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2,4-Dichloro-*N*-(2-chlorophenyl)-benzenesulfonamide

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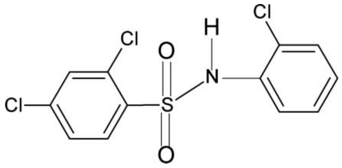
Key indicators: single-crystal X-ray study; *T* = 299 K; mean σ(C–C) = 0.004 Å; *R* factor = 0.042; *wR* factor = 0.107; data-to-parameter ratio = 16.8.

In the title compound, $C_{12}H_8Cl_3NO_2S$, the conformation of the N—H bond in the C—SO$_2$—NH—C segment is *syn* to the *ortho*-Cl in the aniline ring. The dihedral angle between the two benzene rings is 74.3 (1)°. An intramolecular N—H⋯Cl hydrogen bond occurs. In the crystal, pairs of N—H⋯O hydrogen bonds link the molecules into dimers.

Related literature

For the preparation of the title compound, see: Savitha & Gowda (2006). For our studies of the effect of substituents on the structures of *N*-(aryl)arylsulfonamides, see: Gowda *et al.* (2010*a*,*b*); Nirmala *et al.* (2010). For related structures, see: Gelbrich *et al.* (2007); Perlovich *et al.* (2006).



Experimental

Crystal data

$C_{12}H_8Cl_3NO_2S$
M_r = 336.60
Orthorhombic, *Pbcn*
a = 10.5639 (8) Å
b = 16.279 (1) Å
c = 16.693 (1) Å
V = 2870.7 (3) Å^3
Z = 8
Mo *K*α radiation
μ = 0.78 mm^{-1}
T = 299 K
0.30 × 0.24 × 0.22 mm

Data collection

Oxford Diffraction Xcalibur diffractometer with a Sapphire CCD detector
Absorption correction: multi-scan (*CrysAlis RED*; Oxford Diffraction, 2009)
T_{min} = 0.800, T_{max} = 0.847
8054 measured reflections
2938 independent reflections
2054 reflections with $I > 2\sigma(I)$
R_{int} = 0.019

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.042
$wR(F^2)$ = 0.107
S = 1.01
2938 reflections
175 parameters
1 restraint
H atoms treated by a mixture of independent and constrained refinement
$\Delta\rho_{max}$ = 0.36 e Å^{-3}
$\Delta\rho_{min}$ = −0.35 e Å^{-3}

Table 1
Hydrogen-bond geometry (Å, °).

D—H⋯*A*	*D*—H	H⋯*A*	*D*⋯*A*	*D*—H⋯*A*
N1—H1*N*⋯O1^{i}	0.86 (1)	2.23 (2)	3.007 (3)	152 (2)
N1—H1*N*⋯Cl3	0.86 (1)	2.57 (3)	2.975 (2)	110 (2)

Symmetry code: (i) $-x, y, -z + \frac{1}{2}$.

Data collection: *CrysAlis CCD* (Oxford Diffraction, 2009); cell refinement: *CrysAlis RED* (Oxford Diffraction, 2009); data reduction: *CrysAlis RED*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *PLATON* (Spek, 2009); software used to prepare material for publication: *SHELXL97*.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: BQ2220).

References

Gelbrich, T., Hursthouse, M. B. & Threlfall, T. L. (2007). *Acta Cryst.* B**63**, 621–632.
Gowda, B. T., Foro, S., Nirmala, P. G. & Fuess, H. (2010*a*). *Acta Cryst.* E**66**, o1552.
Gowda, B. T., Foro, S., Nirmala, P. G. & Fuess, H. (2010*b*). Private communication (refcode CCDC 740692). CCDC, Union Road, Cambridge, England.
Nirmala, P. G., Gowda, B. T., Foro, S. & Fuess, H. (2010). *Acta Cryst.* E**66**, o918.
Oxford Diffraction (2009). *CrysAlis CCD* and *CrysAlis RED*. Oxford Diffraction Ltd, Yarnton, England.
Perlovich, G. L., Tkachev, V. V., Schaper, K.-J. & Raevsky, O. A. (2006). *Acta Cryst.* E**62**, o780–o782.
Savitha, M. B. & Gowda, B. T. (2006). *Z. Naturforsch. Teil A*, **60**, 600–606.
Sheldrick, G. M. (2008). *Acta Cryst.* A**64**, 112–122.
Spek, A. L. (2009). *Acta Cryst.* D**65**, 148–155.

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supporting information

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2,4-Dichloro-*N*-(2-chlorophenyl)benzenesulfonamide

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S1. Comment

As a part of studying the substituent effects on the structures of *N*-(aryl)arylsulfonamides (Gowda *et al.* , 2010*a,b*; Nirmala *et al.*, 2010), the structure of 2,4-dichloro-*N*-(2-chlorophenyl)-benzenesulfonamide (I) has been determined (Fig. 1). The conformation of the N—H bond in the C—SO_2—NH—C segment is *syn* to the *ortho*-Cl in the aniline ring (Fig. 1), contrary to the *anti* conformation observed between the N—H bond and *meta*-Cl in the aniline ring of 2,4-dichloro-*N*-(3-chlorophenyl)-benzenesulfonamide (II) (Gowda *et al.*, 2010*a*).

The molecule is twisted at the S atom with the C1—SO_2—NH—C7 torsion angle of 54.5 (3)°, compared to the values of 62.3 (2)° in (II), 67.8 (2)° in 2,4-dichloro-*N*-(4-chlorophenyl)benzenesulfonamide (III)(Nirmala *et al.*, 2010), 55.1 (3)° (molecule 1) and -48.3 (3)° (molecule 2) in 2,4-dichloro-*N*-(phenyl)-benzenesulfonamide (IV) (Gowda *et al.*, 2010*b*). The sulfonyl benzene and the aniline benzene rings in (I) are tilted relative to each other by 74.3 (1)°, compared to the values of 69.3 (1)° in (II), 65.0 (1)° in (III), 80.5 (2)° in the molecule 1 and 64.9 (1)° in molecule 2 of (IV).

The other bond parameters in (I) are similar to those observed in (II), (III), (IV) and other aryl sulfonamides (Perlovich *et al.*, 2006; Gelbrich *et al.*, 2007).

In the crystal structure, the pairs of intermolecular N–H···O hydrogen bonds (Table 1) link the molecules *via* dimers into infinite sequences running parallel to the *c*-axis. Part of the crystal structure is shown in Fig. 2.

S2. Experimental

The solution of 1,3-dichlorobenzene (10 cc) in chloroform (40 cc) was treated dropwise with chlorosulfonic acid (25 cc) at 0 ° C. After the initial evolution of hydrogen chloride subsided, the reaction mixture was brought to room temperature and poured into crushed ice in a beaker. The chloroform layer was separated, washed with cold water and allowed to evaporate slowly. The residual 2,4-dichlorobenzenesulfonylchloride was treated with 2-chloroaniline in the stoichiometric ratio and boiled for ten minutes. The reaction mixture was then cooled to room temperature and added to ice cold water (100 cc). The resultant solid 2,4-dichloro-*N*-(2-chlorophenyl)benzenesulfonamide was filtered under suction and washed thoroughly with cold water. It was then recrystallized to constant melting point from dilute ethanol. The purity of the compound was checked and characterized by recording its infrared and NMR spectra (Savitha & Gowda, 2006). Prism like colorless single crystals used in X-ray diffraction studies were grown in ethanolic solution by slow evaporation at room temperature.

S3. Refinement

The H atom of the NH group was located in a difference map and later restrained to N—H = 0.86 (1)Å. The other H atoms were positioned with idealized geometry using a riding model with C—H = 0.93 Å. All H atoms were refined with isotropic displacement parameters (set to 1.2 times of the U_{eq} of the parent atom).

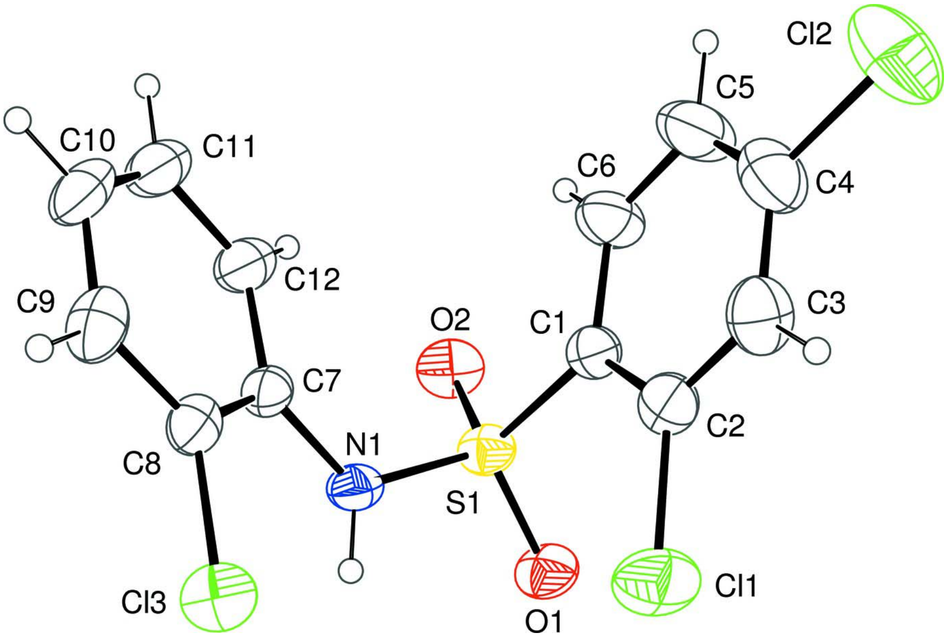

Figure 1

Molecular structure of (I), showing the atom labelling scheme and displacement ellipsoids are drawn at the 50% probability level.

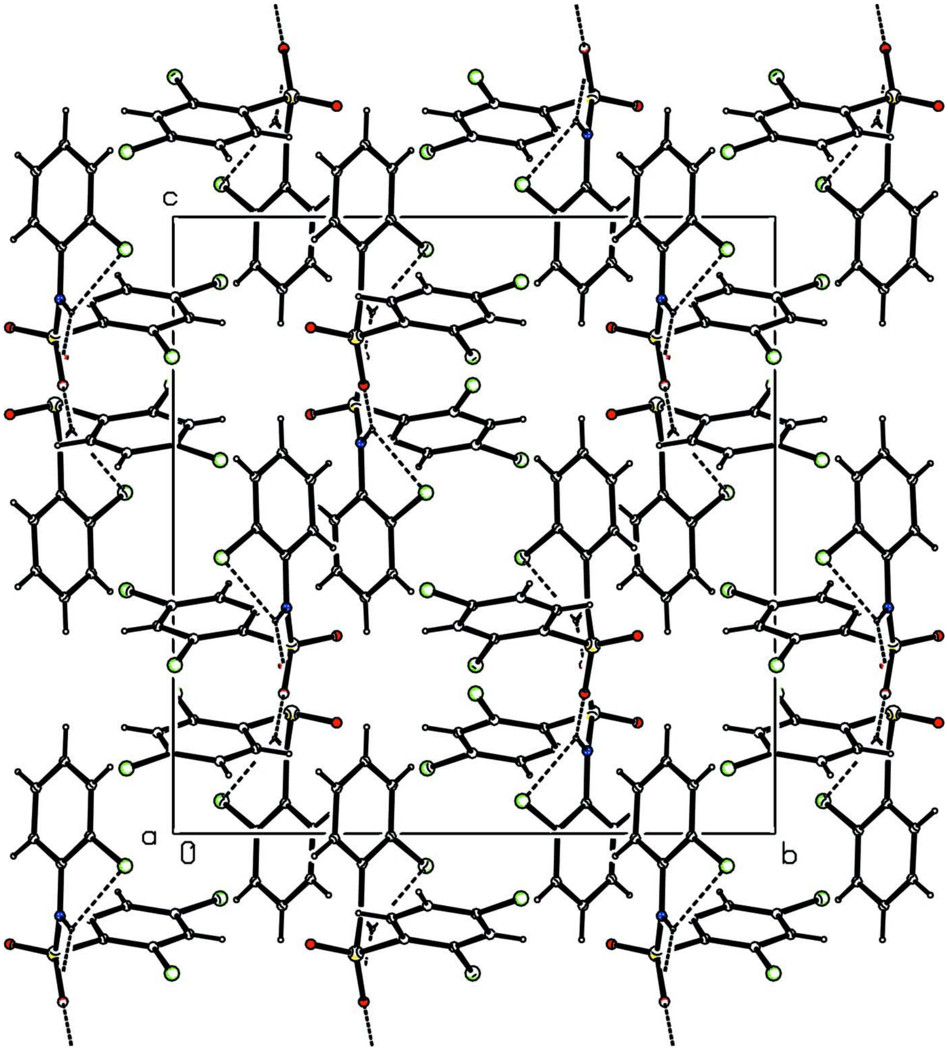

Figure 2

Molecular packing of (I) with hydrogen bonding shown as dashed lines.

2,4-Dichloro-*N*-(2-chlorophenyl)benzenesulfonamide

Crystal data

$C_{12}H_8Cl_3NO_2S$
M_r = 336.60
Orthorhombic, *Pbcn*
Hall symbol: -P 2n 2ab
a = 10.5639 (8) Å
b = 16.279 (1) Å
c = 16.693 (1) Å
V = 2870.7 (3) Å^3
Z = 8

F(000) = 1360
D_x = 1.558 Mg m^{-3}
Mo *Kα* radiation, *λ* = 0.71073 Å
Cell parameters from 2957 reflections
θ = 2.5–27.8°
μ = 0.78 mm^{-1}
T = 299 K
Prism, colourless
0.30 × 0.24 × 0.22 mm

Data collection

Oxford Diffraction Xcalibur diffractometer with a Sapphire CCD detector
Radiation source: fine-focus sealed tube
Graphite monochromator
Rotation method data acquisition using ω and phi scans
Absorption correction: multi-scan (*CrysAlis RED*; Oxford Diffraction, 2009)
$T_{\min} = 0.800$, $T_{\max} = 0.847$
8054 measured reflections
2938 independent reflections
2054 reflections with $I > 2\sigma(I)$
$R_{\text{int}} = 0.019$
$\theta_{\max} = 26.4°$, $\theta_{\min} = 2.5°$
$h = -9 \rightarrow 13$
$k = -14 \rightarrow 20$
$l = -11 \rightarrow 20$

Refinement

Refinement on F^2
Least-squares matrix: full
$R[F^2 > 2\sigma(F^2)] = 0.042$
$wR(F^2) = 0.107$
$S = 1.01$
2938 reflections
175 parameters
1 restraint
Primary atom site location: structure-invariant direct methods
Secondary atom site location: difference Fourier map
Hydrogen site location: inferred from neighbouring sites
H atoms treated by a mixture of independent and constrained refinement
$w = 1/[\sigma^2(F_o^2) + (0.0458P)^2 + 1.5152P]$ where $P = (F_o^2 + 2F_c^2)/3$
$(\Delta/\sigma)_{\max} < 0.001$
$\Delta\rho_{\max} = 0.36$ e Å^{-3}
$\Delta\rho_{\min} = -0.35$ e Å^{-3}

Special details

Experimental. CrysAlis RED (Oxford Diffraction, 2009) Empirical absorption correction using spherical harmonics, implemented in SCALE3 ABSPACK scaling algorithm.

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R-factor wR and goodness of fit S are based on F^2, conventional R-factors R are based on F, with F set to zero for negative F^2. The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R-factors(gt) *etc*. and is not relevant to the choice of reflections for refinement. R-factors based on F^2 are statistically about twice as large as those based on F, and R- factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	*y*	*z*	U_{iso}*/U_{eq}
C1	0.3134 (2)	0.61639 (15)	0.32900 (15)	0.0497 (6)
C2	0.2866 (3)	0.53340 (17)	0.31598 (16)	0.0579 (7)
C3	0.3728 (3)	0.47403 (19)	0.3366 (2)	0.0753 (9)
H3	0.3545	0.4188	0.3282	0.090*
C4	0.4862 (3)	0.4967 (2)	0.3697 (2)	0.0850 (10)
C5	0.5150 (3)	0.5775 (2)	0.3827 (2)	0.0919 (11)
H5	0.5923	0.5920	0.4052	0.110*
C6	0.4285 (3)	0.6371 (2)	0.36218 (19)	0.0713 (8)
H6	0.4480	0.6921	0.3708	0.086*
C7	0.0989 (2)	0.68618 (15)	0.45102 (14)	0.0477 (6)
C8	0.0161 (2)	0.63852 (15)	0.49537 (15)	0.0509 (6)
C9	0.0238 (3)	0.63637 (17)	0.57821 (16)	0.0664 (8)
H9	−0.0334	0.6050	0.6076	0.080*

C10	0.1164 (3)	0.68083 (19)	0.61698 (17)	0.0730 (9)
H10	0.1232	0.6786	0.6725	0.088*
C11	0.1978 (3)	0.7281 (2)	0.57371 (17)	0.0703 (8)
H11	0.2600	0.7581	0.6000	0.084*
C12	0.1892 (3)	0.73195 (18)	0.49182 (16)	0.0629 (7)
H12	0.2444	0.7655	0.4633	0.075*
N1	0.08883 (19)	0.68827 (15)	0.36630 (12)	0.0542 (6)
H1N	0.0212 (16)	0.6693 (15)	0.3452 (15)	0.065*
O1	0.15449 (18)	0.68354 (12)	0.22707 (9)	0.0648 (5)
O2	0.27286 (18)	0.77149 (11)	0.32041 (12)	0.0675 (5)
Cl1	0.14598 (8)	0.50245 (5)	0.27304 (6)	0.0914 (3)
Cl2	0.59587 (12)	0.42218 (8)	0.39456 (10)	0.1453 (5)
Cl3	−0.09873 (7)	0.58020 (5)	0.44734 (5)	0.0691 (2)
S1	0.20683 (6)	0.69660 (4)	0.30490 (4)	0.05119 (19)

Atomic displacement parameters (Å^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
C1	0.0463 (14)	0.0562 (15)	0.0466 (14)	−0.0060 (12)	0.0030 (11)	0.0018 (12)
C2	0.0561 (16)	0.0593 (16)	0.0584 (16)	−0.0085 (14)	0.0062 (13)	−0.0016 (13)
C3	0.082 (2)	0.0563 (17)	0.087 (2)	0.0018 (17)	0.0103 (18)	0.0005 (17)
C4	0.075 (2)	0.083 (2)	0.097 (3)	0.024 (2)	−0.0010 (19)	0.006 (2)
C5	0.0572 (19)	0.099 (3)	0.119 (3)	0.0074 (19)	−0.0242 (18)	−0.003 (2)
C6	0.0537 (17)	0.0678 (19)	0.092 (2)	−0.0065 (15)	−0.0125 (16)	−0.0029 (17)
C7	0.0517 (14)	0.0510 (14)	0.0403 (13)	0.0108 (12)	−0.0058 (11)	−0.0019 (11)
C8	0.0630 (16)	0.0433 (13)	0.0463 (14)	0.0131 (12)	−0.0027 (12)	0.0000 (12)
C9	0.093 (2)	0.0583 (17)	0.0482 (16)	0.0183 (17)	0.0063 (15)	0.0079 (14)
C10	0.110 (3)	0.069 (2)	0.0401 (15)	0.0277 (19)	−0.0159 (16)	−0.0087 (15)
C11	0.086 (2)	0.0712 (19)	0.0535 (17)	0.0103 (17)	−0.0189 (15)	−0.0123 (15)
C12	0.0659 (17)	0.0709 (19)	0.0518 (16)	−0.0033 (15)	−0.0088 (13)	−0.0053 (14)
N1	0.0456 (12)	0.0765 (16)	0.0404 (11)	−0.0022 (11)	−0.0071 (9)	−0.0013 (11)
O1	0.0674 (11)	0.0888 (14)	0.0382 (9)	−0.0078 (11)	−0.0027 (8)	0.0105 (9)
O2	0.0745 (13)	0.0518 (10)	0.0763 (13)	−0.0140 (10)	−0.0064 (10)	0.0104 (10)
Cl1	0.0830 (6)	0.0772 (5)	0.1140 (7)	−0.0306 (5)	−0.0171 (5)	−0.0062 (5)
Cl2	0.1218 (9)	0.1269 (10)	0.1873 (13)	0.0642 (8)	−0.0140 (9)	0.0199 (9)
Cl3	0.0743 (5)	0.0665 (5)	0.0665 (5)	−0.0100 (4)	−0.0044 (4)	0.0089 (4)
S1	0.0528 (4)	0.0574 (4)	0.0434 (3)	−0.0060 (3)	−0.0025 (3)	0.0086 (3)

Geometric parameters (Å, °)

C1—C6	1.379 (3)	C7—N1	1.419 (3)
C1—C2	1.397 (3)	C8—C9	1.386 (4)
C1—S1	1.770 (3)	C8—Cl3	1.737 (3)
C2—C3	1.372 (4)	C9—C10	1.378 (4)
C2—Cl1	1.725 (3)	C9—H9	0.9300
C3—C4	1.369 (5)	C10—C11	1.361 (4)
C3—H3	0.9300	C10—H10	0.9300
C4—C5	1.367 (5)	C11—C12	1.372 (4)

C4—Cl2	1.728 (3)	C11—H11	0.9300
C5—C6	1.377 (4)	C12—H12	0.9300
C5—H5	0.9300	N1—S1	1.620 (2)
C6—H6	0.9300	N1—H1N	0.855 (10)
C7—C8	1.383 (4)	O1—S1	1.4279 (17)
C7—C12	1.389 (3)	O2—S1	1.4282 (18)

C6—C1—C2	118.5 (3)	C9—C8—Cl3	119.2 (2)
C6—C1—S1	118.1 (2)	C10—C9—C8	119.8 (3)
C2—C1—S1	123.3 (2)	C10—C9—H9	120.1
C3—C2—C1	120.5 (3)	C8—C9—H9	120.1
C3—C2—Cl1	118.1 (2)	C11—C10—C9	119.7 (3)
C1—C2—Cl1	121.4 (2)	C11—C10—H10	120.1
C4—C3—C2	119.5 (3)	C9—C10—H10	120.1
C4—C3—H3	120.3	C10—C11—C12	120.9 (3)
C2—C3—H3	120.3	C10—C11—H11	119.6
C5—C4—C3	121.2 (3)	C12—C11—H11	119.6
C5—C4—Cl2	119.2 (3)	C11—C12—C7	120.6 (3)
C3—C4—Cl2	119.6 (3)	C11—C12—H12	119.7
C4—C5—C6	119.5 (3)	C7—C12—H12	119.7
C4—C5—H5	120.3	C7—N1—S1	125.13 (17)
C6—C5—H5	120.3	C7—N1—H1N	117.7 (19)
C5—C6—C1	120.8 (3)	S1—N1—H1N	114.3 (19)
C5—C6—H6	119.6	O1—S1—O2	118.76 (12)
C1—C6—H6	119.6	O1—S1—N1	105.39 (11)
C8—C7—C12	118.2 (2)	O2—S1—N1	109.43 (12)
C8—C7—N1	120.0 (2)	O1—S1—C1	110.08 (12)
C12—C7—N1	121.8 (2)	O2—S1—C1	106.14 (12)
C7—C8—C9	120.7 (3)	N1—S1—C1	106.48 (12)
C7—C8—Cl3	120.06 (19)		

C6—C1—C2—C3	0.7 (4)	Cl3—C8—C9—C10	178.5 (2)
S1—C1—C2—C3	−179.4 (2)	C8—C9—C10—C11	1.5 (4)
C6—C1—C2—Cl1	−178.8 (2)	C9—C10—C11—C12	−0.2 (5)
S1—C1—C2—Cl1	1.1 (3)	C10—C11—C12—C7	−1.5 (4)
C1—C2—C3—C4	−0.4 (5)	C8—C7—C12—C11	1.7 (4)
Cl1—C2—C3—C4	179.0 (3)	N1—C7—C12—C11	−179.2 (3)
C2—C3—C4—C5	0.1 (5)	C8—C7—N1—S1	−145.6 (2)
C2—C3—C4—Cl2	−179.1 (2)	C12—C7—N1—S1	35.3 (4)
C3—C4—C5—C6	0.0 (6)	C7—N1—S1—O1	171.4 (2)
Cl2—C4—C5—C6	179.3 (3)	C7—N1—S1—O2	−59.8 (2)
C4—C5—C6—C1	0.2 (5)	C7—N1—S1—C1	54.5 (2)
C2—C1—C6—C5	−0.5 (4)	C6—C1—S1—O1	134.6 (2)
S1—C1—C6—C5	179.6 (3)	C2—C1—S1—O1	−45.3 (2)
C12—C7—C8—C9	−0.3 (4)	C6—C1—S1—O2	4.9 (2)
N1—C7—C8—C9	−179.4 (2)	C2—C1—S1—O2	−175.0 (2)
C12—C7—C8—Cl3	179.94 (19)	C6—C1—S1—N1	−111.6 (2)
N1—C7—C8—Cl3	0.8 (3)	C2—C1—S1—N1	68.5 (2)

C7—C8—C9—C10	−1.3 (4)

Hydrogen-bond geometry (Å, º)

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
N1—H1*N*···O1^{i}	0.86 (1)	2.23 (2)	3.007 (3)	152 (2)
N1—H1*N*···Cl3	0.86 (1)	2.57 (3)	2.975 (2)	110 (2)

Symmetry code: (i) −*x*, *y*, −*z*+1/2.