MS27-04 | SUB-MESOSCALE OXYGEN ORDERING IN NON-STOICHIOMETRIC OXYGEN ION

CONDUCTOR $PR_2NIO_{4+\delta}$

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Non-stoichiometric oxides can undergo important variations in the oxygen stoichiometry, enabling to tune physical and chemical properties. Combining neutron diffraction, inelastic neutron scattering and *ab initio* lattice dynamics calculations, we have recently evidenced the importance of lattice dynamics, i.e. soft phonon modes, triggering low temperature oxygen mobility in Brownmillerite type (Ca/Sr)Fe/CoO_{2.5} and Ruddlesden Popper type oxides, e.g. (Pr/Nd)₂NiO_{4+d}. [1-4]. This new concept has technological relevance for the optimization of oxygen membranes and electrolytes in SOFCs.

We report here on single-crystal synchrotron diffraction experiments on non-stoichiometric Pr_2NiO_{4+d} , uncovering unprecedented oxygen ordering up to the sub meso-scale. Complex oxygen ordering is established during a topotactic solid-state reaction already proceeding at ambient temperature, following small oxygen release. The resulting 3D-incommensurate modulated structure is described in terms of modulation vectors and related twin domain structures. Melting of the oxygen ordering around 365°C strongly amplifies the oxygen mobility, evidenced by ${}^{18}O/{}^{16}O$ oxygen isotope exchange, and associated with phonon softening manifested in a pronounced increase of thermal diffuse scattering. Our results thus strengthen the idea of a phonon assisted oxygen diffusion mechanism, which can be more generally applied to understand cooperative diffusion mechanisms.

- [1] Paulus, W., et al., J. Am. Chem. Soc, 2008. 130(47): p. 16080-16085.
- [2] Perrichon, A., et al., J. Phys. Chem. C, 2015. 119(3): p. 1557-1564.
- [3] M. Ceretti et al. J. Mater. Chem. A, 3, 42 (2015) p21140-48.
- [4] M. Ceretti et al., InorganicChemistry, 2018, 57 (8), pp.4657-4666.