

# A monoclinic polymorph of 1-(4-chlorophenyl)-3-(4-methoxyphenyl)prop-2-en-1-one

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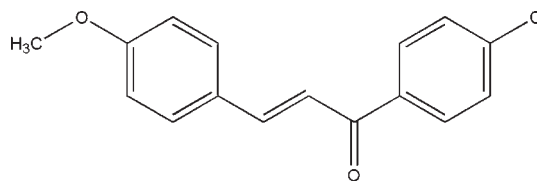
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Key indicators: single-crystal X-ray study;  $T = 110$  K; mean  $\sigma(\text{C}-\text{C}) = 0.002$  Å;  $R$  factor = 0.035;  $wR$  factor = 0.096; data-to-parameter ratio = 14.7.

The crystal structure of the title compound,  $\text{C}_{16}\text{H}_{13}\text{ClO}_2$  (II), (space group  $P2_1/c$ ), is a polymorph of the structure, (I), reported by Harrison, Yathirajan, Sarojini, Narayana & Indira [*Acta Cryst.* (2006), **E62**, o1647–o1649] in the orthorhombic space group  $Pna2_1$ . The dihedral angle between the mean planes of the 4-chloro- and 4-methoxy-substituted benzene rings is  $52.9$  (1)° in (II) compared to  $21.82$  (6)° for polymorph (I). The dihedral angles between the mean planes of the prop-2-en-1-one group and those of the 4-chlorophenyl and 4-methoxyphenyl rings are  $23.3$  (3) and  $33.7$  (1)°, respectively, in (II). The corresponding values are  $17.7$  (1) and  $6.0$  (3)°, respectively, in polymorph (I). In the crystal, weak  $\text{C}-\text{H}\cdots\pi$  interactions are observed.

## Related literature

For the orthorhombic polymorph, see: Harrison *et al.* (2006). For the biological activity of chalcones and flavonoids, see: Dimmock *et al.* (1999); Opletalova & Sedivy (1999); Lin *et al.* (2002); Nowakowska (2007). For the synthesis and biological activity of some fluorinated chalcone derivatives, see: Nakamura *et al.* (2002). For non-linear optical studies of chalcones and their derivatives, see: Sarojini *et al.* (2006); Poornesh *et al.* (2009); Shettigar *et al.* (2006, 2008). For our studies of chalcones, see: Jasinski *et al.* (2009).



## Experimental

### Crystal data

$\text{C}_{16}\text{H}_{13}\text{ClO}_2$   
 $M_r = 272.71$   
Monoclinic,  $P2_1/c$   
 $a = 15.6695$  (7) Å  
 $b = 14.1235$  (8) Å  
 $c = 5.8455$  (3) Å  
 $\beta = 90.771$  (5)°

$V = 1293.53$  (12) Å<sup>3</sup>  
 $Z = 4$   
Cu  $K\alpha$  radiation  
 $\mu = 2.57$  mm<sup>-1</sup>  
 $T = 110$  K  
 $0.54 \times 0.13 \times 0.08$  mm

### Data collection

Oxford Diffraction Xcalibur diffractometer with a Ruby (Gemini Cu) detector  
Absorption correction: multi-scan (*CrysAlis RED*; Oxford

Diffraction, 2007)  
 $T_{\min} = 0.483$ ,  $T_{\max} = 0.558$   
5083 measured reflections  
2537 independent reflections  
2223 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.021$

### Refinement

$R[F^2 > 2\sigma(F^2)] = 0.035$   
 $wR(F^2) = 0.096$   
 $S = 1.04$   
2537 reflections

173 parameters  
H-atom parameters constrained  
 $\Delta\rho_{\max} = 0.25$  e Å<sup>-3</sup>  
 $\Delta\rho_{\min} = -0.22$  e Å<sup>-3</sup>

**Table 1**

Hydrogen-bond geometry (Å, °).

Cg1 and Cg2 are the centroids of the C1–C6 and C10–C15 rings, respectively.

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{C2}-\text{H2}\cdots\text{Cg1}^{\text{i}}$	0.95	2.85	3.4675 (15)	124
$\text{C12}-\text{H12}\cdots\text{Cg2}^{\text{ii}}$	0.95	2.92	3.6616 (17)	136

Symmetry codes: (i)  $x, -y + \frac{3}{2}, z + \frac{1}{2}$ ; (ii)  $x, -y + \frac{3}{2}, z - \frac{1}{2}$ .

Data collection: *CrysAlis PRO* (Oxford Diffraction, 2007); cell refinement: *CrysAlis RED* (Oxford Diffraction, 2007); data reduction: *CrysAlis RED*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXL97*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: OM2306).

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## supporting information

*Acta Cryst.* (2010). E66, o269–o270 [https://doi.org/10.1107/S1600536809054956]

## A monoclinic polymorph of 1-(4-chlorophenyl)-3-(4-methoxyphenyl)prop-2-en-1-one

Jerry P. Jasinski, Ray J. Butcher, B. Narayana, S. Samshuddin and H. S. Yathirajan

### S1. Comment

Chalcone is an unique template molecule that is associated with several biological activities. A review on the bioactivities of chalcones is described (Dimmock *et al.* 1999). Chalcones and their heterocyclic analogs as potential antifungal chemotherapeutic agents is published (Opletalova & Sedivy, 1999). Chalcones and flavonoids as anti-tuberculosis agents has been reported (Lin *et al.* 2002) and a review of anti-infective and anti-inflammatory chalcones is also described (Nowakowska, 2007) as well as the synthesis and biological activities of some fluorinated chalcone derivatives (Nakamura *et al.* 2002). In addition, chalcones are finding applications as organic non-linear optical materials (NLO) due to their good SHG conversion efficiencies (Sarojini *et al.* 2006). Recently, non-linear optical studies on a few chalcones and their derivatives were reported (Poornesh *et al.* 2009; Shettigar *et al.* 2006; 2008). In continuation of our work on chalcones (Jasinski *et al.* 2009) and in view of the importance of chloro chalcones, this paper describes a new polymorphic form of (I), C<sub>16</sub>H<sub>13</sub>ClO<sub>2</sub>, 1-(4-chlorophenyl)-3-(4-methoxyphenyl)-prop-2-en-1-one, first reported by Harrison *et al.* (2006). Substantial changes in the cell parameters provides solid support for the recognition of this new polymorphic form for (I).

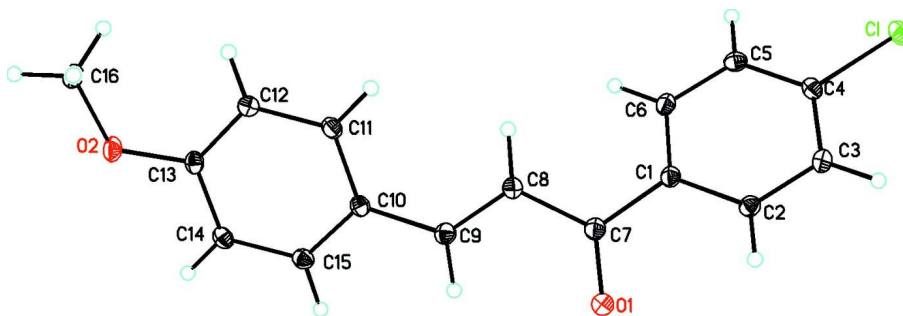
The title compound, (II), is a chalcone derivative with 4-chlorophenyl and 4-methoxyphenyl rings bonded at the opposite ends of a propenone group, the biologically active region (Fig.1). The dihedral angle between mean planes of the 4-chloro and 4-methoxy substituted benzene rings in (II) is 52.9 (1)° compared to 21.82 (6)° (Harrison *et al.* (2006); 4-chlorophenyl & 4-methoxyphenyl) for polymorph (I) in the orthorhombic, *Pna2<sub>1</sub>*, space group. The angles between the mean plane of the prop-2-ene-1-one group and those of the 4-chlorophenyl and 4-methoxyphenyl rings in (II) are 23.3 (3)° and 33.7 (1)°, respectively. This compares to 17.7 (1)° and 6.0 (3)° in polymorph (I). A weak intramolecular C9–H9···O1 hydrogen bond interaction is present which may help to maintain the molecular conformation of the molecule (Table 1) and similar to that observed in (I). While no classical hydrogen bonds are present, weak intermolecular C–H···Cg  $\pi$ -ring interactions are observed, Cg1 = C1–C6 and Cg2 = C10–C15, see Table 1.

### S2. Experimental

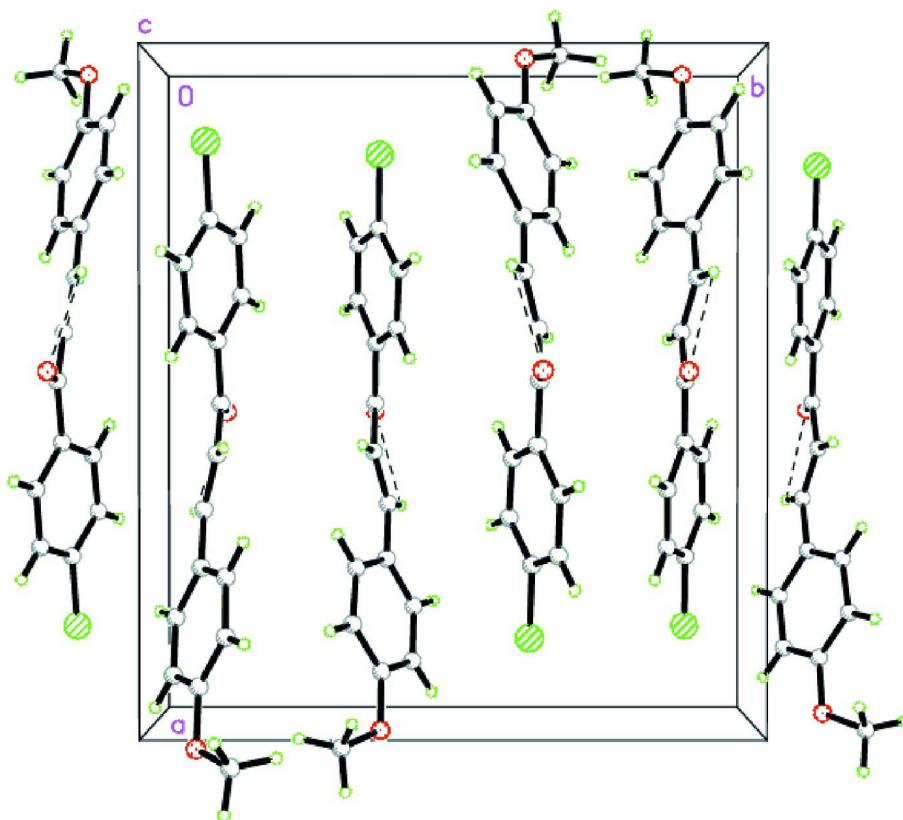
In (II), 4-chloroacetophenone in ethanol (1.54 g, 0.01 mol) (25 ml) was mixed with 4-methoxybenzaldehyde (1.36 g, 0.01 mol) in ethanol (25 ml) and the mixture was treated with an aqueous solution of potassium hydroxide (20 ml, 5%). This mixture was stirred well and left to stand for 24 hr. The resulting crude solid mass was collected by filtration and recrystallized from ethanol, yielding clear blocks of (II). Yield: 90%, m.p.: 391–393 K, analysis found (calculated) for C<sub>16</sub>H<sub>13</sub>ClO<sub>2</sub>: C: 70.5 (70.4%); H: 4.72 (4.76%). The preparation and crystallization procedure for (I) was identical to that described above for (II). However, in (I) the m.p. measured 380 K, a difference of 12 K. The samples of (I) and (II) were not independently tested for concomitant polymorphism.

## S3. Refinement

All of the H atoms were placed in their calculated positions and then refined using the riding model with C—H = 0.95–0.98 Å, and with  $U_{\text{iso}}(\text{H}) = 1.18\text{--}1.48U_{\text{eq}}(\text{C})$ .

**Figure 1**

Molecular structure of the title compound,  $\text{C}_{16}\text{H}_{13}\text{ClO}_2$ , (II), showing the atom labeling scheme and 50% probability displacement ellipsoids.

**Figure 2**

Packing diagram of the title compound, (II), viewed down the  $c$  axis. Weak C—H...O intramolecular hydrogen bond interactions are shown as dashed lines.

## 1-(4-chlorophenyl)-3-(4-methoxyphenyl)prop-2-en-1-one

## Crystal data

C<sub>16</sub>H<sub>13</sub>ClO<sub>2</sub> $M_r = 272.71$ Monoclinic,  $P2_1/c$ 

Hall symbol: -P 2ybc

 $a = 15.6695$  (7) Å $b = 14.1235$  (8) Å $c = 5.8455$  (3) Å $\beta = 90.771$  (5)° $V = 1293.53$  (12) Å<sup>3</sup> $Z = 4$  $F(000) = 568$  $D_x = 1.400$  Mg m<sup>-3</sup>Cu  $K\alpha$  radiation,  $\lambda = 1.54184$  Å

Cell parameters from 3121 reflections

 $\theta = 4.2$ – $73.8$ ° $\mu = 2.57$  mm<sup>-1</sup> $T = 110$  K

Needle, colorless

 $0.54 \times 0.13 \times 0.08$  mm

## Data collection

Oxford Diffraction Xcalibur

diffractometer with a Ruby (Gemini Cu)

detector

Radiation source: Enhance (Cu) X-ray Source

Graphite monochromator

Detector resolution: 10.5081 pixels mm<sup>-1</sup> $\omega$  scans

Absorption correction: multi-scan

(CrysAlis RED; Oxford Diffraction, 2007)

 $T_{\min} = 0.483$ ,  $T_{\max} = 0.558$ 

5083 measured reflections

2537 independent reflections

2223 reflections with  $I > 2\sigma(I)$  $R_{\text{int}} = 0.021$  $\theta_{\max} = 74.0$ °,  $\theta_{\min} = 4.2$ ° $h = -19 \rightarrow 17$  $k = -16 \rightarrow 17$  $l = -5 \rightarrow 7$ 

## Refinement

Refinement on  $F^2$ 

Least-squares matrix: full

 $R[F^2 > 2\sigma(F^2)] = 0.035$  $wR(F^2) = 0.096$  $S = 1.04$ 

2537 reflections

173 parameters

0 restraints

Primary atom site location: structure-invariant

direct methods

Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring sites

H-atom parameters constrained

 $w = 1/[\sigma^2(F_o^2) + (0.0574P)^2 + 0.4041P]$ where  $P = (F_o^2 + 2F_c^2)/3$  $(\Delta/\sigma)_{\max} = 0.001$  $\Delta\rho_{\max} = 0.25$  e Å<sup>-3</sup> $\Delta\rho_{\min} = -0.22$  e Å<sup>-3</sup>

## Special details

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) etc. and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å<sup>2</sup>)

	$x$	$y$	$z$	$U_{\text{iso}}^*/U_{\text{eq}}$
Cl1	0.84683 (2)	0.61936 (3)	0.17788 (6)	0.02523 (13)
O1	0.47093 (7)	0.63259 (9)	0.70069 (19)	0.0289 (3)
O2	0.01821 (7)	0.61488 (8)	-0.07719 (19)	0.0241 (3)
C1	0.57310 (9)	0.62821 (10)	0.4074 (3)	0.0181 (3)

C2	0.63878 (9)	0.66090 (10)	0.5520 (2)	0.0188 (3)
H2	0.6255	0.6849	0.6990	0.023*
C3	0.72286 (9)	0.65859 (10)	0.4828 (3)	0.0198 (3)
H3	0.7672	0.6817	0.5800	0.024*
C4	0.74128 (9)	0.62174 (10)	0.2682 (3)	0.0193 (3)
C5	0.67775 (10)	0.58747 (11)	0.1233 (3)	0.0205 (3)
H5	0.6916	0.5612	-0.0212	0.025*
C6	0.59317 (9)	0.59211 (11)	0.1929 (3)	0.0202 (3)
H6	0.5489	0.5705	0.0934	0.024*
C7	0.48352 (9)	0.63066 (10)	0.4945 (3)	0.0200 (3)
C8	0.41225 (9)	0.63059 (11)	0.3255 (3)	0.0211 (3)
H8	0.4223	0.6467	0.1704	0.025*
C9	0.33359 (9)	0.60773 (11)	0.3930 (3)	0.0200 (3)
H9	0.3284	0.5885	0.5479	0.024*
C10	0.25468 (9)	0.60893 (10)	0.2562 (3)	0.0181 (3)
C11	0.24715 (9)	0.65496 (11)	0.0440 (3)	0.0202 (3)
H11	0.2960	0.6839	-0.0206	0.024*
C12	0.16938 (9)	0.65893 (11)	-0.0730 (3)	0.0200 (3)
H12	0.1648	0.6915	-0.2149	0.024*
C13	0.09809 (9)	0.61476 (10)	0.0193 (3)	0.0187 (3)
C14	0.10492 (9)	0.56679 (11)	0.2274 (3)	0.0202 (3)
H14	0.0567	0.5354	0.2882	0.024*
C15	0.18181 (9)	0.56513 (11)	0.3443 (2)	0.0199 (3)
H15	0.1856	0.5336	0.4877	0.024*
C16	0.00342 (10)	0.67344 (13)	-0.2722 (3)	0.0263 (3)
H16A	0.0140	0.7397	-0.2311	0.039*
H16B	-0.0559	0.6662	-0.3248	0.039*
H16C	0.0420	0.6547	-0.3948	0.039*

*Atomic displacement parameters (Å<sup>2</sup>)*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
C11	0.0180 (2)	0.0276 (2)	0.0302 (2)	0.00214 (13)	0.00556 (14)	0.00128 (15)
O1	0.0214 (6)	0.0441 (7)	0.0212 (6)	-0.0035 (5)	0.0006 (4)	0.0008 (5)
O2	0.0158 (5)	0.0266 (6)	0.0298 (6)	-0.0025 (4)	-0.0041 (4)	0.0048 (5)
C1	0.0174 (7)	0.0170 (7)	0.0198 (7)	-0.0009 (5)	-0.0019 (5)	0.0039 (6)
C2	0.0197 (7)	0.0188 (7)	0.0180 (7)	0.0019 (5)	-0.0009 (5)	-0.0002 (6)
C3	0.0185 (7)	0.0188 (7)	0.0221 (7)	0.0000 (5)	-0.0031 (5)	-0.0009 (6)
C4	0.0167 (7)	0.0172 (7)	0.0239 (7)	0.0018 (5)	0.0022 (6)	0.0034 (6)
C5	0.0251 (7)	0.0189 (7)	0.0176 (7)	0.0019 (6)	0.0016 (6)	-0.0002 (5)
C6	0.0194 (7)	0.0206 (7)	0.0206 (7)	-0.0015 (5)	-0.0038 (5)	0.0008 (6)
C7	0.0196 (7)	0.0193 (7)	0.0212 (7)	-0.0006 (5)	-0.0008 (6)	0.0012 (6)
C8	0.0178 (7)	0.0248 (8)	0.0207 (7)	0.0005 (6)	0.0000 (6)	0.0021 (6)
C9	0.0206 (7)	0.0189 (7)	0.0204 (7)	0.0017 (5)	-0.0001 (6)	0.0010 (6)
C10	0.0167 (7)	0.0170 (7)	0.0207 (7)	0.0023 (5)	0.0015 (5)	-0.0012 (6)
C11	0.0164 (7)	0.0222 (7)	0.0220 (7)	-0.0007 (5)	0.0032 (5)	0.0011 (6)
C12	0.0195 (7)	0.0215 (8)	0.0190 (7)	0.0001 (6)	0.0012 (5)	0.0008 (6)
C13	0.0161 (7)	0.0171 (7)	0.0228 (7)	0.0009 (5)	-0.0013 (5)	-0.0033 (6)

C14	0.0177 (7)	0.0187 (7)	0.0245 (7)	-0.0015 (5)	0.0039 (5)	0.0012 (6)
C15	0.0212 (7)	0.0186 (7)	0.0201 (7)	0.0016 (6)	0.0030 (5)	0.0015 (6)
C16	0.0213 (7)	0.0341 (9)	0.0235 (8)	0.0027 (6)	-0.0036 (6)	0.0014 (7)

*Geometric parameters (Å, °)*

C11—C4	1.7431 (15)	C8—H8	0.9500
O1—C7	1.2240 (19)	C9—C10	1.464 (2)
O2—C13	1.3659 (18)	C9—H9	0.9500
O2—C16	1.4248 (19)	C10—C15	1.403 (2)
C1—C6	1.394 (2)	C10—C11	1.404 (2)
C1—C2	1.401 (2)	C11—C12	1.391 (2)
C1—C7	1.500 (2)	C11—H11	0.9500
C2—C3	1.384 (2)	C12—C13	1.395 (2)
C2—H2	0.9500	C12—H12	0.9500
C3—C4	1.392 (2)	C13—C14	1.395 (2)
C3—H3	0.9500	C14—C15	1.377 (2)
C4—C5	1.386 (2)	C14—H14	0.9500
C5—C6	1.393 (2)	C15—H15	0.9500
C5—H5	0.9500	C16—H16A	0.9800
C6—H6	0.9500	C16—H16B	0.9800
C7—C8	1.481 (2)	C16—H16C	0.9800
C8—C9	1.339 (2)		
C13—O2—C16	118.00 (12)	C8—C9—H9	116.2
C6—C1—C2	119.36 (14)	C10—C9—H9	116.2
C6—C1—C7	122.52 (13)	C15—C10—C11	118.00 (13)
C2—C1—C7	118.09 (13)	C15—C10—C9	118.70 (14)
C3—C2—C1	120.69 (14)	C11—C10—C9	123.25 (13)
C3—C2—H2	119.7	C12—C11—C10	121.06 (13)
C1—C2—H2	119.7	C12—C11—H11	119.5
C2—C3—C4	118.82 (13)	C10—C11—H11	119.5
C2—C3—H3	120.6	C11—C12—C13	119.52 (14)
C4—C3—H3	120.6	C11—C12—H12	120.2
C5—C4—C3	121.70 (14)	C13—C12—H12	120.2
C5—C4—C11	119.01 (12)	O2—C13—C12	125.02 (13)
C3—C4—C11	119.29 (11)	O2—C13—C14	114.83 (13)
C4—C5—C6	118.91 (14)	C12—C13—C14	120.15 (13)
C4—C5—H5	120.5	C15—C14—C13	119.82 (14)
C6—C5—H5	120.5	C15—C14—H14	120.1
C5—C6—C1	120.48 (13)	C13—C14—H14	120.1
C5—C6—H6	119.8	C14—C15—C10	121.41 (14)
C1—C6—H6	119.8	C14—C15—H15	119.3
O1—C7—C8	121.77 (14)	C10—C15—H15	119.3
O1—C7—C1	119.92 (13)	O2—C16—H16A	109.5
C8—C7—C1	118.31 (13)	O2—C16—H16B	109.5
C9—C8—C7	119.51 (14)	H16A—C16—H16B	109.5
C9—C8—H8	120.2	O2—C16—H16C	109.5

C7—C8—H8	120.2	H16A—C16—H16C	109.5
C8—C9—C10	127.64 (14)	H16B—C16—H16C	109.5
C6—C1—C2—C3	0.7 (2)	C7—C8—C9—C10	176.11 (14)
C7—C1—C2—C3	178.69 (13)	C8—C9—C10—C15	167.66 (15)
C1—C2—C3—C4	-1.0 (2)	C8—C9—C10—C11	-14.8 (2)
C2—C3—C4—C5	-0.1 (2)	C15—C10—C11—C12	1.4 (2)
C2—C3—C4—C11	179.45 (11)	C9—C10—C11—C12	-176.08 (14)
C3—C4—C5—C6	1.4 (2)	C10—C11—C12—C13	-1.4 (2)
C11—C4—C5—C6	-178.06 (11)	C16—O2—C13—C12	-7.9 (2)
C4—C5—C6—C1	-1.8 (2)	C16—O2—C13—C14	171.26 (13)
C2—C1—C6—C5	0.7 (2)	C11—C12—C13—O2	178.95 (14)
C7—C1—C6—C5	-177.17 (14)	C11—C12—C13—C14	-0.2 (2)
C6—C1—C7—O1	155.92 (15)	O2—C13—C14—C15	-177.58 (13)
C2—C1—C7—O1	-22.0 (2)	C12—C13—C14—C15	1.6 (2)
C6—C1—C7—C8	-24.0 (2)	C13—C14—C15—C10	-1.6 (2)
C2—C1—C7—C8	158.05 (14)	C11—C10—C15—C14	0.0 (2)
O1—C7—C8—C9	-17.5 (2)	C9—C10—C15—C14	177.68 (13)
C1—C7—C8—C9	162.41 (14)		

#### Hydrogen-bond geometry ( $\text{\AA}$ , $^\circ$ )

Cg1 and Cg2 are the centroids of the C1—C6 and C10—C15 rings, respectively.

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
C9—H9 $\cdots$ O1	0.95	2.47	2.8080 (19)	101
C2—H2 $\cdots$ Cg1 <sup>i</sup>	0.95	2.85	3.4675 (15)	124
C12—H12 $\cdots$ Cg2 <sup>ii</sup>	0.95	2.92	3.6616 (17)	136

Symmetry codes: (i)  $x, -y+3/2, z+1/2$ ; (ii)  $x, -y+3/2, z-1/2$ .