

Bis[*N*-(pyridin-2-ylcarbonyl)pyridine-2-carboxamidato]iron(III) perchlorate monohydrate

Xi-Feng Li,* Tian-Bao Qiu, Li-Xin Hu and Chun-Yue Hu

Pharmacy College, Henan University of Traditional Chinese Medicine, Zhengzhou 450008, People's Republic of China

Correspondence e-mail: lxf_52@yeah.net

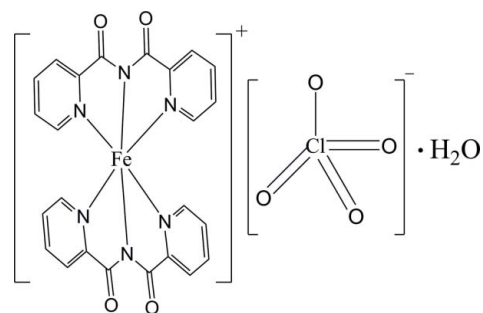
Received 19 September 2011; accepted 10 November 2011

 Key indicators: single-crystal X-ray study; $T = 298$ K; mean $\sigma(\text{C}-\text{C}) = 0.005$ Å; disorder in main residue; R factor = 0.053; wR factor = 0.165; data-to-parameter ratio = 14.4.

The structure of the title salt complex, $[\text{Fe}(\text{C}_{12}\text{H}_8\text{N}_3\text{O}_2)_2]\text{ClO}_4 \cdot \text{H}_2\text{O}$, contains one Fe^{III} cation, two *N*-(pyridin-2-ylcarbonyl)pyridine-2-carboxamidate (bpca^-) anions, one perchlorate anion and one water molecule. The Fe^{III} cation has an approximate octahedral geometry, defined by six N atoms from two bpca^- anions. The nearly parallel [dihedral angle = $1.50(1)^\circ$] bpca^- anions form two-dimensional supermolecules along the a axis by the way of weak $\pi-\pi$ stacking interactions [centroid-centroid distances = 3.948 (2), 4.000 (2), 3.948 (2), 3.911 (2), 3.897 (2), 3.984 (2) and 3.929 (2) Å]. Intra- and intermolecular C—H \cdots O hydrogen bonding occurs. The water molecule [occupancies 0.520 (5) and 0.480 (5)], two carbonyl O atoms [occupancies 0.622 (7) and 0.378 (7)] and the four perchlorate O atoms [occupancies 0.887 (4) and 0.113 (4)] are each disordered over two positions.

Related literature

For the use of organic ligands containing nitrogen heteroatoms in the synthesis of transition-metal complexes, see: Feng *et al.* (2006); Wu *et al.* (2009); Xie & Huang (2011); Yu *et al.* (2010). For the N-donor tridentate rigid ligand Hbpca, see: Casellas *et al.* (2005); Kajiwaru *et al.* (2002). For mononuclear complexes of the tridentate ligand bpca^- , see: Madariaga *et al.* (1991); Marcos *et al.* (1989, 1990); Wocadlo & Massa (1993).



Experimental

Crystal data

 $[\text{Fe}(\text{C}_{12}\text{H}_8\text{N}_3\text{O}_2)_2]\text{ClO}_4 \cdot \text{H}_2\text{O}$
 $M_r = 625.74$

 Triclinic, $P\bar{1}$
 $a = 8.8828(8)$ Å

 $b = 11.7228(3)$ Å

 $c = 14.4551(9)$ Å

 $\alpha = 109.931(3)^\circ$
 $\beta = 103.585(4)^\circ$
 $\gamma = 99.456(3)^\circ$
 $V = 1325.39(15)$ Å³
 $Z = 2$

 Mo $K\alpha$ radiation

 $\mu = 0.73$ mm⁻¹
 $T = 298$ K

 $0.20 \times 0.14 \times 0.10$ mm

Data collection

Bruker APEXII CCD area-detector diffractometer

Absorption correction: multi-scan

 (*SADABS*; Bruker, 2007)

 $T_{\text{min}} = 0.867$, $T_{\text{max}} = 0.930$

8065 measured reflections

5880 independent reflections

 4516 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.021$

Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.053$
 $wR(F^2) = 0.165$
 $S = 1.06$

5880 reflections

407 parameters

50 restraints

H atoms treated by a mixture of independent and constrained refinement

 $\Delta\rho_{\text{max}} = 0.70$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.40$ e Å⁻³
Table 1

Selected bond lengths (Å).

Fe1—N5	1.902 (2)	Fe1—N3	1.974 (2)
Fe1—N2	1.914 (2)	Fe1—N6	1.976 (2)
Fe1—N4	1.964 (2)	Fe1—N1	1.979 (2)

Table 2

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
C12—H12 \cdots O3 ⁱ	0.93	2.30	3.084 (3)	142
C12—H12 \cdots O4 ⁱ	0.93	2.55	3.288 (4)	136
C2—H2 \cdots O5	0.93	2.52	3.292 (5)	140
C2—H2 \cdots O5A	0.93	2.41	3.116 (8)	133
C11—H11 \cdots O6 ⁱ	0.93	2.62	3.385 (6)	140
C11—H11 \cdots O8A ⁱ	0.93	2.32	3.191 (11)	156
C16—H16 \cdots O6A ⁱⁱ	0.93	2.58	3.431 (10)	153

 Symmetry codes: (i) $-x + 1, -y, -z + 1$; (ii) $x + 1, y, z$.

Data collection: *APEX2* (Bruker, 2007); cell refinement: *SAINT* (Bruker, 2007); data reduction: *SAINT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXS97* (Sheldrick, 2008); molecular graphics:

metal-organic compounds

SHELXTL (Sheldrick, 2008); software used to prepare material for publication: *pubCIF* (Westrip, 2010).

This study was supported by the Science and Technology Department of Henan Province (102102310321) and the Doctoral Research Fund of Henan Chinese Medicine (BSJJ2009–38)

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: JJ2105).

References

- Bruker (2007). *APEX2*, *SAINT* and *SADABS*. Bruker AXS Inc., Madison, Wisconsin, USA.
- Casellas, H., Costantino, F., Mandonnet, A., Caneschi, A. & Gatteschi, D. (2005). *Inorg. Chim. Acta*, **358**, 177–185.
- Feng, W.-J., Zhou, G.-P., Zheng, X.-F., Liu, Y.-G. & Xu, Y. (2006). *Acta Cryst. E* **62**, m2033–m2035.
- Kajiwara, T., Sensui, R., Noguchi, T., Kamiyama, A. & Ito, T. (2002). *Inorg. Chim. Acta*, **337**, 299–307.
- Madariaga, G., Zúñiga, F.-J., Rojo, T. & Folgado, J.-V. (1991). *Acta Cryst. C* **47**, 1632–1634.
- Marcos, D., Folgado, J.-V., Beltran-Porter, D., do Prado-Gambardella, M. T., Pulcinelli, S. H. & de Almeida-Santos, R. H. (1990). *Polyhedron*, **9**, 2699–2704.
- Marcos, D., Martinez-Manez, R., Folgado, J.-V., Beltran-Porter, A. Beltran-Porter, D. & Fuertes, A. (1989). *Inorg. Chim. Acta*, **159**, 11–18.
- Sheldrick, G. M. (2008). *Acta Cryst. A* **64**, 112–122.
- Westrip, S. P. (2010). *J. Appl. Cryst.* **43**, 920–925.
- Wocadlo, S. & Massa, W. (1993). *Inorg. Chim. Acta*, **207**, 199–206.
- Wu, D. Y., Sato, O. & Duan, C. Y. (2009). *Inorg. Chem. Commun.* **12**, 325–327.
- Xie, J.-Y. & Huang, F. (2011). *Acta Cryst. E* **67**, m1326.
- Yu, X. Y., Ye, L., Zhang, X., Cui, X. B., Zhang, J. P., Xu, J. Q., Hou, Q. & Wang, T. G. (2010). *Dalton Trans.* **39**, 10617–10625.

supporting information

Acta Cryst. (2011). E67, m1765–m1766 [https://doi.org/10.1107/S1600536811047684]

Bis[*N*-(pyridin-2-ylcarbonyl)pyridine-2-carboxamidato]iron(III) perchlorate monohydrate

Xi-Feng Li, Tian-Bao Qiu, Li-Xin Hu and Chun-Yue Hu

S1. Comment

Organic ligands containing nitrogen heteroatoms play an important role in the assembling process of transition-metal complexes (Wu *et al.*, 2009; Xie *et al.*, 2011; Feng *et al.*, 2006; Yu *et al.*, 2010). The N-donor tridentate rigid ligand Hbpca (Hbpca = bis(2-pyridylcarbonyl)amine) has been attracting widespread attention (Kajiwara *et al.*, 2002; Casellas *et al.*, 2005). Mononuclear complexes of the tridentate ligand bpca[−], [M(bpca)]⁺ (*M* = divalent metal ions) exemplify these types of complexed ligands (Marcos *et al.*, 1989, 1990); Madariaga *et al.*, 1991; Wocadlo *et al.*, 1993).

As a continuation of our studies related to the behaviour of the bpca ligand with Fe^{III} ions, we have synthesized the title salt complex [Fe(bpca)₂(ClO₄) · H₂O]. As shown in Fig. 1, each asymmetric unit contains one Fe^{III} cation, two bpca[−] anions, one perchlorate anion and one lattice water molecule. The bpca ligands are nearly parallel in the structure and form two-dimensional supermolecules by the way of weak π – π stacking interactions [centroid to centroid distances = 3.9475 Å, 3.9997 Å, 3.9488 Å, 3.9107 Å, 3.8973 Å, 3.9838 Å and 3.9287 Å, respectively]. The bond lengths for Fe–N vary from 1.902 (2) Å to 1.979 (2) Å, and the angles for N–Fe–N are between 81.65 (9)° and 178.54 (10)°. These bond lengths and angles are in agreement with those found in similar related Fe^{III} compounds (Wocadlo *et al.*, 1993).

S2. Experimental

A mixture of methanol and water (1:1, 2 ml) was gently layered on the top of a solution of Fe(ClO₄)₃·6H₂O (46.2 mg, 0.1 mmol) in water (3 ml). A solution of Hbpca = bis(2-pyridylcarbonyl)amine (22.7 mg, 0.1 mmol), in methanol (10 ml) was added carefully as the third layer. Red crystals suitable for X-ray diffraction were obtained after 3 weeks, washed with ethanol and ether, and dried in air.

S3. Refinement

Lattice water OW1 and OW1' atoms were disordered over two positions with approximate part occupancies of 0.520 (5) and 0.480 (5). Hydrogen atoms on OW1 and OW1' were also treated using the above two part model. The carbonyl oxygen atoms, O1A and O1B, were disordered over two positions with approximate part occupancies of 0.622 (7) and 0.378 (7). The four perchlorate oxygen atoms O5, O6, O7 and O8 were also disordered with approximate part occupancies of 0.887 (4) and 0.113 (4). All of the remaining H atoms were placed in calculated positions and then refined using the riding model with Atom–H lengths of 0.93 Å (CH). Isotropic displacement parameters for these atoms were set to 1.2 (CH) times U_{eq} of the parent atom.

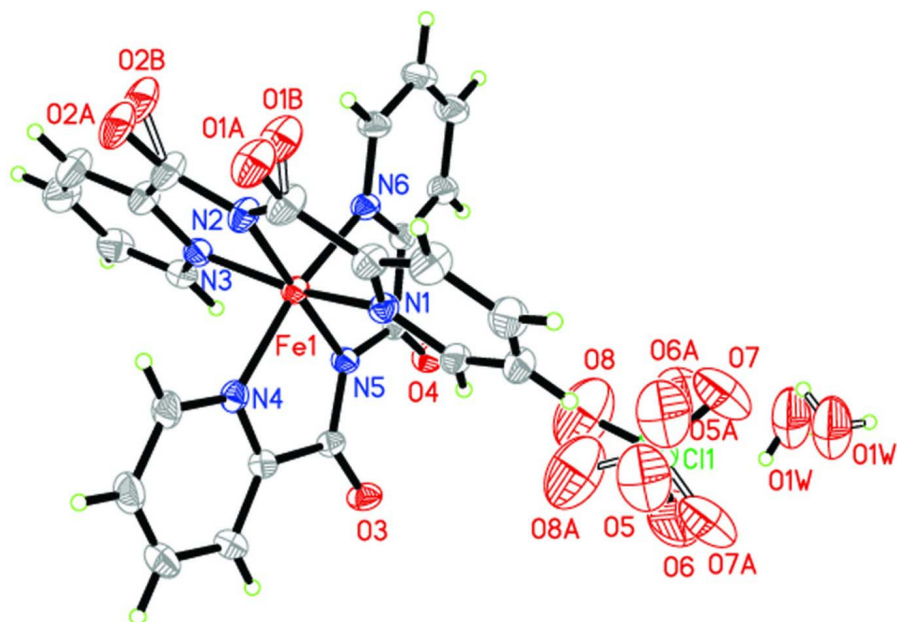


Figure 1

The molecular structure for title salt with displacement ellipsoids at the 30% probability level. OW1 and OW1' atoms were disordered over two positions with approximate part occupancies of 0.520 (5) and 0.480 (5). Hydrogen atoms on OW1 and OW1' were also treated using the above two part model. O1A and O1B, were disordered over two positions with approximate part occupancies of 0.622 (7) and 0.378 (7). O5, O6, O7 and O8 were disordered with approximate part occupancies of 0.887 (4) and 0.113 (4).

Bis[*N*-(pyridin-2-ylcarbonyl)pyridine-2-carboxamidato]iron(III) perchlorate monohydrate

Crystal data

[Fe(C₁₂H₈N₃O₂)₂]ClO₄·H₂O

M_r = 625.74

Triclinic, *P*1̄

Hall symbol: -P 1

a = 8.8828 (8) Å

b = 11.7228 (3) Å

c = 14.4551 (9) Å

α = 109.931 (3)°

β = 103.585 (4)°

γ = 99.456 (3)°

V = 1325.39 (15) Å³

Z = 2

F(000) = 638

D_x = 1.568 Mg m⁻³

Mo *K*α radiation, λ = 0.71073 Å

Cell parameters from 3005 reflections

θ = 2.7–23.6°

μ = 0.73 mm⁻¹

T = 298 K

Block, red

0.20 × 0.14 × 0.10 mm

Data collection

Bruker APEXII CCD area-detector
diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

φ and ω scans

Absorption correction: multi-scan

(*SADABS*; Sheldrick, 2008)

T_{min} = 0.867, *T_{max}* = 0.930

8065 measured reflections

5880 independent reflections

4516 reflections with *I* > 2σ(*I*)

R_{int} = 0.021

θ_{max} = 27.6°, θ_{min} = 1.6°

h = -9→11

k = -15→10

l = -18→18

Refinement

Refinement on F^2
 Least-squares matrix: full
 $R[F^2 > 2\sigma(F^2)] = 0.053$
 $wR(F^2) = 0.165$
 $S = 1.06$
 5880 reflections
 407 parameters
 50 restraints
 Primary atom site location: structure-invariant
 direct methods

Secondary atom site location: difference Fourier
 map
 Hydrogen site location: inferred from
 neighbouring sites
 H atoms treated by a mixture of independent
 and constrained refinement
 $w = 1/[\sigma^2(F_o^2) + (0.0989P)^2 + 0.1594P]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\max} = 0.005$
 $\Delta\rho_{\max} = 0.70 \text{ e } \text{\AA}^{-3}$
 $\Delta\rho_{\min} = -0.40 \text{ e } \text{\AA}^{-3}$

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to z 56b8 ero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$	Occ. (<1)
Fe1	0.58172 (4)	0.31372 (3)	0.72928 (3)	0.03208 (11)	
Cl1	0.08364 (11)	0.27277 (9)	0.31279 (8)	0.0707 (3)	
O1W	-0.3468 (9)	-0.0077 (8)	0.1161 (7)	0.138 (2)	0.520 (5)
H1W1	-0.260 (3)	-0.007 (9)	0.100 (5)	0.208*	0.520 (5)
H1W2	-0.409 (6)	-0.071 (7)	0.064 (6)	0.208*	0.520 (5)
O1W'	-0.3403 (10)	0.1088 (9)	0.1280 (8)	0.138 (2)	0.480 (5)
H1W3	-0.434 (4)	0.075 (7)	0.089 (6)	0.208*	0.480 (5)
H1W4	-0.338 (13)	0.074 (5)	0.171 (4)	0.208*	0.480 (5)
O1A	0.6988 (7)	0.6549 (5)	0.9646 (4)	0.0645 (14)	0.622 (7)
O2A	0.8091 (7)	0.4809 (5)	1.0372 (4)	0.0649 (14)	0.622 (7)
O1B	0.6349 (13)	0.6377 (9)	0.9628 (8)	0.0645 (14)	0.378 (7)
O2B	0.7488 (13)	0.4655 (10)	1.0392 (8)	0.0649 (14)	0.378 (7)
O3	0.5582 (2)	0.1375 (2)	0.43325 (14)	0.0472 (5)	
O4	0.2819 (2)	0.02485 (19)	0.47266 (14)	0.0438 (5)	
O5	0.1944 (4)	0.3923 (3)	0.3485 (3)	0.1119 (15)	0.887 (4)
O6	0.0713 (5)	0.2023 (4)	0.2094 (2)	0.153 (2)	0.887 (4)
O7	-0.0680 (3)	0.2885 (4)	0.3206 (3)	0.1280 (16)	0.887 (4)
O8	0.1436 (6)	0.2088 (3)	0.3751 (3)	0.155 (2)	0.887 (4)
O5A	0.1129 (12)	0.3944 (5)	0.3903 (6)	0.1119 (15)	0.113 (4)
O6A	-0.0561 (7)	0.1927 (8)	0.3105 (8)	0.153 (2)	0.113 (4)
O7A	0.0631 (15)	0.2805 (10)	0.2154 (4)	0.1280 (16)	0.113 (4)
O8A	0.2168 (8)	0.2240 (9)	0.3367 (10)	0.155 (2)	0.113 (4)
N1	0.5132 (3)	0.4529 (2)	0.69872 (17)	0.0360 (5)	

N2	0.6593 (3)	0.4400 (2)	0.86762 (17)	0.0415 (6)
N3	0.6736 (3)	0.2127 (2)	0.80085 (17)	0.0361 (5)
N4	0.7769 (3)	0.3479 (2)	0.69072 (17)	0.0366 (5)
N5	0.5024 (2)	0.1912 (2)	0.59060 (16)	0.0335 (5)
N6	0.3654 (3)	0.2409 (2)	0.73032 (16)	0.0355 (5)
C1	0.4350 (3)	0.4494 (3)	0.6055 (2)	0.0417 (6)
H1	0.4102	0.3760	0.5466	0.050*
C2	0.3910 (4)	0.5550 (3)	0.5964 (3)	0.0526 (8)
H2	0.3361	0.5510	0.5317	0.063*
C3	0.4283 (4)	0.6644 (3)	0.6823 (3)	0.0615 (9)
H3	0.3998	0.7352	0.6766	0.074*
C4	0.5083 (5)	0.6674 (3)	0.7770 (3)	0.0612 (9)
H4	0.5347	0.7402	0.8367	0.073*
C5	0.5489 (4)	0.5612 (3)	0.7822 (2)	0.0449 (7)
C6	0.6359 (4)	0.5582 (3)	0.8833 (2)	0.0559 (9)
C7	0.7348 (4)	0.4089 (3)	0.9468 (2)	0.0538 (8)
C8	0.7429 (4)	0.2746 (3)	0.9043 (2)	0.0470 (7)
C9	0.8163 (5)	0.2199 (4)	0.9657 (3)	0.0657 (10)
H9	0.8666	0.2659	1.0363	0.079*
C10	0.8141 (5)	0.0945 (4)	0.9200 (3)	0.0707 (10)
H10	0.8604	0.0540	0.9599	0.085*
C11	0.7430 (4)	0.0316 (3)	0.8159 (3)	0.0568 (8)
H11	0.7417	-0.0523	0.7840	0.068*
C12	0.6729 (3)	0.0915 (3)	0.7576 (2)	0.0424 (6)
H12	0.6241	0.0470	0.6866	0.051*
C13	0.9227 (3)	0.4269 (3)	0.7516 (2)	0.0467 (7)
H13	0.9385	0.4723	0.8215	0.056*
C14	1.0488 (4)	0.4418 (3)	0.7125 (3)	0.0570 (9)
H14	1.1484	0.4963	0.7559	0.068*
C15	1.0272 (4)	0.3770 (4)	0.6108 (3)	0.0609 (9)
H15	1.1114	0.3880	0.5839	0.073*
C16	0.8785 (3)	0.2938 (3)	0.5464 (3)	0.0492 (7)
H16	0.8617	0.2477	0.4765	0.059*
C17	0.7584 (3)	0.2822 (3)	0.5893 (2)	0.0373 (6)
C18	0.5952 (3)	0.1934 (2)	0.5270 (2)	0.0356 (6)
C19	0.3515 (3)	0.1088 (2)	0.55652 (19)	0.0330 (5)
C20	0.2785 (3)	0.1384 (2)	0.64330 (19)	0.0345 (6)
C21	0.1325 (3)	0.0665 (3)	0.6333 (2)	0.0464 (7)
H21	0.0782	-0.0058	0.5739	0.056*
C22	0.0673 (4)	0.1040 (4)	0.7139 (3)	0.0570 (9)
H22	-0.0321	0.0577	0.7089	0.068*
C23	0.1517 (4)	0.2100 (4)	0.8003 (3)	0.0559 (8)
H23	0.1089	0.2379	0.8542	0.067*
C24	0.3021 (3)	0.2759 (3)	0.8075 (2)	0.0429 (7)
H24	0.3603	0.3464	0.8677	0.052*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Fe1	0.03236 (19)	0.03226 (19)	0.02527 (17)	0.00582 (14)	0.00651 (14)	0.00684 (14)
Cl1	0.0706 (5)	0.0586 (5)	0.0855 (6)	0.0197 (4)	0.0324 (5)	0.0253 (5)
O1W	0.112 (4)	0.135 (4)	0.175 (6)	0.048 (4)	0.013 (4)	0.084 (5)
O1W'	0.112 (4)	0.135 (4)	0.175 (6)	0.048 (4)	0.013 (4)	0.084 (5)
O1A	0.088 (4)	0.0408 (16)	0.0411 (13)	0.015 (2)	0.003 (2)	0.0007 (12)
O2A	0.076 (4)	0.0571 (17)	0.0343 (12)	0.018 (2)	-0.009 (2)	0.0030 (12)
O1B	0.088 (4)	0.0408 (16)	0.0411 (13)	0.015 (2)	0.003 (2)	0.0007 (12)
O2B	0.076 (4)	0.0571 (17)	0.0343 (12)	0.018 (2)	-0.009 (2)	0.0030 (12)
O3	0.0503 (11)	0.0521 (12)	0.0297 (9)	0.0037 (9)	0.0138 (8)	0.0086 (9)
O4	0.0392 (10)	0.0434 (11)	0.0308 (9)	0.0001 (8)	0.0061 (8)	0.0017 (8)
O5	0.084 (2)	0.071 (2)	0.164 (4)	0.0011 (19)	0.025 (2)	0.045 (2)
O6	0.146 (4)	0.202 (5)	0.093 (3)	0.092 (4)	0.040 (3)	0.014 (3)
O7	0.089 (2)	0.135 (4)	0.187 (4)	0.047 (2)	0.084 (2)	0.061 (3)
O8	0.213 (5)	0.107 (3)	0.169 (5)	0.055 (3)	0.043 (4)	0.089 (3)
O5A	0.084 (2)	0.071 (2)	0.164 (4)	0.0011 (19)	0.025 (2)	0.045 (2)
O6A	0.146 (4)	0.202 (5)	0.093 (3)	0.092 (4)	0.040 (3)	0.014 (3)
O7A	0.089 (2)	0.135 (4)	0.187 (4)	0.047 (2)	0.084 (2)	0.061 (3)
O8A	0.213 (5)	0.107 (3)	0.169 (5)	0.055 (3)	0.043 (4)	0.089 (3)
N1	0.0355 (11)	0.0365 (11)	0.0351 (11)	0.0074 (9)	0.0103 (9)	0.0147 (9)
N2	0.0504 (13)	0.0337 (11)	0.0282 (10)	0.0090 (10)	0.0025 (10)	0.0052 (9)
N3	0.0358 (11)	0.0370 (11)	0.0333 (11)	0.0090 (9)	0.0107 (9)	0.0115 (9)
N4	0.0338 (10)	0.0342 (11)	0.0350 (11)	0.0042 (9)	0.0040 (9)	0.0125 (9)
N5	0.0301 (10)	0.0370 (11)	0.0265 (10)	0.0049 (9)	0.0088 (8)	0.0060 (9)
N6	0.0361 (11)	0.0378 (11)	0.0284 (10)	0.0075 (9)	0.0101 (8)	0.0092 (9)
C1	0.0352 (13)	0.0517 (16)	0.0361 (13)	0.0072 (12)	0.0089 (11)	0.0184 (12)
C2	0.0504 (16)	0.0660 (19)	0.0548 (17)	0.0216 (15)	0.0160 (13)	0.0375 (15)
C3	0.070 (2)	0.0612 (19)	0.075 (2)	0.0331 (16)	0.0305 (17)	0.0400 (17)
C4	0.083 (2)	0.0467 (17)	0.0556 (19)	0.0259 (17)	0.0209 (17)	0.0192 (15)
C5	0.0553 (16)	0.0360 (14)	0.0368 (14)	0.0100 (13)	0.0106 (12)	0.0103 (12)
C6	0.080 (2)	0.0402 (16)	0.0366 (15)	0.0173 (16)	0.0085 (15)	0.0087 (13)
C7	0.070 (2)	0.0501 (17)	0.0300 (14)	0.0170 (16)	0.0021 (13)	0.0107 (13)
C8	0.0546 (17)	0.0466 (16)	0.0360 (14)	0.0142 (14)	0.0079 (12)	0.0159 (12)
C9	0.086 (2)	0.069 (2)	0.0423 (17)	0.0322 (19)	0.0092 (16)	0.0248 (16)
C10	0.090 (2)	0.076 (2)	0.071 (2)	0.046 (2)	0.0302 (19)	0.0451 (19)
C11	0.0686 (19)	0.0473 (16)	0.069 (2)	0.0259 (15)	0.0326 (16)	0.0277 (15)
C12	0.0433 (14)	0.0413 (14)	0.0449 (15)	0.0119 (12)	0.0204 (12)	0.0153 (12)
C13	0.0346 (14)	0.0465 (16)	0.0463 (16)	0.0011 (12)	0.0009 (12)	0.0155 (13)
C14	0.0333 (14)	0.064 (2)	0.066 (2)	0.0034 (14)	0.0050 (14)	0.0280 (17)
C15	0.0365 (15)	0.077 (2)	0.075 (2)	0.0094 (15)	0.0228 (15)	0.0366 (19)
C16	0.0407 (14)	0.0607 (18)	0.0533 (16)	0.0136 (13)	0.0218 (12)	0.0261 (14)
C17	0.0361 (13)	0.0408 (13)	0.0377 (13)	0.0106 (11)	0.0131 (10)	0.0177 (11)
C18	0.0386 (13)	0.0373 (13)	0.0305 (12)	0.0100 (11)	0.0114 (10)	0.0125 (10)
C19	0.0328 (12)	0.0337 (12)	0.0271 (11)	0.0074 (10)	0.0054 (10)	0.0088 (10)
C20	0.0319 (12)	0.0367 (13)	0.0303 (12)	0.0081 (10)	0.0071 (10)	0.0100 (10)
C21	0.0363 (14)	0.0549 (17)	0.0410 (15)	0.0045 (13)	0.0113 (12)	0.0146 (13)

C22	0.0376 (15)	0.077 (2)	0.0498 (17)	0.0029 (15)	0.0177 (13)	0.0208 (16)
C23	0.0475 (16)	0.077 (2)	0.0464 (16)	0.0181 (15)	0.0254 (13)	0.0199 (16)
C24	0.0480 (15)	0.0450 (15)	0.0330 (13)	0.0125 (12)	0.0173 (11)	0.0091 (12)

Geometric parameters (Å, °)

Fe1—N5	1.902 (2)	N6—C20	1.353 (3)
Fe1—N2	1.914 (2)	C1—C2	1.395 (4)
Fe1—N4	1.964 (2)	C1—H1	0.9300
Fe1—N3	1.974 (2)	C2—C3	1.371 (5)
Fe1—N6	1.976 (2)	C2—H2	0.9300
Fe1—N1	1.979 (2)	C3—C4	1.372 (5)
Cl1—O6	1.407 (3)	C3—H3	0.9300
Cl1—O7A	1.412 (4)	C4—C5	1.372 (4)
Cl1—O6A	1.414 (4)	C4—H4	0.9300
Cl1—O7	1.414 (3)	C5—C6	1.499 (4)
Cl1—O5A	1.417 (4)	C7—C8	1.503 (4)
Cl1—O8A	1.418 (4)	C8—C9	1.370 (5)
Cl1—O5	1.419 (3)	C9—C10	1.386 (5)
Cl1—O8	1.422 (3)	C9—H9	0.9300
O1W—H1W1	0.85 (2)	C10—C11	1.361 (5)
O1W—H1W2	0.84 (2)	C10—H10	0.9300
O1W—H1W3	1.43 (8)	C11—C12	1.375 (4)
O1W—H1W4	0.99 (7)	C11—H11	0.9300
O1W'—H1W3	0.828 (19)	C12—H12	0.9300
O1W'—H1W4	0.85 (2)	C13—C14	1.381 (5)
O1A—C6	1.246 (5)	C13—H13	0.9300
O2A—C7	1.230 (6)	C14—C15	1.354 (5)
O1B—C6	1.211 (10)	C14—H14	0.9300
O2B—C7	1.237 (11)	C15—C16	1.394 (4)
O3—C18	1.220 (3)	C15—H15	0.9300
O4—C19	1.206 (3)	C16—C17	1.365 (4)
N1—C5	1.346 (3)	C16—H16	0.9300
N1—C1	1.347 (3)	C17—C18	1.499 (4)
N2—C7	1.378 (4)	C19—C20	1.508 (4)
N2—C6	1.385 (4)	C20—C21	1.372 (4)
N3—C12	1.342 (4)	C21—C22	1.392 (4)
N3—C8	1.350 (4)	C21—H21	0.9300
N4—C13	1.351 (3)	C22—C23	1.364 (5)
N4—C17	1.358 (3)	C22—H22	0.9300
N5—C18	1.375 (3)	C23—C24	1.391 (4)
N5—C19	1.383 (3)	C23—H23	0.9300
N6—C24	1.337 (3)	C24—H24	0.9300
N5—Fe1—N2	178.54 (10)	C2—C3—C4	118.6 (3)
N5—Fe1—N4	82.29 (9)	C2—C3—H3	120.7
N2—Fe1—N4	97.79 (10)	C4—C3—H3	120.7
N5—Fe1—N3	99.81 (9)	C3—C4—C5	118.9 (3)

N2—Fe1—N3	81.65 (9)	C3—C4—H4	120.6
N4—Fe1—N3	90.49 (9)	C5—C4—H4	120.6
N5—Fe1—N6	82.06 (9)	N1—C5—C4	123.3 (3)
N2—Fe1—N6	97.90 (10)	N1—C5—C6	115.2 (3)
N4—Fe1—N6	164.25 (9)	C4—C5—C6	121.5 (3)
N3—Fe1—N6	90.42 (9)	O1B—C6—N2	127.9 (6)
N5—Fe1—N1	96.77 (9)	O1A—C6—N2	126.1 (4)
N2—Fe1—N1	81.76 (10)	O1B—C6—C5	118.5 (6)
N4—Fe1—N1	91.62 (9)	O1A—C6—C5	122.9 (4)
N3—Fe1—N1	163.41 (9)	N2—C6—C5	110.1 (2)
N6—Fe1—N1	91.98 (9)	O2A—C7—N2	127.3 (4)
O6—C11—O6A	94.3 (5)	O2B—C7—N2	124.6 (6)
O7A—C11—O6A	110.0 (4)	O2A—C7—C8	121.7 (4)
O6—C11—O7	110.5 (2)	O2B—C7—C8	121.9 (6)
O7A—C11—O7	96.5 (6)	N2—C7—C8	110.1 (2)
O6A—C11—O7	46.1 (4)	N3—C8—C9	122.7 (3)
O6—C11—O5A	146.3 (5)	N3—C8—C7	114.7 (3)
O7A—C11—O5A	109.7 (4)	C9—C8—C7	122.5 (3)
O6A—C11—O5A	109.4 (4)	C8—C9—C10	118.5 (3)
O7—C11—O5A	73.5 (4)	C8—C9—H9	120.7
O6—C11—O8A	84.0 (6)	C10—C9—H9	120.7
O7A—C11—O8A	109.5 (4)	C11—C10—C9	118.8 (3)
O6A—C11—O8A	109.3 (4)	C11—C10—H10	120.6
O7—C11—O8A	150.4 (5)	C9—C10—H10	120.6
O5A—C11—O8A	109.0 (4)	C10—C11—C12	120.4 (3)
O6—C11—O5	109.3 (2)	C10—C11—H11	119.8
O7A—C11—O5	82.8 (4)	C12—C11—H11	119.8
O6A—C11—O5	152.3 (4)	N3—C12—C11	121.4 (3)
O7—C11—O5	109.7 (2)	N3—C12—H12	119.3
O8A—C11—O5	87.9 (4)	C11—C12—H12	119.3
O6—C11—O8	108.8 (2)	N4—C13—C14	121.4 (3)
O7A—C11—O8	144.5 (5)	N4—C13—H13	119.3
O6A—C11—O8	75.9 (4)	C14—C13—H13	119.3
O7—C11—O8	110.1 (2)	C15—C14—C13	119.9 (3)
O5A—C11—O8	100.3 (4)	C15—C14—H14	120.0
O5—C11—O8	108.3 (2)	C13—C14—H14	120.0
H1W1—O1W—H1W2	99 (3)	C14—C15—C16	119.8 (3)
H1W1—O1W—H1W3	115 (9)	C14—C15—H15	120.1
H1W2—O1W—H1W3	92 (8)	C16—C15—H15	120.1
H1W1—O1W—H1W4	111 (8)	C17—C16—C15	117.8 (3)
H1W2—O1W—H1W4	145 (8)	C17—C16—H16	121.1
H1W3—O1W—H1W4	61 (4)	C15—C16—H16	121.1
H1W3—O1W'—H1W4	100 (3)	N4—C17—C16	123.3 (3)
C5—N1—C1	118.1 (3)	N4—C17—C18	115.0 (2)
C5—N1—Fe1	114.70 (19)	C16—C17—C18	121.8 (3)
C1—N1—Fe1	127.2 (2)	O3—C18—N5	128.6 (2)
C7—N2—C6	123.1 (2)	O3—C18—C17	121.5 (2)
C7—N2—Fe1	118.6 (2)	N5—C18—C17	109.9 (2)

C6—N2—Fe1	118.30 (19)	O4—C19—N5	128.7 (3)
C12—N3—C8	118.1 (3)	O4—C19—C20	122.2 (2)
C12—N3—Fe1	126.92 (19)	N5—C19—C20	109.0 (2)
C8—N3—Fe1	114.95 (19)	N6—C20—C21	122.7 (3)
C13—N4—C17	117.8 (2)	N6—C20—C19	115.6 (2)
C13—N4—Fe1	128.3 (2)	C21—C20—C19	121.7 (2)
C17—N4—Fe1	113.95 (17)	C20—C21—C22	118.7 (3)
C18—N5—C19	122.9 (2)	C20—C21—H21	120.6
C18—N5—Fe1	117.98 (16)	C22—C21—H21	120.6
C19—N5—Fe1	119.04 (17)	C23—C22—C21	118.8 (3)
C24—N6—C20	118.1 (2)	C23—C22—H22	120.6
C24—N6—Fe1	127.95 (19)	C21—C22—H22	120.6
C20—N6—Fe1	113.76 (17)	C22—C23—C24	119.7 (3)
N1—C1—C2	120.7 (3)	C22—C23—H23	120.1
N1—C1—H1	119.7	C24—C23—H23	120.1
C2—C1—H1	119.7	N6—C24—C23	121.8 (3)
C3—C2—C1	120.4 (3)	N6—C24—H24	119.1
C3—C2—H2	119.8	C23—C24—H24	119.1
C1—C2—H2	119.8		
N5—Fe1—N1—C5	179.5 (2)	Fe1—N2—C6—O1A	-168.7 (5)
N2—Fe1—N1—C5	-0.5 (2)	C7—N2—C6—C5	-179.3 (3)
N4—Fe1—N1—C5	-98.1 (2)	Fe1—N2—C6—C5	0.3 (4)
N3—Fe1—N1—C5	-0.9 (4)	N1—C5—C6—O1B	-160.9 (7)
N6—Fe1—N1—C5	97.2 (2)	C4—C5—C6—O1B	18.5 (8)
N5—Fe1—N1—C1	0.3 (2)	N1—C5—C6—O1A	168.7 (4)
N2—Fe1—N1—C1	-179.7 (2)	C4—C5—C6—O1A	-11.9 (7)
N4—Fe1—N1—C1	82.7 (2)	N1—C5—C6—N2	-0.7 (4)
N3—Fe1—N1—C1	179.9 (3)	C4—C5—C6—N2	178.8 (3)
N6—Fe1—N1—C1	-82.0 (2)	C6—N2—C7—O2A	-10.2 (7)
N4—Fe1—N2—C7	-89.8 (3)	Fe1—N2—C7—O2A	170.2 (5)
N3—Fe1—N2—C7	-0.4 (2)	C6—N2—C7—O2B	21.4 (8)
N6—Fe1—N2—C7	88.8 (3)	Fe1—N2—C7—O2B	-158.2 (7)
N1—Fe1—N2—C7	179.7 (3)	C6—N2—C7—C8	-179.3 (3)
N4—Fe1—N2—C6	90.6 (3)	Fe1—N2—C7—C8	1.1 (4)
N3—Fe1—N2—C6	180.0 (3)	C12—N3—C8—C9	2.1 (5)
N6—Fe1—N2—C6	-90.8 (3)	Fe1—N3—C8—C9	-178.2 (3)
N1—Fe1—N2—C6	0.1 (2)	C12—N3—C8—C7	-178.6 (3)
N5—Fe1—N3—C12	-0.7 (2)	Fe1—N3—C8—C7	1.2 (4)
N2—Fe1—N3—C12	179.3 (2)	O2A—C7—C8—N3	-171.2 (4)
N4—Fe1—N3—C12	-82.9 (2)	O2B—C7—C8—N3	158.5 (6)
N6—Fe1—N3—C12	81.3 (2)	N2—C7—C8—N3	-1.4 (4)
N1—Fe1—N3—C12	179.7 (3)	O2A—C7—C8—C9	8.1 (7)
N5—Fe1—N3—C8	179.6 (2)	O2B—C7—C8—C9	-22.1 (8)
N2—Fe1—N3—C8	-0.5 (2)	N2—C7—C8—C9	177.9 (3)
N4—Fe1—N3—C8	97.3 (2)	N3—C8—C9—C10	-2.4 (6)
N6—Fe1—N3—C8	-98.4 (2)	C7—C8—C9—C10	178.3 (4)
N1—Fe1—N3—C8	0.0 (4)	C8—C9—C10—C11	1.7 (6)

N5—Fe1—N4—C13	-174.4 (3)	C9—C10—C11—C12	-0.8 (6)
N2—Fe1—N4—C13	7.1 (3)	C8—N3—C12—C11	-1.1 (4)
N3—Fe1—N4—C13	-74.6 (3)	Fe1—N3—C12—C11	179.2 (2)
N6—Fe1—N4—C13	-167.9 (3)	C10—C11—C12—N3	0.5 (5)
N1—Fe1—N4—C13	89.0 (3)	C17—N4—C13—C14	0.6 (4)
N5—Fe1—N4—C17	5.74 (19)	Fe1—N4—C13—C14	-179.3 (2)
N2—Fe1—N4—C17	-172.78 (19)	N4—C13—C14—C15	0.5 (5)
N3—Fe1—N4—C17	105.6 (2)	C13—C14—C15—C16	-1.1 (6)
N6—Fe1—N4—C17	12.3 (5)	C14—C15—C16—C17	0.7 (5)
N1—Fe1—N4—C17	-90.9 (2)	C13—N4—C17—C16	-1.0 (4)
N4—Fe1—N5—C18	-9.1 (2)	Fe1—N4—C17—C16	178.9 (2)
N3—Fe1—N5—C18	-98.3 (2)	C13—N4—C17—C18	178.2 (2)
N6—Fe1—N5—C18	172.7 (2)	Fe1—N4—C17—C18	-1.9 (3)
N1—Fe1—N5—C18	81.6 (2)	C15—C16—C17—N4	0.4 (5)
N4—Fe1—N5—C19	174.3 (2)	C15—C16—C17—C18	-178.8 (3)
N3—Fe1—N5—C19	85.2 (2)	C19—N5—C18—O3	8.5 (5)
N6—Fe1—N5—C19	-3.9 (2)	Fe1—N5—C18—O3	-167.9 (2)
N1—Fe1—N5—C19	-94.9 (2)	C19—N5—C18—C17	-173.5 (2)
N5—Fe1—N6—C24	-178.0 (3)	Fe1—N5—C18—C17	10.0 (3)
N2—Fe1—N6—C24	0.5 (3)	N4—C17—C18—O3	173.2 (3)
N4—Fe1—N6—C24	175.4 (3)	C16—C17—C18—O3	-7.6 (4)
N3—Fe1—N6—C24	82.1 (3)	N4—C17—C18—N5	-4.9 (3)
N1—Fe1—N6—C24	-81.4 (2)	C16—C17—C18—N5	174.3 (3)
N5—Fe1—N6—C20	6.44 (19)	C18—N5—C19—O4	4.4 (5)
N2—Fe1—N6—C20	-175.04 (19)	Fe1—N5—C19—O4	-179.2 (2)
N4—Fe1—N6—C20	-0.1 (5)	C18—N5—C19—C20	-175.6 (2)
N3—Fe1—N6—C20	-93.42 (19)	Fe1—N5—C19—C20	0.8 (3)
N1—Fe1—N6—C20	103.00 (19)	C24—N6—C20—C21	-3.0 (4)
C5—N1—C1—C2	-0.7 (4)	Fe1—N6—C20—C21	173.0 (2)
Fe1—N1—C1—C2	178.4 (2)	C24—N6—C20—C19	176.1 (2)
N1—C1—C2—C3	0.7 (5)	Fe1—N6—C20—C19	-7.8 (3)
C1—C2—C3—C4	-0.5 (5)	O4—C19—C20—N6	-175.3 (2)
C2—C3—C4—C5	0.3 (6)	N5—C19—C20—N6	4.7 (3)
C1—N1—C5—C4	0.6 (5)	O4—C19—C20—C21	3.9 (4)
Fe1—N1—C5—C4	-178.6 (3)	N5—C19—C20—C21	-176.1 (3)
C1—N1—C5—C6	-180.0 (3)	N6—C20—C21—C22	3.3 (5)
Fe1—N1—C5—C6	0.8 (4)	C19—C20—C21—C22	-175.8 (3)
C3—C4—C5—N1	-0.4 (5)	C20—C21—C22—C23	-0.8 (5)
C3—C4—C5—C6	-179.8 (3)	C21—C22—C23—C24	-1.8 (5)
C7—N2—C6—O1B	-21.4 (9)	C20—N6—C24—C23	0.3 (4)
Fe1—N2—C6—O1B	158.2 (7)	Fe1—N6—C24—C23	-175.1 (2)
C7—N2—C6—O1A	11.7 (7)	C22—C23—C24—N6	2.1 (5)

Hydrogen-bond geometry (\AA , $^\circ$)

<i>D</i> —H \cdots <i>A</i>	<i>D</i> —H	H \cdots <i>A</i>	<i>D</i> \cdots <i>A</i>	<i>D</i> —H \cdots <i>A</i>
C12—H12 \cdots O3 ⁱ	0.93	2.30	3.084 (3)	142
C12—H12 \cdots O4 ⁱ	0.93	2.55	3.288 (4)	136

C2—H2···O5	0.93	2.52	3.292 (5)	140
C2—H2···O5A	0.93	2.41	3.116 (8)	133
C11—H11···O6 ⁱ	0.93	2.62	3.385 (6)	140
C11—H11···O8A ⁱ	0.93	2.32	3.191 (11)	156
C16—H16···O6A ⁱⁱ	0.93	2.58	3.431 (10)	153

Symmetry codes: (i) $-x+1, -y, -z+1$; (ii) $x+1, y, z$.