In (1) the proposed hydrogen-bonding scheme consists of closed loops, with the donor-acceptor sequence:

O(6)H
$$\rightarrow$$
 O(1)H \rightarrow O(3)
H
HO(2) \leftarrow HO(4)

The hydrogen bond $H \cdots O$ distances range from 1.70 to 2.13 Å, the longest being that from the anomeric hydroxyl O(1)H to O(3), and the shortest from O(6)H to the anomeric O(1). The $O-H\cdots O$ angles are exceptionally linear, ranging from 179.6 to 179.9° (a rather unlikely circumstance because of the well-known geometrical factor).

In (2) the position of one hydrogen, H(O4), was poorly defined and placed in an unlikely position which was only 1.797 Å from H(O3). A more reasonable position is that deduced from the structure (1) which places it close to the line between O(4) and O(2) at a hydrogen-bonding distance of 1.89 Å from O(2). The hydrogen-bond scheme for structure (2) consists of infinite chains with two-link side chains, i.e.

The $H\cdots O$ bond distances then span a narrower range, from 1.89 to 2.02 Å, and the $O-H\cdots O$ angles have values from 114 to 180°. The anomeric $O(1)H\cdots O(6)$ remains the longest hydrogen bond, making an unusually small $O-H\cdots O$ angle. The anomeric O is not a hydrogen-bond acceptor.

Our interest in this problem arises from the prediction by Tse & Newton (1977) and the observations by Jeffrey, Gress & Takagi (1977) and Jeffrey & Lewis (1977) that anomeric hydroxyls in simple pyranoses are generally strong hydrogen-bond donors, but very weak acceptors. The hydrogen-bonding proposed in (1) provided one of the few strong exceptions to this hypothesis, whereas that in (2) is in better agreement, especially if the true position for H(O1) is closer to the line between O(1) and O(6), thereby shortening the $H\cdots O$ bond distance. For this reason, we favor the hydrogen-bonding scheme (2). Final resolution of this question must come from a neutron diffraction study, which must await the growth of suitable crystals, an experiment which hitherto has been unsuccessful.

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