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# 2,2'-{(1*E*,1'*E*)-[Ethane-1,2-diylbis(azanylylidene)]-bis(methanylylidene)}bis(4-iodophenol)

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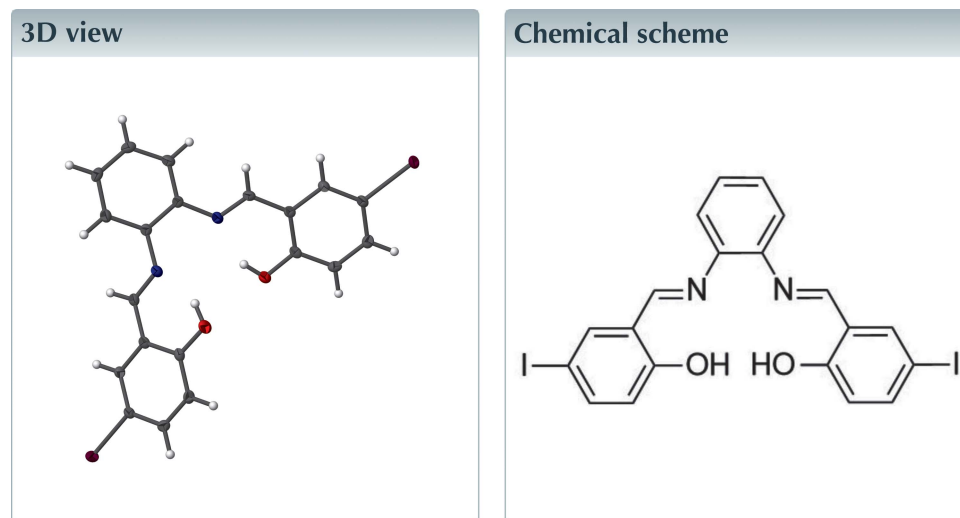
Edited by J. F. Gallagher, Dublin City University, Ireland

Keywords: crystal structure; Schiff base; hydrogen bonding;  $\pi$ - $\pi$  stacking.

CCDC reference: 2205660

Structural data: full structural data are available from [iucrdata.iucr.org](http://iucrdata.iucr.org)

The title compound, C<sub>20</sub>H<sub>14</sub>I<sub>2</sub>N<sub>2</sub>O<sub>2</sub>, a diiodo-Schiff base, crystallizes in space group *Pbca* with one molecule per asymmetric unit. The molecular structure reveals two intramolecular O—H...N hydrogen bonds that give the molecule a twisted structure with non-coplanar rings. In the crystal structure, the molecular packing is stabilized by  $\pi$ - $\pi$  stacking, hydrogen- and halogen-bonding (C—H...I; O...I) interactions.



## Structure description

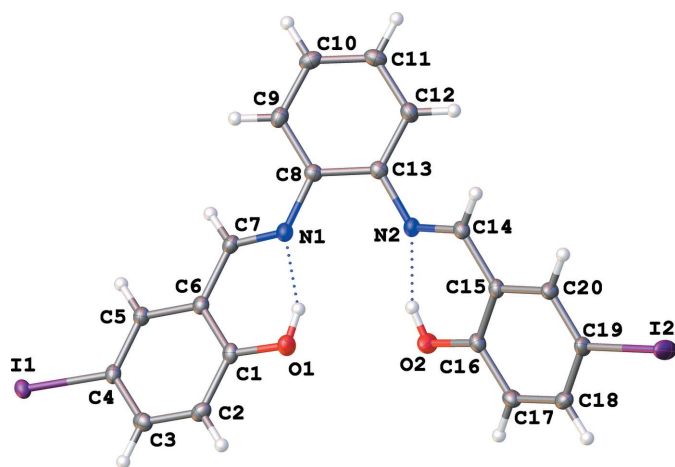
2,2'-{(1*E*,1'*E*)-[Ethane-1,2-diylbis(azanylylidene)]bis(methanylylidene)}bis(4-iodophenol) (I) (C<sub>20</sub>H<sub>14</sub>I<sub>2</sub>N<sub>2</sub>O<sub>2</sub>, Fig. 1) is a salen-type ligand under investigation in our laboratory for possible antimicrobial effects (Ceramella *et al.*, 2022) and the ability to bind to lanthanide metals, which could have other important medicinal properties (Kaczmarek *et al.*, 2018). As a result of the wide range of medical applications for such compounds, and in a continuation of our work in this area (Reimann *et al.*, 2019), the title compound (I) was prepared and its crystal structure is reported here.

Each molecule of (I) consists of a central ring (C8–C13) with *ortho* imine nitrogen atoms (N1 and N2) covalently bound through the imine C atoms, C7 and C14 respectively, to I1–Ar(O1–H1) and I2–Ar(O2–H2) rings (Fig. 1). Two intramolecular hydrogen bonds (Table 1), O1–H1...N1 [1.842 (15) Å, 152 (3)°] and O2–H2...N2 [1.806 (15) Å, 153 (3)°], result in an overall non-planar molecule with the I–Ar(OH) ring planes twisted with respect to the central ring (C8–13) plane [24.97 (7)° versus I1–Ar(O1–H1), and 39.37 (5)° versus I2–Ar(O2–H2)] (Fig. 1). The intramolecular hydrogen bonds of (I) show similarity to those of the dichlorinated Schiff base 3,5-dichloro-*N*-[2-(methylthio)phenyl]salicylaldehyde (Hamaker *et al.*, 2010). Halogen bonding (I1...O2) and slipped  $\pi$ - $\pi$  stacking interactions stabilize the packing pattern. Along the *a*-axis direction, adjacent molecules related by 2<sub>1</sub> screw-axis symmetry form



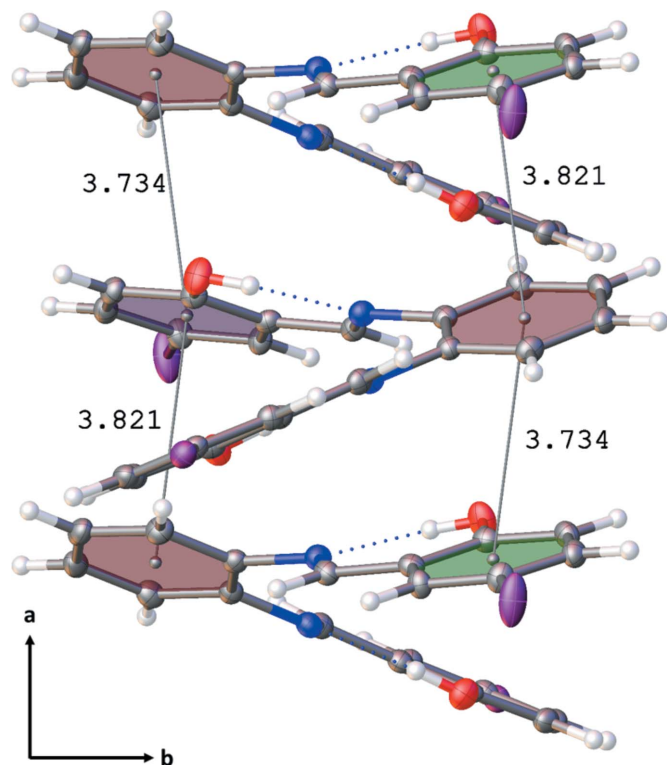
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**Figure 1**  
Molecular structure of (I) depicting the two intramolecular hydrogen bonds (O1–H1···N1 and O2–H2···N2). Displacement ellipsoids are shown at the 50% probability level

2-stacks/molecule of non-coplanar and alternating central ring (C8–C13) to I2–Ar(O2–H2) ring interactions (Fig. 2). This results in alternating longer [3.821 (13) Å] and shorter [3.734 (13) Å] central ring (C8–C13) centroid to I2–Ar(O2–H2) ring centroid distances with the corresponding alternating centroids slipped by 1.583 (4) and 1.548 (3) Å with respect to each other (Fig. 2). The additional hydrogen bonding provides three shorter and one longer C–H···I type interactions in which I2 has a larger displacement ellipsoid than I1 [C18–



**Figure 2**  
Diagram of the basic packing motif along the *a*-axis direction depicting the slipped ring  $\pi$ - $\pi$  interactions.

**Table 1**  
Hydrogen-bond geometry (Å, °).

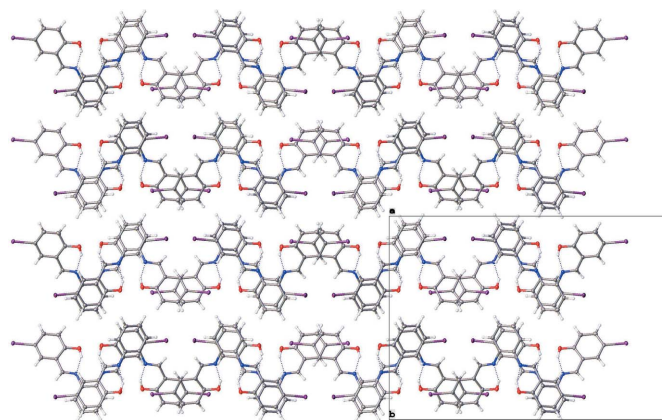
<i>D</i> –H··· <i>A</i>	<i>D</i> –H	H··· <i>A</i>	<i>D</i> ··· <i>A</i>	<i>D</i> –H··· <i>A</i>
C2–H2 <i>A</i> ···I2 <sup>i</sup>	0.95	3.32	4.009 (2)	131
C5–H5···I2 <sup>ii</sup>	0.95	3.17	4.061 (2)	156
C7–H7···I2 <sup>ii</sup>	0.95	3.23	4.094 (2)	152
C18–H18···I1 <sup>iii</sup>	0.95	3.16	4.066 (2)	159
O1–H1···N1	0.84 (1)	1.84 (2)	2.609 (3)	152 (3)
O2–H2···N2	0.84 (1)	1.81 (2)	2.580 (3)	153 (3)

Symmetry codes: (i)  $-x + 1, -y, -z + 1$ ; (ii)  $x, -y + \frac{1}{2}, z + \frac{1}{2}$ ; (iii)  $-x + \frac{1}{2}, -y, z - \frac{1}{2}$ .

H18···I1 = 3.16 Å (159°), C5–H5···I2 = 3.17 Å (156°), C7–H7···I2 = 3.23 Å (152°), and C2–H2···I2 3.32 Å (131°)]. The two C–I bond lengths of (I) are similar to the C–I bond of 2,3,5,6-tetrafluoro-1,4-diiodobenzene (Tan & Tiekink, 2019). Along the *b*-axis direction, adjacent molecules appear to be arranged in a H (head, central C8–C13 ring) to T [tail, I1–Ar(O1–H1) and I2–Ar(O2–H2) rings] repeating pattern with the I1 and I2 atoms of the I1–Ar(O1–H1) and I2–Ar(O2–H2) rings interdigitated along the *c*-axis direction (Fig. 3).

### Synthesis and crystallization

To a solution of 2-hydroxy-5-iodobenzaldehyde (4.60 g, 18.5 mmol) in ethanol (200 ml) was added an ethanol (10 ml) solution of *o*-phenylenediamine (1.00 g, 9.20 mmol), and the reaction mixture brought to reflux with vigorous stirring for 2 h. Upon cooling, the title compound (I) precipitated as an orange solid, and was filtered, washed with ethanol and dried under vacuum. Crystals of (I) suitable for single-crystal X-ray diffraction were grown from acetone layered with hexane. Yield: 4.37 g (83%), m.p. 212–214°C. <sup>1</sup>H NMR (500 MHz, CDCl<sub>3</sub>)  $\delta$  8.54 (s, 2H), 7.73–7.54 (m, 4H), 7.37 (q, *J* = 3.2 Hz, 2H), 7.22 (q, *J* = 3.2 Hz, 2H), 6.84 (d, *J* = 9.0 Hz, 2H). [Note: phenolic H atoms (2Hs) undergo rapid exchange thus their NMR signals are broadened into the baseline beyond recognition as reported (Charisiadis *et al.*, 2014).] <sup>13</sup>C NMR



**Figure 3**  
Molecular packing view of (I) along the *a*-axis direction. Displacement ellipsoids are shown at the 50% probability level.

**Table 2**  
Experimental details.

Crystal data	
Chemical formula	C <sub>20</sub> H <sub>14</sub> I <sub>2</sub> N <sub>2</sub> O <sub>2</sub>
<i>M<sub>r</sub></i>	568.13
Crystal system, space group	Orthorhombic, <i>Pbca</i>
Temperature (K)	112
<i>a</i> , <i>b</i> , <i>c</i> (Å)	7.4776 (1), 19.1040 (2), 26.0919 (3)
<i>V</i> (Å <sup>3</sup> )	3727.28 (8)
<i>Z</i>	8
Radiation type	Mo <i>K</i> α
μ (mm <sup>-1</sup> )	3.39
Crystal size (mm)	0.30 × 0.26 × 0.12
Data collection	
Diffractometer	XtaLAB Synergy, Single source at offset/far, HyPix3000
Absorption correction	Gaussian ( <i>CrysAlis PRO</i> ; Rigaku OD, 2019)
<i>T<sub>min</sub></i> , <i>T<sub>max</sub></i>	0.300, 1.000
No. of measured, independent and observed [ <i>I</i> > 2σ( <i>I</i> )] reflections	34066, 4062, 3774
<i>R<sub>int</sub></i>	0.033
(sin θ/λ) <sub>max</sub> (Å <sup>-1</sup> )	0.648
Refinement	
<i>R</i> [ <i>F</i> <sup>2</sup> > 2σ( <i>F</i> <sup>2</sup> )], <i>wR</i> ( <i>F</i> <sup>2</sup> ), <i>S</i>	0.020, 0.044, 1.09
No. of reflections	4062
No. of parameters	242
No. of restraints	2
H-atom treatment	H atoms treated by a mixture of independent and constrained refinement
Δρ <sub>max</sub> , Δρ <sub>min</sub> (e Å <sup>-3</sup> )	0.66, -0.92

Computer programs: *CrysAlis PRO* (Rigaku OD, 2019), *SHELXT2014/5* (Sheldrick, 2015a), *SHELXL* (Sheldrick, 2015b), and *OLEX2* (Dolomanov *et al.*, 2009).

(500 MHz, CDCl<sub>3</sub>) δ 162.30, 161.11, 142.25, 141.81, 140.46, 128.32, 121.46, 120.15, 119.63, 79.69. MALDI-TOF MS:

monoisotopic *m/z* calculated for [*M* + H]<sup>+</sup>: 568.9; observed: 568.8.

## Refinement

Crystal data, data collection and structure refinement details of (I) are summarized in Table 2.

## Funding information

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## full crystallographic data

*IUCrData* (2022). 7, x220895 [https://doi.org/10.1107/S2414314622008951]

## 2,2'-{(1*E*,1'*E*)-[Ethane-1,2-diylbis(azanylylidene)]bis(methanylylidene)}bis(4-iodophenol)

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(I)

### Crystal data

$C_{20}H_{14}I_2N_2O_2$

$M_r = 568.13$

Orthorhombic, *Pbca*

$a = 7.4776$  (1) Å

$b = 19.1040$  (2) Å

$c = 26.0919$  (3) Å

$V = 3727.28$  (8) Å<sup>3</sup>

$Z = 8$

$F(000) = 2160$

$D_x = 2.025$  Mg m<sup>-3</sup>

Mo  $K\alpha$  radiation,  $\lambda = 0.71073$  Å

Cell parameters from 25879 reflections

$\theta = 1.5$ – $27.4^\circ$

$\mu = 3.39$  mm<sup>-1</sup>

$T = 112$  K

Rect. Prism, orange

$0.30 \times 0.26 \times 0.12$  mm

### Data collection

XtaLAB Synergy, Single source at offset/far,

HyPix3000

diffractometer

Radiation source: micro-focus sealed X-ray

tube, PhotonJet (Mo) X-ray Source

Mirror monochromator

Detector resolution: 10.0000 pixels mm<sup>-1</sup>

$\omega$  scans

Absorption correction: gaussian

(CrysAlisPro; Rigaku OD, 2019)

$T_{\min} = 0.300$ ,  $T_{\max} = 1.000$

34066 measured reflections

4062 independent reflections

3774 reflections with  $I > 2\sigma(I)$

$R_{\text{int}} = 0.033$

$\theta_{\max} = 27.4^\circ$ ,  $\theta_{\min} = 1.6^\circ$

$h = -9 \rightarrow 9$

$k = -23 \rightarrow 24$

$l = -32 \rightarrow 32$

### Refinement

Refinement on  $F^2$

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.020$

$wR(F^2) = 0.044$

$S = 1.09$

4062 reflections

242 parameters

2 restraints

Primary atom site location: dual

Hydrogen site location: mixed

H atoms treated by a mixture of independent and constrained refinement

$w = 1/[\sigma^2(F_o^2) + (0.0133P)^2 + 5.3948P]$

where  $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{\max} = 0.002$

$\Delta\rho_{\max} = 0.66$  e Å<sup>-3</sup>

$\Delta\rho_{\min} = -0.92$  e Å<sup>-3</sup>

Extinction correction: SHELXL-2016/6

(Sheldrick 2015b,

$F_c^* = kF_c[1 + 0.001x F_c^2 \lambda^3 / \sin(2\theta)]^{-1/4}$

Extinction coefficient: 0.000120 (17)

*Special details*

**Geometry.** All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
I1	0.48524 (2)	0.12444 (2)	0.85138 (2)	0.01709 (5)
I2	0.28304 (3)	0.11138 (2)	0.30070 (2)	0.03584 (7)
O1	0.4758 (2)	0.15083 (9)	0.61326 (6)	0.0205 (4)
O2	0.1200 (2)	0.13607 (9)	0.53361 (6)	0.0233 (4)
N1	0.3343 (3)	0.27475 (10)	0.62428 (7)	0.0162 (4)
N2	0.1963 (3)	0.26760 (10)	0.52825 (7)	0.0163 (4)
C1	0.4718 (3)	0.14643 (12)	0.66467 (9)	0.0159 (5)
C2	0.5299 (3)	0.08468 (12)	0.68798 (9)	0.0183 (5)
H2A	0.569184	0.046666	0.667387	0.022*
C3	0.5307 (3)	0.07839 (12)	0.74077 (9)	0.0181 (5)
H3	0.571539	0.036324	0.756275	0.022*
C4	0.4721 (3)	0.13336 (12)	0.77109 (9)	0.0158 (5)
C5	0.4097 (3)	0.19424 (12)	0.74922 (9)	0.0162 (5)
H5	0.367238	0.231184	0.770347	0.019*
C6	0.4087 (3)	0.20174 (12)	0.69545 (8)	0.0154 (5)
C7	0.3358 (3)	0.26551 (12)	0.67317 (8)	0.0167 (5)
H7	0.288625	0.300786	0.694995	0.020*
C8	0.2667 (3)	0.33732 (12)	0.60266 (8)	0.0152 (5)
C9	0.2710 (3)	0.40210 (13)	0.62775 (9)	0.0190 (5)
H9	0.316488	0.405051	0.661683	0.023*
C10	0.2096 (3)	0.46185 (12)	0.60350 (9)	0.0201 (5)
H10	0.212521	0.505533	0.620872	0.024*
C11	0.1437 (3)	0.45812 (12)	0.55393 (9)	0.0199 (5)
H11	0.100415	0.499165	0.537540	0.024*
C12	0.1408 (3)	0.39464 (12)	0.52821 (9)	0.0192 (5)
H12	0.095597	0.392368	0.494231	0.023*
C13	0.2040 (3)	0.33413 (12)	0.55195 (8)	0.0154 (5)
C14	0.2275 (3)	0.26073 (12)	0.47991 (8)	0.0161 (5)
H14	0.260311	0.300477	0.460144	0.019*
C15	0.2129 (3)	0.19260 (12)	0.45542 (8)	0.0156 (5)
C16	0.1605 (3)	0.13285 (12)	0.48342 (9)	0.0165 (5)
C17	0.1499 (3)	0.06804 (12)	0.45880 (9)	0.0180 (5)
H17	0.115967	0.027658	0.477692	0.022*
C18	0.1884 (3)	0.06224 (12)	0.40725 (9)	0.0186 (5)
H18	0.181371	0.018057	0.390663	0.022*
C19	0.2375 (3)	0.12147 (12)	0.37979 (9)	0.0183 (5)
C20	0.2510 (3)	0.18604 (12)	0.40303 (8)	0.0168 (5)
H20	0.285889	0.225866	0.383654	0.020*
H1	0.428 (3)	0.1892 (8)	0.6061 (10)	0.020*

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H2                    0.139 (3)                    0.1776 (7)                    0.5419 (9)                    0.020\*

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*Atomic displacement parameters (Å<sup>2</sup>)*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
I1	0.01938 (8)	0.01833 (8)	0.01356 (8)	−0.00151 (6)	−0.00173 (6)	0.00283 (5)
I2	0.07494 (16)	0.01979 (9)	0.01279 (9)	0.00557 (9)	0.00604 (8)	−0.00114 (6)
O1	0.0249 (9)	0.0219 (9)	0.0145 (8)	0.0034 (7)	0.0003 (7)	−0.0015 (7)
O2	0.0358 (10)	0.0204 (9)	0.0136 (8)	−0.0054 (8)	0.0043 (7)	0.0029 (7)
N1	0.0167 (10)	0.0177 (10)	0.0142 (9)	−0.0021 (8)	−0.0008 (8)	0.0004 (8)
N2	0.0172 (10)	0.0182 (10)	0.0135 (9)	−0.0003 (8)	−0.0012 (8)	−0.0005 (8)
C1	0.0121 (11)	0.0218 (12)	0.0138 (11)	−0.0023 (9)	−0.0012 (9)	−0.0008 (9)
C2	0.0162 (11)	0.0177 (12)	0.0211 (12)	0.0006 (9)	0.0006 (9)	−0.0046 (9)
C3	0.0149 (11)	0.0180 (12)	0.0215 (12)	0.0000 (9)	−0.0022 (9)	0.0022 (10)
C4	0.0161 (11)	0.0187 (11)	0.0127 (11)	−0.0035 (9)	−0.0009 (9)	0.0001 (9)
C5	0.0167 (11)	0.0181 (11)	0.0138 (10)	−0.0004 (10)	−0.0006 (9)	−0.0010 (9)
C6	0.0145 (11)	0.0170 (11)	0.0148 (11)	−0.0020 (9)	−0.0013 (9)	−0.0004 (9)
C7	0.0180 (11)	0.0171 (11)	0.0150 (11)	−0.0018 (9)	−0.0005 (9)	−0.0033 (9)
C8	0.0152 (11)	0.0155 (11)	0.0149 (11)	0.0001 (9)	0.0018 (9)	0.0008 (9)
C9	0.0203 (12)	0.0209 (12)	0.0157 (11)	−0.0029 (10)	0.0003 (9)	−0.0025 (10)
C10	0.0192 (12)	0.0163 (11)	0.0248 (12)	−0.0009 (10)	0.0035 (10)	−0.0042 (10)
C11	0.0195 (12)	0.0161 (11)	0.0241 (12)	0.0030 (10)	0.0025 (10)	0.0041 (10)
C12	0.0205 (12)	0.0206 (12)	0.0165 (11)	0.0016 (10)	−0.0017 (10)	0.0020 (9)
C13	0.0148 (11)	0.0165 (11)	0.0149 (11)	−0.0004 (9)	0.0020 (9)	0.0010 (9)
C14	0.0163 (11)	0.0170 (11)	0.0150 (11)	−0.0010 (9)	−0.0016 (9)	0.0023 (9)
C15	0.0147 (11)	0.0186 (11)	0.0135 (11)	−0.0010 (9)	−0.0017 (9)	0.0006 (9)
C16	0.0165 (11)	0.0193 (12)	0.0136 (11)	−0.0004 (9)	−0.0006 (9)	0.0029 (9)
C17	0.0196 (12)	0.0165 (11)	0.0179 (11)	−0.0027 (10)	−0.0009 (10)	0.0036 (9)
C18	0.0202 (12)	0.0157 (11)	0.0199 (12)	−0.0005 (9)	−0.0027 (10)	−0.0017 (9)
C19	0.0244 (12)	0.0188 (12)	0.0117 (11)	0.0023 (10)	0.0005 (9)	−0.0001 (9)
C20	0.0215 (12)	0.0151 (11)	0.0138 (11)	−0.0014 (9)	0.0009 (9)	0.0015 (9)

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*Geometric parameters (Å, °)*

I1—C4	2.104 (2)	C8—C9	1.401 (3)
I2—C19	2.100 (2)	C8—C13	1.405 (3)
O1—C1	1.344 (3)	C9—H9	0.9500
O1—H1	0.835 (10)	C9—C10	1.384 (3)
O2—C16	1.345 (3)	C10—H10	0.9500
O2—H2	0.836 (10)	C10—C11	1.386 (3)
N1—C7	1.288 (3)	C11—H11	0.9500
N1—C8	1.415 (3)	C11—C12	1.386 (3)
N2—C13	1.415 (3)	C12—H12	0.9500
N2—C14	1.290 (3)	C12—C13	1.394 (3)
C1—C2	1.397 (3)	C14—H14	0.9500
C1—C6	1.409 (3)	C14—C15	1.454 (3)
C2—H2A	0.9500	C15—C16	1.411 (3)
C2—C3	1.383 (3)	C15—C20	1.402 (3)

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C3—H3	0.9500	C16—C17	1.397 (3)
C3—C4	1.386 (3)	C17—H17	0.9500
C4—C5	1.377 (3)	C17—C18	1.380 (3)
C5—H5	0.9500	C18—H18	0.9500
C5—C6	1.410 (3)	C18—C19	1.389 (3)
C6—C7	1.456 (3)	C19—C20	1.378 (3)
C7—H7	0.9500	C20—H20	0.9500
C1—O1—H1	105.6 (18)	C9—C10—C11	120.2 (2)
C16—O2—H2	104.9 (18)	C11—C10—H10	119.9
C7—N1—C8	120.9 (2)	C10—C11—H11	119.9
C14—N2—C13	120.8 (2)	C10—C11—C12	120.1 (2)
O1—C1—C2	118.7 (2)	C12—C11—H11	119.9
O1—C1—C6	122.0 (2)	C11—C12—H12	119.8
C2—C1—C6	119.3 (2)	C11—C12—C13	120.3 (2)
C1—C2—H2A	119.7	C13—C12—H12	119.8
C3—C2—C1	120.6 (2)	C8—C13—N2	117.7 (2)
C3—C2—H2A	119.7	C12—C13—N2	122.5 (2)
C2—C3—H3	120.0	C12—C13—C8	119.7 (2)
C2—C3—C4	120.1 (2)	N2—C14—H14	119.8
C4—C3—H3	120.0	N2—C14—C15	120.5 (2)
C3—C4—I1	119.49 (17)	C15—C14—H14	119.8
C5—C4—I1	119.79 (17)	C16—C15—C14	121.2 (2)
C5—C4—C3	120.7 (2)	C20—C15—C14	119.6 (2)
C4—C5—H5	120.0	C20—C15—C16	119.3 (2)
C4—C5—C6	120.0 (2)	O2—C16—C15	122.0 (2)
C6—C5—H5	120.0	O2—C16—C17	118.4 (2)
C1—C6—C5	119.3 (2)	C17—C16—C15	119.6 (2)
C1—C6—C7	121.7 (2)	C16—C17—H17	119.8
C5—C6—C7	119.0 (2)	C18—C17—C16	120.5 (2)
N1—C7—C6	120.9 (2)	C18—C17—H17	119.8
N1—C7—H7	119.6	C17—C18—H18	120.2
C6—C7—H7	119.6	C17—C18—C19	119.5 (2)
C9—C8—N1	123.5 (2)	C19—C18—H18	120.2
C9—C8—C13	119.1 (2)	C18—C19—I2	118.35 (17)
C13—C8—N1	117.3 (2)	C20—C19—I2	120.16 (17)
C8—C9—H9	119.7	C20—C19—C18	121.5 (2)
C10—C9—C8	120.5 (2)	C15—C20—H20	120.2
C10—C9—H9	119.7	C19—C20—C15	119.6 (2)
C9—C10—H10	119.9	C19—C20—H20	120.2
I1—C4—C5—C6	176.96 (17)	C8—N1—C7—C6	-178.7 (2)
I2—C19—C20—C15	177.30 (17)	C8—C9—C10—C11	-0.3 (4)
O1—C1—C2—C3	178.9 (2)	C9—C8—C13—N2	-178.1 (2)
O1—C1—C6—C5	-179.3 (2)	C9—C8—C13—C12	-2.5 (3)
O1—C1—C6—C7	3.2 (3)	C9—C10—C11—C12	-0.7 (4)
O2—C16—C17—C18	179.4 (2)	C10—C11—C12—C13	0.0 (4)
N1—C8—C9—C10	177.2 (2)	C11—C12—C13—N2	177.0 (2)

N1—C8—C13—N2	6.2 (3)	C11—C12—C13—C8	1.6 (3)
N1—C8—C13—C12	-178.1 (2)	C13—N2—C14—C15	-177.7 (2)
N2—C14—C15—C16	1.5 (3)	C13—C8—C9—C10	1.8 (3)
N2—C14—C15—C20	-178.7 (2)	C14—N2—C13—C8	-145.5 (2)
C1—C2—C3—C4	0.6 (4)	C14—N2—C13—C12	38.9 (3)
C1—C6—C7—N1	-3.3 (3)	C14—C15—C16—O2	0.7 (4)
C2—C1—C6—C5	1.6 (3)	C14—C15—C16—C17	-179.2 (2)
C2—C1—C6—C7	-176.0 (2)	C14—C15—C20—C19	179.9 (2)
C2—C3—C4—I1	-177.29 (17)	C15—C16—C17—C18	-0.7 (4)
C2—C3—C4—C5	1.1 (3)	C16—C15—C20—C19	-0.3 (3)
C3—C4—C5—C6	-1.5 (3)	C16—C17—C18—C19	-0.2 (4)
C4—C5—C6—C1	0.1 (3)	C17—C18—C19—I2	-177.10 (17)
C4—C5—C6—C7	177.7 (2)	C17—C18—C19—C20	0.8 (4)
C5—C6—C7—N1	179.2 (2)	C18—C19—C20—C15	-0.6 (4)
C6—C1—C2—C3	-1.9 (3)	C20—C15—C16—O2	-179.1 (2)
C7—N1—C8—C9	29.0 (3)	C20—C15—C16—C17	1.0 (3)
C7—N1—C8—C13	-155.5 (2)		

Hydrogen-bond geometry ( $\text{\AA}$ ,  $^\circ$ )

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
C2—H2A $\cdots$ I2 <sup>i</sup>	0.95	3.32	4.009 (2)	131
C5—H5 $\cdots$ I2 <sup>ii</sup>	0.95	3.17	4.061 (2)	156
C7—H7 $\cdots$ I2 <sup>ii</sup>	0.95	3.23	4.094 (2)	152
C18—H18 $\cdots$ I1 <sup>iii</sup>	0.95	3.16	4.066 (2)	159
O1—H1 $\cdots$ N1	0.84 (1)	1.84 (2)	2.609 (3)	152 (3)
O2—H2 $\cdots$ N2	0.84 (1)	1.81 (2)	2.580 (3)	153 (3)

Symmetry codes: (i)  $-x+1, -y, -z+1$ ; (ii)  $x, -y+1/2, z+1/2$ ; (iii)  $-x+1/2, -y, z-1/2$ .