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2-Chloro-*N*-{3-cyano-1-[2,6-dichloro-4-(trifluoromethyl)phenyl]-1*H*-pyrazol-5-yl}acetamide

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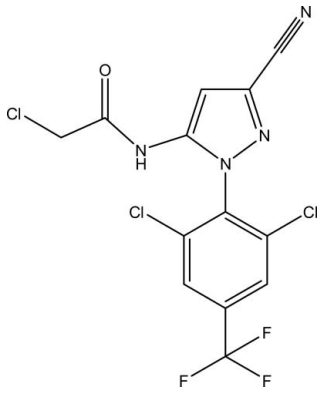
Key indicators: single-crystal X-ray study; T = 293 K; mean σ(C–C) = 0.005 Å; disorder in main residue; R factor = 0.053; wR factor = 0.155; data-to-parameter ratio = 11.9.

The title compound, $C_{13}H_6Cl_3F_3N_4O$, was synthesized by the reaction of 5-amino-1-[2,6-dichloro-4-(trifluoromethyl)phenyl]-1*H*-pyrazole-3-carbonitrile and 2-chloroacetyl chloride. The five-membered pyrazole ring makes a dihedral angle of 71.5 (3)° with the benzene ring. The $-CF_3$ group is disordered by rotation, and the F atoms are split over two sets of sites with occupancies of 0.59 (2) and 0.41 (2). The crystal structure features weak C—H···O and N—H···N interactions involving the carbonyl and cyano groups as acceptors.

Related literature

For biological properties of *N*-pyrazole derivatives, see: Cheng *et al.* (2008); Liu *et al.* (2010); Hatton *et al.* (1993). For related structures, see: Yang *et al.* (2004); Zhang *et al.* (2005); Zhong *et al.* (2004).



Experimental

Crystal data

$C_{13}H_6Cl_3F_3N_4O$
M_r = 397.57
Triclinic, $P\overline{1}$
a = 8.4190 (17) Å
b = 9.2650 (19) Å
c = 11.944 (2) Å
α = 69.77 (3)°
β = 76.74 (3)°
γ = 66.10 (3)°
V = 794.9 (3) Å^3
Z = 2
Mo $K\alpha$ radiation
μ = 0.62 mm^{-1}
T = 293 K
0.30 × 0.20 × 0.20 mm

Data collection

Enraf–Nonius CAD-4 diffractometer
Absorption correction: ψ scan (North *et al.*, 1968)
T_{min} = 0.837, T_{max} = 0.887
3133 measured reflections
2921 independent reflections
2313 reflections with $I > 2\sigma(I)$
R_{int} = 0.030
3 standard reflections every 200 reflections
intensity decay: 1%

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.053
$wR(F^2)$ = 0.155
S = 1.01
2921 reflections
246 parameters
H-atom parameters constrained
$\Delta\rho_{max}$ = 0.60 e Å^{-3}
$\Delta\rho_{min}$ = −0.58 e Å^{-3}

Table 1
Hydrogen-bond geometry (Å, °).

D—H···A	D—H	H···A	D···A	D—H···A
N4—H4A···N3^{i}	0.86	2.49	3.280 (5)	153
C4—H4B···O^{ii}	0.93	2.53	3.349 (5)	148

Symmetry codes: (i) $x+1, y, z$; (ii) $x, y+1, z$.

Data collection: *CAD-4 Software* (Enraf–Nonius, 1989); cell refinement: *CAD-4 Software*; data reduction: *XCAD4* (Harms & Wocadlo, 1995); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXL97*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: BH2393).

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supporting information

Acta Cryst. (2012). E**68**, o104 [doi:10.1107/S1600536811052743]

2-Chloro-*N*-{3-cyano-1-[2,6-dichloro-4-(trifluoromethyl)phenyl]-1*H*-pyrazol-5-yl}acetamide

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S1. Comment

In a variety of biological heterocyclic compounds, *N*-pyrazole derivatives are of great interest because of their chemical and pharmaceutical properties (Cheng *et al.*, 2008). Some X-ray structures of *N*-pyrazole compounds have already been reported (Zhang *et al.*, 2005; Zhong *et al.*, 2004; Yang *et al.*, 2004), and they have been found to exhibit good insecticidal activities against diamond-back moth, mustard beetle, vetch aphid and so on (Hatton *et al.*, 1993). Besides, some other *N*-pyrazole derivatives are known to have antifungal activities (Liu *et al.*, 2010). Herein we report the crystal structure of a new derivative (Fig. 1). In this structure, the pyrazole ring N1/N2/C8/C9/C10 is a planar five-membered ring and the mean deviation from plane is 0.0063 Å. The dihedral angle between the pyrazole and benzene rings is 71.5 (3)°. In the crystal structure, weak intermolecular C—H···O and N—H···N hydrogen bonds (Table 1) link symmetry-related molecules, to form a trimeric unit (Fig. 2), which may be effective in the stabilization of the crystal.

S2. Experimental

To a stirred solution of 5-amino-1-[2,6-dichloro-4-(trifluoromethyl)phenyl]-1*H*-pyrazole-3-carbonitrile (5 mmol) in THF (20 ml) was added 2-chloroacetyl chloride (5 mmol) dropwise at 0–5 °C. After the addition, the reaction mixture was allowed to rise to room temperature and was stirred for 2 h. The crude product precipitated and was filtered. Pure compound was obtained by crystallization from ethanol. Crystals suitable for X-ray diffraction were obtained by slow evaporation of an acetone solution.

S3. Refinement

All H atoms bonded to the C atoms were placed geometrically at the distances of 0.93–0.97 Å and included in the refinement in riding motion approximation with U_{iso}(H) = 1.2 or 1.5U_{eq} of the carrier atom. F atoms were disordered over two sites, occupancies were refined and converged to 0.565 (12) and 0.435 (12), respectively.

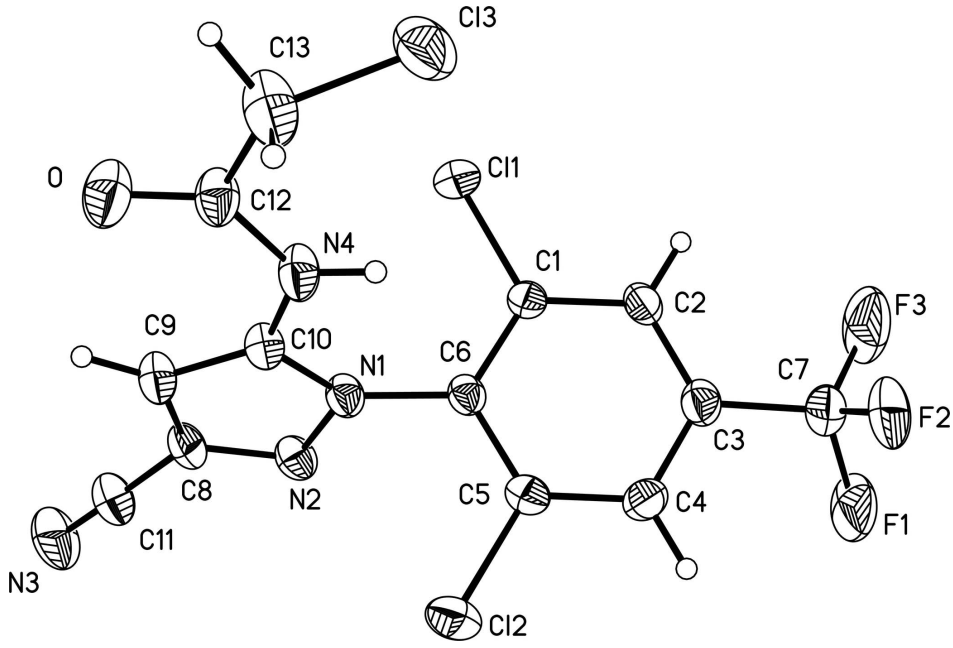

Figure 1

A view of the molecular structure of the title compound. Displacement ellipsoids are drawn at the 50% probability level.

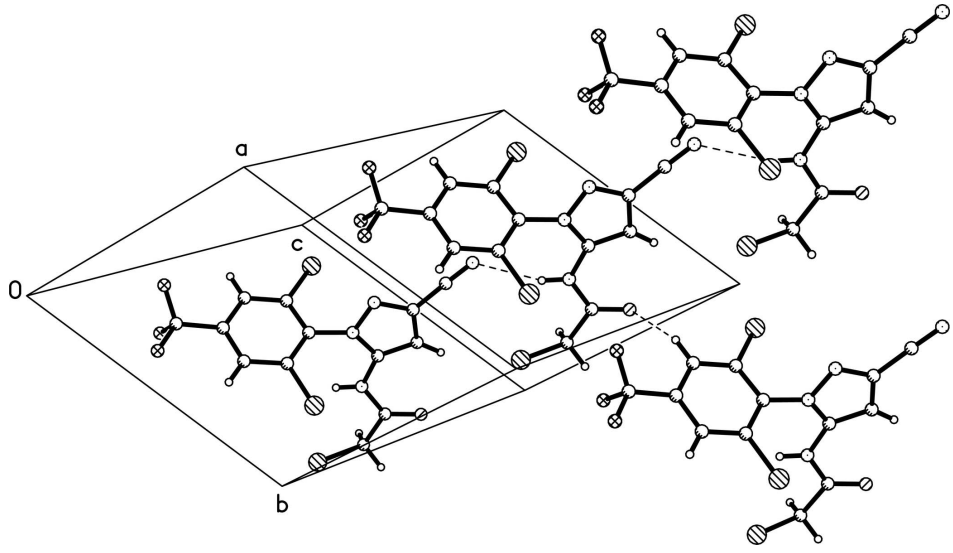

Figure 2

Partial packing view showing the hydrogen-bonded network. Dashed lines indicate intermolecular N—H···N and C—H···O hydrogen bonds.

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γ = 66.10 (3)°

V = 794.9 (3) Å^3
Z = 2
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Melting point: 483 K
Mo $K\alpha$ radiation, λ = 0.71073 Å

Cell parameters from 25 reflections
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T = 293 K
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0.30 × 0.20 × 0.20 mm

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Radiation source: fine-focus sealed tube
Graphite monochromator
$\omega/2\theta$ scans
Absorption correction: ψ scan (North *et al.*, 1968)
T_{min} = 0.837, T_{max} = 0.887
3133 measured reflections

2921 independent reflections
2313 reflections with $I > 2\sigma(I)$
R_{int} = 0.030
θ_{max} = 25.4°, θ_{min} = 1.8°
h = 0→10
k = −10→11
l = −14→14
3 standard reflections every 200 reflections
intensity decay: 1%

Refinement

Refinement on F^2
Least-squares matrix: full
$R[F^2 > 2\sigma(F^2)]$ = 0.053
$wR(F^2)$ = 0.155
S = 1.01
2921 reflections
246 parameters
0 restraints
Primary atom site location: structure-invariant direct methods
Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring sites
H-atom parameters constrained
$w = 1/[\sigma^2(F_o^2) + (0.1P)^2 + 0.320P]$ where $P = (F_o^2 + 2F_c^2)/3$
$(\Delta/\sigma)_{max} < 0.001$
$\Delta\rho_{max}$ = 0.60 e Å^{-3}
$\Delta\rho_{min}$ = −0.58 e Å^{-3}
Extinction correction: *SHELXL97* (Sheldrick, 2008), $Fc^*=kFc[1+0.001xFc^2\lambda^3/\sin(2\theta)]^{-1/4}$
Extinction coefficient: 0.195 (13)

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	y	z	U_{iso}*/U_{eq}	Occ. (<1)
O	0.3070 (3)	0.0918 (3)	0.0994 (3)	0.0713 (8)	
Cl1	0.20070 (11)	0.26994 (9)	0.46773 (7)	0.0529 (3)	
Cl2	0.04187 (13)	0.81751 (11)	0.09593 (7)	0.0636 (3)	
Cl3	0.66610 (14)	0.09291 (16)	0.24189 (12)	0.0861 (4)	
N1	0.0425 (3)	0.4846 (3)	0.2382 (2)	0.0414 (6)	
C1	0.1882 (4)	0.4703 (3)	0.3992 (3)	0.0380 (6)	
N2	−0.1343 (3)	0.5475 (3)	0.2373 (2)	0.0462 (6)	
C2	0.2458 (4)	0.5460 (4)	0.4546 (3)	0.0444 (7)	
H2B	0.2910	0.4894	0.5282	0.053*	
C3	0.2352 (4)	0.7066 (4)	0.3993 (3)	0.0480 (7)	
N3	−0.4748 (4)	0.5210 (5)	0.1439 (4)	0.0816 (11)	
N4	0.3004 (3)	0.2913 (4)	0.1692 (3)	0.0510 (7)	
H4A	0.3625	0.3317	0.1884	0.061*	
C4	0.1707 (4)	0.7917 (4)	0.2891 (3)	0.0482 (7)	
H4B	0.1657	0.8996	0.2520	0.058*	
C5	0.1144 (4)	0.7150 (4)	0.2353 (3)	0.0431 (7)	
C6	0.1175 (3)	0.5552 (3)	0.2906 (2)	0.0371 (6)	

C7	0.2921 (6)	0.7937 (5)	0.4592 (5)	0.0759 (12)	
C8	−0.1616 (4)	0.4631 (4)	0.1792 (3)	0.0468 (7)	
C9	−0.0098 (4)	0.3489 (4)	0.1413 (3)	0.0504 (8)	
H9A	0.0011	0.2764	0.0998	0.060*	
C10	0.1202 (4)	0.3680 (4)	0.1793 (3)	0.0419 (7)	
C11	−0.3373 (5)	0.4970 (5)	0.1602 (3)	0.0597 (9)	
C12	0.3826 (4)	0.1554 (4)	0.1304 (3)	0.0511 (8)	
C13	0.5789 (5)	0.0841 (6)	0.1261 (4)	0.0791 (13)	
H13A	0.6268	0.1419	0.0503	0.095*	
H13B	0.6178	−0.0302	0.1271	0.095*	
F1	0.193 (2)	0.9401 (14)	0.447 (3)	0.155 (10)	0.59 (2)
F2	0.4449 (12)	0.8100 (17)	0.3936 (8)	0.104 (3)	0.59 (2)
F3	0.3470 (18)	0.7085 (13)	0.5612 (6)	0.111 (4)	0.59 (2)
F1′	0.1496 (12)	0.851 (3)	0.5471 (17)	0.111 (6)	0.41 (2)
F2′	0.327 (6)	0.913 (4)	0.4027 (11)	0.169 (13)	0.41 (2)
F3′	0.3932 (17)	0.6929 (17)	0.5431 (17)	0.135 (8)	0.41 (2)

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
O	0.0581 (15)	0.0791 (18)	0.097 (2)	−0.0240 (13)	−0.0006 (14)	−0.0539 (16)
Cl1	0.0595 (5)	0.0366 (4)	0.0579 (5)	−0.0150 (3)	−0.0143 (4)	−0.0052 (3)
Cl2	0.0760 (6)	0.0575 (5)	0.0474 (5)	−0.0184 (4)	−0.0215 (4)	0.0000 (4)
Cl3	0.0623 (6)	0.0942 (8)	0.1011 (8)	−0.0006 (5)	−0.0307 (6)	−0.0455 (7)
N1	0.0354 (13)	0.0464 (14)	0.0460 (13)	−0.0133 (11)	−0.0082 (10)	−0.0167 (11)
C1	0.0340 (14)	0.0352 (14)	0.0432 (15)	−0.0093 (11)	−0.0075 (11)	−0.0105 (12)
N2	0.0336 (13)	0.0546 (15)	0.0508 (15)	−0.0133 (11)	−0.0082 (11)	−0.0159 (12)
C2	0.0419 (16)	0.0485 (17)	0.0454 (16)	−0.0113 (13)	−0.0148 (13)	−0.0157 (13)
C3	0.0396 (16)	0.0499 (18)	0.0607 (19)	−0.0134 (14)	−0.0091 (14)	−0.0236 (15)
N3	0.0492 (19)	0.114 (3)	0.097 (3)	−0.0287 (19)	−0.0217 (17)	−0.040 (2)
N4	0.0399 (14)	0.0645 (17)	0.0633 (17)	−0.0206 (12)	−0.0029 (12)	−0.0346 (14)
C4	0.0478 (18)	0.0361 (15)	0.0593 (19)	−0.0151 (13)	−0.0093 (14)	−0.0093 (13)
C5	0.0390 (15)	0.0418 (16)	0.0433 (15)	−0.0090 (12)	−0.0095 (12)	−0.0089 (12)
C6	0.0320 (14)	0.0402 (15)	0.0407 (15)	−0.0103 (11)	−0.0059 (11)	−0.0150 (12)
C7	0.075 (3)	0.058 (2)	0.112 (4)	−0.018 (2)	−0.040 (3)	−0.031 (2)
C8	0.0410 (17)	0.0599 (19)	0.0465 (16)	−0.0212 (14)	−0.0118 (13)	−0.0149 (14)
C9	0.0486 (18)	0.063 (2)	0.0535 (18)	−0.0246 (15)	−0.0079 (14)	−0.0256 (15)
C10	0.0407 (16)	0.0487 (16)	0.0420 (15)	−0.0174 (13)	−0.0047 (12)	−0.0174 (13)
C11	0.053 (2)	0.077 (2)	0.058 (2)	−0.0258 (18)	−0.0128 (16)	−0.0231 (18)
C12	0.0449 (18)	0.063 (2)	0.0556 (19)	−0.0200 (16)	0.0018 (14)	−0.0334 (16)
C13	0.044 (2)	0.116 (4)	0.095 (3)	−0.013 (2)	−0.0016 (19)	−0.072 (3)
F1	0.120 (7)	0.087 (7)	0.31 (3)	0.014 (6)	−0.107 (12)	−0.124 (12)
F2	0.106 (6)	0.127 (7)	0.122 (5)	−0.082 (5)	−0.039 (4)	−0.020 (5)
F3	0.200 (10)	0.138 (7)	0.059 (4)	−0.118 (8)	−0.023 (4)	−0.027 (4)
F1′	0.075 (5)	0.149 (12)	0.155 (11)	−0.020 (6)	−0.009 (5)	−0.124 (10)
F2′	0.35 (4)	0.16 (2)	0.100 (7)	−0.21 (3)	−0.044 (17)	−0.007 (11)
F3′	0.074 (5)	0.128 (9)	0.230 (19)	0.043 (7)	−0.106 (8)	−0.119 (12)

Geometric parameters (Å, º)

O—C12	1.205 (4)	N4—H4A	0.8600
Cl1—C1	1.720 (3)	C4—C5	1.368 (4)
Cl2—C5	1.717 (3)	C4—H4B	0.9300
Cl3—C13	1.747 (4)	C5—C6	1.390 (4)
N1—C10	1.353 (4)	C7—F2′	1.192 (11)
N1—N2	1.362 (3)	C7—F1	1.249 (8)
N1—C6	1.420 (4)	C7—F3	1.273 (9)
C1—C2	1.382 (4)	C7—F3′	1.304 (13)
C1—C6	1.387 (4)	C7—F2	1.381 (9)
N2—C8	1.319 (4)	C7—F1′	1.454 (11)
C2—C3	1.378 (5)	C8—C9	1.389 (5)
C2—H2B	0.9300	C8—C11	1.437 (4)
C3—C4	1.383 (5)	C9—C10	1.369 (4)
C3—C7	1.504 (5)	C9—H9A	0.9300
N3—C11	1.137 (5)	C12—C13	1.506 (5)
N4—C12	1.353 (4)	C13—H13A	0.9700
N4—C10	1.387 (4)	C13—H13B	0.9700
C10—N1—N2	111.9 (2)	F2′—C7—F1′	103.7 (13)
C10—N1—C6	130.1 (2)	F3′—C7—F1′	91.6 (9)
N2—N1—C6	117.8 (2)	F2′—C7—F3′	114.9 (16)
C2—C1—C6	120.8 (3)	F1—C7—C3	112.7 (5)
C2—C1—Cl1	119.3 (2)	F2—C7—C3	106.0 (6)
C6—C1—Cl1	119.9 (2)	F3—C7—C3	115.0 (5)
C8—N2—N1	103.3 (2)	F1′—C7—C3	107.6 (5)
C3—C2—C1	119.0 (3)	F2′—C7—C3	120.9 (7)
C3—C2—H2B	120.5	F3′—C7—C3	113.0 (7)
C1—C2—H2B	120.5	N2—C8—C9	113.8 (3)
C2—C3—C4	121.2 (3)	N2—C8—C11	119.2 (3)
C2—C3—C7	120.2 (3)	C9—C8—C11	127.1 (3)
C4—C3—C7	118.6 (3)	C10—C9—C8	103.8 (3)
C12—N4—C10	122.7 (3)	C10—C9—H9A	128.1
C12—N4—H4A	118.6	C8—C9—H9A	128.1
C10—N4—H4A	118.6	N1—C10—C9	107.2 (3)
C5—C4—C3	119.0 (3)	N1—C10—N4	120.4 (3)
C5—C4—H4B	120.5	C9—C10—N4	132.4 (3)
C3—C4—H4B	120.5	N3—C11—C8	178.4 (4)
C4—C5—C6	121.2 (3)	O—C12—N4	123.4 (3)
C4—C5—Cl2	119.2 (2)	O—C12—C13	120.0 (3)
C6—C5—Cl2	119.6 (2)	N4—C12—C13	116.7 (3)
C1—C6—C5	118.7 (3)	C12—C13—Cl3	115.8 (2)
C1—C6—N1	121.4 (3)	C12—C13—H13A	108.3
C5—C6—N1	119.8 (3)	Cl3—C13—H13A	108.3
F1—C7—F2	102.2 (9)	C12—C13—H13B	108.3
F1—C7—F3	118.1 (10)	Cl3—C13—H13B	108.3
F3—C7—F2	100.1 (6)	H13A—C13—H13B	107.4

C10—N1—N2—C8	1.6 (3)	C4—C3—C7—F1	39.4 (16)
C6—N1—N2—C8	177.3 (3)	C2—C3—C7—F3	−0.4 (9)
C6—C1—C2—C3	−1.3 (4)	C4—C3—C7—F3	178.8 (8)
Cl1—C1—C2—C3	−179.9 (2)	C2—C3—C7—F3′	18.7 (11)
C1—C2—C3—C4	−1.0 (5)	C4—C3—C7—F3′	−162.0 (10)
C1—C2—C3—C7	178.2 (3)	C2—C3—C7—F2	109.2 (6)
C2—C3—C4—C5	0.9 (5)	C4—C3—C7—F2	−71.6 (6)
C7—C3—C4—C5	−178.3 (3)	C2—C3—C7—F1′	−80.8 (11)
C3—C4—C5—C6	1.4 (5)	C4—C3—C7—F1′	98.4 (11)
C3—C4—C5—Cl2	−177.6 (2)	N1—N2—C8—C9	−0.6 (4)
C2—C1—C6—C5	3.5 (4)	N1—N2—C8—C11	−179.8 (3)
Cl1—C1—C6—C5	−177.8 (2)	N2—C8—C9—C10	−0.6 (4)
C2—C1—C6—N1	−173.9 (2)	C11—C8—C9—C10	178.6 (3)
Cl1—C1—C6—N1	4.8 (4)	N2—N1—C10—C9	−2.0 (3)
C4—C5—C6—C1	−3.6 (4)	C6—N1—C10—C9	−177.0 (3)
Cl2—C5—C6—C1	175.4 (2)	N2—N1—C10—N4	179.1 (3)
C4—C5—C6—N1	173.8 (3)	C6—N1—C10—N4	4.0 (5)
Cl2—C5—C6—N1	−7.1 (4)	C8—C9—C10—N1	1.5 (4)
C10—N1—C6—C1	−76.2 (4)	C8—C9—C10—N4	−179.8 (3)
N2—N1—C6—C1	109.0 (3)	C12—N4—C10—N1	168.1 (3)
C10—N1—C6—C5	106.4 (4)	C12—N4—C10—C9	−10.5 (6)
N2—N1—C6—C5	−68.3 (3)	C10—N4—C12—O	1.4 (6)
C2—C3—C7—F2′	161 (3)	C10—N4—C12—C13	−179.0 (3)
C4—C3—C7—F2′	−20 (3)	O—C12—C13—Cl3	−143.6 (4)
C2—C3—C7—F1	−139.8 (16)	N4—C12—C13—Cl3	36.7 (5)

Hydrogen-bond geometry (Å, º)

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
N4—H4*A*···N3^{i}	0.86	2.49	3.280 (5)	153
C4—H4*B*···O^{ii}	0.93	2.53	3.349 (5)	148

Symmetry codes: (i) *x*+1, *y*, *z*; (ii) *x*, *y*+1, *z*.