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Creatinium hydrogen oxalate

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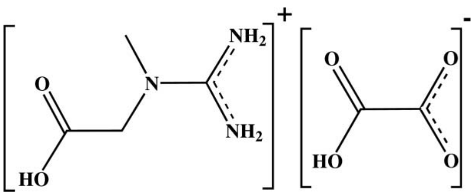
Key indicators: single-crystal X-ray study; T = 293 K; mean σ(C–C) = 0.002 Å; R factor = 0.037; wR factor = 0.102; data-to-parameter ratio = 10.1.

The crystal structure of the title compound, $C_4H_{10}N_3O_2^+\cdot C_2HO_4^-$, is stabilized by N—H⋯O and O—H⋯O hydrogen bonds. The anions are connected by an O—H⋯O hydrogen bond, leading to $C(5)$ chain extending along c axis. The cations are dimerized around the corners of the unit cell, leading to an $R_2^2(14)$ ring motif. This leads to a cationic molecular aggregation at x = 0 or 1 and an anionic molecular aggregation at x = 1/2.

Related literature

For related structures see: Ali *et al.* (2011*a*,*b*); Bahadur, Kannan *et al.* (2007); Bahadur, Sivapragasam *et al.* (2007); Bahadur, Rajalakshmi *et al.* (2007). For hydrogen-bonding motifs, see Bernstein *et al.* (1995). For the biological importance of creatine, see: Cannan & Shore (1928); Greenhaff *et al.* (1993).



Experimental

Crystal data

$C_4H_{10}N_3O_2^+\cdot C_2HO_4^-$
M_r = 221.18
Monoclinic, $P2_1/c$
a = 7.1545 (4) Å
b = 12.3681 (7) Å
c = 10.5151 (6) Å
β = 94.18 (1)°
V = 927.98 (9) Å^3
Z = 4
Mo $K\alpha$ radiation
μ = 0.14 mm^{-1}
T = 293 K
0.24 × 0.22 × 0.18 mm

Data collection

Bruker SMART APEX CCD area-detector diffractometer
8631 measured reflections
1641 independent reflections
1587 reflections with $I > 2\sigma(I)$
$R_{\rm int}$ = 0.018

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.037
$wR(F^2)$ = 0.102
S = 1.08
1641 reflections
162 parameters
1 restraint
H atoms treated by a mixture of independent and constrained refinement
$\Delta\rho_{\rm max}$ = 0.25 e Å^{-3}
$\Delta\rho_{\rm min}$ = −0.23 e Å^{-3}

Table 1
Hydrogen-bond geometry (Å, °).

D—H⋯A	D—H	H⋯A	D⋯A	D—H⋯A
N2—H1N⋯O1$^{\rm i}$	0.82 (3)	2.37 (3)	3.094 (2)	148 (2)
N2—H2N⋯O11$^{\rm ii}$	0.85 (2)	2.08 (2)	2.910 (2)	168 (2)
N3—H3N⋯O1$^{\rm iii}$	0.85 (2)	2.18 (2)	2.985 (2)	157 (2)
N3—H4N⋯O14	0.86 (2)	2.04 (2)	2.903 (2)	174 (2)
O2—H2⋯O12$^{\rm iv}$	0.94 (3)	1.60 (3)	2.538 (2)	173 (2)
O13—H13O⋯O11$^{\rm v}$	0.90 (3)	1.72 (3)	2.605 (1)	168 (2)

Symmetry codes: (i) $-x+2, -y+2, -z$; (ii) $x+1, -y+\frac{3}{2}, z-\frac{1}{2}$; (iii) $-x+2, y-\frac{1}{2}, -z+\frac{1}{2}$; (iv) $x, y, z-1$; (v) $x, -y+\frac{3}{2}, z-\frac{1}{2}$.

Data collection: *SMART* (Bruker, 2001); cell refinement: *SAINT* (Bruker, 2001); data reduction: *SAINT*; program(s) used to solve structure: *SHELXTL/PC* (Sheldrick, 2008); program(s) used to refine structure: *SHELXTL/PC*; molecular graphics: *PLATON* (Spek, 2009); software used to prepare material for publication: *SHELXTL/PC*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HG5159).

References

Ali, A. J., Athimoolam, S. & Bahadur, S. A. (2011*a*). *Acta Cryst.* E**67**, o1376.
Ali, A. J., Athimoolam, S. & Bahadur, S. A. (2011*b*). *Acta Cryst.* E**67**, o2905.
Bahadur, S. A., Kannan, R. S. & Sridhar, B. (2007). *Acta Cryst.* E**63**, o2387–o2389.
Bahadur, S. A., Rajalakshmi, M., Athimoolam, S., Kannan, R. S. & Ramakrishnan, V. (2007). *Acta Cryst.* E**63**, o4195.
Bahadur, S. A., Sivapragasam, S., Kannan, R. S. & Sridhar, B. (2007). *Acta Cryst.* E**63**, o1714–o1716.
Bernstein, J., Davis, R. E., Shimoni, L. & Chang, N. L. (1995). *Angew. Chem. Int. Ed. Engl.* **34**, 1555–1573.
Bruker (2001). *SAINT* and *SMART*. Bruker AXS Inc., Madison, Wisconsin, USA.
Cannan, R. K. & Shore, A. (1928). *Biochem. J.* **22**, 920–929.
Greenhaff, P. L., Casey, A., Short, A. H., Harris, R., Soderlund, K. & Hultman, E. (1993). *Clin. Sci.* **84**, 565–571.
Sheldrick, G. M. (2008). *Acta Cryst.* A**64**, 112–122.
Spek, A. L. (2009). *Acta Cryst.* D**65**, 148–155.

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Creatinium hydrogen oxalate

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S1. Comment

Creatine is a nitrogenous organic acid that occurs naturally in vertebrates and helps to supply energy to all cells in the body, primarily muscle (Cannan, 1928; Greenhaff *et al.*,1993)). We are interested in the specificity of recognition between organic acids and cretine and creatinine molecules and have reported a number of creatinine related structures (Ali *et al.*, 2011*a*,b; Bahadur, Kannan *et al.*, 2007; Bahadur, Sivapragasam *et al.*, 2007; Bahadur, Rajalakshmi *et al.*, 2007).

The asymmetric part of the title compound, (I), contains one creatinium cation and one hydrogen oxalate anion (Fig. 1). The protonation of the N site of the cation is evident from C—N bond distances. The deprotonation on the one of the –COOH groups of the oxalic acid is confirmed from that $–COO^-$ bond geometry. The planes of –COOH and $–COO^-$ groups are twisted out from each other with an angle of 25.1 (2)°. This twisting of planes may be caused due to the hydrogen bonding association and molecular aggregation. The crystal structure and the molecular aggregations are stabilized through intricate three dimensional hydrogen bonding network (Fig. 2; Table 1). All the N and O atoms of the cation and anion participate in the hydrogen bonding interactions.

Hydrogen oxalate anions are connected themselves through a O—H···O hydrogen bond leading to a linear chain C(5) motif extending along *c* axis of the unit cell (Bernstein *et al.*, 1995). Creatinium cations are dimerized around inversion centres of the unit cell, especially at the corners of the unit cell and making a ring $R_2^2(4)$ motif through N2—H1N···O1 (2 - *x*, 2 - *y*, -*z*) hydrogen bond. Also, these cationic dimers are connected themselves through another N—H···O hydrogen bond leading to a zigzag chain C(7) motif extending along *b* axis of the unit cell [N3—H3N···O1(-*x* + 2, *y* - 1/2, -*z* + 1/2)]. These interconnected cationic dimers are connected with oxalate anion leading to a zigzag chain $C_2^2(11)$ motif extending along *ac*-plane of the unit cell through N2—H2N···O11(1 + *x*, 3/2 - *y*, -1/2 + *z*) and O2—H2···O12(*x*, *y*, -1 + *z*). Another pair of N—H···O and O—H···O hydrogen bonds between cation and anion leading to a linear chain $C_2^2(12)$ motifs extending along *c* axis of the unit cell [N3—H3N···O14 and O2—H2···O12(*x*, *y*, -1 + *z*)]. Dimerization of cations and anionic chain motifs lead to cationic molecular aggregation at *x*=0 or 1 and molecular aggregation of anions at *x*=1/2. These cationic and anionic aggregations are connected further through other N—H···O hydrogen bonds leading to a three dimensional hydrogen bonding network.

S2. Experimental

The title compound was crystallized from an aqueous mixture containing creatine (0.13g) and oxalic acid (0.09g) in the stoichiometric ratio of 1:1 (20 ml of water) at room temperature by slow evaporation technique.

S3. Refinement

All the H atoms except the atoms involved in hydrogen bonds were positioned geometrically and refined using a riding model, with C—H = 0.96 ($–CH_3$) and 0.97 Å ($–CH_2$) and $U_{iso}(H) = 1.2–1.5\ U_{eq}$ (parent atom). H atoms involved in hydrogen bonds were located from differential Fourier maps and refined isotropically.

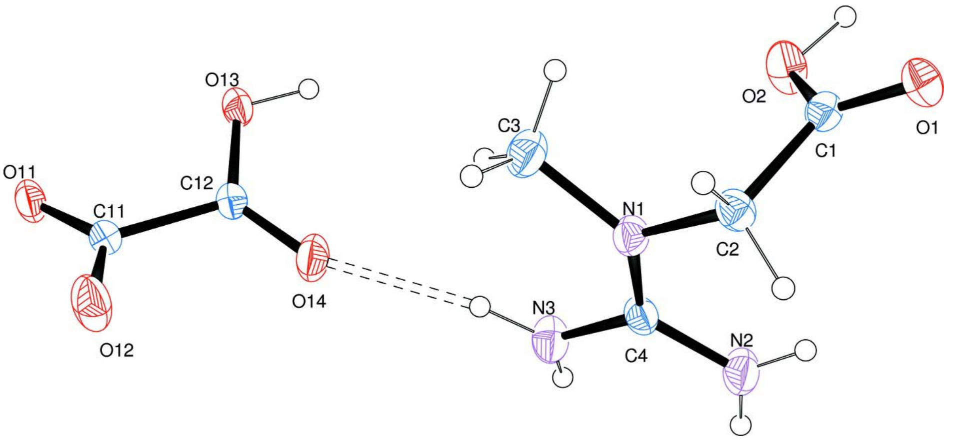

Figure 1

The molecular structure of the title compound (I) with the numbering scheme for the atoms and 50% probability displacement ellipsoids. H bonds are drawn as dashed lines.

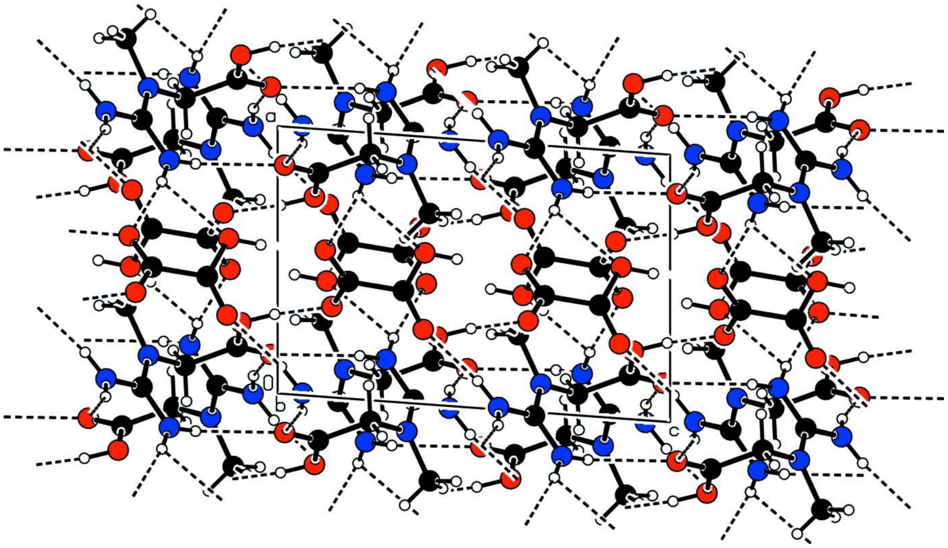

Figure 2

Packing diagram of the molecules viewed down the *b*-axis. H atoms not involved in the H-bonds (dashed lines) are omitted for clarity.

{amino[(carboxymethyl)(methyl)amino]methylidene}azanium hydrogen oxalate

Crystal data

$C_4H_{10}N_3O_2^{+}\cdot C_2HO_4^{-}$
$M_r = 221.18$
Monoclinic, $P2_1/c$
Hall symbol: -P 2ybc
$a = 7.1545$ (4) Å
$b = 12.3681$ (7) Å
$c = 10.5151$ (6) Å
$\beta = 94.18$ (1)°

$V = 927.98$ (9) Å^3
$Z = 4$
$F(000) = 464$
$D_x = 1.583$ Mg m^{-3}
Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å
Cell parameters from 3216 reflections
$\theta = 2.1$–24.7°
$\mu = 0.14$ mm^{-1}

T = 293 K
Block, colourless

0.24 × 0.22 × 0.18 mm

Data collection

Bruker SMART APEX CCD area-detector diffractometer
Radiation source: fine-focus sealed tube
Graphite monochromator
ω scans
8631 measured reflections
1641 independent reflections

1587 reflections with $I > 2\sigma(I)$
$R_{\rm int}$ = 0.018
$\theta_{\rm max}$ = 25.0°, $\theta_{\rm min}$ = 2.6°
$h = -8 \rightarrow 8$
$k = -14 \rightarrow 14$
$l = -12 \rightarrow 12$

Refinement

Refinement on F^2
Least-squares matrix: full
$R[F^2 > 2\sigma(F^2)] = 0.037$
$wR(F^2) = 0.102$
S = 1.08
1641 reflections
162 parameters
1 restraint
Primary atom site location: structure-invariant direct methods
Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring sites
H atoms treated by a mixture of independent and constrained refinement
$w = 1/[\sigma^2(F_o^2) + (0.0629P)^2 + 0.2614P]$ where $P = (F_o^2 + 2F_c^2)/3$
$(\Delta/\sigma)_{\rm max} < 0.001$
$\Delta\rho_{\rm max}$ = 0.25 e Å^{-3}
$\Delta\rho_{\rm min}$ = −0.23 e Å^{-3}
Extinction correction: *SHELXTL/PC* (Sheldrick, 2008), $Fc^*=kFc[1+0.001xFc^2\lambda^3/\sin(2\theta)]^{-1/4}$
Extinction coefficient: 0.040 (6)

Special details

Geometry. All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R-factor wR and goodness of fit S are based on F^2, conventional R-factors R are based on F, with F set to zero for negative F^2. The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R-factors(gt) *etc*. and is not relevant to the choice of reflections for refinement. R-factors based on F^2 are statistically about twice as large as those based on F, and R- factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	y	z	$U_{\rm iso}$*/$U_{\rm eq}$
C1	0.83544 (18)	1.03293 (11)	0.10352 (12)	0.0299 (3)
C2	0.92061 (19)	1.05574 (11)	0.23633 (13)	0.0326 (3)
H2A	1.0542	1.0671	0.2321	0.039*
H2B	0.8676	1.1225	0.2661	0.039*
C3	0.7146 (2)	0.97346 (13)	0.38693 (16)	0.0440 (4)
H3A	0.6828	0.9015	0.4120	0.066*
H3B	0.6190	1.0005	0.3264	0.066*
H3C	0.7240	1.0195	0.4606	0.066*
C4	1.01280 (19)	0.88981 (11)	0.34586 (12)	0.0316 (3)
N1	0.89312 (16)	0.97171 (9)	0.32919 (11)	0.0323 (3)
N2	1.1567 (2)	0.87994 (13)	0.27569 (14)	0.0482 (4)
N3	0.9923 (2)	0.81801 (10)	0.43697 (13)	0.0403 (3)
O1	0.85571 (15)	1.09547 (8)	0.01761 (9)	0.0385 (3)

O2	0.74475 (17)	0.94241 (9)	0.09369 (11)	0.0460 (3)
H1N	1.163 (4)	0.912 (2)	0.208 (3)	0.083 (8)*
H2N	1.238 (3)	0.8321 (15)	0.2955 (19)	0.051 (5)*
H3N	1.061 (3)	0.7617 (18)	0.4380 (19)	0.053 (5)*
H4N	0.914 (2)	0.8210 (14)	0.495 (2)	0.056 (6)*
H2	0.705 (4)	0.9278 (19)	0.008 (3)	0.084 (7)*
C11	0.56186 (18)	0.81947 (11)	0.82074 (12)	0.0286 (3)
C12	0.58163 (18)	0.79173 (11)	0.67945 (12)	0.0284 (3)
O11	0.45079 (14)	0.76467 (9)	0.87863 (8)	0.0384 (3)
O12	0.66030 (19)	0.89487 (11)	0.86180 (11)	0.0566 (4)
O13	0.43397 (14)	0.74502 (9)	0.62504 (9)	0.0385 (3)
O14	0.72071 (16)	0.81332 (10)	0.62893 (10)	0.0490 (4)
H13O	0.449 (3)	0.7329 (19)	0.542 (2)	0.071 (6)*

Atomic displacement parameters (Å^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
C1	0.0297 (7)	0.0292 (7)	0.0313 (7)	0.0010 (5)	0.0054 (5)	0.0002 (5)
C2	0.0375 (7)	0.0294 (7)	0.0306 (7)	−0.0027 (5)	0.0004 (5)	0.0042 (5)
C3	0.0418 (8)	0.0448 (9)	0.0472 (9)	0.0102 (7)	0.0160 (7)	0.0084 (7)
C4	0.0357 (7)	0.0355 (7)	0.0236 (6)	0.0037 (6)	0.0024 (5)	−0.0009 (5)
N1	0.0341 (6)	0.0333 (6)	0.0303 (6)	0.0037 (5)	0.0064 (5)	0.0058 (4)
N2	0.0442 (8)	0.0637 (9)	0.0384 (8)	0.0205 (7)	0.0147 (6)	0.0151 (7)
N3	0.0507 (8)	0.0353 (7)	0.0364 (7)	0.0121 (6)	0.0128 (6)	0.0077 (5)
O1	0.0498 (6)	0.0348 (5)	0.0308 (5)	−0.0016 (4)	0.0019 (4)	0.0058 (4)
O2	0.0615 (7)	0.0432 (6)	0.0335 (6)	−0.0218 (5)	0.0044 (5)	−0.0033 (5)
C11	0.0283 (6)	0.0346 (7)	0.0230 (6)	0.0007 (5)	0.0013 (5)	−0.0007 (5)
C12	0.0317 (7)	0.0304 (7)	0.0237 (6)	0.0000 (5)	0.0045 (5)	0.0012 (5)
O11	0.0429 (6)	0.0515 (6)	0.0215 (5)	−0.0093 (5)	0.0066 (4)	−0.0017 (4)
O12	0.0687 (8)	0.0683 (8)	0.0339 (6)	−0.0311 (7)	0.0101 (5)	−0.0161 (5)
O13	0.0363 (6)	0.0575 (7)	0.0220 (5)	−0.0070 (5)	0.0042 (4)	−0.0083 (4)
O14	0.0447 (6)	0.0712 (8)	0.0330 (6)	−0.0182 (5)	0.0148 (5)	−0.0060 (5)

Geometric parameters (Å, º)

C1—O1	1.2060 (17)	C4—N1	1.3296 (18)
C1—O2	1.2943 (17)	N2—H1N	0.82 (3)
C1—C2	1.5090 (19)	N2—H2N	0.85 (2)
C2—N1	1.4493 (17)	N3—H3N	0.85 (2)
C2—H2A	0.9700	N3—H4N	0.86 (2)
C2—H2B	0.9700	O2—H2	0.94 (3)
C3—N1	1.4540 (18)	C11—O12	1.2279 (17)
C3—H3A	0.9600	C11—O11	1.2377 (17)
C3—H3B	0.9600	C11—C12	1.5413 (18)
C3—H3C	0.9600	C12—O14	1.1920 (17)
C4—N2	1.3152 (19)	C12—O13	1.2996 (17)
C4—N3	1.3223 (19)	O13—H13O	0.90 (3)

O1—C1—O2	125.59 (13)	C4—N1—C2	121.16 (12)
O1—C1—C2	120.72 (12)	C4—N1—C3	122.23 (12)
O2—C1—C2	113.69 (11)	C2—N1—C3	115.92 (11)
N1—C2—C1	115.09 (11)	C4—N2—H1N	122.7 (19)
N1—C2—H2A	108.5	C4—N2—H2N	118.7 (14)
C1—C2—H2A	108.5	H1N—N2—H2N	118 (2)
N1—C2—H2B	108.5	C4—N3—H3N	117.5 (14)
C1—C2—H2B	108.5	C4—N3—H4N	126.7 (10)
H2A—C2—H2B	107.5	H3N—N3—H4N	115.7 (17)
N1—C3—H3A	109.5	C1—O2—H2	110.9 (15)
N1—C3—H3B	109.5	O12—C11—O11	127.98 (12)
H3A—C3—H3B	109.5	O12—C11—C12	114.67 (12)
N1—C3—H3C	109.5	O11—C11—C12	117.35 (11)
H3A—C3—H3C	109.5	O14—C12—O13	125.56 (12)
H3B—C3—H3C	109.5	O14—C12—C11	121.20 (12)
N2—C4—N3	118.47 (14)	O13—C12—C11	113.24 (11)
N2—C4—N1	121.30 (13)	C12—O13—H13O	110.6 (15)
N3—C4—N1	120.19 (13)		

Hydrogen-bond geometry (Å, º)

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
N2—H1*N*···O1^{i}	0.82 (3)	2.37 (3)	3.094 (2)	148 (2)
N2—H2*N*···O11^{ii}	0.85 (2)	2.08 (2)	2.910 (2)	168 (2)
N3—H3*N*···O1^{iii}	0.85 (2)	2.18 (2)	2.985 (2)	157 (2)
N3—H4*N*···O14	0.86 (2)	2.04 (2)	2.903 (2)	174 (2)
O2—H2···O12^{iv}	0.94 (3)	1.60 (3)	2.538 (2)	173 (2)
O13—H13*O*···O11^{v}	0.90 (3)	1.72 (3)	2.605 (1)	168 (2)

Symmetry codes: (i) −*x*+2, −*y*+2, −*z*; (ii) *x*+1, −*y*+3/2, *z*−1/2; (iii) −*x*+2, *y*−1/2, −*z*+1/2; (iv) *x*, *y*, *z*−1; (v) *x*, −*y*+3/2, *z*−1/2.