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The quinternary thiophosphate $Cs_{0.5}Ag_{0.5}Nb_2PS_{10}$

Sojeong Park and Hoseop Yun*

Division of Energy Systems Research and Department of Chemistry, Ajou University, Suwon 443-749, Republic of Korea
Correspondence e-mail: hsyun@ajou.ac.kr

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Key indicators: single-crystal X-ray study; T = 290 K; mean σ(S–P) = 0.002 Å; disorder in solvent or counterion; R factor = 0.034; wR factor = 0.075; data-to-parameter ratio = 22.5.

The quinternary thiophosphate $Cs_{0.5}Ag_{0.5}Nb_2PS_{10}$, cesium silver tris(disulfido)[tetrathiophosphato(V)]diniobate(IV), has been prepared from the elements using a CsCl flux. The crystal structure is made up of ${}_{\infty}^{1}[Nb_2PS_{10}]$ chains expanding along [010]. These chains are built up from bicapped trigonal-prismatic $[Nb_2S_{12}]$ units and tetrahedral $[PS_4]$ groups and are linked through a linear S—Ag—S bridge, forming a two-dimensional layer. These layers then stack on top of each other, completing the three-dimensional structure with an undulating van der Waals gap. The disordered Cs^+ ions reside on sites with half-occupation in the voids of this arrangement. Short [2.8843 (5) Å] and long [3.7316 (4) Å] Nb—Nb distances alternate along the chains, and anionic S_2^{2-} and S^{2-} species are observed. The charge balance of the compound can be represented by the formula $[Cs^+]_{0.5}[Ag^+]_{0.5}[Nb^{4+}]_2[PS_4^{3-}][S_2^{2-}]_3$.

Related literature

For Nb_2PS_{10}-related quaternary thiophosphates, see: Do & Yun (1996) for KNb_2PS_{10}, Kim & Yun (2002) for $RbNb_2PS_{10}$, Kwak *et al.* (2007) for $CsNb_2PS_{10}$, Bang *et al.* (2008) for $TlNb_2PS_{10}$, and Do & Yun (2009) for $Ag_{0.88}Nb_2PS_{10}$. For quintenary thiophosphates, see: Kwak & Yun (2008) for $K_{0.34}Cu_{0.5}Nb_2PS_{10}$, Dong *et al.* (2005*a*) for $K_{0.5}Ag_{0.5}Nb_2PS_{10}$, and Dong *et al.* (2005*b*) for $Rb_{0.38}Ag_{0.5}Nb_2PS_{10}$. *PLATON* (Spek, 2009) was used for structure validation. For typical Nb—P and P—S bond length, see: Brec *et al.* (1983), and for typical Nb^{4+}–Nb^{4+} bond lengths, see: Angenault *et al.* (2000). For general background, see: Lee *et al.* (1988).

Experimental

Crystal data

$Cs_{0.5}Ag_{0.5}Nb_2PS_{10}$
M_r = 657.78
Monoclinic, $P2_1/c$
a = 7.3594 (3) Å
b = 12.8534 (4) Å
c = 13.7788 (6) Å
β = 91.0886 (12)°
V = 1303.15 (8) Å^3
Z = 4
Mo $K\alpha$ radiation
μ = 5.54 mm^{-1}
T = 290 K
0.30 × 0.06 × 0.04 mm

Data collection

Rigaku R-AXIS RAPID diffractometer
Absorption correction: multi-scan (*ABSCOR*; Higashi, 1995)
T_{min} = 0.602, T_{max} = 1.000
12389 measured reflections
2991 independent reflections
2430 reflections with $I > 2\sigma(I)$
R_{int} = 0.049

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.034
$wR(F^2)$ = 0.075
S = 1.08
2991 reflections
133 parameters
$\Delta\rho_{max}$ = 1.18 e Å^{-3}
$\Delta\rho_{min}$ = −1.27 e Å^{-3}

Table 1
Selected geometric parameters (Å, °).

Ag—S1	2.4625 (13)	Nb2—S4	2.4985 (12)
Nb1—S4^{i}	2.4953 (13)	Nb2—S8	2.5075 (12)
Nb1—S7^{i}	2.4958 (12)	Nb2—S5	2.5643 (13)
Nb1—S8^{i}	2.5055 (13)	Nb2—S9	2.5670 (12)
Nb1—S10^{i}	2.5231 (13)	Nb2—S3	2.5920 (12)
Nb1—S9	2.5658 (12)	Nb2—S6	2.6250 (12)
Nb1—S2	2.5895 (12)	P—S1	1.9962 (18)
Nb1—S5	2.5993 (13)	P—S3	2.0391 (17)
Nb1—S6	2.6103 (12)	P—S2	2.0527 (17)
Nb2—S10	2.4910 (13)	P—S6	2.0851 (17)
Nb2—S7	2.4932 (13)		
S1—Ag—S1^{ii}	180		

Symmetry codes: (i) $-x+1, y-\frac{1}{2}, -z+\frac{1}{2}$; (ii) $-x, -y, -z$.

Data collection: *RAPID-AUTO* (Rigaku, 2006); cell refinement: *RAPID-AUTO*; data reduction: *RAPID-AUTO*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: locally modified version of *ORTEP* (Johnson, 1965); software used to prepare material for publication: *WinGX* (Farrugia, 1999).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: WM2357).

References

Angenault, J., Cieren, X. & Quarton, M. (2000). *J. Solid State Chem.* **153**, 55–65.
Bang, H., Kim, Y., Kim, S. & Kim, S. (2008). *J. Solid State Chem.* **181**, 1978–1802.
Brec, R., Grenouilleau, P., Evain, M. & Rouxel, J. (1983). *Rev. Chim. Mineral.* **20**, 295–304.
Do, J. & Yun, H. (1996). *Inorg. Chem.* **35**, 3729–3730.
Do, J. & Yun, H. (2009). *Acta Cryst.* E**65**, i56–i57.
Dong, Y., Kim, S. & Yun, H. (2005*b*). *Acta Cryst.* C**61**, i25–i26.
Dong, Y., Kim, S., Yun, H. & Lim, H. (2005*a*). *Bull. Kor. Chem. Soc.* **26**, 309–311.
Farrugia, L. J. (1999). *J. Appl. Cryst.* **32**, 837–838.
Higashi, T. (1995). *ABSCOR*. Rigaku Corporation, Tokyo, Japan.

Johnson, C. K. (1965). *ORTEP*. Report ORNL-3794. Oak Ridge National Laboratory, Tennessee, USA.
Kim, C. & Yun, H. (2002). *Acta* Cryst. **C58**, i53.
Kwak, J., Kim, C., Yun, H. & Do, J. (2007). *Bull. Kor. Chem. Soc.* **28**, 701–704.
Kwak, J. & Yun, H. (2008). *Bull. Kor. Chem. Soc.* **29**, 273–275.
Lee, S., Colombet, P., Ouvrard, G. & Brec, R. (1988). *Inorg. Chem.* **27**, 1291–1294.
Rigaku (2006). *RAPID-AUTO*. Rigaku Corporation, Tokyo, Japan.
Sheldrick, G. M. (2008). *Acta Cryst.* A**64**, 112–122.
Spek, A. L. (2009). *Acta Cryst.* D**65**, 148–155.

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The quinternary thiophosphate $Cs_{0.5}Ag_{0.5}Nb_2PS_{10}$

Sojeong Park and Hoseop Yun

S1. Comment

During an effort to expand representatives of group 5 transition metal thiophosphates by substituting various monovalent cations, we were able to prepare a new derivative in this system. Here we report the synthesis and characterization of the new layered quinternary thiophosphate, $Cs_{0.5}Ag_{0.5}Nb_2PS_{10}$.

The title compound is isostructural with the previously reported $K_{0.34}Cu_{0.5}Nb_2PS_{10}$ (Kwak & Yun, 2008). The ${}_{\infty}{}^{1}[Nb_2PS_{10}]$ chains found in this structure are composed of the typical biprismatic $[Nb_2S_{12}]$ and tetrahedral $[PS_4]$ units. The Nb atoms are surrounded by 8 S atoms in a bicapped trigonal-prismatic fashion. Two prisms are sharing a rectangular face to form the $[Nb_2S_{12}]$ unit. These units are bound through the S—S prism edges and through one of the capping sulfur atoms to make ${}_{\infty}{}^{1}[Nb_2S_9]$ chains. One of the S atoms at the prism edge and two other capping S atoms are bound to the P atom to which an additional S atom (S1) is attached to complete the ${}_{\infty}{}^{1}[Nb_2PS_{10}]$ chains. These anionic chains propagate parallel to [010] and are linked through the linear S—Ag—S bridge to form a two-dimensional layer along ($\overline{2}01$). These layers then stack on top of each other to complete the three-dimensional structure with an undulating van der Waals gap. The disordered Cs^+ cations reside in the voids of this arrangement.

The Nb—S and P—S distances are in agreement with those found in other related phases (Brec *et al.*, 1983). Along the chain, The Nb(1)···Nb(2) interactions alternate in the sequence of one short (2.8843 (5) Å) and one long (3.7316 (4) Å) distance. The short distance is close to that of the typical Nb^{4+}—Nb^{4+} bond (Angenault *et al.*, 2000), and the long Nb···Nb distance shows that there is no significant intermetallic bonding interaction. Such an arrangement is consistent with the high electric resistivity of the crystal along the needle axis (*b* axis).

The coordination around the Ag atom ($\overline{1}$ symmetry) can be described as a [2 + 4] interaction. Four S atoms are bound to the Ag atoms in the plane (Ag—S6, 3.139 (3) Å, Ag—S9, 3.232 (3) Å), whereas two *trans* S atoms are coordinated to the Ag atom at short distances of Ag—S1 = 2.4625 (13) Å. The large ADPs of Ag could be explained by the second-order Jahn-Teller coupling between the filled Ag e_g and the empty *s* orbitals (Lee *et al.*, 1988), which is a common trend of d^{10} elements. The charge balance of the compound can be represented by the formula $[Cs^+]_{0.5}[Ag^+]_{0.5}[Nb^{4+}]_2[PS_4^{3-}][S_2^{2-}]_3$.

For Nb_2PS_{10}-related quaternary thiophosphates, see: Do & Yun (1996) for KNb_2PS_{10}, Kim & Yun (2002) for $RbNb_2PS_{10}$, Kwak *et al.* (2007) for $CsNb_2PS_{10}$, Bang *et al.* (2008) for $TlNb_2PS_{10}$, and Do & Yun (2009) for $Ag_{0.88}Nb_2PS_{10}$; for quinternary thiophosphates, see: Kwak & Yun (2008) for $K_{0.34}Cu_{0.5}Nb_2PS_{10}$, Dong *et al.* (2005*a*) for $K_{0.5}Ag_{0.5}Nb_2PS_{10}$, and Dong *et al.* (2005*b*) for $Rb_{0.38}Ag_{0.5}Nb_2PS_{10}$.

S2. Experimental

$Cs_{0.5}Ag_{0.5}Nb_2PS_{10}$ was prepared by the reaction of elemental powders, using the reactive halide-flux technique. Ag powder (CERAC 99.999%), Nb powder (CERAC 99.8%), P powder (CERAC 99.5%) and S powder (Aldrich 99.999%) were mixed in a fused silica tube in a molar ratio of Ag:Nb:P:S=1:2:1:10 and then CsCl was added in a weight ratio of $AgNb_2PS_{10}$:CsCl=1:3. The tube was evacuated to 0.133 Pa, sealed and heated gradually (50 K/h) to 973 K, where it was kept for 72 h. The tube was cooled to room temperature at the rate of 4 K/h. The excess halide was removed with distilled water and dark red needle-shaped crystals were obtained. The crystals are stable in air and water. A qualitative X-ray fluorescence analysis of the needles indicated the presence of Cs, Ag, Nb, P, and S. The composition of the compound was determined by single-crystal X-ray diffraction.

S3. Refinement

Refinement went smoothly but the anisotropic displacement parameters (ADPs) of the Cs (Wyckoff position 4*e*) and Ag (2*a*) atoms were large compared with those of the other atoms. Because non-stoichiometry in these phases is sometimes observed and the distance between Cs atoms is too short if full occupancy is assumed, the occupancies of each metal atom were checked by refining the site occupation factors (SOFs) while those of the other atoms were fixed. With the non-stoichiometric model, the SOF of the Cs site was reduced significantly from 1 to 0.49 and the residuals improved also. As no evidence was found for ordering of the Cs site at Wyckoff position 2*c*, a statistically disordered structure was finally modelled. The final difference Fourier map showed that the highest residual electron density (1.18 e/Å^3) is 0.94 Å from the Nb2 site and the deepest hole (-1.27 e/Å^3) is 0.84 Å from the Nb2 site. No additional symmetry, as tested by *PLATON* (Spek, 2009), has been detected in this structure.

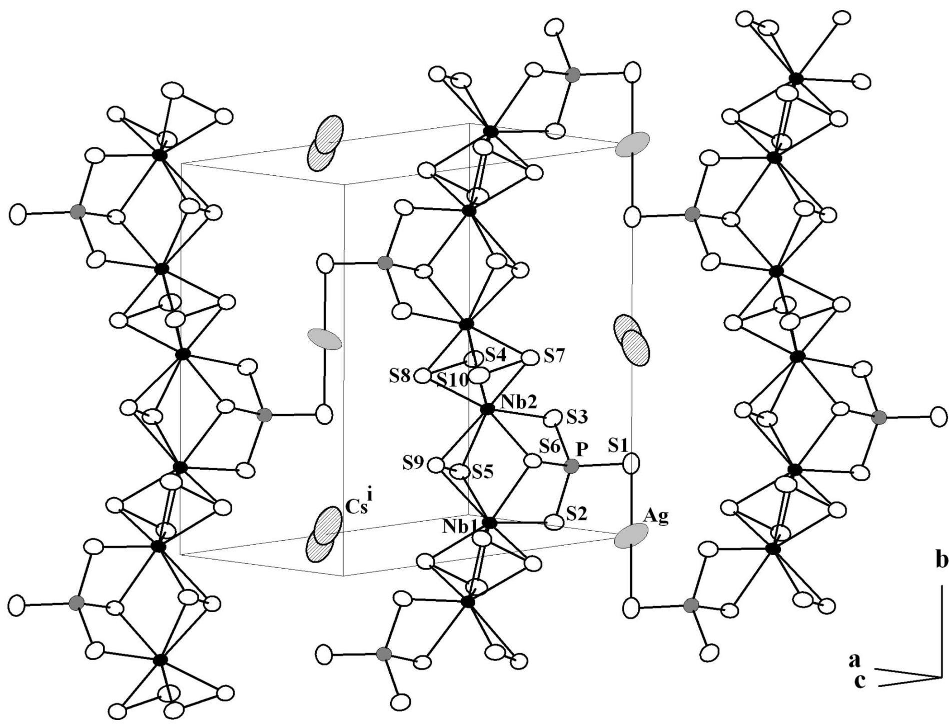

Figure 1

A view of the $Cs_{0.5}Ag_{0.5}Nb_2PS_{10}$ structure. Anisotropic displacement ellipsoids are drawn at the 90% probability level. Symmetry codes are given in Table 1.

cesium silver tris(disulfido)[tetrathiophosphato(V)]diniobate(IV)

Crystal data

$Cs_{0.5}Ag_{0.5}Nb_2PS_{10}$	a = 7.3594 (3) Å
M_r = 657.78	b = 12.8534 (4) Å
Monoclinic, $P2_1/c$	c = 13.7788 (6) Å
Hall symbol: -P 2ybc	β = 91.0886 (12)°

V = 1303.15 (8) Å^3
Z = 4
F(000) = 1232
D_x = 3.353 Mg m^{-3}
Mo $K\alpha$ radiation, λ = 0.71073 Å
Cell parameters from 8832 reflections
θ = 3.2–27.5°
μ = 5.54 mm^{-1}
T = 290 K
Needle, dark brown
0.30 × 0.06 × 0.04 mm

Data collection

Rigaku R-AXIS RAPID diffractometer
Graphite monochromator
ω scans
Absorption correction: multi-scan (*ABSCOR*; Higashi, 1995)
T_{min} = 0.602, T_{max} = 1.000
12389 measured reflections
2991 independent reflections
2430 reflections with $I > 2\sigma(I)$
R_{int} = 0.049
θ_{max} = 27.5°, θ_{min} = 3.2°
h = −9→9
k = −16→14
l = −17→17

Refinement

Refinement on F^2
Least-squares matrix: full
$R[F^2 > 2\sigma(F^2)]$ = 0.034
$wR(F^2)$ = 0.075
S = 1.08
2991 reflections
133 parameters
0 restraints
$w = 1/[\sigma^2(F_o^2) + (0.0248P)^2 + 3.4157P]$ where $P = (F_o^2 + 2F_c^2)/3$
$(\Delta/\sigma)_{max}$ < 0.001
$\Delta\rho_{max}$ = 1.18 e Å^{-3}
$\Delta\rho_{min}$ = −1.27 e Å^{-3}

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	y	z	U_{iso}*/U_{eq}	Occ. (<1)
Cs	−0.0009 (4)	0.02093 (15)	0.48715 (19)	0.0512 (5)	0.5
Ag	0	0	0	0.0539 (2)	
Nb1	0.42651 (5)	0.03565 (3)	0.24995 (3)	0.01333 (11)	
Nb2	0.43445 (5)	0.32590 (3)	0.25353 (3)	0.01280 (11)	
P	0.11427 (16)	0.18631 (9)	0.14370 (10)	0.0170 (3)	
S1	−0.03805 (19)	0.18878 (10)	0.02229 (11)	0.0292 (3)	
S2	0.07865 (16)	0.05281 (9)	0.22287 (10)	0.0210 (3)	
S3	0.08568 (16)	0.31790 (9)	0.22464 (10)	0.0207 (3)	
S4	0.33249 (16)	0.47121 (9)	0.35994 (9)	0.0187 (3)	
S5	0.38477 (17)	0.18131 (9)	0.37848 (9)	0.0190 (3)	
S6	0.39304 (16)	0.18214 (8)	0.11963 (9)	0.0161 (2)	
S7	0.42828 (17)	0.44324 (9)	0.10942 (9)	0.0194 (3)	
S8	0.58572 (16)	0.41695 (9)	0.39426 (9)	0.0184 (3)	
S9	0.63765 (16)	0.17797 (9)	0.31795 (9)	0.0189 (3)	
S10	0.67972 (16)	0.38737 (9)	0.14538 (9)	0.0204 (3)	

Atomic displacement parameters (Å^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Cs	0.0270 (3)	0.0779 (16)	0.0489 (13)	−0.0003 (11)	0.0060 (7)	−0.0283 (10)
Ag	0.0518 (4)	0.0298 (3)	0.0800 (6)	−0.0028 (3)	0.0017 (4)	−0.0287 (4)
Nb1	0.0150 (2)	0.00814 (18)	0.0168 (2)	0.00081 (16)	−0.00185 (16)	0.00082 (16)
Nb2	0.0145 (2)	0.00802 (19)	0.0158 (2)	−0.00071 (16)	−0.00210 (16)	−0.00048 (16)
P	0.0157 (5)	0.0115 (5)	0.0235 (7)	0.0014 (5)	−0.0057 (5)	−0.0025 (5)
S1	0.0325 (7)	0.0228 (6)	0.0316 (8)	0.0023 (6)	−0.0168 (6)	−0.0038 (6)
S2	0.0181 (5)	0.0140 (5)	0.0310 (7)	−0.0014 (5)	−0.0023 (5)	0.0030 (5)
S3	0.0168 (5)	0.0146 (5)	0.0305 (7)	0.0023 (5)	−0.0034 (5)	−0.0073 (5)
S4	0.0190 (5)	0.0146 (5)	0.0225 (7)	−0.0007 (5)	0.0017 (5)	−0.0031 (5)
S5	0.0256 (6)	0.0126 (5)	0.0189 (6)	0.0001 (5)	0.0008 (5)	0.0007 (5)
S6	0.0186 (5)	0.0104 (5)	0.0193 (6)	0.0019 (5)	−0.0015 (5)	−0.0007 (5)
S7	0.0269 (6)	0.0131 (5)	0.0180 (6)	−0.0009 (5)	−0.0044 (5)	0.0004 (5)
S8	0.0233 (6)	0.0124 (5)	0.0195 (6)	−0.0013 (5)	−0.0043 (5)	0.0000 (5)
S9	0.0185 (5)	0.0128 (5)	0.0251 (7)	−0.0001 (5)	−0.0054 (5)	0.0009 (5)
S10	0.0217 (6)	0.0152 (5)	0.0244 (7)	−0.0003 (5)	0.0048 (5)	−0.0021 (5)

Geometric parameters (Å, º)

Cs—Cs^{i}	0.644 (3)	Nb2—S8	2.5075 (12)
Cs—S10^{ii}	3.444 (3)	Nb2—S5	2.5643 (13)
Cs—S10^{iii}	3.469 (3)	Nb2—S9	2.5670 (12)
Cs—S7^{iv}	3.536 (3)	Nb2—S3	2.5920 (12)
Cs—S7^{v}	3.581 (3)	Nb2—S6	2.6250 (12)
Cs—S2	3.722 (3)	Nb2—Nb1^{viii}	2.8843 (5)
Cs—S1^{v}	3.773 (2)	P—S1	1.9962 (18)
Cs—S5	3.835 (3)	P—S3	2.0391 (17)
Cs—S3^{v}	3.915 (3)	P—S2	2.0527 (17)
Cs—S3^{iv}	3.956 (3)	P—S6	2.0851 (17)
Cs—S9^{vi}	4.044 (3)	S1—Cs^{ix}	3.773 (2)
Cs—S2^{i}	4.157 (3)	S2—Cs^{i}	4.157 (3)
Ag—S1	2.4625 (13)	S3—Cs^{ix}	3.915 (3)
Ag—S1^{vii}	2.4625 (13)	S3—Cs^{x}	3.956 (3)
Nb1—S4^{iii}	2.4953 (13)	S4—S8	2.0371 (17)
Nb1—S7^{iii}	2.4958 (12)	S4—Nb1^{viii}	2.4953 (13)
Nb1—S8^{iii}	2.5055 (13)	S5—S9	2.0542 (18)
Nb1—S10^{iii}	2.5231 (13)	S7—S10	2.0372 (17)
Nb1—S9	2.5658 (12)	S7—Nb1^{viii}	2.4958 (12)
Nb1—S2	2.5895 (12)	S7—Cs^{x}	3.536 (3)
Nb1—S5	2.5993 (13)	S7—Cs^{ix}	3.581 (3)
Nb1—S6	2.6103 (12)	S8—Nb1^{viii}	2.5055 (13)
Nb1—Nb2^{iii}	2.8843 (5)	S9—Cs^{xi}	4.044 (3)
Nb2—S10	2.4910 (13)	S10—Nb1^{viii}	2.5231 (13)
Nb2—S7	2.4932 (13)	S10—Cs^{xii}	3.444 (3)
Nb2—S4	2.4985 (12)	S10—Cs^{viii}	3.469 (3)

Cs^{i}—Cs—$S10^{ii}$	86.8 (5)	S9—Nb1—$Nb2^{iii}$	117.38 (3)
Cs^{i}—Cs—$S10^{iii}$	82.5 (5)	S2—Nb1—$Nb2^{iii}$	115.30 (3)
$S10^{ii}$—Cs—$S10^{iii}$	169.32 (5)	S5—Nb1—$Nb2^{iii}$	136.98 (3)
Cs^{i}—Cs—$S7^{iv}$	88.8 (5)	S6—Nb1—$Nb2^{iii}$	133.76 (3)
$S10^{ii}$—Cs—$S7^{iv}$	73.85 (7)	S10—Nb2—S7	48.25 (4)
$S10^{iii}$—Cs—$S7^{iv}$	105.80 (8)	S10—Nb2—S4	110.06 (4)
Cs^{i}—Cs—$S7^{v}$	80.9 (5)	S7—Nb2—S4	90.81 (4)
$S10^{ii}$—Cs—$S7^{v}$	105.35 (8)	S10—Nb2—S8	89.91 (4)
$S10^{iii}$—Cs—$S7^{v}$	73.00 (7)	S7—Nb2—S8	109.55 (4)
$S7^{iv}$—Cs—$S7^{v}$	169.64 (5)	S4—Nb2—S8	48.02 (4)
Cs^{i}—Cs—S2	128.9 (5)	S10—Nb2—S5	138.22 (4)
$S10^{ii}$—Cs—S2	134.71 (8)	S7—Nb2—S5	166.52 (4)
$S10^{iii}$—Cs—S2	54.41 (5)	S4—Nb2—S5	95.72 (4)
$S7^{iv}$—Cs—S2	79.52 (6)	S8—Nb2—S5	83.45 (4)
$S7^{v}$—Cs—S2	107.01 (8)	S10—Nb2—S9	91.02 (4)
Cs^{i}—Cs—$S1^{v}$	139.1 (6)	S7—Nb2—S9	136.48 (4)
$S10^{ii}$—Cs—$S1^{v}$	61.86 (5)	S4—Nb2—S9	122.02 (4)
$S10^{iii}$—Cs—$S1^{v}$	127.50 (7)	S8—Nb2—S9	80.28 (4)
$S7^{iv}$—Cs—$S1^{v}$	105.15 (7)	S5—Nb2—S9	47.20 (4)
$S7^{v}$—Cs—$S1^{v}$	82.99 (6)	S10—Nb2—S3	130.35 (5)
S2—Cs—$S1^{v}$	91.71 (4)	S7—Nb2—S3	84.19 (4)
Cs^{i}—Cs—S5	131.1 (6)	S4—Nb2—S3	79.18 (4)
$S10^{ii}$—Cs—S5	125.61 (7)	S8—Nb2—S3	124.13 (4)
$S10^{iii}$—Cs—S5	62.86 (5)	S5—Nb2—S3	85.47 (4)
$S7^{iv}$—Cs—S5	131.64 (7)	S9—Nb2—S3	126.33 (4)
$S7^{v}$—Cs—S5	57.45 (5)	S10—Nb2—S6	83.03 (4)
S2—Cs—S5	55.08 (4)	S7—Nb2—S6	82.29 (4)
$S1^{v}$—Cs—S5	64.82 (4)	S4—Nb2—S6	155.03 (4)
Cs^{i}—Cs—$S3^{v}$	88.9 (5)	S8—Nb2—S6	156.36 (4)
$S10^{ii}$—Cs—$S3^{v}$	52.55 (5)	S5—Nb2—S6	86.88 (4)
$S10^{iii}$—Cs—$S3^{v}$	126.89 (9)	S9—Nb2—S6	77.33 (4)
$S7^{iv}$—Cs—$S3^{v}$	126.40 (9)	S3—Nb2—S6	76.27 (4)
$S7^{v}$—Cs—$S3^{v}$	53.89 (5)	S10—Nb2—$Nb1^{viii}$	55.41 (3)
S2—Cs—$S3^{v}$	137.18 (6)	S7—Nb2—$Nb1^{viii}$	54.72 (3)
$S1^{v}$—Cs—$S3^{v}$	51.72 (4)	S4—Nb2—$Nb1^{viii}$	54.67 (3)
S5—Cs—$S3^{v}$	86.11 (6)	S8—Nb2—$Nb1^{viii}$	54.84 (3)
Cs^{i}—Cs—$S3^{iv}$	81.7 (5)	S5—Nb2—$Nb1^{viii}$	138.05 (3)
$S10^{ii}$—Cs—$S3^{iv}$	126.35 (9)	S9—Nb2—$Nb1^{viii}$	119.58 (3)
$S10^{iii}$—Cs—$S3^{iv}$	52.02 (5)	S3—Nb2—$Nb1^{viii}$	112.65 (3)
$S7^{iv}$—Cs—$S3^{iv}$	53.79 (5)	S6—Nb2—$Nb1^{viii}$	133.06 (3)
$S7^{v}$—Cs—$S3^{iv}$	123.87 (8)	S1—P—S3	112.52 (8)
S2—Cs—$S3^{iv}$	51.42 (5)	S1—P—S2	112.54 (8)
$S1^{v}$—Cs—$S3^{iv}$	137.41 (7)	S3—P—S2	112.78 (8)
S5—Cs—$S3^{iv}$	100.03 (7)	S1—P—S6	113.93 (9)
$S3^{v}$—Cs—$S3^{iv}$	170.63 (4)	S3—P—S6	102.73 (7)
Cs^{i}—Cs—$S9^{vi}$	137.7 (6)	S2—P—S6	101.48 (7)
$S10^{ii}$—Cs—$S9^{vi}$	75.21 (6)	P—S1—Ag	91.46 (6)
$S10^{iii}$—Cs—$S9^{vi}$	113.01 (7)	P—S1—Cs^{ix}	94.72 (7)

S7^{iv}—Cs—S9^{vi}	49.70 (4)	Ag—S1—Cs^{ix}	161.76 (7)
S7^{v}—Cs—S9^{vi}	140.51 (6)	P—S2—Nb1	90.68 (5)
S2—Cs—S9^{vi}	59.66 (4)	P—S2—Cs	129.50 (7)
S1^{v}—Cs—S9^{vi}	62.08 (4)	Nb1—S2—Cs	91.32 (6)
S5—Cs—S9^{vi}	89.44 (4)	P—S2—Cs^{i}	136.28 (7)
S3^{v}—Cs—S9^{vi}	108.21 (6)	Nb1—S2—Cs^{i}	89.76 (5)
S3^{iv}—Cs—S9^{vi}	79.11 (6)	P—S3—Nb2	90.27 (5)
Cs^{i}—Cs—S2^{i}	44.2 (5)	P—S3—Cs^{ix}	89.92 (6)
S10^{ii}—Cs—S2^{i}	50.32 (4)	Nb2—S3—Cs^{ix}	104.61 (6)
S10^{iii}—Cs—S2^{i}	120.06 (6)	P—S3—Cs^{x}	99.22 (6)
S7^{iv}—Cs—S2^{i}	99.18 (6)	Nb2—S3—Cs^{x}	103.24 (6)
S7^{v}—Cs—S2^{i}	73.35 (5)	S8—S4—Nb1^{viii}	66.22 (5)
S2—Cs—S2^{i}	173.07 (6)	S8—S4—Nb2	66.22 (5)
S1^{v}—Cs—S2^{i}	95.19 (7)	Nb1^{viii}—S4—Nb2	70.56 (3)
S5—Cs—S2^{i}	127.91 (8)	S9—S5—Nb2	66.47 (5)
S3^{v}—Cs—S2^{i}	48.73 (4)	S9—S5—Nb1	65.71 (5)
S3^{iv}—Cs—S2^{i}	122.41 (5)	Nb2—S5—Nb1	92.55 (4)
S9^{vi}—Cs—S2^{i}	124.54 (8)	S9—S5—Cs	146.11 (7)
S1—Ag—S1^{vii}	180.00 (10)	Nb2—S5—Cs	140.15 (6)
S4^{iii}—Nb1—S7^{iii}	90.82 (4)	Nb1—S5—Cs	88.67 (5)
S4^{iii}—Nb1—S8^{iii}	48.08 (4)	P—S6—Nb1	89.39 (5)
S7^{iii}—Nb1—S8^{iii}	109.54 (4)	P—S6—Nb2	88.37 (5)
S4^{iii}—Nb1—S10^{iii}	109.13 (4)	Nb1—S6—Nb2	90.92 (4)
S7^{iii}—Nb1—S10^{iii}	47.89 (4)	S10—S7—Nb2	65.82 (5)
S8^{iii}—Nb1—S10^{iii}	89.23 (4)	S10—S7—Nb1^{viii}	66.76 (5)
S4^{iii}—Nb1—S9	91.47 (4)	Nb2—S7—Nb1^{viii}	70.64 (3)
S7^{iii}—Nb1—S9	78.95 (4)	S10—S7—Cs^{x}	171.29 (7)
S8^{iii}—Nb1—S9	137.40 (4)	Nb2—S7—Cs^{x}	118.24 (6)
S10^{iii}—Nb1—S9	121.46 (4)	Nb1^{viii}—S7—Cs^{x}	121.50 (5)
S4^{iii}—Nb1—S2	130.77 (4)	S10—S7—Cs^{ix}	161.56 (7)
S7^{iii}—Nb1—S2	124.06 (4)	Nb2—S7—Cs^{ix}	117.07 (6)
S8^{iii}—Nb1—S2	85.24 (4)	Nb1^{viii}—S7—Cs^{ix}	131.67 (5)
S10^{iii}—Nb1—S2	80.24 (4)	S4—S8—Nb1^{viii}	65.70 (5)
S9—Nb1—S2	125.60 (4)	S4—S8—Nb2	65.76 (5)
S4^{iii}—Nb1—S5	138.34 (4)	Nb1^{viii}—S8—Nb2	70.25 (3)
S7^{iii}—Nb1—S5	82.43 (4)	S5—S9—Nb1	67.42 (5)
S8^{iii}—Nb1—S5	167.42 (4)	S5—S9—Nb2	66.33 (5)
S10^{iii}—Nb1—S5	96.47 (4)	Nb1—S9—Nb2	93.27 (4)
S9—Nb1—S5	46.87 (4)	S5—S9—Cs^{xi}	111.56 (7)
S2—Nb1—S5	84.69 (4)	Nb1—S9—Cs^{xi}	104.00 (5)
S4^{iii}—Nb1—S6	83.16 (4)	Nb2—S9—Cs^{xi}	160.34 (6)
S7^{iii}—Nb1—S6	155.62 (4)	S7—S10—Nb2	65.93 (5)
S8^{iii}—Nb1—S6	83.80 (4)	S7—S10—Nb1^{viii}	65.35 (5)
S10^{iii}—Nb1—S6	155.76 (4)	Nb2—S10—Nb1^{viii}	70.23 (3)
S9—Nb1—S6	77.61 (4)	S7—S10—Cs^{xii}	110.68 (8)
S2—Nb1—S6	76.07 (4)	Nb2—S10—Cs^{xii}	176.58 (7)
S5—Nb1—S6	86.46 (4)	Nb1^{viii}—S10—Cs^{xii}	109.02 (5)
S4^{iii}—Nb1—Nb2^{iii}	54.77 (3)	S7—S10—Cs^{viii}	108.80 (8)

S7^{iii}—Nb1—Nb2^{iii}	54.64 (3)	Nb2—S10—Cs^{viii}	168.71 (5)
S8^{iii}—Nb1—Nb2^{iii}	54.91 (3)	Nb1^{viii}—S10—Cs^{viii}	98.55 (4)
S10^{iii}—Nb1—Nb2^{iii}	54.37 (3)		

Symmetry codes: (i) −*x*, −*y*, −*z*+1; (ii) *x*−1, −*y*+1/2, *z*+1/2; (iii) −*x*+1, *y*−1/2, −*z*+1/2; (iv) −*x*, *y*−1/2, −*z*+1/2; (v) *x*, −*y*+1/2, *z*+1/2; (vi) *x*−1, *y*, *z*; (vii) −*x*, −*y*, −*z*; (viii) −*x*+1, *y*+1/2, −*z*+1/2; (ix) *x*, −*y*+1/2, *z*−1/2; (x) −*x*, *y*+1/2, −*z*+1/2; (xi) *x*+1, *y*, *z*; (xii) *x*+1, −*y*+1/2, *z*−1/2.