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## Structure Reports

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# Ethyl 5-amino-1-(6-chloropyridazin-3-yl)-1H-pyrazole-4-carboxylate

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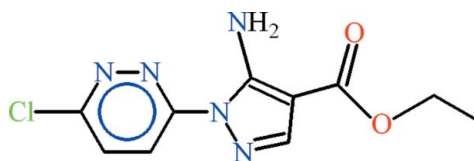
Received 23 August 2010; accepted 25 August 2010

Key indicators: single-crystal X-ray study;  $T = 296$  K; mean  $\sigma(\text{C}-\text{C}) = 0.002$  Å;  $R$  factor = 0.035;  $wR$  factor = 0.101; data-to-parameter ratio = 13.0.

In the title compound,  $\text{C}_{10}\text{H}_{10}\text{ClN}_5\text{O}_2$ , the dihedral angle between the aromatic rings is  $0.16$  ( $9^\circ$ ). Two  $S(6)$  ring motifs are formed due to intramolecular  $\text{N}-\text{H}\cdots\text{N}$  and  $\text{N}-\text{H}\cdots\text{O}$  hydrogen bonds. In the crystal, inversion dimers linked by pairs of  $\text{N}-\text{H}\cdots\text{N}$  hydrogen bonds generate  $R_2^2(14)$  [or  $R_4^4(10)$  via the intramolecular hydrogen bonds] ring motifs. Polymeric chains propagating in  $[210]$  are formed as a result of interlinking the dimers by pairs of  $\text{C}-\text{H}\cdots\text{N}$  interactions, completing  $R_2^2(6)$  ring motifs.

## Related literature

For biochemical background and related structures, see: Ather *et al.* (2010*a,b,c*). For graph-set notation, see: Bernstein *et al.* (1995).



## Experimental

### Crystal data

$\text{C}_{10}\text{H}_{10}\text{ClN}_5\text{O}_2$   
 $M_r = 267.68$   
 Triclinic,  $P\bar{1}$   
 $a = 5.3618$  (3) Å  
 $b = 8.6168$  (4) Å

$c = 13.1585$  (7) Å  
 $\alpha = 77.734$  ( $2^\circ$ )  
 $\beta = 82.928$  ( $1^\circ$ )  
 $\gamma = 86.722$  ( $2^\circ$ )  
 $V = 589.24$  ( $5$ ) Å<sup>3</sup>

$Z = 2$   
 Mo  $K\alpha$  radiation  
 $\mu = 0.33$  mm<sup>-1</sup>

$T = 296$  K  
 $0.25 \times 0.20 \times 0.08$  mm

### Data collection

Bruker Kappa APEXII CCD diffractometer  
 Absorption correction: multi-scan (*SADABS*; Bruker, 2005)  
 $T_{\min} = 0.982$ ,  $T_{\max} = 0.988$

8832 measured reflections  
 2125 independent reflections  
 1721 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.032$

### Refinement

$R[F^2 > 2\sigma(F^2)] = 0.035$   
 $wR(F^2) = 0.101$   
 $S = 1.06$   
 2125 reflections

164 parameters  
 H-atom parameters constrained  
 $\Delta\rho_{\max} = 0.22$  e Å<sup>-3</sup>  
 $\Delta\rho_{\min} = -0.16$  e Å<sup>-3</sup>

**Table 1**  
 Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{N5}-\text{H5A}\cdots\text{N1}$	0.86	2.17	2.775 (2)	127
$\text{N5}-\text{H5B}\cdots\text{O2}$	0.86	2.40	2.942 (2)	122
$\text{N5}-\text{H5B}\cdots\text{N2}^i$	0.86	2.41	3.017 (2)	128
$\text{C5}-\text{H5}\cdots\text{N4}^{ii}$	0.93	2.53	3.313 (2)	142

Symmetry codes: (i)  $-x + 2, -y + 1, -z$ ; (ii)  $-x, -y + 2, -z$ .

Data collection: *APEX2* (Bruker, 2009); cell refinement: *SAINT* (Bruker, 2009); data reduction: *SAINT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEP-3* (Farrugia, 1997) and *PLATON* (Spek, 2009); software used to prepare material for publication: *WinGX* (Farrugia, 1999) and *PLATON*.

The authors acknowledge the provision of funds for the purchase of the diffractometer and encouragement by Dr Muhammad Akram Chaudhary, Vice Chancellor, University of Sargodha, Pakistan. They also acknowledge the technical support provided by Bana International, Karachi, Pakistan.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HB5619).

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## supporting information

*Acta Cryst.* (2010). E66, o2445 [https://doi.org/10.1107/S1600536810034240]

**Ethyl 5-amino-1-(6-chloropyridazin-3-yl)-1H-pyrazole-4-carboxylate**

**Abdul Qayyum Ather, M. Nawaz Tahir, Misbahul Ain Khan, Muhammad Makshoof Athar and Eliana Aparecida Silicz Bueno**

**S1. Comment**

In continuation of our studies of pyrazolylpyridazine derivatives (Ather *et al.*, 2010*a, b, c*), the title compound (I, Fig. 1) is being reported here.

In (I), the 1-(6-chloropyridazin-3-yl)-1H-pyrazol-5-amine moiety A (C1—C7/N1—N5/CL1) and ethyl formate group B (C8—C10/O1/O2) are planar with r. m. s. deviations of 0.0026 and 0.0293 Å, respectively. The dihedral angle between A/B is 3.09 (12)°. There exist two S(6) ring motifs (Bernstein *et al.*, 1995) due to N—H⋯N and N—H⋯O types of intramolecular H-bondings (Table 1, Fig. 1). The molecules are dimerized due to N—H⋯N type of H-bonding (Table 2, Fig. 2) with  $R_4^4(10)$  ring motifs. The dimers are interlinked in the form of polymeric chains due to H-bondings of C—H⋯N type with  $R_2^2(6)$  ring motifs (Table 2, Fig. 2).

**S2. Experimental**

3-Chloro-6-hydrazinylpyridazine (2 g, 13.84 mmol) and ethylethoxymethylene cyanoacetate (2.35 g, 13.84 mmol) were dissolved in acetic acid (10 ml). The obtained reaction mixture was refluxed for 4 h and cooled to room temperature. The resulting product was poured in 100 ml of distilled water and the precipitates were formed. The precipitates obtained by filtration were washed three times by water. The crude material obtained was dried and purified by column chromatography. The final product was re-crystallized in benzene to obtain light brown plates of (I).

**S3. Refinement**

The H-atoms were positioned geometrically (N—H = 0.86, C—H = 0.93–0.97 Å) and were included in the refinement in the riding model approximation, with  $U_{\text{iso}}(\text{H}) = xU_{\text{eq}}(\text{C, N})$ , where  $x = 1.5$  for methyl and  $x = 1.2$  for all other H-atoms.

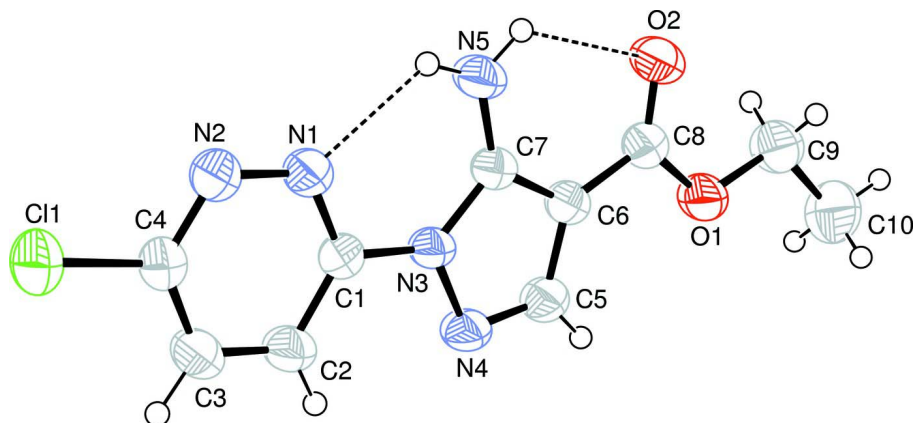


Figure 1

View of (I) with displacement ellipsoids drawn at the 50% probability level. H-atoms are shown as small spheres of arbitrary radius. The dotted lines indicate the intramolecular H-bonds.

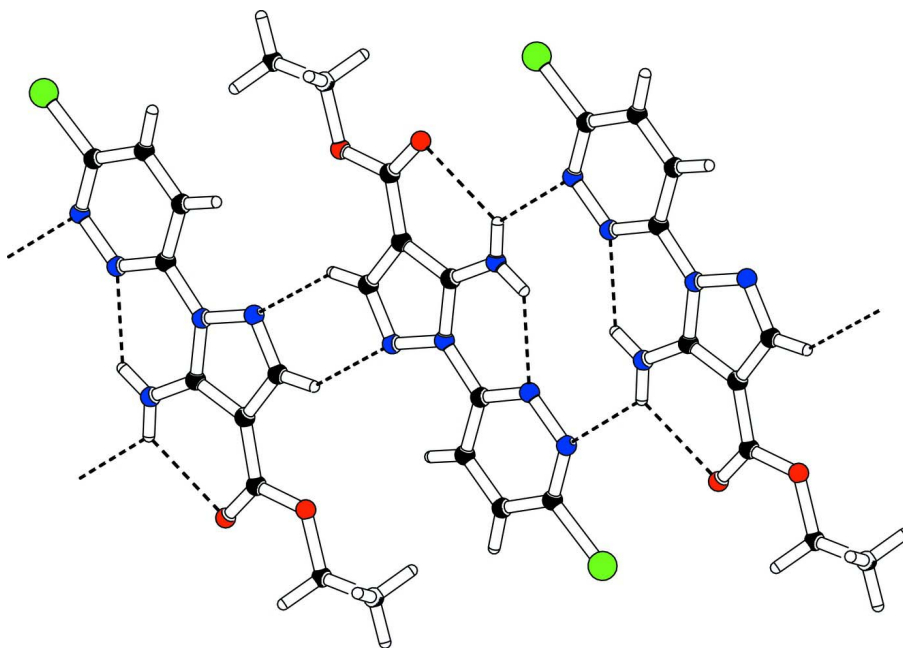


Figure 2

Packing diagram of (I) showing that the molecules form dimers, which are interlinked in the form of polymeric chains.

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#### Crystal data

$C_{10}H_{10}ClN_5O_2$

$M_r = 267.68$

Triclinic,  $P\bar{1}$

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$\alpha = 77.734$  (2)°

$\beta = 82.928$  (1)°

$\gamma = 86.722$  (2)°

$V = 589.24$  (5) Å<sup>3</sup>

$Z = 2$

$F(000) = 276$

$D_x = 1.509$  Mg m<sup>-3</sup>

Mo  $K\alpha$  radiation,  $\lambda = 0.71073$  Å

Cell parameters from 1721 reflections

$\theta = 2.4$ – $25.2$ °

$\mu = 0.33$  mm<sup>-1</sup>

$T = 296$  K  $0.25 \times 0.20 \times 0.08$  mm  
 Plate, light brown

*Data collection*

Bruker Kappa APEXII CCD diffractometer Radiation source: fine-focus sealed tube Graphite monochromator Detector resolution: 8.10 pixels mm <sup>-1</sup> $\omega$ scans Absorption correction: multi-scan (SADABS; Bruker, 2005) $T_{\min} = 0.982$ , $T_{\max} = 0.988$	8832 measured reflections 2125 independent reflections 1721 reflections with $I > 2\sigma(I)$ $R_{\text{int}} = 0.032$ $\theta_{\text{max}} = 25.2^\circ$ , $\theta_{\text{min}} = 2.4^\circ$ $h = -6 \rightarrow 6$ $k = -10 \rightarrow 10$ $l = -15 \rightarrow 15$
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*Refinement*

Refinement on $F^2$ Least-squares matrix: full $R[F^2 > 2\sigma(F^2)] = 0.035$ $wR(F^2) = 0.101$ $S = 1.06$ 2125 reflections 164 parameters 0 restraints Primary atom site location: structure-invariant direct methods	Secondary atom site location: difference Fourier map Hydrogen site location: inferred from neighbouring sites H-atom parameters constrained $w = 1/[\sigma^2(F_o^2) + (0.049P)^2 + 0.1261P]$ where $P = (F_o^2 + 2F_c^2)/3$ $(\Delta/\sigma)_{\text{max}} = 0.001$ $\Delta\rho_{\text{max}} = 0.22 \text{ e } \text{\AA}^{-3}$ $\Delta\rho_{\text{min}} = -0.16 \text{ e } \text{\AA}^{-3}$
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*Special details*

**Geometry.** Bond distances, angles *etc.* have been calculated using the rounded fractional coordinates. All su's are estimated from the variances of the (full) variance-covariance matrix. The cell e.s.d.'s are taken into account in the estimation of distances, angles and torsion angles

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	$x$	$y$	$z$	$U_{\text{iso}}^*/U_{\text{eq}}$
C11	1.21255 (10)	0.81971 (7)	-0.39777 (4)	0.0677 (2)
O1	0.0151 (3)	0.69934 (15)	0.27701 (10)	0.0604 (4)
O2	0.3413 (3)	0.52161 (16)	0.27988 (10)	0.0640 (5)
N1	0.8615 (3)	0.69752 (18)	-0.12132 (12)	0.0510 (5)
N2	1.0286 (3)	0.70752 (18)	-0.20741 (12)	0.0542 (5)
N3	0.4983 (2)	0.78156 (16)	-0.03130 (10)	0.0431 (5)
N4	0.2872 (3)	0.88448 (17)	-0.03037 (12)	0.0502 (5)
N5	0.6846 (3)	0.56546 (17)	0.08430 (12)	0.0555 (5)
C1	0.6639 (3)	0.79680 (19)	-0.12305 (13)	0.0415 (5)
C2	0.6174 (3)	0.9141 (2)	-0.21063 (15)	0.0533 (6)
C3	0.7854 (3)	0.9223 (2)	-0.29680 (15)	0.0561 (6)
C4	0.9885 (3)	0.8155 (2)	-0.29040 (14)	0.0481 (6)
C5	0.1714 (3)	0.8415 (2)	0.06357 (14)	0.0500 (6)
C6	0.2931 (3)	0.7149 (2)	0.12671 (13)	0.0445 (5)

C7	0.5052 (3)	0.67718 (19)	0.06360 (13)	0.0422 (5)
C8	0.2250 (3)	0.6341 (2)	0.23350 (14)	0.0486 (6)
C9	-0.0707 (5)	0.6236 (3)	0.38284 (17)	0.0732 (8)
C10	-0.2913 (4)	0.7155 (3)	0.42191 (19)	0.0806 (9)
H2	0.47768	0.98310	-0.20953	0.0639*
H3	0.76539	0.99675	-0.35797	0.0673*
H5	0.02210	0.89012	0.08673	0.0600*
H5A	0.80622	0.55512	0.03686	0.0665*
H5B	0.67887	0.50366	0.14512	0.0665*
H9A	0.06306	0.61935	0.42685	0.0879*
H9B	-0.11731	0.51568	0.38514	0.0879*
H10A	-0.42248	0.71987	0.37778	0.1208*
H10B	-0.24295	0.82139	0.42099	0.1208*
H10C	-0.35108	0.66460	0.49222	0.1208*

*Atomic displacement parameters (Å<sup>2</sup>)*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
Cl1	0.0694 (3)	0.0765 (4)	0.0499 (3)	0.0052 (3)	0.0126 (2)	-0.0098 (3)
O1	0.0675 (8)	0.0591 (8)	0.0434 (7)	0.0132 (6)	0.0109 (6)	0.0005 (6)
O2	0.0691 (9)	0.0651 (9)	0.0483 (8)	0.0172 (7)	-0.0037 (7)	0.0029 (6)
N1	0.0510 (8)	0.0536 (9)	0.0424 (8)	0.0121 (7)	0.0004 (7)	-0.0037 (7)
N2	0.0530 (9)	0.0583 (10)	0.0457 (9)	0.0122 (7)	0.0024 (7)	-0.0064 (8)
N3	0.0442 (8)	0.0407 (8)	0.0398 (8)	0.0096 (6)	-0.0023 (6)	-0.0027 (6)
N4	0.0469 (8)	0.0485 (8)	0.0486 (9)	0.0168 (7)	-0.0014 (7)	-0.0022 (7)
N5	0.0544 (9)	0.0579 (10)	0.0445 (9)	0.0196 (7)	-0.0021 (7)	0.0032 (7)
C1	0.0422 (9)	0.0420 (9)	0.0395 (9)	0.0041 (7)	-0.0040 (7)	-0.0085 (7)
C2	0.0535 (10)	0.0515 (10)	0.0480 (11)	0.0137 (8)	-0.0039 (8)	-0.0003 (8)
C3	0.0608 (11)	0.0570 (11)	0.0428 (10)	0.0081 (9)	-0.0026 (9)	0.0026 (9)
C4	0.0512 (10)	0.0512 (10)	0.0403 (10)	0.0011 (8)	-0.0008 (8)	-0.0093 (8)
C5	0.0472 (9)	0.0500 (10)	0.0479 (10)	0.0106 (8)	0.0013 (8)	-0.0064 (8)
C6	0.0462 (9)	0.0451 (9)	0.0390 (9)	0.0058 (7)	-0.0027 (7)	-0.0050 (8)
C7	0.0441 (9)	0.0404 (9)	0.0400 (9)	0.0047 (7)	-0.0058 (7)	-0.0050 (7)
C8	0.0526 (10)	0.0482 (10)	0.0429 (10)	0.0042 (8)	-0.0034 (8)	-0.0073 (8)
C9	0.0894 (15)	0.0670 (13)	0.0492 (12)	0.0089 (11)	0.0189 (11)	0.0010 (10)
C10	0.0777 (15)	0.0906 (17)	0.0658 (15)	0.0008 (13)	0.0197 (12)	-0.0161 (13)

*Geometric parameters (Å, °)*

Cl1—C4	1.7337 (18)	C2—C3	1.351 (3)
O1—C8	1.348 (2)	C3—C4	1.384 (2)
O1—C9	1.438 (3)	C5—C6	1.405 (2)
O2—C8	1.214 (2)	C6—C7	1.389 (2)
N1—N2	1.346 (2)	C6—C8	1.442 (2)
N1—C1	1.323 (2)	C9—C10	1.486 (4)
N2—C4	1.307 (2)	C2—H2	0.9300
N3—N4	1.398 (2)	C3—H3	0.9300
N3—C1	1.395 (2)	C5—H5	0.9300

N3—C7	1.378 (2)	C9—H9A	0.9700
N4—C5	1.301 (2)	C9—H9B	0.9700
N5—C7	1.332 (2)	C10—H10A	0.9600
N5—H5A	0.8600	C10—H10B	0.9600
N5—H5B	0.8600	C10—H10C	0.9600
C1—C2	1.400 (2)		
C8—O1—C9	115.59 (16)	N3—C7—N5	124.20 (15)
N2—N1—C1	119.35 (15)	N3—C7—C6	105.71 (14)
N1—N2—C4	118.42 (16)	O1—C8—O2	123.39 (16)
N4—N3—C1	118.30 (13)	O1—C8—C6	111.74 (15)
N4—N3—C7	111.47 (12)	O2—C8—C6	124.87 (16)
C1—N3—C7	130.23 (13)	O1—C9—C10	109.1 (2)
N3—N4—C5	104.02 (14)	C1—C2—H2	122.00
H5A—N5—H5B	120.00	C3—C2—H2	122.00
C7—N5—H5A	120.00	C2—C3—H3	121.00
C7—N5—H5B	120.00	C4—C3—H3	121.00
N1—C1—N3	116.76 (15)	N4—C5—H5	123.00
N1—C1—C2	123.38 (16)	C6—C5—H5	123.00
N3—C1—C2	119.86 (15)	O1—C9—H9A	110.00
C1—C2—C3	116.75 (16)	O1—C9—H9B	110.00
C2—C3—C4	117.32 (17)	C10—C9—H9A	110.00
N2—C4—C3	124.78 (17)	C10—C9—H9B	110.00
C11—C4—C3	119.87 (14)	H9A—C9—H9B	108.00
C11—C4—N2	115.35 (13)	C9—C10—H10A	109.00
N4—C5—C6	113.70 (15)	C9—C10—H10B	109.00
C5—C6—C7	105.11 (15)	C9—C10—H10C	109.00
C5—C6—C8	130.39 (16)	H10A—C10—H10B	109.00
C7—C6—C8	124.51 (15)	H10A—C10—H10C	109.00
N5—C7—C6	130.09 (16)	H10B—C10—H10C	109.00
C9—O1—C8—O2	-2.0 (3)	C1—N3—C7—C6	-179.82 (16)
C9—O1—C8—C6	178.45 (17)	N3—N4—C5—C6	0.2 (2)
C8—O1—C9—C10	176.03 (17)	N1—C1—C2—C3	0.6 (3)
C1—N1—N2—C4	-0.3 (2)	N3—C1—C2—C3	179.95 (16)
N2—N1—C1—N3	-179.52 (15)	C1—C2—C3—C4	-0.6 (2)
N2—N1—C1—C2	-0.1 (3)	C2—C3—C4—C11	-179.69 (14)
N1—N2—C4—C11	-179.85 (13)	C2—C3—C4—N2	0.2 (3)
N1—N2—C4—C3	0.2 (3)	N4—C5—C6—C7	-0.2 (2)
C1—N3—N4—C5	179.72 (14)	N4—C5—C6—C8	-179.72 (17)
C7—N3—N4—C5	0.00 (18)	C5—C6—C7—N3	0.21 (18)
N4—N3—C1—N1	-179.93 (15)	C5—C6—C7—N5	-179.42 (18)
N4—N3—C1—C2	0.6 (2)	C8—C6—C7—N3	179.73 (16)
C7—N3—C1—N1	-0.3 (3)	C8—C6—C7—N5	0.1 (3)
C7—N3—C1—C2	-179.69 (16)	C5—C6—C8—O1	-2.9 (3)
N4—N3—C7—N5	179.53 (15)	C5—C6—C8—O2	177.57 (18)
N4—N3—C7—C6	-0.13 (18)	C7—C6—C8—O1	177.73 (16)
C1—N3—C7—N5	-0.2 (3)	C7—C6—C8—O2	-1.8 (3)

*Hydrogen-bond geometry (Å, °)*

<i>D</i> —H $\cdots$ <i>A</i>	<i>D</i> —H	H $\cdots$ <i>A</i>	<i>D</i> $\cdots$ <i>A</i>	<i>D</i> —H $\cdots$ <i>A</i>
N5—H5A $\cdots$ N1	0.86	2.17	2.775 (2)	127
N5—H5B $\cdots$ O2	0.86	2.40	2.942 (2)	122
N5—H5B $\cdots$ N2 <sup>i</sup>	0.86	2.41	3.017 (2)	128
C5—H5 $\cdots$ N4 <sup>ii</sup>	0.93	2.53	3.313 (2)	142

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