

# catena-Poly[silver(I)-bis[ $\mu$ -4-methyl-1*H*-1,2,4-triazole-3(4*H*)-thione- $\kappa^2$ S:S]-silver(I)-di- $\mu$ -thiocyanato- $\kappa^2$ S:N; $\kappa^2$ N:S]

Kultida Kodcharat, Chaveng Pakawatchai and Saowanit Saithong\*

Department of Chemistry and Center of Excellence for Innovation in Chemistry, Faculty of Science, Prince of Songkla University, Hat Yai, Songkhla 90112, Thailand  
Correspondence e-mail: saowanit.sa@psu.ac.th

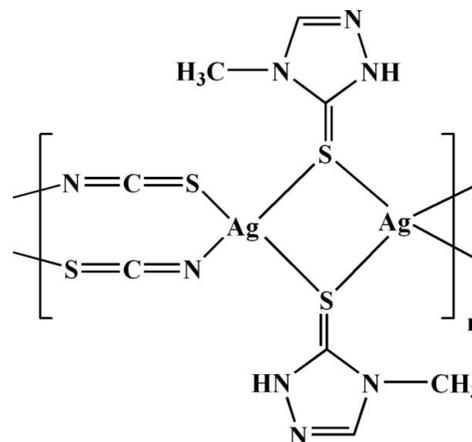
Received 5 April 2013; accepted 7 April 2013

Key indicators: single-crystal X-ray study;  $T = 293$  K; mean  $\sigma(\text{N-N}) = 0.005$  Å;  $R$  factor = 0.029;  $wR$  factor = 0.070; data-to-parameter ratio = 20.0.

In the title one-dimensional coordination polymer,  $[\text{Ag}_2(\text{NCS})_2(\text{C}_3\text{H}_5\text{N}_3\text{S})_2]_n$ , the  $\text{Ag}^{\text{I}}$  atom adopts a distorted tetrahedral  $\text{AgNS}_3$  geometry. Adjacent  $\text{Ag}^{\text{I}}$  atoms in the [001] chain are alternately linked by pairs of bridging 4-methyl-1*H*-1,2,4-triazole-3(4*H*)-thione (Hmptrz) ligands (*via* their S atoms) and double thiocyanate bridges linking through both S and N atoms ( $\mu$ -1,3-SCN). An intrachain  $\text{N-H}\cdots\text{N}$  hydrogen bond occurs between the NH group of the triazole ring and the N atom of the thiocyanate bridging ligand. A (101) sheet structure arises from interchain  $\text{S}\cdots\text{N}$  short contacts [3.239 (3) Å] involving the thiocyanate S atom and the triazole-ring N atom and possible very weak  $\pi$ - $\pi$  stacking [centroid-centroid separation = 4.0762 (18) Å] between the triazole rings.

## Related literature

For examples of complexes with multifunctional ligand donors, see: Zhang *et al.* (2009); Wang *et al.* (2011). For background to complexes containing derivatives of the 1,2,4-triazole ligand, see: Zhang *et al.* (1999); Jiang *et al.* (2011). For the thiocyanate bridging ligand, end-to-end mode, see: Vicente *et al.* (1997); Chen *et al.* (1999); Diaz *et al.* (1999); Goher *et al.* (2000); Song *et al.* (2000); Cai *et al.* (2007); Saithong *et al.* (2007).



## Experimental

### Crystal data

$[\text{Ag}_2(\text{NCS})_2(\text{C}_3\text{H}_5\text{N}_3\text{S})_2]$   
 $M_r = 562.22$   
Triclinic,  $P\bar{1}$   
 $a = 7.4842$  (6) Å  
 $b = 7.5420$  (6) Å  
 $c = 8.4262$  (7) Å  
 $\alpha = 79.985$  (2)°  
 $\beta = 84.329$  (2)°

$\gamma = 64.508$  (1)°  
 $V = 422.62$  (6) Å<sup>3</sup>  
 $Z = 1$   
Mo  $K\alpha$  radiation  
 $\mu = 2.82$  mm<sup>-1</sup>  
 $T = 293$  K  
0.31 × 0.12 × 0.05 mm

### Data collection

Bruker APEX CCD diffractometer  
Absorption correction: multi-scan  
(SADABS; Bruker, 2003)  
 $T_{\text{min}} = 0.682$ ,  $T_{\text{max}} = 0.879$

5887 measured reflections  
2083 independent reflections  
1904 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.026$

### Refinement

$R[F^2 > 2\sigma(F^2)] = 0.029$   
 $wR(F^2) = 0.070$   
 $S = 1.05$   
2083 reflections  
104 parameters  
1 restraint

H atoms treated by a mixture of independent and constrained refinement  
 $\Delta\rho_{\text{max}} = 0.79$  e Å<sup>-3</sup>  
 $\Delta\rho_{\text{min}} = -0.63$  e Å<sup>-3</sup>

**Table 1**

Selected bond lengths (Å).

Ag1—N4 <sup>i</sup>	2.354 (3)	Ag1—S1	2.5554 (8)
Ag1—S2	2.4987 (8)	Ag1—S1 <sup>ii</sup>	2.6688 (8)

Symmetry codes: (i)  $-x + 1, -y, -z + 1$ ; (ii)  $-x + 1, -y, -z$ .

**Table 2**

Hydrogen-bond geometry (Å, °).

$D\cdots H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
N1—H1 <sup>i</sup> ⋯N4 <sup>i</sup>	0.86 (2)	2.10 (2)	2.954 (4)	171 (3)

Symmetry code: (i)  $-x + 1, -y, -z + 1$ .

Data collection: SMART (Bruker, 2003); cell refinement: SAINT (Bruker, 2003); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: Mercury (Macrae *et al.*, 2008); software used to prepare material for publication: SHELXTL (Sheldrick, 2008), PLATON (Spek, 2009) and publCIF (Westrip, 2010).

Financial support from the Center of Excellent for Innovation in Chemistry (PERCH-CIC), Office of the Higher Education Commission, Ministry of Education, and Graduate School, Prince of Songkla University, are gratefully acknowledged.

---

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HB7066).

---

## References

- Bruker (2003). *SMART, SAINT and SADABS*. Bruker AXS Inc., Madison, Wisconsin, USA.
- Cai, H., Guo, Y., Li, Y. & Li, J.-G. (2007). *Acta Cryst.* **E63**, m936–m938.
- Chen, H.-J., Yang, G. & Chen, X.-M. (1999). *Acta Cryst.* **C55**, 2012–2014.
- Diaz, C., Ribas, J., Sanz, N., Solans, X. & Font-Bardi'a, M. (1999). *Inorg. Chim. Acta*, **286**, 169–174.
- Goher, M. A. S., Yang, Q.-C. & Mak, T. C. W. (2000). *Polyhedron*, **19**, 615–621.
- Jiang, Y.-L., Wang, Y.-L., Lin, J.-X., Liu, Q.-Y., Lu, Z.-H., Zhang, N., Wei, J.-J. & Li, L.-Q. (2011). *CrystEngComm*, **13**, 1697–1706.
- Macrae, C. F., Bruno, I. J., Chisholm, J. A., Edgington, P. R., McCabe, P., Pidcock, E., Rodriguez-Monge, L., Taylor, R., van de Streek, J. & Wood, P. A. (2008). *J. Appl. Cryst.* **41**, 466–470.
- Saithong, S., Pakawatchai, C. & Charmant, J. P. H. (2007). *Acta Cryst.* **E63**, m857–m858.
- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
- Song, Y., Zhu, D.-R., Zhang, K.-L., Xu, Y., Duan, C.-Y. & You, X.-Z. (2000). *Polyhedron*, **19**, 1461–1464.
- Spek, A. L. (2009). *Acta Cryst.* **D65**, 148–155.
- Vicente, R., Escuer, A., Penalba, E., Solans, X. & Bardia, M. F. (1997). *Inorg. Chim. Acta*, **255**, 7–12.
- Wang, Y.-L., Zhang, N., Liu, Q.-Y., Shan, Z.-M., Cao, R., Wang, M.-S., Luo, J.-J. & Yang, E.-L. (2011). *Cryst. Growth Des.* **11**, 130–138.
- Westrip, S. P. (2010). *J. Appl. Cryst.* **43**, 920–925.
- Zhang, H., Wang, X., Zhang, K. & Teo, B. K. (1999). *Coord. Chem. Rev.* **183**, 157–195.
- Zhang, W., Ye, H.-Y. & Xiong, R.-G. (2009). *Coord. Chem. Rev.* **253**, 2980–2997.

## supporting information

*Acta Cryst.* (2013). E69, m265–m266 [https://doi.org/10.1107/S160053681300946X]

**catena-Poly[silver(I)-bis[ $\mu$ -4-methyl-1*H*-1,2,4-triazole-3(4*H*)-thione- $\kappa^2$ S:S]-silver(I)-di- $\mu$ -thiocyanato- $\kappa^2$ S:N; $\kappa^2$ N:S]**

**Kultida Kodcharat, Chaveng Pakawatchai and Saowanit Saithong**

**S1. Comment**

One of the active areas of material research is the coordination compounds of the metal ions with the multifunctional ligands leading to the structural diversities and properties (Zhang *et al.*, 2009; Wang *et al.*, 2011). For this work, we report the mixed ligands Ag(I) complex containing multidonor atoms, 4-methyl-1,2,4-triazole-3-thiol (Hmptrz) and thiocyanate ligands. The Hmptrz is one of 1,2,4-triazole derivative ligands - based heterocyclic thioamide containing thiol group which has three potential donor atoms. Both Hmptrz and thiocyanate group are amphidentate ligands, which can bind to the metal center with either the N or S atom or both of them (Zhang *et al.*, 1999; Jiang *et al.*, 2011).

The title complex exhibits a one-dimensional chain polymeric structure and the asymmetric unit consists of one Ag(I) atom, one Hmptrz molecule and one SCN<sup>-</sup> anion. The chemical structure of this complex is shown in Scheme I and the crystal structure is depicted in Figure 1.

The Ag atom features a distorted tetrahedral environment with the range of angles from 101.00 (2) to 124.52 (3)°. Each Ag is bonded by two  $\mu_2$ -S-bridging atoms of two Hmptrz molecules with the distances of 2.5554 (8) and 2.6688 (8) Å. The other two coordination sites are occupied by S and N atoms from the different  $\mu_2$ -1,3-SCN bridges coordinated as a pair alternating bidentate end to end fashion similar to those complexes of the same thiocyanate bridges (Chen *et al.*, 1999; Diaz *et al.*, 1999; Goher *et al.*, 2000; Song *et al.*, 2000; Cai *et al.*, 2007; Saithong *et al.* 2007). The Cu—S<sub>thiocyanato</sub> and Cu—N<sub>thiocyanato</sub> bond distances are 2.4987 (8) and 2.2354 (4) Å, respectively. The SCN bond angle is almost perfectly linear [178.6 (3)°] as compare with the same  $\mu_2$ -1,3-SCN configuration mode of those complexes (Vicente *et al.*, 1997; Song *et al.*, 2000). The S2—C4 and C4—N4 distances [1.646 (3) and 1.150 (4) Å] refer to thiocyanate resonance form which indicate to a  $\pi$ -delocalized system along the metal-thiocyanate chain (Zhang *et al.*, 1999).

An infinite one-dimensional structure of this complex is based on [Ag( $\mu_2$ -Hmptrz)( $\mu_2$ -1,3-SCN)] double-bridges, in which both Hmptrz and SCN<sup>-</sup> ligands adopt the  $\mu_2$ -end-on and  $\mu_2$ -end-to-end bridging mode, respectively. As illustrated in Figure 2, the Hmptrz and thiocyanato ligands interconnect the Ag(I) ions into an infinite chain generated by the unit *c* translation running parallel to *c* axis, which consist of four-membered ring [—Ag—S—Ag—S—] and eight-membered ring [—Ag—S=C=N—Ag—S=C=N—]. In addition, The Ag...Ag separation with the distances of 3.3241 (5) Å in the four-membered ring is slightly shorter than the sum of the van der Waals radii of Ag atoms (3.44 Å), which indicates that there is the Ag...Ag interaction.

The weak intra-molecular hydrogen bonding interaction [N1—H1...N4<sup>i</sup>, (i) = -x + 1, -y, -z + 1] is found between N(1) of triazole ring and N(4) of thiocyanate bridging ligand at 2.954 (4) Å. The inter-short contact at 3.239 (3) Å arises from S2 donor of triazole ring with N2 acceptor from the thiocyanate bridge of the neighbouring adjacent chain which is smaller than the sum of S and N van der Waals radii (1.80 + 1.55 Å). In addition, the  $\pi$ ... $\pi$  stacking between the triazole rings of the neighbouring chain is observed with the centroid-centroid distance of 4.0762 (18) Å. Both of these interactions

generate the supramolecular layer interactions related by *ac*-plane. A view of intra-molecular hydrogen bonding is depicted in Figure 3 and The layered network interactions in crystal packing are shown in Figure 4.

## S2. Experimental

A mixture of  $\text{AgNO}_3$  (0.15 g, 0.88 mmol),  $\text{KSCN}$  (0.09 g, 0.87 mmol) in EtOH 30 ml was heated and stirred to 75 °C for 1 h. After that, the Hmptrz ligand (0.1 g, 0.087 mmol) was added to the mixture and further continuous stirring for 12 h. The colorless crystals of the complex were obtained after the colorless filtrate was kept to stand at room temperature for a day. The complex melts at 130–132 °C.

## S3. Refinement

All carbon H-atom of the triazole ring and the methyl group were placed in calculated positions ( $\text{C}-sp^2\text{—H} = 0.93$  and  $\text{C}-sp^3 = 0.96$  Å) and were included in the refinement in the riding-model approximation, with  $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$  and  $U_{\text{iso}}(\text{H}) = 1.5U_{\text{eq}}(\text{C})$ , respectively. The H atom of triazole ring N atom is located in a difference map and restrained,  $\text{N—H} = 0.86$  Å with  $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{N})$ .

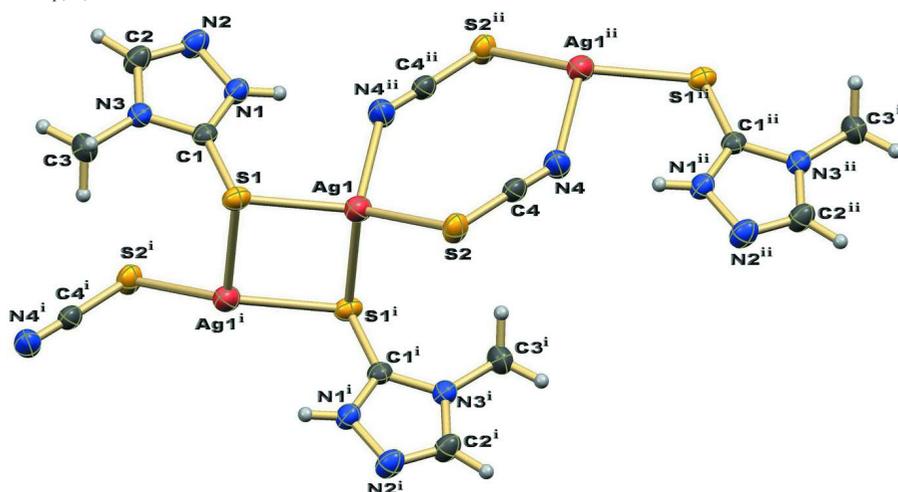


Figure 1

The structure of the title complex with displacement ellipsoids plotted at the 30% probability level.

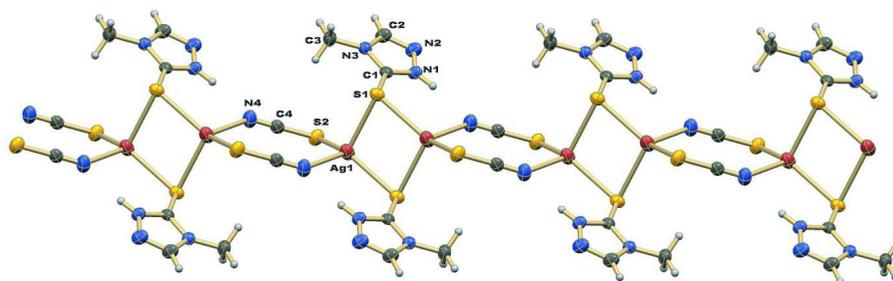


Figure 2

The one-dimensional chain of the title complex.

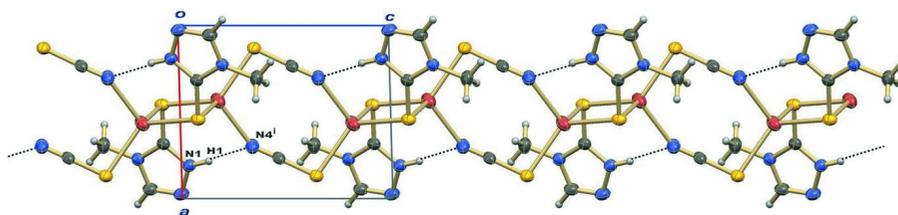


Figure 3

The intra-chain hydrogen-bonding interactions of the title complex.

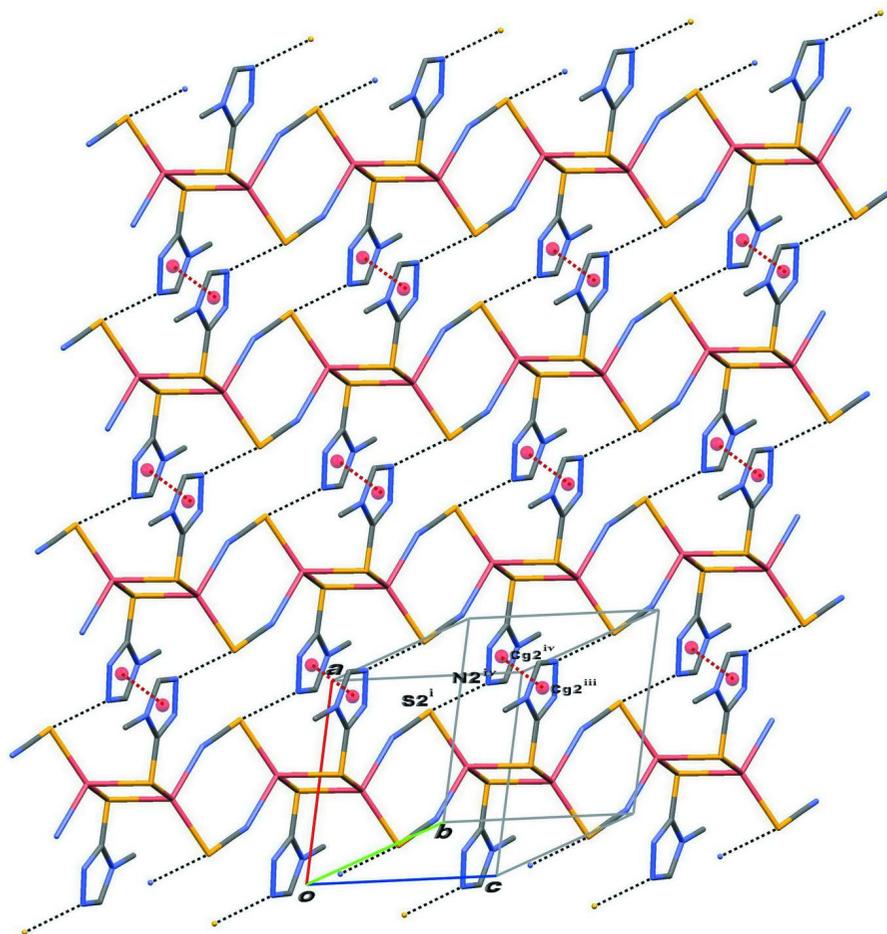


Figure 4

The two-dimensional-layer inter-interactions of the title complex. All H atoms not involving the interactions are omitted.

*catena*-Poly[silver(I)-bis[ $\mu$ -4-methyl-1*H*-1,2,4-triazole-3(4*H*)-thione- $\kappa^2$ S:S]-silver(I)-di- $\mu$ -thiocyanato- $\kappa^2$ S:N; $\kappa^2$ N:S]

#### Crystal data

[Ag<sub>2</sub>(NCS)<sub>2</sub>(C<sub>3</sub>H<sub>5</sub>N<sub>3</sub>S)<sub>2</sub>]

$M_r = 562.22$

Triclinic,  $P\bar{1}$

Hall symbol: -P 1

$a = 7.4842$  (6) Å

$b = 7.5420$  (6) Å

$c = 8.4262$  (7) Å

$\alpha = 79.985$  (2)°

$\beta = 84.329$  (2)°

$\gamma = 64.508$  (1)°

$V = 422.62$  (6) Å<sup>3</sup>

$Z = 1$

$F(000) = 272$   
 $D_x = 2.209 \text{ Mg m}^{-3}$   
 Mo  $K\alpha$  radiation,  $\lambda = 0.71073 \text{ \AA}$   
 Cell parameters from 3160 reflections  
 $\theta = 2.5\text{--}28.1^\circ$

$\mu = 2.82 \text{ mm}^{-1}$   
 $T = 293 \text{ K}$   
 Block, colourless  
 $0.31 \times 0.12 \times 0.05 \text{ mm}$

*Data collection*

Bruker APEX CCD  
 diffractometer  
 Radiation source: fine-focus sealed tube  
 Graphite monochromator  
 Frames, each covering  $0.3^\circ$  in  $\omega$  scans  
 Absorption correction: multi-scan  
 (SADABS; Bruker, 2003)  
 $T_{\min} = 0.682$ ,  $T_{\max} = 0.879$

5887 measured reflections  
 2083 independent reflections  
 1904 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.026$   
 $\theta_{\text{max}} = 28.3^\circ$ ,  $\theta_{\text{min}} = 2.5^\circ$   
 $h = -9 \rightarrow 9$   
 $k = -10 \rightarrow 10$   
 $l = -11 \rightarrow 11$

*Refinement*

Refinement on  $F^2$   
 Least-squares matrix: full  
 $R[F^2 > 2\sigma(F^2)] = 0.029$   
 $wR(F^2) = 0.070$   
 $S = 1.05$   
 2083 reflections  
 104 parameters  
 1 restraint  
 Primary atom site location: structure-invariant  
 direct methods

Secondary atom site location: difference Fourier  
 map  
 Hydrogen site location: inferred from  
 neighbouring sites  
 H atoms treated by a mixture of independent  
 and constrained refinement  
 $w = 1/[\sigma^2(F_o^2) + (0.0262P)^2 + 0.3652P]$   
 where  $P = (F_o^2 + 2F_c^2)/3$   
 $(\Delta/\sigma)_{\text{max}} < 0.001$   
 $\Delta\rho_{\text{max}} = 0.79 \text{ e \AA}^{-3}$   
 $\Delta\rho_{\text{min}} = -0.63 \text{ e \AA}^{-3}$

*Special details*

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) etc. and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	$x$	$y$	$z$	$U_{\text{iso}}^*/U_{\text{eq}}$
Ag1	0.43439 (4)	0.06942 (4)	0.18148 (3)	0.06373 (11)
S1	0.45768 (10)	0.28186 (11)	-0.08435 (10)	0.04902 (17)
N1	0.8089 (4)	0.2218 (4)	0.0367 (3)	0.0471 (5)
H1	0.784 (5)	0.168 (5)	0.130 (3)	0.057*
N2	0.9737 (4)	0.2569 (4)	-0.0040 (3)	0.0580 (7)
N3	0.7724 (3)	0.3598 (3)	-0.2066 (3)	0.0443 (5)
C1	0.6839 (4)	0.2840 (4)	-0.0829 (3)	0.0399 (5)
C2	0.9452 (4)	0.3410 (5)	-0.1505 (4)	0.0548 (7)
H2	1.0329	0.3844	-0.2120	0.066*
C3	0.6939 (6)	0.4444 (6)	-0.3672 (4)	0.0614 (8)
H3A	0.6677	0.3493	-0.4116	0.092*

H3B	0.7889	0.4779	-0.4352	0.092*
H3C	0.5733	0.5619	-0.3608	0.092*
S2	0.13129 (11)	0.14847 (14)	0.36025 (10)	0.0576 (2)
C4	0.2341 (4)	0.0469 (5)	0.5367 (4)	0.0479 (6)
N4	0.3024 (4)	-0.0245 (5)	0.6610 (3)	0.0661 (8)

*Atomic displacement parameters (Å<sup>2</sup>)*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
Ag1	0.06481 (18)	0.0881 (2)	0.05807 (16)	-0.04888 (16)	0.01183 (12)	-0.02344 (13)
S1	0.0410 (4)	0.0516 (4)	0.0602 (4)	-0.0237 (3)	-0.0010 (3)	-0.0116 (3)
N1	0.0437 (12)	0.0499 (13)	0.0494 (13)	-0.0240 (11)	-0.0036 (10)	0.0014 (10)
N2	0.0431 (13)	0.0646 (16)	0.0680 (17)	-0.0281 (12)	-0.0086 (12)	0.0049 (13)
N3	0.0451 (12)	0.0457 (12)	0.0454 (12)	-0.0225 (10)	0.0004 (9)	-0.0069 (10)
C1	0.0400 (13)	0.0357 (12)	0.0466 (13)	-0.0172 (10)	0.0015 (10)	-0.0107 (10)
C2	0.0427 (15)	0.0599 (18)	0.0645 (19)	-0.0275 (14)	-0.0007 (13)	-0.0007 (15)
C3	0.073 (2)	0.072 (2)	0.0447 (15)	-0.0374 (18)	-0.0050 (15)	-0.0028 (14)
S2	0.0427 (4)	0.0772 (5)	0.0541 (4)	-0.0308 (4)	-0.0013 (3)	0.0023 (4)
C4	0.0431 (14)	0.0621 (17)	0.0496 (15)	-0.0326 (13)	0.0082 (12)	-0.0135 (13)
N4	0.0619 (17)	0.100 (2)	0.0501 (15)	-0.0486 (17)	-0.0010 (13)	-0.0081 (15)

*Geometric parameters (Å, °)*

Ag1—N4 <sup>i</sup>	2.354 (3)	N3—C1	1.354 (3)
Ag1—S2	2.4987 (8)	N3—C2	1.363 (4)
Ag1—S1	2.5554 (8)	N3—C3	1.455 (4)
Ag1—S1 <sup>ii</sup>	2.6688 (8)	C2—H2	0.9300
Ag1—Ag1 <sup>ii</sup>	3.3241 (5)	C3—H3A	0.9600
S1—C1	1.701 (3)	C3—H3B	0.9600
S1—Ag1 <sup>ii</sup>	2.6688 (8)	C3—H3C	0.9600
N1—C1	1.325 (3)	S2—C4	1.646 (3)
N1—N2	1.369 (3)	C4—N4	1.150 (4)
N1—H1	0.860 (18)	N4—Ag1 <sup>i</sup>	2.354 (3)
N2—C2	1.278 (4)		
N4 <sup>i</sup> —Ag1—S2	107.73 (7)	C1—N3—C3	125.5 (2)
N4 <sup>i</sup> —Ag1—S1	106.91 (7)	C2—N3—C3	127.7 (3)
S2—Ag1—S1	124.52 (3)	N1—C1—N3	104.4 (2)
N4 <sup>i</sup> —Ag1—S1 <sup>ii</sup>	104.55 (9)	N1—C1—S1	129.2 (2)
S2—Ag1—S1 <sup>ii</sup>	110.41 (3)	N3—C1—S1	126.3 (2)
S1—Ag1—S1 <sup>ii</sup>	101.00 (2)	N2—C2—N3	112.6 (3)
N4 <sup>i</sup> —Ag1—Ag1 <sup>ii</sup>	115.17 (8)	N2—C2—H2	123.7
S2—Ag1—Ag1 <sup>ii</sup>	135.70 (2)	N3—C2—H2	123.7
S1—Ag1—Ag1 <sup>ii</sup>	52.010 (19)	N3—C3—H3A	109.5
S1 <sup>ii</sup> —Ag1—Ag1 <sup>ii</sup>	48.991 (18)	N3—C3—H3B	109.5
C1—S1—Ag1	104.45 (10)	H3A—C3—H3B	109.5
C1—S1—Ag1 <sup>ii</sup>	99.55 (9)	N3—C3—H3C	109.5
Ag1—S1—Ag1 <sup>ii</sup>	79.00 (2)	H3A—C3—H3C	109.5

C1—N1—N2	113.0 (2)	H3B—C3—H3C	109.5
C1—N1—H1	122 (2)	C4—S2—Ag1	100.04 (10)
N2—N1—H1	125 (2)	N4—C4—S2	178.6 (3)
C2—N2—N1	103.3 (2)	C4—N4—Ag1 <sup>i</sup>	142.0 (2)
C1—N3—C2	106.7 (2)		
N4 <sup>i</sup> —Ag1—S1—C1	11.91 (13)	C3—N3—C1—S1	-3.5 (4)
S2—Ag1—S1—C1	138.45 (9)	Ag1—S1—C1—N1	-10.7 (3)
S1 <sup>ii</sup> —Ag1—S1—C1	-97.15 (9)	Ag1 <sup>ii</sup> —S1—C1—N1	-91.7 (3)
Ag1 <sup>ii</sup> —Ag1—S1—C1	-97.15 (9)	Ag1—S1—C1—N3	172.3 (2)
N4 <sup>i</sup> —Ag1—S1—Ag1 <sup>ii</sup>	109.06 (9)	Ag1 <sup>ii</sup> —S1—C1—N3	91.3 (2)
S2—Ag1—S1—Ag1 <sup>ii</sup>	-124.40 (3)	N1—N2—C2—N3	-0.6 (4)
S1 <sup>ii</sup> —Ag1—S1—Ag1 <sup>ii</sup>	0.0	C1—N3—C2—N2	1.2 (4)
C1—N1—N2—C2	-0.2 (4)	C3—N3—C2—N2	-178.9 (3)
N2—N1—C1—N3	0.9 (3)	N4 <sup>i</sup> —Ag1—S2—C4	-24.14 (14)
N2—N1—C1—S1	-176.7 (2)	S1—Ag1—S2—C4	-150.33 (11)
C2—N3—C1—N1	-1.2 (3)	S1 <sup>ii</sup> —Ag1—S2—C4	89.46 (11)
C3—N3—C1—N1	178.9 (3)	Ag1 <sup>ii</sup> —Ag1—S2—C4	141.07 (11)
C2—N3—C1—S1	176.4 (2)		

Symmetry codes: (i)  $-x+1, -y, -z+1$ ; (ii)  $-x+1, -y, -z$ .

*Hydrogen-bond geometry* ( $\text{\AA}, ^\circ$ )

<i>D</i> —H $\cdots$ <i>A</i>	<i>D</i> —H	H $\cdots$ <i>A</i>	<i>D</i> $\cdots$ <i>A</i>	<i>D</i> —H $\cdots$ <i>A</i>
N1—H1 $\cdots$ N4 <sup>i</sup>	0.86 (2)	2.10 (2)	2.954 (4)	171 (3)

Symmetry code: (i)  $-x+1, -y, -z+1$ .