.20 \times 0.18 \times 0.15 \text{ mm}$

*Data collection*

Bruker APEXII area-detector diffractometer	3439 independent reflections
Radiation source: fine-focus sealed tube	3085 reflections with $I > 2\sigma(I)$
Monochromator: graphite	$R_{\text{int}} = 0.027$
$T = 296$ K	$\theta_{\text{max}} = 25.2^\circ$
$\varphi$ and $\omega$ scan	$\theta_{\text{min}} = 1.9^\circ$
Absorption correction: multi-scan (SADABS; Sheldrick, 1996)	$h = -18 \rightarrow 19$
$T_{\text{min}} = 0.357$ , $T_{\text{max}} = 0.449$	$k = -17 \rightarrow 17$
9735 measured reflections	$l = -7 \rightarrow 9$

*Refinement*

Refinement on $F^2$	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.023$	H atoms treated by a mixture of independent and constrained refinement
$wR(F^2) = 0.054$	$w = 1/[\sigma^2(F_o^2) + (0.0229P)^2 + 2.5125P]$
$S = 1.04$	where $P = (F_o^2 + 2F_c^2)/3$
3439 reflections	$(\Delta/\sigma)_{\text{max}} = 0.001$
320 parameters	$\Delta\rho_{\text{max}} = 0.64 \text{ e } \text{\AA}^{-3}$
158 restraints	$\Delta\rho_{\text{min}} = -0.71 \text{ e } \text{\AA}^{-3}$
Primary atom site location: structure-invariant direct methods	Extinction correction: none

*Special details*

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	$x$	$y$	$z$	$U_{\text{iso}}^*/U_{\text{eq}}$	Occ. (<1)
Ho1	0.454984 (11)	0.383812 (12)	0.04996 (2)	0.01844 (7)	
Ag1	-0.02677 (3)	0.24033 (4)	0.10033 (7)	0.06287 (16)	
C1	-0.3677 (3)	0.0808 (3)	0.3622 (6)	0.0313 (10)	

## supplementary materials

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C2	-0.2892 (3)	0.1236 (3)	0.3041 (6)	0.0289 (10)	
C3	-0.2191 (3)	0.0720 (3)	0.2964 (7)	0.0410 (12)	
H10	-0.2199	0.0117	0.3268	0.049*	
C4	-0.1474 (3)	0.1114 (4)	0.2427 (8)	0.0519 (15)	
H8	-0.1002	0.0762	0.2380	0.062*	
C5	-0.2111 (3)	0.2470 (3)	0.2064 (8)	0.0481 (14)	
H7	-0.2086	0.3072	0.1761	0.058*	
C6	-0.2852 (3)	0.2128 (3)	0.2588 (7)	0.0409 (12)	
H9	-0.3315	0.2493	0.2634	0.049*	
C7	0.1336 (3)	0.1719 (4)	-0.0405 (7)	0.0442 (13)	
H5	0.1066	0.1176	-0.0251	0.053*	
C8	0.2120 (3)	0.1707 (3)	-0.1043 (6)	0.0351 (11)	
H4	0.2372	0.1166	-0.1302	0.042*	
C9	0.2521 (3)	0.2510 (3)	-0.1289 (5)	0.0230 (9)	
C10	0.3375 (2)	0.2510 (3)	-0.1983 (5)	0.0206 (9)	
C11	0.2125 (3)	0.3295 (3)	-0.0872 (6)	0.0325 (11)	
H3	0.2382	0.3846	-0.1017	0.039*	
C12	0.1343 (3)	0.3254 (3)	-0.0238 (6)	0.0376 (12)	
H6	0.1080	0.3787	0.0034	0.045*	
N1	-0.1427 (2)	0.1969 (3)	0.1974 (6)	0.0458 (11)	
N2	0.0950 (2)	0.2480 (3)	0.0001 (5)	0.0394 (10)	
O1	-0.4242 (2)	0.1332 (2)	0.4072 (5)	0.0387 (9)	
O2	-0.3699 (2)	-0.0029 (2)	0.3616 (4)	0.0397 (8)	
O3	0.35906 (18)	0.18441 (19)	-0.2810 (4)	0.0303 (7)	
O4	0.38257 (17)	0.31880 (19)	-0.1691 (4)	0.0260 (7)	
O5	0.54736 (19)	0.5035 (2)	0.1583 (4)	0.0312 (7)	
O6	0.5344 (2)	0.3847 (2)	0.3129 (4)	0.0352 (8)	
O1W	0.4985 (2)	0.2318 (2)	0.0763 (4)	0.0353 (8)	
H1W	0.478 (3)	0.197 (2)	0.143 (5)	0.053*	
H2W	0.516 (3)	0.200 (3)	0.002 (4)	0.053*	
C13	0.5691 (3)	0.4589 (3)	0.2876 (5)	0.0272 (10)	
C14	0.6256 (13)	0.4991 (15)	0.423 (2)	0.040 (3)	0.49 (4)
H14A	0.5956	0.5074	0.5247	0.060*	0.49 (4)
H14B	0.6457	0.5560	0.3852	0.060*	0.49 (4)
H14C	0.6713	0.4594	0.4455	0.060*	0.49 (4)
C14'	0.6450 (11)	0.4886 (16)	0.388 (3)	0.040 (3)	0.51 (4)
H14D	0.6545	0.4486	0.4821	0.060*	0.51 (4)
H14E	0.6368	0.5484	0.4297	0.060*	0.51 (4)
H14F	0.6918	0.4877	0.3171	0.060*	0.51 (4)
C11	0.91931 (10)	0.45924 (11)	0.2560 (2)	0.0625 (4)	0.539 (12)
O7	0.8512 (8)	0.4643 (8)	0.3717 (18)	0.126 (5)	0.539 (12)
O8	0.8847 (9)	0.4462 (8)	0.0955 (12)	0.111 (5)	0.539 (12)
O9	0.9703 (9)	0.3896 (8)	0.300 (2)	0.130 (6)	0.539 (12)
O10	0.9589 (10)	0.5442 (8)	0.259 (2)	0.086 (6)	0.539 (12)
C11'	0.91931 (10)	0.45924 (11)	0.2560 (2)	0.0625 (4)	0.461 (12)
O7'	0.8366 (6)	0.4707 (8)	0.210 (2)	0.094 (5)	0.461 (12)
O8'	0.9624 (10)	0.4225 (10)	0.1137 (18)	0.132 (6)	0.461 (12)
O9'	0.9275 (11)	0.3943 (10)	0.3831 (18)	0.131 (7)	0.461 (12)
O10'	0.9581 (11)	0.5392 (9)	0.304 (2)	0.082 (6)	0.461 (12)

Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
Ho1	0.01607 (11)	0.01592 (11)	0.02369 (12)	-0.00070 (7)	0.00617 (7)	-0.00010 (7)
Ag1	0.0259 (2)	0.0859 (4)	0.0783 (3)	-0.0142 (2)	0.0260 (2)	0.0008 (3)
C1	0.027 (2)	0.023 (2)	0.044 (3)	-0.0049 (19)	0.016 (2)	-0.004 (2)
C2	0.027 (2)	0.027 (2)	0.033 (2)	-0.0057 (19)	0.0118 (19)	-0.0029 (19)
C3	0.027 (3)	0.028 (3)	0.069 (4)	0.002 (2)	0.017 (2)	0.004 (2)
C4	0.027 (3)	0.046 (4)	0.083 (4)	0.000 (2)	0.017 (3)	-0.006 (3)
C5	0.033 (3)	0.034 (3)	0.078 (4)	-0.009 (2)	0.020 (3)	0.008 (3)
C6	0.026 (3)	0.033 (3)	0.064 (4)	-0.001 (2)	0.017 (2)	0.005 (2)
C7	0.032 (3)	0.041 (3)	0.060 (3)	-0.013 (2)	0.014 (2)	0.003 (3)
C8	0.028 (3)	0.028 (3)	0.050 (3)	-0.0040 (19)	0.012 (2)	-0.001 (2)
C9	0.019 (2)	0.027 (2)	0.023 (2)	-0.0016 (17)	0.0029 (17)	-0.0019 (17)
C10	0.018 (2)	0.022 (2)	0.023 (2)	0.0006 (16)	0.0015 (17)	0.0003 (17)
C11	0.025 (2)	0.028 (3)	0.045 (3)	-0.0027 (19)	0.012 (2)	-0.002 (2)
C12	0.028 (3)	0.035 (3)	0.051 (3)	0.005 (2)	0.015 (2)	-0.001 (2)
N1	0.025 (2)	0.046 (3)	0.067 (3)	-0.0090 (19)	0.018 (2)	-0.001 (2)
N2	0.025 (2)	0.047 (3)	0.047 (3)	-0.0028 (18)	0.0136 (18)	-0.001 (2)
O1	0.0304 (18)	0.0221 (17)	0.065 (2)	-0.0017 (13)	0.0281 (17)	-0.0023 (15)
O2	0.0346 (18)	0.0219 (17)	0.064 (2)	-0.0025 (14)	0.0279 (17)	-0.0039 (15)
O3	0.0263 (17)	0.0225 (16)	0.0431 (19)	-0.0017 (13)	0.0144 (14)	-0.0104 (14)
O4	0.0185 (15)	0.0259 (16)	0.0338 (17)	-0.0060 (12)	0.0058 (12)	-0.0054 (13)
O5	0.0382 (18)	0.0269 (17)	0.0279 (16)	-0.0076 (14)	-0.0067 (14)	0.0043 (13)
O6	0.044 (2)	0.0298 (18)	0.0315 (18)	-0.0093 (15)	-0.0022 (15)	0.0102 (14)
O1W	0.049 (2)	0.0215 (16)	0.037 (2)	0.0061 (15)	0.0189 (16)	0.0013 (14)
C13	0.028 (2)	0.029 (2)	0.025 (2)	-0.0030 (19)	-0.0014 (18)	0.0027 (19)
C14	0.039 (6)	0.044 (5)	0.036 (6)	-0.003 (5)	-0.005 (5)	-0.003 (4)
C14'	0.039 (6)	0.044 (5)	0.036 (6)	-0.003 (5)	-0.005 (5)	-0.003 (4)
C11	0.0599 (10)	0.0510 (9)	0.0761 (11)	0.0018 (7)	-0.0046 (8)	-0.0019 (8)
O7	0.119 (9)	0.127 (8)	0.133 (9)	-0.027 (7)	0.053 (7)	-0.013 (7)
O8	0.143 (10)	0.114 (8)	0.072 (6)	-0.050 (7)	-0.042 (6)	-0.005 (6)
O9	0.142 (9)	0.098 (7)	0.148 (10)	0.072 (7)	-0.037 (7)	-0.016 (7)
O10	0.077 (8)	0.084 (9)	0.097 (9)	-0.030 (6)	0.007 (6)	-0.010 (6)
C11'	0.0599 (10)	0.0510 (9)	0.0761 (11)	0.0018 (7)	-0.0046 (8)	-0.0019 (8)
O7'	0.057 (6)	0.077 (7)	0.145 (10)	0.001 (5)	-0.021 (6)	0.011 (7)
O8'	0.147 (11)	0.131 (9)	0.124 (9)	0.016 (8)	0.068 (8)	-0.035 (7)
O9'	0.143 (11)	0.133 (10)	0.117 (9)	-0.011 (8)	-0.019 (8)	0.065 (8)
O10'	0.069 (9)	0.076 (9)	0.103 (9)	-0.026 (7)	0.022 (7)	-0.045 (7)

Geometric parameters ( $\text{\AA}$ ,  $^\circ$ )

Ho1—O4	2.278 (3)	C9—C11	1.380 (6)
Ho1—O2 <sup>i</sup>	2.303 (3)	C9—C10	1.504 (5)
Ho1—O1 <sup>ii</sup>	2.306 (3)	C10—O3	1.245 (5)
Ho1—O3 <sup>iii</sup>	2.318 (3)	C10—O4	1.264 (5)
Ho1—O5 <sup>iv</sup>	2.352 (3)	C11—C12	1.379 (6)

## supplementary materials

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Ho1—O1W	2.380 (3)	C11—H3	0.9300
Ho1—O6	2.410 (3)	C12—N2	1.335 (6)
Ho1—O5	2.465 (3)	C12—H6	0.9300
Ag1—N1	2.152 (4)	O1—Ho1 <sup>v</sup>	2.306 (3)
Ag1—N2	2.154 (4)	O2—Ho1 <sup>vi</sup>	2.303 (3)
C1—O2	1.248 (5)	O3—Ho1 <sup>vii</sup>	2.318 (3)
C1—O1	1.264 (5)	O5—C13	1.260 (5)
C1—C2	1.509 (6)	O5—Ho1 <sup>iv</sup>	2.352 (3)
C2—C3	1.376 (6)	O6—C13	1.259 (5)
C2—C6	1.379 (7)	O1W—H1W	0.82 (4)
C3—C4	1.380 (7)	O1W—H2W	0.81 (4)
C3—H10	0.9300	C13—C14'	1.510 (9)
C4—N1	1.327 (7)	C13—C14	1.511 (9)
C4—H8	0.9300	C14—H14A	0.9600
C5—N1	1.341 (7)	C14—H14B	0.9600
C5—C6	1.380 (6)	C14—H14C	0.9600
C5—H7	0.9300	C14'—H14D	0.9600
C6—H9	0.9300	C14'—H14E	0.9600
C7—N2	1.340 (7)	C14'—H14F	0.9600
C7—C8	1.382 (6)	C11—O9	1.364 (8)
C7—H5	0.9300	C11—O8	1.385 (8)
C8—C9	1.380 (6)	C11—O10	1.419 (9)
C8—H4	0.9300	C11—O7	1.458 (8)
O4—Ho1—O2 <sup>i</sup>	104.08 (12)	C2—C6—H9	120.6
O4—Ho1—O1 <sup>ii</sup>	90.36 (11)	C5—C6—H9	120.6
O2 <sup>i</sup> —Ho1—O1 <sup>ii</sup>	139.11 (10)	N2—C7—C8	122.7 (5)
O4—Ho1—O3 <sup>iii</sup>	84.99 (11)	N2—C7—H5	118.7
O2 <sup>i</sup> —Ho1—O3 <sup>iii</sup>	74.14 (10)	C8—C7—H5	118.7
O1 <sup>ii</sup> —Ho1—O3 <sup>iii</sup>	146.11 (10)	C9—C8—C7	119.0 (4)
O4—Ho1—O5 <sup>iv</sup>	76.99 (10)	C9—C8—H4	120.5
O2 <sup>i</sup> —Ho1—O5 <sup>iv</sup>	72.10 (11)	C7—C8—H4	120.5
O1 <sup>ii</sup> —Ho1—O5 <sup>iv</sup>	74.39 (11)	C11—C9—C8	118.4 (4)
O3 <sup>iii</sup> —Ho1—O5 <sup>iv</sup>	136.13 (11)	C11—C9—C10	121.9 (4)
O4—Ho1—O1W	78.79 (12)	C8—C9—C10	119.7 (4)
O2 <sup>i</sup> —Ho1—O1W	148.20 (11)	O3—C10—O4	124.2 (4)
O1 <sup>ii</sup> —Ho1—O1W	71.57 (11)	O3—C10—C9	118.0 (3)
O3 <sup>iii</sup> —Ho1—O1W	74.60 (10)	O4—C10—C9	117.8 (3)
O5 <sup>iv</sup> —Ho1—O1W	137.66 (11)	C12—C11—C9	119.4 (4)
O4—Ho1—O6	155.03 (10)	C12—C11—H3	120.3
O2 <sup>i</sup> —Ho1—O6	92.45 (12)	C9—C11—H3	120.3
O1 <sup>ii</sup> —Ho1—O6	89.10 (12)	N2—C12—C11	122.5 (4)
O3 <sup>iii</sup> —Ho1—O6	81.64 (11)	N2—C12—H6	118.7
O5 <sup>iv</sup> —Ho1—O6	126.62 (10)	C11—C12—H6	118.7
O1W—Ho1—O6	77.37 (12)	C4—N1—C5	117.7 (4)

O4—Ho1—O5	149.88 (10)	C4—N1—Ag1	116.4 (3)
O2 <sup>i</sup> —Ho1—O5	74.19 (11)	C5—N1—Ag1	125.8 (4)
O1 <sup>ii</sup> —Ho1—O5	74.55 (11)	C12—N2—C7	118.0 (4)
O3 <sup>iii</sup> —Ho1—O5	121.89 (11)	C12—N2—Ag1	123.1 (3)
O5 <sup>iv</sup> —Ho1—O5	73.91 (11)	C7—N2—Ag1	118.9 (3)
O1W—Ho1—O5	118.90 (12)	C1—O1—Ho1 <sup>v</sup>	134.8 (3)
O6—Ho1—O5	52.72 (10)	C1—O2—Ho1 <sup>vi</sup>	138.5 (3)
O4—Ho1—C13	170.02 (11)	C10—O3—Ho1 <sup>vii</sup>	149.1 (3)
O2 <sup>i</sup> —Ho1—C13	83.92 (13)	C10—O4—Ho1	139.5 (3)
O1 <sup>ii</sup> —Ho1—C13	79.67 (13)	C13—O5—Ho1 <sup>iv</sup>	160.5 (3)
O3 <sup>iii</sup> —Ho1—C13	103.15 (12)	C13—O5—Ho1	92.9 (2)
O5 <sup>iv</sup> —Ho1—C13	100.29 (11)	Ho1 <sup>iv</sup> —O5—Ho1	106.09 (11)
O1W—Ho1—C13	97.63 (13)	C13—O6—Ho1	95.6 (3)
O6—Ho1—C13	26.34 (11)	Ho1—O1W—H1W	122 (3)
O5—Ho1—C13	26.45 (11)	Ho1—O1W—H2W	127 (3)
O4—Ho1—Ho1 <sup>iv</sup>	114.60 (7)	H1W—O1W—H2W	105 (4)
O2 <sup>i</sup> —Ho1—Ho1 <sup>iv</sup>	68.77 (7)	O6—C13—O5	118.5 (4)
O1 <sup>ii</sup> —Ho1—Ho1 <sup>iv</sup>	70.43 (7)	O6—C13—C14'	122.3 (10)
O3 <sup>iii</sup> —Ho1—Ho1 <sup>iv</sup>	141.14 (7)	O5—C13—C14'	118.3 (10)
O5 <sup>iv</sup> —Ho1—Ho1 <sup>iv</sup>	37.97 (7)	O6—C13—C14	119.9 (10)
O1W—Ho1—Ho1 <sup>iv</sup>	139.55 (8)	O5—C13—C14	120.9 (10)
O6—Ho1—Ho1 <sup>iv</sup>	88.65 (7)	C14'—C13—C14	17.2 (11)
O5—Ho1—Ho1 <sup>iv</sup>	35.94 (7)	O6—C13—Ho1	58.1 (2)
C13—Ho1—Ho1 <sup>iv</sup>	62.34 (9)	O5—C13—Ho1	60.6 (2)
N1—Ag1—N2	165.55 (17)	C14'—C13—Ho1	166.4 (11)
O2—C1—O1	126.7 (4)	C14—C13—Ho1	176.2 (10)
O2—C1—C2	116.5 (4)	C13—C14—H14A	109.5
O1—C1—C2	116.8 (4)	C13—C14—H14B	109.5
C3—C2—C6	118.8 (4)	C13—C14—H14C	109.5
C3—C2—C1	118.9 (4)	C13—C14'—H14D	109.5
C6—C2—C1	122.3 (4)	C13—C14'—H14E	109.5
C2—C3—C4	118.7 (5)	H14D—C14'—H14E	109.5
C2—C3—H10	120.6	C13—C14'—H14F	109.5
C4—C3—H10	120.6	H14D—C14'—H14F	109.5
N1—C4—C3	123.3 (5)	H14E—C14'—H14F	109.5
N1—C4—H8	118.4	O9—C11—O8	110.5 (7)
C3—C4—H8	118.4	O9—C11—O10	113.8 (8)
N1—C5—C6	122.7 (5)	O8—C11—O10	108.2 (8)
N1—C5—H7	118.7	O9—C11—O7	110.2 (8)
C6—C5—H7	118.7	O8—C11—O7	106.8 (7)
C2—C6—C5	118.9 (5)	O10—C11—O7	107.0 (7)

Symmetry codes: (i)  $-x, y+1/2, -z+1/2$ ; (ii)  $x+1, -y+1/2, z-1/2$ ; (iii)  $x, -y+1/2, z+1/2$ ; (iv)  $-x+1, -y+1, -z$ ; (v)  $x-1, -y+1/2, z+1/2$ ; (vi)  $-x, y-1/2, -z+1/2$ ; (vii)  $x, -y+1/2, z-1/2$ .

## supplementary materials

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### Hydrogen-bond geometry (Å, °)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
O1W—H1W $\cdots$ O4 <sup>iii</sup>	0.82 (4)	2.19 (3)	2.898 (4)	145 (4)
O1W—H2W $\cdots$ O6 <sup>vii</sup>	0.81 (4)	1.99 (4)	2.787 (4)	168 (5)

Symmetry codes: (iii)  $x, -y+1/2, z+1/2$ ; (vii)  $x, -y+1/2, z-1/2$ .



Fig. 2

