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Bis(1,10-phenanthroline-κ^2N,N')(sulfato-κ^2O,O')zinc(II) propane-1,3-diol solvate

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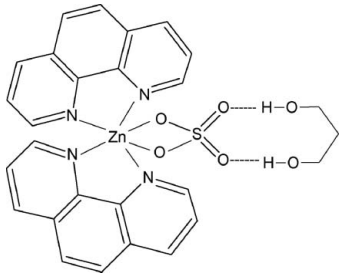
Key indicators: single-crystal X-ray study; T = 223 K; mean σ(C–C) = 0.005 Å; R factor = 0.044; wR factor = 0.104; data-to-parameter ratio = 12.5.

In the title compound, $[Zn(SO_4)(C_{12}H_8N_2)_2]\cdot C_3H_8O_2$, the Zn^{2+} ion (site symmetry 2) is coordinated by two chelating 1,10-phenanthroline ligands and an O,O'-bidentate sulfate ion (S site symmetry 2), resulting in a distorted *cis*-ZnO_2N_4 octahedral geometry for the metal ion. The complete propane-1,3-diol molecule is generated by crystallographic twofold symmetry and two O—H···O hydrogen bonds are formed with the uncoordinated O atoms of the sulfate group.

Related literature

For related structures and background references, see: Zhong (2010*a*,*b*).



Experimental

Crystal data

$[Zn(SO_4)(C_{12}H_8N_2)_2]\cdot C_3H_8O_2$
M_r = 597.96
Monoclinic, $C2/c$
a = 18.330 (4) Å
b = 12.406 (3) Å
c = 13.215 (3) Å
β = 121.78 (3)°
V = 2554.6 (13) Å^3
Z = 4
Mo $K\alpha$ radiation
μ = 1.10 mm^{-1}
T = 223 K
0.25 × 0.20 × 0.12 mm

Data collection

Rigaku Mercury CCD diffractometer
Absorption correction: multi-scan (*REQAB*; Jacobson, 1998)
T_{min} = 0.790, T_{max} = 1.000
7464 measured reflections
2241 independent reflections
1932 reflections with $I > 2\sigma(I)$
R_{int} = 0.039

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.044
$wR(F^2)$ = 0.104
S = 1.08
2241 reflections
179 parameters
H-atom parameters constrained
$\Delta\rho_{max}$ = 0.63 e Å^{-3}
$\Delta\rho_{min}$ = −0.34 e Å^{-3}

Table 1
Selected geometric parameters (Å, °).

Zn1—N2	2.145 (3)	Zn1—O1	2.174 (2)
Zn1—N1	2.147 (3)		
N2—Zn1—N1	77.87 (10)	O1—Zn1—O1^{i}	65.58 (11)

Symmetry code: (i) $-x, y, -z+\frac{1}{2}$.

Table 2
Hydrogen-bond geometry (Å, °).

D—H···A	D—H	H···A	D···A	D—H···A
O3—H3B···O2	0.82	1.95	2.727 (4)	157

Data collection: *CrystalClear* (Rigaku, 2007); cell refinement: *CrystalClear*; data reduction: *CrystalClear*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *XP* in *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HB5408).

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Zhong, K.-L. (2010*a*). *Acta Cryst.* E**66**, m247.
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supporting information

Acta Cryst. (2010). E**66**, m564 [https://doi.org/10.1107/S1600536810014194]

Bis(1,10-phenanthroline-κ^2N,N')(sulfato-κ^2O,O')zinc(II) propane-1,3-diol solvate

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S1. Comment

The title compound, (I), was obtained unitentionally during an attempt to synthesize coordination polymers of Zn(II) with 1,10-phenanthroline as second ligand *via* a solvothermal reaction. It is isomorphous with the recently reported cobalt(II) structure (Zhong 2010a).

In this study, the structure of Zn^{II} complexe with bidentate-chelating sulfate ligand, *viz*. $[ZnSO_4(phen)_2].C_3H_8O_2$, has been characterized. each Zn^{II} metal ion is six-coordinated in a distorted octahedral manner by four N atoms from two chelating phen ligands and two O atoms from a bidentate-chelating sulfate ligand. The formula unit lies on a twofold rotation axis [symmetry code: -*x*, *y*, - *z* + 1/2] passes through the Zn^{II} and S atoms, and also through the central carbon of the propane-1,3-diol solvent molecule, in C/2c . Around the twofold axis two planar phen ligands are arranged in a propeller manner. Intermolecular O—H···O hydrogen bonds help to further stabilize the crystal structure(see Fig. 1). Selected coordination bond distances and angles in Table 1 and intermolecular hydrogen bond see Table 2.

We discuss the title complexe and compare it with the previously reported compound $[ZnSO_4(C_{10}H_8N_2)_2].C_2H_6O_2$, (II) ($C_{10}H_8N_2$ is 2,2′-bipyridine; Zhong, 2010b). In (I), the Zn^{II} metal ions has an octahedral coordinaiton environment is in good agreement with that observed in (II), The Zn—O bond distance [2.174 (2) Å] and the Zn—N bond distances [2.145 (3)-2.147 (3) Å] are close to those found in (II) [2.1811 (15)Å and 2.1287 (17)-2.1452 (17) Å; respectively]. The N—Zn—N angle [77.87 (10)°] and the O—Zn—O angle [65.58 (11)°] in (I) are also comparable with values reported in (II) [76.61 (7)° and 65.64 (8)° respectively], The dihedral angle (79.8°) between the two chelating NCCN groups is slightly less than that found in (II), 81.1 (1)°.

S2. Experimental

0.2 mmol phen, 0.1 mmol melamine, 0.1 mmol $ZnO_4.7H_2O$, 2.0 ml propane-1,3-diol and 1.0 ml water were mixed and placed in a thick Pyrex tube, which was sealed and heated to 413 K for 96 h. After cooling, colorless blocks of (I) were obtained.

S3. Refinement

The H atoms of phen were positioned geometrically and allowed to ride on their parent atoms, with C—H = 0.93 Å and U_{iso}(H) = 1.2U_{eq}(C). The H atoms of central carbon of propane-1,3-diol were located in difference Fourier syntheses and were freely refined [C—H = 0.97 Å] and U_{iso}(H) = 1.2U_{eq}(C), whereas other H atoms were placed in geometrically idealized positions and refined as riding atoms, with C—H = 0.97 Å and O—H = 0.82 Å; U_{iso}(H) = 1.2U_{eq}(C) and 1.5U_{eq}(O).

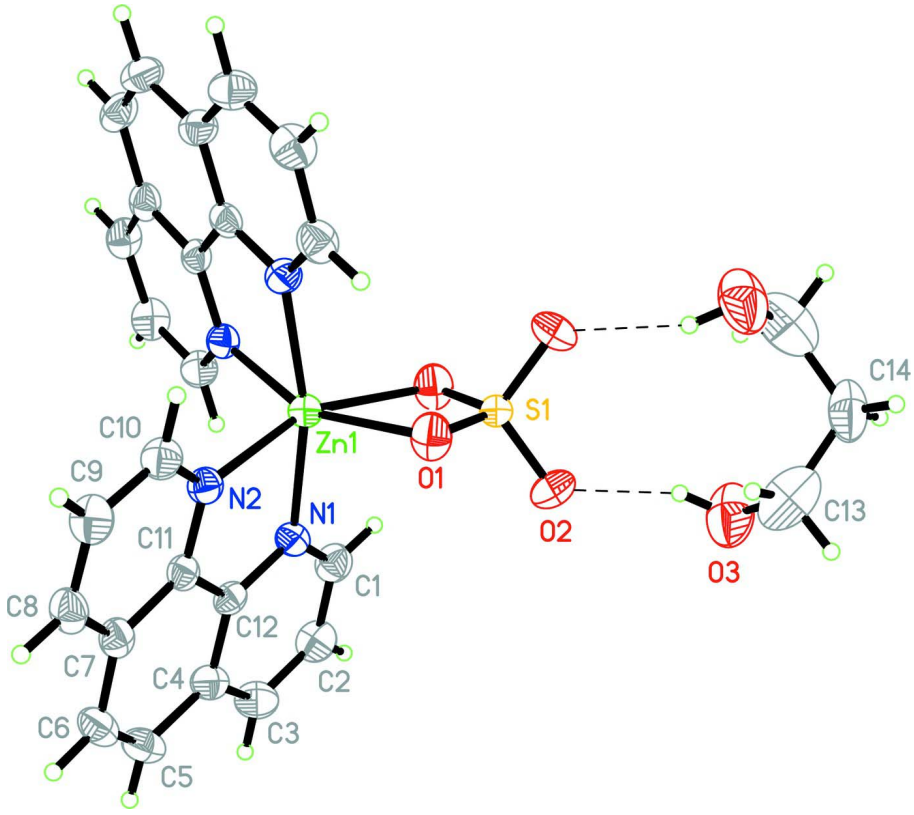

Figure 1

The molecular structure of (I) with displacement ellipsoids drawn at the 50% probability level. The dashed lines represent O—H···O interactions. Unlabeled atoms are related to the labelled atoms by the symmetry operator(-*x*, *y*, - *z* + 1/2).

Bis(1,10-phenanthroline-κ^2N,N')(sulfato- κ^2O,O')zinc(II) propane-1,3-diol solvate

Crystal data

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Monoclinic, *C*2/*c*
Hall symbol: -C 2yc
$a = 18.330\ (4)$ Å
$b = 12.406\ (3)$ Å
$c = 13.215\ (3)$ Å
$\beta = 121.78\ (3)°$
$V = 2554.6\ (13)$ Å^3
$Z = 4$
$F(000) = 1232$
$D_x = 1.555$ Mg m^{-3}
Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å
Cell parameters from 4259 reflections
$\theta = 3.2–27.5°$
$\mu = 1.10$ mm^{-1}
$T = 223$ K
Block, colourless
0.25 × 0.20 × 0.12 mm

Data collection

Rigaku Mercury CCD diffractometer
Radiation source: fine-focus sealed tube
Graphite Monochromator monochromator
Detector resolution: 28.5714 pixels mm^{-1}
ω scans
Absorption correction: multi-scan (REQAB; Jacobson, 1998)
$T_{min} = 0.790$, $T_{max} = 1.000$
7464 measured reflections
2241 independent reflections
1932 reflections with $I > 2\sigma(I)$
$R_{int} = 0.039$
$\theta_{max} = 25.0°$, $\theta_{min} = 3.2°$
$h = -21\rightarrow15$
$k = -14\rightarrow13$
$l = -14\rightarrow15$

Refinement

Refinement on F^2
Least-squares matrix: full
$R[F^2 > 2\sigma(F^2)] = 0.044$
$wR(F^2) = 0.104$
$S = 1.08$
2241 reflections
179 parameters
0 restraints
Primary atom site location: structure-invariant direct methods
Secondary atom site location: difference Fourier map
Hydrogen site location: inferred from neighbouring sites
H-atom parameters constrained
$w = 1/[\sigma^2(F_o^2) + (0.0559P)^2 + 1.7701P]$ where $P = (F_o^2 + 2F_c^2)/3$
$(\Delta/\sigma)_{max} < 0.001$
$\Delta\rho_{max} = 0.63$ e Å^{-3}
$\Delta\rho_{min} = -0.34$ e Å^{-3}
Extinction correction: *SHELXL97* (Sheldrick, 2008), $Fc^*=kFc[1+0.001xFc^2\lambda^3/\sin(2\theta)]^{-1/4}$
Extinction coefficient: 0.0021 (4)

Special details

Geometry. All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted *R*-factor wR and goodness of fit *S* are based on F^2, conventional *R*-factors *R* are based on *F*, with *F* set to zero for negative F^2. The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating *R*-factors(gt) etc. and is not relevant to the choice of reflections for refinement. *R*-factors based on F^2 are statistically about twice as large as those based on *F*, and *R*- factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	*y*	*z*	U_{iso}*/U_{eq}
Zn1	0.0000	0.31613 (4)	0.2500	0.0275 (2)
S1	0.0000	0.53713 (8)	0.2500	0.0261 (3)
O2	0.05644 (17)	0.6038 (2)	0.2306 (2)	0.0501 (7)
O1	0.05155 (15)	0.46343 (17)	0.35278 (18)	0.0386 (6)
N1	0.08295 (17)	0.29637 (19)	0.1822 (2)	0.0279 (6)
N2	0.09735 (17)	0.21096 (19)	0.3797 (2)	0.0271 (6)
C9	0.1742 (2)	0.1058 (3)	0.5575 (3)	0.0365 (8)
H9A	0.1784	0.0812	0.6268	0.044*
C5	0.2775 (2)	0.1280 (3)	0.2899 (3)	0.0354 (8)
H5A	0.3165	0.1084	0.2684	0.042*
C1	0.0755 (2)	0.3403 (3)	0.0849 (3)	0.0347 (8)
H1A	0.0310	0.3887	0.0408	0.042*
C7	0.2267 (2)	0.1149 (2)	0.4258 (2)	0.0287 (7)
C2	0.1314 (2)	0.3164 (3)	0.0470 (3)	0.0370 (8)
H2A	0.1238	0.3479	−0.0218	0.044*
C4	0.2084 (2)	0.1990 (2)	0.2154 (3)	0.0302 (7)
C3	0.1975 (2)	0.2465 (3)	0.1113 (3)	0.0374 (8)
H3A	0.2353	0.2301	0.0867	0.045*
C6	0.2869 (2)	0.0887 (3)	0.3919 (3)	0.0348 (8)
H6A	0.3331	0.0440	0.4405	0.042*
C8	0.2340 (2)	0.0767 (2)	0.5317 (3)	0.0336 (8)
H8A	0.2794	0.0321	0.5831	0.040*
C12	0.14832 (19)	0.2268 (2)	0.2462 (2)	0.0249 (6)

C10	0.1065 (2)	0.1729 (3)	0.4796 (3)	0.0325 (7)
H10A	0.0660	0.1918	0.4985	0.039*
C11	0.15675 (19)	0.1826 (2)	0.3525 (2)	0.0246 (6)
O3	0.0629 (3)	0.8156 (2)	0.1807 (3)	0.0736 (10)
H3B	0.0462	0.7553	0.1845	0.110*
C14	0.0000	0.9448 (4)	0.2500	0.0615 (17)
C13	0.0766 (4)	0.8753 (4)	0.2756 (5)	0.0860 (17)
H13A	0.1258	0.9217	0.3005	0.103*
H13B	0.0905	0.8272	0.3412	0.103*
H14B	−0.0159	0.9906	0.1817	0.103*

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Zn1	0.0250 (3)	0.0253 (3)	0.0328 (3)	0.000	0.0157 (2)	0.000
S1	0.0229 (6)	0.0234 (5)	0.0330 (6)	0.000	0.0153 (5)	0.000
O2	0.0474 (16)	0.0435 (14)	0.0763 (18)	−0.0092 (13)	0.0442 (14)	0.0034 (13)
O1	0.0360 (14)	0.0343 (12)	0.0306 (12)	0.0007 (11)	0.0073 (10)	0.0020 (9)
N1	0.0247 (14)	0.0276 (13)	0.0297 (13)	0.0026 (11)	0.0133 (11)	0.0027 (10)
N2	0.0284 (15)	0.0251 (13)	0.0289 (13)	−0.0018 (11)	0.0160 (11)	−0.0031 (10)
C9	0.045 (2)	0.0368 (18)	0.0274 (16)	0.0002 (16)	0.0188 (15)	0.0027 (13)
C5	0.0276 (18)	0.0409 (19)	0.0379 (17)	0.0037 (15)	0.0175 (14)	−0.0052 (14)
C1	0.0340 (19)	0.0337 (18)	0.0326 (17)	−0.0015 (15)	0.0150 (14)	0.0038 (13)
C7	0.0282 (17)	0.0241 (15)	0.0268 (15)	−0.0022 (14)	0.0096 (13)	−0.0039 (12)
C2	0.044 (2)	0.0414 (18)	0.0292 (16)	−0.0030 (17)	0.0218 (15)	0.0013 (14)
C4	0.0275 (17)	0.0332 (17)	0.0303 (16)	−0.0019 (14)	0.0155 (14)	−0.0057 (12)
C3	0.037 (2)	0.047 (2)	0.0347 (17)	−0.0002 (17)	0.0231 (15)	−0.0012 (15)
C6	0.0260 (17)	0.0337 (17)	0.0347 (17)	0.0062 (15)	0.0091 (14)	−0.0053 (13)
C8	0.0318 (18)	0.0299 (17)	0.0280 (16)	0.0039 (15)	0.0080 (14)	0.0048 (12)
C12	0.0239 (16)	0.0236 (14)	0.0245 (14)	−0.0060 (13)	0.0110 (12)	−0.0051 (12)
C10	0.0377 (19)	0.0332 (17)	0.0318 (16)	−0.0006 (15)	0.0219 (15)	−0.0003 (13)
C11	0.0258 (16)	0.0211 (14)	0.0254 (14)	−0.0021 (13)	0.0124 (12)	−0.0042 (11)
O3	0.122 (3)	0.0507 (17)	0.089 (2)	−0.0088 (18)	0.084 (2)	−0.0004 (16)
C14	0.092 (5)	0.032 (3)	0.070 (4)	0.000	0.049 (4)	0.000
C13	0.084 (4)	0.086 (4)	0.084 (3)	−0.031 (3)	0.041 (3)	−0.005 (3)

Geometric parameters (Å, º)

Zn1—N2^{i}	2.145 (3)	C1—H1A	0.9300
Zn1—N2	2.145 (3)	C7—C11	1.407 (4)
Zn1—N1	2.147 (3)	C7—C8	1.415 (4)
Zn1—N1^{i}	2.147 (3)	C7—C6	1.429 (5)
Zn1—O1	2.174 (2)	C2—C3	1.363 (5)
Zn1—O1^{i}	2.174 (2)	C2—H2A	0.9300
S1—O2	1.449 (2)	C4—C12	1.403 (5)
S1—O2^{i}	1.449 (2)	C4—C3	1.410 (4)
S1—O1^{i}	1.491 (2)	C3—H3A	0.9300
S1—O1	1.491 (2)	C6—H6A	0.9300

N1—C1	1.335 (4)	C8—H8A	0.9300
N1—C12	1.352 (4)	C12—C11	1.439 (4)
N2—C10	1.327 (4)	C10—H10A	0.9300
N2—C11	1.360 (4)	O3—C13	1.361 (6)
C9—C8	1.357 (5)	O3—H3B	0.8200
C9—C10	1.395 (5)	C14—C13^{i}	1.527 (7)
C9—H9A	0.9300	C14—C13	1.527 (7)
C5—C6	1.356 (5)	C14—H14B	0.9728
C5—C4	1.427 (5)	C13—H13A	0.9700
C5—H5A	0.9300	C13—H13B	0.9700
C1—C2	1.391 (5)		

N2^{i}—Zn1—N2	105.08 (13)	N1—C1—H1A	118.7
N2^{i}—Zn1—N1	94.08 (10)	C2—C1—H1A	118.7
N2—Zn1—N1	77.87 (10)	C11—C7—C8	117.3 (3)
N2^{i}—Zn1—N1^{i}	77.87 (10)	C11—C7—C6	119.6 (3)
N2—Zn1—N1^{i}	94.08 (10)	C8—C7—C6	123.1 (3)
N1—Zn1—N1^{i}	166.89 (13)	C3—C2—C1	119.5 (3)
N2^{i}—Zn1—O1	156.26 (9)	C3—C2—H2A	120.3
N2—Zn1—O1	96.15 (9)	C1—C2—H2A	120.3
N1—Zn1—O1	100.74 (9)	C12—C4—C3	116.7 (3)
N1^{i}—Zn1—O1	90.32 (9)	C12—C4—C5	119.9 (3)
N2^{i}—Zn1—O1^{i}	96.15 (9)	C3—C4—C5	123.4 (3)
N2—Zn1—O1^{i}	156.26 (9)	C2—C3—C4	119.9 (3)
N1—Zn1—O1^{i}	90.32 (9)	C2—C3—H3A	120.1
N1^{i}—Zn1—O1^{i}	100.74 (9)	C4—C3—H3A	120.1
O1—Zn1—O1^{i}	65.58 (11)	C5—C6—C7	121.0 (3)
O2—S1—O2^{i}	110.4 (2)	C5—C6—H6A	119.5
O2—S1—O1^{i}	110.96 (14)	C7—C6—H6A	119.5
O2^{i}—S1—O1^{i}	110.01 (15)	C9—C8—C7	119.4 (3)
O2—S1—O1	110.01 (14)	C9—C8—H8A	120.3
O2^{i}—S1—O1	110.96 (14)	C7—C8—H8A	120.3
O1^{i}—S1—O1	104.33 (18)	N1—C12—C4	123.2 (3)
O2—S1—Zn1	124.79 (11)	N1—C12—C11	117.2 (3)
O2^{i}—S1—Zn1	124.79 (11)	C4—C12—C11	119.5 (3)
O1^{i}—S1—Zn1	52.17 (9)	N2—C10—C9	123.0 (3)
O1—S1—Zn1	52.17 (9)	N2—C10—H10A	118.5
S1—O1—Zn1	95.04 (11)	C9—C10—H10A	118.5
C1—N1—C12	118.0 (3)	N2—C11—C7	122.6 (3)
C1—N1—Zn1	128.2 (2)	N2—C11—C12	118.1 (3)
C12—N1—Zn1	113.7 (2)	C7—C11—C12	119.3 (3)
C10—N2—C11	118.1 (3)	C13—O3—H3B	109.5
C10—N2—Zn1	128.9 (2)	C13^{i}—C14—C13	111.2 (5)
C11—N2—Zn1	113.02 (19)	C13^{i}—C14—H14B	109.5
C8—C9—C10	119.7 (3)	C13—C14—H14B	109.1
C8—C9—H9A	120.2	O3—C13—C14	113.7 (4)
C10—C9—H9A	120.2	O3—C13—H13A	108.8
C6—C5—C4	120.7 (3)	C14—C13—H13A	108.8

C6—C5—H5A	119.7	O3—C13—H13B	108.8
C4—C5—H5A	119.7	C14—C13—H13B	108.8
N1—C1—C2	122.6 (3)	H13A—C13—H13B	107.7

$N2^{i}$—Zn1—S1—O2	−110.63 (16)	$O1^{i}$—Zn1—N2—C10	114.9 (3)
N2—Zn1—S1—O2	69.37 (16)	S1—Zn1—N2—C10	87.6 (3)
N1—Zn1—S1—O2	−10.36 (14)	$N2^{i}$—Zn1—N2—C11	88.39 (19)
$N1^{i}$—Zn1—S1—O2	169.64 (14)	N1—Zn1—N2—C11	−2.61 (19)
O1—Zn1—S1—O2	89.31 (18)	$N1^{i}$—Zn1—N2—C11	166.92 (19)
$O1^{i}$—Zn1—S1—O2	−90.69 (18)	O1—Zn1—N2—C11	−102.3 (2)
$N2^{i}$—Zn1—S1—$O2^{i}$	69.37 (16)	$O1^{i}$—Zn1—N2—C11	−64.3 (3)
N2—Zn1—S1—$O2^{i}$	−110.63 (16)	S1—Zn1—N2—C11	−91.61 (19)
N1—Zn1—S1—$O2^{i}$	169.64 (14)	C12—N1—C1—C2	0.6 (5)
$N1^{i}$—Zn1—S1—$O2^{i}$	−10.36 (14)	Zn1—N1—C1—C2	−176.5 (2)
O1—Zn1—S1—$O2^{i}$	−90.69 (18)	N1—C1—C2—C3	−0.8 (5)
$O1^{i}$—Zn1—S1—$O2^{i}$	89.31 (18)	C6—C5—C4—C12	−1.4 (5)
$N2^{i}$—Zn1—S1—$O1^{i}$	−19.95 (15)	C6—C5—C4—C3	176.9 (3)
N2—Zn1—S1—$O1^{i}$	160.05 (15)	C1—C2—C3—C4	0.1 (5)
N1—Zn1—S1—$O1^{i}$	80.33 (14)	C12—C4—C3—C2	0.6 (5)
$N1^{i}$—Zn1—S1—$O1^{i}$	−99.67 (14)	C5—C4—C3—C2	−177.8 (3)
O1—Zn1—S1—$O1^{i}$	180.0	C4—C5—C6—C7	1.7 (5)
$N2^{i}$—Zn1—S1—O1	160.05 (15)	C11—C7—C6—C5	−0.2 (5)
N2—Zn1—S1—O1	−19.95 (15)	C8—C7—C6—C5	−179.5 (3)
N1—Zn1—S1—O1	−99.67 (14)	C10—C9—C8—C7	−0.1 (5)
$N1^{i}$—Zn1—S1—O1	80.33 (14)	C11—C7—C8—C9	0.4 (4)
$O1^{i}$—Zn1—S1—O1	180.0	C6—C7—C8—C9	179.7 (3)
O2—S1—O1—Zn1	−119.08 (14)	C1—N1—C12—C4	0.2 (4)
$O2^{i}$—S1—O1—Zn1	118.43 (14)	Zn1—N1—C12—C4	177.7 (2)
$O1^{i}$—S1—O1—Zn1	0.0	C1—N1—C12—C11	178.3 (3)
$N2^{i}$—Zn1—O1—S1	−42.3 (3)	Zn1—N1—C12—C11	−4.2 (3)
N2—Zn1—O1—S1	164.20 (12)	C3—C4—C12—N1	−0.8 (4)
N1—Zn1—O1—S1	85.42 (13)	C5—C4—C12—N1	177.7 (3)
$N1^{i}$—Zn1—O1—S1	−101.66 (13)	C3—C4—C12—C11	−178.9 (3)
$O1^{i}$—Zn1—O1—S1	0.0	C5—C4—C12—C11	−0.4 (4)
$N2^{i}$—Zn1—N1—C1	76.3 (3)	C11—N2—C10—C9	0.2 (5)
N2—Zn1—N1—C1	−179.1 (3)	Zn1—N2—C10—C9	−179.0 (2)
$N1^{i}$—Zn1—N1—C1	127.8 (3)	C8—C9—C10—N2	−0.2 (5)
O1—Zn1—N1—C1	−85.1 (3)	C10—N2—C11—C7	0.2 (4)
$O1^{i}$—Zn1—N1—C1	−19.9 (3)	Zn1—N2—C11—C7	179.5 (2)
S1—Zn1—N1—C1	−52.2 (3)	C10—N2—C11—C12	−178.0 (3)
$N2^{i}$—Zn1—N1—C12	−100.9 (2)	Zn1—N2—C11—C12	1.3 (3)
N2—Zn1—N1—C12	3.66 (19)	C8—C7—C11—N2	−0.5 (4)
$N1^{i}$—Zn1—N1—C12	−49.37 (19)	C6—C7—C11—N2	−179.8 (3)
O1—Zn1—N1—C12	97.7 (2)	C8—C7—C11—C12	177.7 (3)
$O1^{i}$—Zn1—N1—C12	162.9 (2)	C6—C7—C11—C12	−1.7 (4)
S1—Zn1—N1—C12	130.63 (19)	N1—C12—C11—N2	2.0 (4)
$N2^{i}$—Zn1—N2—C10	−92.4 (3)	C4—C12—C11—N2	−179.8 (3)
N1—Zn1—N2—C10	176.6 (3)	N1—C12—C11—C7	−176.3 (2)

N1^{i}—Zn1—N2—C10	−13.9 (3)	C4—C12—C11—C7	1.9 (4)
O1—Zn1—N2—C10	76.9 (3)	C13^{i}—C14—C13—O3	65.1 (3)

Symmetry code: (i) −*x*, *y*, −*z*+1/2.

Hydrogen-bond geometry (Å, º)

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
O3—H3*B*···O2	0.82	1.95	2.727 (4)	157