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Tetrakis(acetonitrile)copper(I) hydrogen oxalate–oxalic acid–acetonitrile (1/0.5/0.5)

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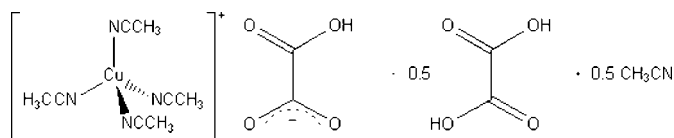
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Key indicators: single-crystal X-ray study; $T = 100$ K; mean $\sigma(\text{C}-\text{C}) = 0.004$ Å; disorder in main residue; R factor = 0.036; wR factor = 0.098; data-to-parameter ratio = 12.4.

In the title compound, $[\text{Cu}(\text{CH}_3\text{CN})_4](\text{C}_2\text{HO}_4) \cdot 0.5\text{C}_2\text{H}_2\text{O}_4 \cdot 0.5\text{CH}_3\text{CN}$, the Cu^{I} ion is coordinated by the N atoms of four acetonitrile ligands in a slightly distorted tetrahedral environment. The oxalic acid molecule lies across an inversion center. The acetonitrile solvent molecule is disordered across an inversion center and was refined with half occupancy. In the crystal, the hydrogen oxalate anions and oxalic acid molecules are linked via $\text{O}-\text{H} \cdots \text{O}$ hydrogen bonds, forming chains along $[010]$.

Related literature

For background to tetrakis(acetonitrile)copper(I) complexes, see: Morgan (1923); Heckel (1966); Kubas *et al.* (1979). For details of the affinity of nitrile ligands for Cu^{I} ions, see: Cotton *et al.* (1999). For the hard–soft acid–base theory, see: Pearson (1968). For the structure of the closely related tetrakis(acetonitrile)copper(I) tetrafluoroborate, see: Jones & Crespo (1998).



Experimental

Crystal data

$[\text{Cu}(\text{C}_2\text{H}_3\text{N})_4](\text{C}_2\text{HO}_4) \cdot 0.5\text{C}_2\text{H}_2\text{O}_4 \cdot 0.5\text{C}_2\text{H}_3\text{N}$
 $M_r = 382.33$
 Monoclinic, $P2_1/n$
 $a = 9.5637$ (4) Å
 $b = 5.5670$ (2) Å
 $c = 32.0682$ (12) Å
 $\beta = 92.901$ (2)°
 $V = 1705.16$ (11) Å³
 $Z = 4$
 Cu $K\alpha$ radiation
 $\mu = 2.15$ mm⁻¹
 $T = 100$ K
 $0.17 \times 0.14 \times 0.03$ mm

Data collection

Bruker APEXII CCD diffractometer
 Absorption correction: multi-scan (SADABS; Bruker, 2007)
 $T_{\text{min}} = 0.715$, $T_{\text{max}} = 0.938$
 8349 measured reflections
 2928 independent reflections
 2745 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.020$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.036$
 $wR(F^2) = 0.098$
 $S = 1.12$
 2928 reflections
 236 parameters
 18 restraints

H atoms treated by a mixture of independent and constrained refinement

$\Delta\rho_{\text{max}} = 0.89$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.50$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D-\text{H} \cdots A$	$D-\text{H}$	$\text{H} \cdots A$	$D \cdots A$	$D-\text{H} \cdots A$
$\text{O1}-\text{H1} \cdots \text{O4}^{\text{i}}$	0.85 (2)	1.69 (2)	2.538 (2)	176 (3)
$\text{O5}-\text{H5} \cdots \text{O3}$	0.83 (2)	1.74 (2)	2.553 (2)	165 (4)

Symmetry code: (i) $x, y + 1, z$.

Data collection: APEX2 (Bruker, 2007); cell refinement: SAINT (Bruker, 2007); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: OLEX2 (Dolomanov *et al.*, 2009); software used to prepare material for publication: CHEMDRAW (Cambridgesoft, 2003).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: LH5649).

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supporting information

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Tetrakis(acetonitrile)copper(I) hydrogen oxalate–oxalic acid–acetonitrile (1/0.5/0.5)

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S1. Comment

The tetrakis(acetonitrile)copper(I) ion, an important starting material for the synthesis of copper(I) complexes, was first synthesized as the nitrate salt (Morgan, 1923). Commonly available and easily synthesized compounds containing the $[\text{Cu}(\text{CH}_3\text{CN})_4]^+$ cation generally do contain weakly coordinating anions such as BF_4^- (Heckel, 1966) or PF_6^- (Kubas *et al.*, 1979). The structure of tetrakis(acetonitrile)copper(I) tetrafluoroborate already appears in the literature (Jones & Crespo, 1998).

In this work, we report the synthesis and crystal structure of a compound containing the $[\text{Cu}(\text{CH}_3\text{CN})_4]^+$ cation containing a potentially coordinating HC_2O_4^- anion, which does not coordinate to the metal center. This is in keeping with the known affinity of nitrile ligands for Cu^I (Cotton *et al.*, 1999), and also with the hard-soft acid-base theory (Pearson, 1968), which predicts a weak interaction between the "hard" hydrogen oxalate ligand and the "soft" Cu^+ ion.

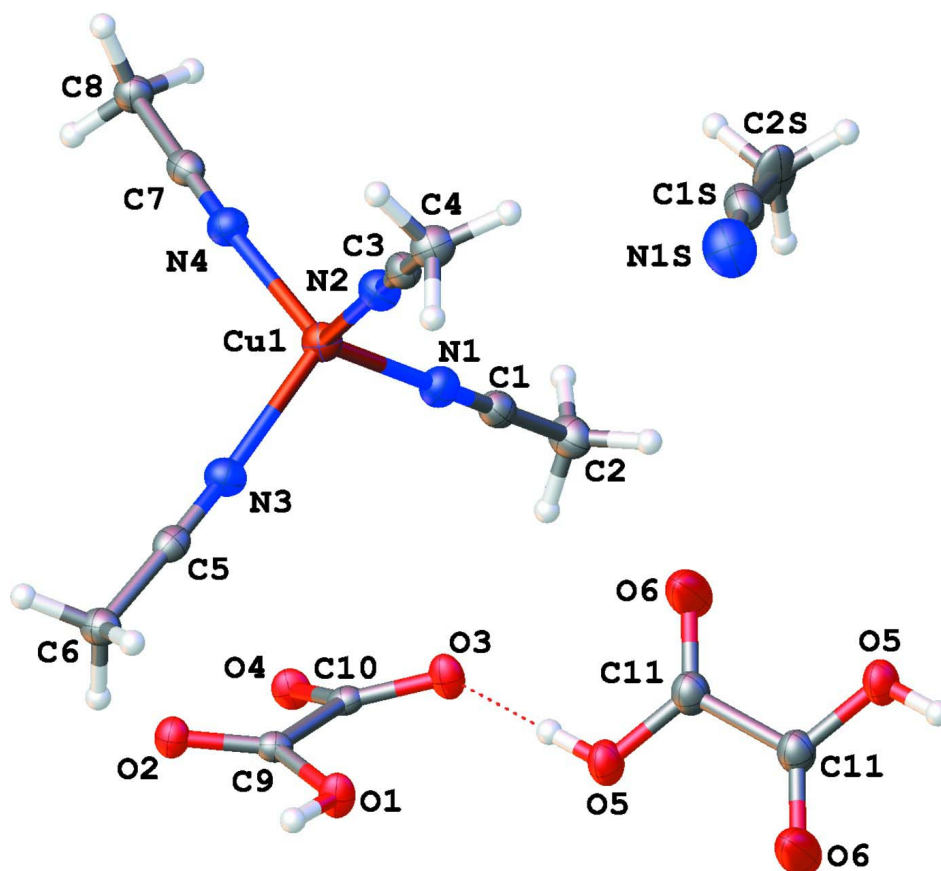
In the title compound, the $[\text{Cu}(\text{CH}_3\text{CN})_4]^+$ cation adopts a slightly distorted tetrahedral geometry. The HC_2O_4^- anion is not coordinated to the Cu^I center and the distances of the nearest O atoms to the Cu^I ion are all greater than 4.7 Å. The anion is non-planar, whereas the oxalic acid molecules are strictly planar, and reside on inversion centers. All the OH groups (in the hydrogen oxalate anion and in oxalic acid) in this structure are involved in intermolecular hydrogen bonding interactions with the carboxylate O atoms of the hydrogen oxalate anion, forming one-dimensional chains along [010]. The acetonitrile solvent molecules present in the structure are disordered, and positioned in linear channels between the $[\text{Cu}(\text{CH}_3\text{CN})_4]^+$ cations, parallel to the *b* axis.

S2. Experimental

All manipulations were carried out under nitrogen. In a 10 ml round-bottom flask, 180 mg anhydrous oxalic acid (2 mmol), 0.143 g copper(I) oxide (1 mmol) and 7 ml degassed, dry acetonitrile were stirred together. All the red Cu_2O powder dissolved in 2 min., forming a white precipitate and a clear, pale blue supernatant. After 15 min. of stirring, a copious amount of white and dark purple solids settled to the bottom of the flask. The dark purple solid was likely copper metal powder. This reaction mixture was stirred for 1 hr., then heated at 313K for 15 min., during which the white solid redissolved. Cooling to room temperature produced mm-sized, colorless, air- and moisture-sensitive, platelike crystals in 2 hrs.

S3. Refinement

H atoms were placed in calculated positions with $\text{C—H} = 0.96$ Å and included in the refinement in a riding-motion approximation with $U_{\text{iso}}(\text{H}) = 1.5U_{\text{eq}}(\text{C})$. H atoms bonded to O atoms were refined independently with $U_{\text{iso}}(\text{H}) = 1.5U_{\text{eq}}(\text{O})$.

**Figure 1**

The molecular structure of the title compound with displacement ellipsoids shown at the 50% probability level. The symmetry complete oxalic acid molecule is shown and the acetonitrile solvent molecule is half occupancy. The dotted line indicates a hydrogen bond.

Tetrakis(acetonitrile)copper(I) hydrogen oxalate–oxalic acid–acetonitrile (1/0.5/0.5)

Crystal data

$[\text{Cu}(\text{C}_2\text{H}_3\text{N})_4](\text{C}_2\text{HO}_4) \cdot 0.5\text{C}_2\text{H}_2\text{O}_4 \cdot 0.5\text{C}_2\text{H}_3\text{N}$

$M_r = 382.33$

Monoclinic, $P2_1/n$

Hall symbol: $-P\ 2_1/n$

$a = 9.5637(4)\ \text{\AA}$

$b = 5.5670(2)\ \text{\AA}$

$c = 32.0682(12)\ \text{\AA}$

$\beta = 92.901(2)^\circ$

$V = 1705.16(11)\ \text{\AA}^3$

$Z = 4$

$F(000) = 784$

$D_x = 1.489\ \text{Mg m}^{-3}$

Melting point: not measured K

Cu $K\alpha$ radiation, $\lambda = 1.54184\ \text{\AA}$

$\theta = 2.8\text{--}66.4^\circ$

$\mu = 2.15\ \text{mm}^{-1}$

$T = 100\ \text{K}$

Plate, colorless

$0.17 \times 0.14 \times 0.03\ \text{mm}$

Data collection

Bruker APEXII CCD
diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

φ and ω scans

Absorption correction: multi-scan
(*SADABS*; Bruker, 2007)

$T_{\min} = 0.715$, $T_{\max} = 0.938$

8349 measured reflections

2928 independent reflections

2745 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.020$
 $\theta_{\text{max}} = 66.5^\circ$, $\theta_{\text{min}} = 2.8^\circ$

$h = -11 \rightarrow 11$
 $k = -4 \rightarrow 6$
 $l = -37 \rightarrow 36$

Refinement

Refinement on F^2
 Least-squares matrix: full
 $R[F^2 > 2\sigma(F^2)] = 0.036$
 $wR(F^2) = 0.098$
 $S = 1.12$
 2928 reflections
 236 parameters
 18 restraints
 Primary atom site location: structure-invariant
 direct methods

Secondary atom site location: difference Fourier
 map
 Hydrogen site location: inferred from
 neighbouring sites
 H atoms treated by a mixture of independent
 and constrained refinement
 $w = 1/[\sigma^2(F_o^2) + (0.0394P)^2 + 2.6242P]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\text{max}} < 0.001$
 $\Delta\rho_{\text{max}} = 0.89 \text{ e } \text{\AA}^{-3}$
 $\Delta\rho_{\text{min}} = -0.50 \text{ e } \text{\AA}^{-3}$

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$	Occ. (<1)
Cu1	0.00357 (4)	0.91194 (7)	0.139464 (11)	0.02214 (14)	
N1	0.1201 (2)	0.6945 (4)	0.10675 (7)	0.0249 (5)	
C1	0.1869 (3)	0.5512 (5)	0.09207 (8)	0.0223 (5)	
C2	0.2716 (3)	0.3648 (5)	0.07389 (8)	0.0270 (6)	
H2A	0.3673	0.3724	0.0863	0.040*	
H2B	0.2735	0.3897	0.0437	0.040*	
H2C	0.2312	0.2070	0.0794	0.040*	
N2	-0.1196 (2)	1.1336 (4)	0.10639 (7)	0.0246 (5)	
C3	-0.1829 (3)	1.2842 (5)	0.09054 (7)	0.0221 (5)	
C4	-0.2621 (3)	1.4787 (5)	0.07048 (8)	0.0261 (6)	
H4A	-0.2024	1.6211	0.0689	0.039*	
H4B	-0.3433	1.5164	0.0868	0.039*	
H4C	-0.2941	1.4297	0.0422	0.039*	
N3	0.1258 (2)	1.1393 (4)	0.17405 (7)	0.0233 (5)	
C5	0.1743 (2)	1.2988 (5)	0.19157 (7)	0.0197 (5)	
C6	0.2332 (3)	1.5037 (5)	0.21462 (8)	0.0228 (5)	
H6A	0.2187	1.6500	0.1979	0.034*	
H6B	0.3337	1.4783	0.2204	0.034*	
H6C	0.1868	1.5207	0.2410	0.034*	
N4	-0.0979 (2)	0.6916 (4)	0.17694 (7)	0.0236 (5)	
C7	-0.1385 (3)	0.5341 (5)	0.19555 (8)	0.0209 (5)	

C8	−0.1875 (3)	0.3283 (5)	0.21895 (8)	0.0227 (5)	
H8A	−0.1808	0.1826	0.2020	0.034*	
H8B	−0.2852	0.3540	0.2257	0.034*	
H8C	−0.1295	0.3102	0.2448	0.034*	
C9	0.5095 (2)	1.0524 (4)	0.17421 (7)	0.0159 (5)	
C10	0.5127 (2)	0.7966 (4)	0.15534 (7)	0.0163 (5)	
O1	0.53399 (19)	1.2210 (3)	0.14699 (5)	0.0218 (4)	
H1	0.533 (3)	1.355 (4)	0.1595 (9)	0.033*	
O2	0.48651 (18)	1.0852 (3)	0.21057 (5)	0.0203 (4)	
O3	0.48900 (19)	0.7711 (3)	0.11715 (5)	0.0230 (4)	
O4	0.53722 (18)	0.6316 (3)	0.18140 (5)	0.0203 (4)	
C11	0.4732 (3)	0.9454 (5)	0.02028 (8)	0.0263 (6)	
O5	0.5469 (2)	1.0158 (3)	0.05293 (5)	0.0268 (4)	
H5	0.519 (3)	0.957 (6)	0.0747 (7)	0.040*	
O6	0.3702 (3)	0.8228 (5)	0.02020 (6)	0.0575 (8)	
N1S	−0.002 (6)	0.4633 (17)	0.0000 (16)	0.0417 (18)	0.50
C1S	−0.0067 (6)	0.2602 (12)	0.00027 (18)	0.0306 (12)	0.50
C2S	0.011 (3)	0.016 (4)	−0.0032 (11)	0.042 (4)	0.50
H2S1	0.1114	−0.0221	−0.0015	0.064*	0.50
H2S2	−0.0304	−0.0391	−0.0301	0.064*	0.50
H2S3	−0.0347	−0.0653	0.0195	0.064*	0.50

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Cu1	0.0241 (2)	0.0165 (2)	0.0258 (2)	0.00096 (15)	0.00143 (15)	−0.00010 (15)
N1	0.0270 (11)	0.0223 (11)	0.0256 (11)	−0.0008 (10)	0.0035 (9)	−0.0003 (9)
C1	0.0234 (13)	0.0219 (13)	0.0215 (12)	−0.0022 (11)	−0.0010 (10)	0.0013 (11)
C2	0.0266 (14)	0.0234 (14)	0.0308 (14)	0.0039 (11)	0.0001 (11)	−0.0054 (11)
N2	0.0261 (12)	0.0204 (11)	0.0271 (11)	0.0013 (10)	−0.0002 (9)	−0.0005 (9)
C3	0.0218 (12)	0.0233 (13)	0.0213 (12)	−0.0020 (11)	0.0017 (10)	−0.0042 (11)
C4	0.0273 (13)	0.0207 (13)	0.0305 (13)	0.0027 (11)	0.0019 (10)	0.0022 (11)
N3	0.0232 (11)	0.0201 (11)	0.0265 (11)	0.0000 (9)	−0.0004 (9)	0.0006 (10)
C5	0.0174 (12)	0.0210 (13)	0.0207 (12)	0.0041 (11)	0.0020 (9)	0.0044 (11)
C6	0.0211 (12)	0.0204 (13)	0.0269 (13)	0.0009 (10)	0.0014 (10)	−0.0028 (11)
N4	0.0238 (11)	0.0204 (11)	0.0267 (11)	−0.0003 (9)	0.0023 (9)	−0.0007 (10)
C7	0.0185 (12)	0.0203 (13)	0.0238 (12)	0.0022 (11)	0.0001 (9)	−0.0048 (11)
C8	0.0242 (13)	0.0199 (13)	0.0241 (12)	−0.0024 (11)	0.0012 (10)	−0.0002 (10)
C9	0.0142 (11)	0.0134 (11)	0.0198 (12)	0.0006 (9)	−0.0005 (9)	0.0016 (9)
C10	0.0131 (11)	0.0147 (12)	0.0215 (12)	0.0001 (9)	0.0027 (9)	0.0004 (10)
O1	0.0346 (10)	0.0102 (8)	0.0208 (8)	−0.0010 (8)	0.0037 (7)	0.0004 (7)
O2	0.0265 (9)	0.0156 (8)	0.0190 (9)	0.0014 (7)	0.0030 (7)	0.0000 (7)
O3	0.0353 (10)	0.0144 (8)	0.0191 (9)	−0.0016 (8)	−0.0005 (7)	−0.0009 (7)
O4	0.0279 (9)	0.0102 (8)	0.0227 (8)	0.0001 (7)	0.0002 (7)	0.0020 (7)
C11	0.0323 (15)	0.0246 (14)	0.0220 (13)	−0.0025 (12)	0.0013 (11)	0.0012 (11)
O5	0.0361 (11)	0.0260 (10)	0.0184 (9)	−0.0037 (8)	0.0004 (7)	0.0029 (8)
O6	0.0642 (16)	0.0815 (19)	0.0265 (11)	−0.0442 (15)	−0.0011 (10)	0.0054 (12)
N1S	0.048 (3)	0.033 (4)	0.043 (2)	0.009 (14)	−0.002 (2)	0.001 (13)

C1S	0.032 (3)	0.036 (3)	0.024 (2)	0.001 (3)	-0.002 (2)	-0.003 (3)
C2S	0.073 (8)	0.026 (5)	0.029 (8)	-0.019 (6)	0.017 (7)	0.007 (6)

Geometric parameters (Å, °)

Cu1—N2	1.977 (2)	C7—C8	1.460 (4)
Cu1—N1	1.981 (2)	C8—H8A	0.9800
Cu1—N4	2.002 (2)	C8—H8B	0.9800
Cu1—N3	2.017 (2)	C8—H8C	0.9800
N1—C1	1.139 (3)	C9—O2	1.211 (3)
C1—C2	1.456 (4)	C9—O1	1.311 (3)
C2—H2A	0.9800	C9—C10	1.548 (3)
C2—H2B	0.9800	C10—O3	1.242 (3)
C2—H2C	0.9800	C10—O4	1.256 (3)
N2—C3	1.139 (3)	O1—H1	0.849 (18)
C3—C4	1.452 (4)	C11—O6	1.198 (4)
C4—H4A	0.9800	C11—O5	1.293 (3)
C4—H4B	0.9800	C11—C11 ⁱ	1.547 (5)
C4—H4C	0.9800	O5—H5	0.829 (18)
N3—C5	1.137 (3)	N1S—C1S	1.132 (12)
C5—C6	1.457 (4)	C1S—C2S	1.38 (3)
C6—H6A	0.9800	C2S—H2S1	0.9800
C6—H6B	0.9800	C2S—H2S2	0.9800
C6—H6C	0.9800	C2S—H2S3	0.9800
N4—C7	1.140 (3)		
N2—Cu1—N1	115.66 (9)	H6A—C6—H6C	109.5
N2—Cu1—N4	114.26 (9)	H6B—C6—H6C	109.5
N1—Cu1—N4	104.25 (9)	C7—N4—Cu1	166.9 (2)
N2—Cu1—N3	102.50 (9)	N4—C7—C8	178.4 (3)
N1—Cu1—N3	110.40 (9)	C7—C8—H8A	109.5
N4—Cu1—N3	109.84 (9)	C7—C8—H8B	109.5
C1—N1—Cu1	171.6 (2)	H8A—C8—H8B	109.5
N1—C1—C2	178.9 (3)	C7—C8—H8C	109.5
C1—C2—H2A	109.5	H8A—C8—H8C	109.5
C1—C2—H2B	109.5	H8B—C8—H8C	109.5
H2A—C2—H2B	109.5	O2—C9—O1	125.5 (2)
C1—C2—H2C	109.5	O2—C9—C10	121.6 (2)
H2A—C2—H2C	109.5	O1—C9—C10	112.99 (19)
H2B—C2—H2C	109.5	O3—C10—O4	126.2 (2)
C3—N2—Cu1	171.1 (2)	O3—C10—C9	119.0 (2)
N2—C3—C4	179.2 (3)	O4—C10—C9	114.79 (19)
C3—C4—H4A	109.5	C9—O1—H1	108 (2)
C3—C4—H4B	109.5	O6—C11—O5	126.1 (2)
H4A—C4—H4B	109.5	O6—C11—C11 ⁱ	122.0 (3)
C3—C4—H4C	109.5	O5—C11—C11 ⁱ	111.8 (3)
H4A—C4—H4C	109.5	C11—O5—H5	112 (2)
H4B—C4—H4C	109.5	N1S—C1S—C2S	169 (3)

C5—N3—Cu1	166.5 (2)	C1S—C2S—H2S1	109.5
N3—C5—C6	178.5 (3)	C1S—C2S—H2S2	109.5
C5—C6—H6A	109.5	H2S1—C2S—H2S2	109.5
C5—C6—H6B	109.5	C1S—C2S—H2S3	109.5
H6A—C6—H6B	109.5	H2S1—C2S—H2S3	109.5
C5—C6—H6C	109.5	H2S2—C2S—H2S3	109.5

Symmetry code: (i) $-x+1, -y+2, -z$.

Hydrogen-bond geometry (Å, °)

<i>D—H...A</i>	<i>D—H</i>	<i>H...A</i>	<i>D...A</i>	<i>D—H...A</i>
O1—H1...O4 ⁱⁱ	0.85 (2)	1.69 (2)	2.538 (2)	176 (3)
O5—H5...O3	0.83 (2)	1.74 (2)	2.553 (2)	165 (4)

Symmetry code: (ii) $x, y+1, z$.