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1*H*-Benzotriazol-3-ium (1*H*-benzotriazole- κ N³)trichloridocobaltate(II) monohydrate: a reformulation

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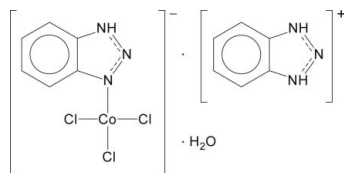
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Key indicators: single-crystal X-ray study; $T = 293$ K; mean $\sigma(\text{C}-\text{C}) = 0.006$ Å; R factor = 0.029; wR factor = 0.078; data-to-parameter ratio = 11.6.

The asymmetric unit of the title compound, $(\text{C}_6\text{H}_6\text{N}_3)\text{[CoCl}_3(\text{C}_6\text{H}_5\text{N}_3)]\cdot\text{H}_2\text{O}$, contains two crystallographically independent cations, two anions and two water molecules. The structure has been reported previously [Zhang, Li, Wang, Xie, Wang & Shen (2004). *Acta Cryst E* **60**, m498–m500] as a neutral cobalt(III) complex accompanied by unprotonated benzotriazole molecules and here has been redetermined as an anionic cobalt(II) complex accompanied by protonated benzotriazole cations. For both complex anions, a tetrahedral CoNCl_3 geometry arises. A network of $\text{N}-\text{H}\cdots\text{O}$, $\text{N}-\text{H}\cdots\text{Cl}$, $\text{O}-\text{H}\cdots\text{O}$ and $\text{O}-\text{H}\cdots\text{Cl}$ hydrogen bonds helps to establish the packing.

Related literature

For the previous structure, see: Zhang *et al.* (2004). For related Co(II)-containing structures, see: Hahn *et al.* (1997); Krawczyk & Gdaniec (2005); Pan & Xu (2004); Zhang *et al.* (2006);



Experimental

Crystal data

 $(\text{C}_6\text{H}_6\text{N}_3)\text{[CoCl}_3(\text{C}_6\text{H}_5\text{N}_3)]\cdot\text{H}_2\text{O}$ $M_r = 422.56$ Triclinic, $P1$ $a = 6.8760$ (8) Å $b = 7.8033$ (12) Å $c = 16.657$ (3) Å $\alpha = 95.450$ (13)° $\beta = 95.962$ (8)° $\gamma = 103.989$ (9)° $V = 855.8$ (2) Å³ $Z = 2$ Mo $K\alpha$ radiation $\mu = 1.48$ mm⁻¹ $T = 293$ (2) K

0.40 × 0.40 × 0.20 mm

Data collection

Bruker P4 diffractometer
Absorption correction: ψ scan
(North *et al.*, 1968)
 $T_{\min} = 0.611$, $T_{\max} = 0.744$
4906 measured reflections
4846 independent reflections

4526 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.078$
3 standard reflections
every 100 reflections
intensity decay: none

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.029$
 $wR(F^2) = 0.078$
 $S = 1.08$
4846 reflections
417 parameters
3 restraints

H-atom parameters constrained
 $\Delta\rho_{\text{max}} = 0.34$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.26$ e Å⁻³
Absolute structure: Flack (1983),
920 Friedel pairs
Flack parameter: -0.008 (13)

Table 1

Selected bond lengths (Å).

Co1—Cl1	2.2649 (13)	Co2—Cl4	2.2441 (13)
Co1—Cl2	2.2616 (13)	Co2—Cl5	2.2450 (15)
Co1—Cl3	2.2195 (16)	Co2—Cl6	2.2584 (16)
Co1—N1	2.027 (3)	Co2—N11	2.022 (3)

Table 2

Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
O1—H1B \cdots Cl6	0.82	2.79	3.279 (5)	120
O2—H2A \cdots O1 ⁱ	0.82	2.41	3.142 (7)	148
O2—H2B \cdots Cl6	0.82	2.52	3.295 (5)	159
N3—H3 \cdots O2 ⁱⁱ	0.86	1.87	2.721 (4)	173
N13—H13 \cdots Cl2 ⁱⁱⁱ	0.86	2.28	3.127 (3)	169
N21—H21 \cdots Cl4	0.86	2.80	3.458 (4)	134
N21—H21 \cdots Cl6 ^{iv}	0.86	2.75	3.316 (4)	125
N23—H23 \cdots Cl1 ^v	0.86	2.30	3.123 (3)	160
N31—H31 \cdots Cl3 ^{vi}	0.86	2.61	3.214 (4)	128
N33—H33 \cdots O1 ⁱ	0.86	1.79	2.646 (5)	170

Symmetry codes: (i) $x, y + 1, z$; (ii) $x, y - 1, z$; (iii) $x, y, z - 1$; (iv) $x + 1, y, z$; (v) $x + 1, y + 1, z$; (vi) $x - 1, y + 1, z - 1$.

Data collection: XSCANS (Bruker, 1997); cell refinement: XSCANS; data reduction: XSCANS; program(s) used to solve structure: SHELXTL (Bruker, 2003); program(s) used to refine structure: SHELXTL; molecular graphics: SHELXTL; software used to prepare material for publication: PLATON (Spek, 2003).

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HB2664).

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supporting information

Acta Cryst. (2008). E64, m58 [https://doi.org/10.1107/S1600536807063131]

1*H*-Benzotriazol-3-ium (1*H*-benzotriazole- κ N³)trichloridocobaltate(II) monohydrate: a reformulation

Lesław Sieroń

S1. Comment

The crystal structure of the salt-like title compound, (I), C₆H₆N₃⁺·[CoCl₃(C₆H₅N₃)]⁻·H₂O, was originally modelled as a [CoCl₃(C₆H₅N₃)]·C₆H₅N₃·H₂O complex (Zhang *et al.*, 2004). A few facts point that the structure was incorrectly described as Co^{III} instead of Co^{II} complex. The blue colour of the crystal and a tetrahedral coordination geometry around the central metal atom is typical for the Co^{II}Cl₃N chromophore, *e.g.* in pyridinium trichloro-pyridine-cobalt(II) (Hahn *et al.*, 1997), quinolinium trichloro-quinoline-cobalt(II) (Pan & Xu, 2004) and 2-aminopyridinium (2-aminopyridine)-trichloro-cobalt(II) (Zhang *et al.*, 2006).

A decrease of the metal oxidation number in (I) from III to II requires the presence of an additional proton for charge balance. In fact, an analysis of difference Fourier maps has shown two peaks in the positions expected for missing benzotriazole H atoms (N–Q = 0.85 and 0.89 Å). Both of these are involved in short, strong N–H \cdots O_{water} hydrogen bonds which confirms the supposition (Fig. 1). The N–N bond lengths in the benzimidazolium cations are practically equalized and range from 1.309 (4) to 1.316 (5) Å, indicating full delocalization of their π electrons, which is not observed in neutral benzotriazole molecules (Krawczyk & Gdaniec, 2005). The metal bond distances and hydrogen bonds are listed in Tables 1 and 2, respectively.

S2. Experimental

Due to the unsuccessful attempts at the synthesis of (I), the deposited structure factors and other data have been taken from the original paper of Zhang *et al.* (2004).

S3. Refinement

All H atoms were located in difference Fourier syntheses but were repositioned with ideal geometry (C–H = 0.93, N–H = 0.86 and O–H = 0.82 Å) and refined as riding with $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C}, \text{N})$ or $1.5U_{\text{eq}}(\text{O})$.

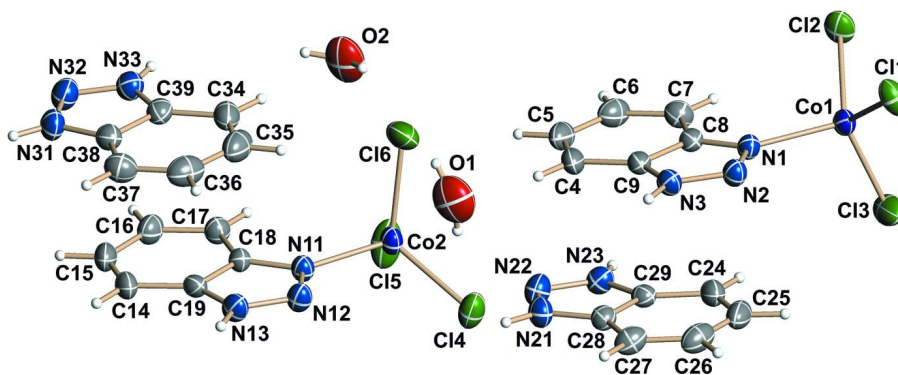


Figure 1

A view of (I), showing displacement ellipsoids for the non-H atoms drawn at the 30% probability level.

1H-Benzotriazol-3-ium (1H-benzotriazole- κ N³)trichloridocobalt(II) monohydrate

Crystal data

(C₆H₆N₃)[CoCl₃(C₆H₅N₃)]·H₂O

$M_r = 422.56$

Triclinic, *P*1

Hall symbol: P 1

$a = 6.8760$ (8) Å

$b = 7.8033$ (12) Å

$c = 16.657$ (3) Å

$\alpha = 95.450$ (13)°

$\beta = 95.962$ (8)°

$\gamma = 103.989$ (9)°

$V = 855.8$ (2) Å³

$Z = 2$

$F(000) = 426$

$D_x = 1.640$ Mg m⁻³

Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å

Cell parameters from 28 reflections

$\theta = 5.3$ – 12.6 °

$\mu = 1.48$ mm⁻¹

$T = 293$ K

Prism, blue

$0.40 \times 0.40 \times 0.20$ mm

Data collection

Bruker P4

diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

ω scans

Absorption correction: ψ scan

(North *et al.*, 1968)

$T_{\min} = 0.611$, $T_{\max} = 0.744$

4906 measured reflections

4846 independent reflections

4526 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.078$

$\theta_{\max} = 27.5$ °, $\theta_{\min} = 2.5$ °

$h = -8 \rightarrow 1$

$k = -9 \rightarrow 10$

$l = -21 \rightarrow 21$

3 standard reflections every 100 reflections

intensity decay: none

Refinement

Refinement on F^2

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.029$

$wR(F^2) = 0.078$

$S = 1.08$

4846 reflections

417 parameters

3 restraints

Primary atom site location: structure-invariant

direct methods

Secondary atom site location: structure-invariant direct methods

Hydrogen site location: difference Fourier map

H-atom parameters constrained

$w = 1/[\sigma^2(F_o^2) + (0.0331P)^2 + 0.2297P]$

where $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{\max} < 0.001$

$\Delta\rho_{\max} = 0.34$ e Å⁻³

$\Delta\rho_{\min} = -0.26$ e Å⁻³

Absolute structure: Flack (1983), 920 Friedel pairs

Absolute structure parameter: -0.008 (13)

Special details

Geometry. Bond distances, angles *etc.* have been calculated using the rounded fractional coordinates. All su's are estimated from the variances of the (full) variance-covariance matrix. The cell e.s.d.'s are taken into account in the estimation of distances, angles and torsion angles

Refinement. Refinement on F^2 for ALL reflections except those flagged by the user for potential systematic errors. Weighted R -factors wR and all goodnesses of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The observed criterion of $F^2 > \sigma(F^2)$ is used only for calculating $-R$ -factor-obs *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
Co1	0.44420 (6)	0.01369 (5)	0.90042 (3)	0.0462 (1)
Cl1	0.1994 (2)	-0.24333 (13)	0.89041 (6)	0.0706 (3)
Cl2	0.3415 (2)	0.22168 (14)	0.97717 (6)	0.0707 (4)
Cl3	0.7529 (2)	-0.0156 (3)	0.93735 (8)	0.0969 (6)
N1	0.4450 (5)	0.0908 (4)	0.78758 (16)	0.0436 (8)
N2	0.3859 (5)	-0.0327 (4)	0.72406 (17)	0.0491 (9)
N3	0.3929 (5)	0.0501 (4)	0.65771 (17)	0.0487 (9)
C4	0.4831 (7)	0.3714 (6)	0.6303 (3)	0.0585 (12)
C5	0.5466 (8)	0.5379 (6)	0.6726 (3)	0.0669 (16)
C6	0.5824 (8)	0.5658 (5)	0.7580 (3)	0.0674 (15)
C7	0.5560 (6)	0.4277 (5)	0.8042 (2)	0.0560 (11)
C8	0.4911 (5)	0.2557 (4)	0.7616 (2)	0.0441 (9)
C9	0.4567 (5)	0.2298 (5)	0.6770 (2)	0.0446 (9)
Co2	0.55479 (7)	0.48473 (5)	0.39999 (3)	0.0484 (1)
Cl4	0.6620 (2)	0.24123 (14)	0.42163 (7)	0.0745 (4)
Cl5	0.7823 (3)	0.72314 (15)	0.46890 (8)	0.0979 (6)
Cl6	0.2402 (2)	0.4745 (2)	0.43067 (8)	0.0891 (5)
N11	0.5282 (5)	0.5107 (4)	0.27999 (16)	0.0445 (8)
N12	0.4780 (6)	0.3689 (4)	0.22487 (17)	0.0514 (9)
N13	0.4662 (6)	0.4279 (4)	0.15251 (17)	0.0542 (9)
C14	0.5165 (7)	0.7309 (6)	0.1018 (2)	0.0570 (11)
C15	0.5618 (8)	0.9044 (6)	0.1336 (3)	0.0630 (16)
C16	0.5992 (7)	0.9600 (5)	0.2186 (3)	0.0606 (13)
C17	0.5933 (6)	0.8414 (5)	0.2737 (2)	0.0512 (10)
C18	0.5479 (5)	0.6616 (4)	0.24206 (18)	0.0409 (9)
C19	0.5082 (6)	0.6077 (5)	0.1583 (2)	0.0453 (10)
N21	0.9897 (6)	0.4698 (4)	0.59051 (19)	0.0554 (10)
N22	1.0593 (6)	0.6201 (4)	0.6387 (2)	0.0608 (11)
N23	1.0793 (5)	0.5710 (4)	0.71148 (18)	0.0531 (9)
C24	1.0184 (7)	0.2821 (6)	0.7735 (2)	0.0581 (11)
C25	0.9546 (7)	0.1017 (6)	0.7485 (3)	0.0670 (16)
C26	0.8995 (7)	0.0348 (6)	0.6655 (3)	0.0695 (16)
C27	0.9026 (7)	0.1394 (5)	0.6052 (3)	0.0605 (14)
C28	0.9658 (6)	0.3237 (5)	0.6305 (2)	0.0462 (10)
C29	1.0242 (5)	0.3911 (5)	0.7118 (2)	0.0439 (10)

N31	0.0523 (6)	0.8803 (5)	0.0749 (2)	0.0609 (13)
N32	0.0951 (7)	1.0044 (5)	0.1374 (2)	0.0673 (14)
N33	0.0884 (6)	0.9186 (5)	0.2018 (2)	0.0581 (11)
C34	0.0247 (7)	0.5974 (7)	0.2283 (3)	0.0671 (16)
C35	-0.0210 (9)	0.4337 (7)	0.1856 (4)	0.0788 (19)
C36	-0.0493 (8)	0.4068 (6)	0.1003 (4)	0.0805 (18)
C37	-0.0293 (8)	0.5435 (6)	0.0545 (3)	0.0688 (14)
C38	0.0182 (7)	0.7128 (5)	0.0974 (2)	0.0531 (11)
C39	0.0438 (6)	0.7406 (5)	0.1818 (2)	0.0514 (11)
O1	0.1150 (10)	0.0529 (7)	0.3557 (2)	0.127 (2)
O2	0.3201 (10)	0.8988 (6)	0.4999 (2)	0.124 (2)
H3	0.36100	-0.00330	0.60880	0.0580*
H4	0.45880	0.35330	0.57380	0.0710*
H5	0.56690	0.63600	0.64400	0.0810*
H6	0.62540	0.68190	0.78410	0.0810*
H7	0.57990	0.44690	0.86070	0.0670*
H13	0.43520	0.35920	0.10710	0.0650*
H14	0.49230	0.69560	0.04600	0.0680*
H15	0.56870	0.99050	0.09830	0.0760*
H16	0.62860	1.08080	0.23720	0.0730*
H17	0.61810	0.87790	0.32940	0.0620*
H21	0.96190	0.46410	0.53860	0.074 (15)*
H23	1.12300	0.64520	0.75490	0.0640*
H24	1.05520	0.32800	0.82800	0.0700*
H25	0.94790	0.02270	0.78710	0.0800*
H26	0.85880	-0.08790	0.65150	0.0840*
H27	0.86550	0.09280	0.55080	0.0730*
H31	0.04630	0.90190	0.02520	0.0730*
H33	0.11010	0.97010	0.25110	0.10 (2)*
H34	0.04250	0.61440	0.28480	0.0810*
H35	-0.03420	0.33520	0.21380	0.0940*
H36	-0.08290	0.29080	0.07420	0.0970*
H37	-0.04650	0.52480	-0.00210	0.0820*
H1A	0.23890	0.07600	0.36260	0.1910*
H1B	0.04880	0.12670	0.36090	0.1910*
H2A	0.30780	0.97330	0.46940	0.1850*
H2B	0.27160	0.79140	0.48810	0.1850*

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Co1	0.0523 (3)	0.0539 (2)	0.0332 (2)	0.0160 (2)	0.0026 (2)	0.0071 (2)
Cl1	0.0959 (8)	0.0581 (5)	0.0467 (4)	-0.0033 (5)	0.0175 (5)	0.0019 (4)
Cl2	0.1015 (9)	0.0690 (6)	0.0423 (4)	0.0284 (6)	0.0096 (5)	-0.0074 (4)
Cl3	0.0682 (7)	0.1819 (15)	0.0594 (6)	0.0585 (9)	0.0058 (5)	0.0423 (7)
N1	0.0503 (16)	0.0458 (13)	0.0372 (12)	0.0171 (13)	0.0039 (12)	0.0076 (10)
N2	0.0596 (18)	0.0496 (14)	0.0389 (13)	0.0173 (14)	0.0029 (13)	0.0050 (11)
N3	0.0563 (18)	0.0541 (15)	0.0365 (12)	0.0166 (15)	0.0054 (13)	0.0034 (11)

C4	0.060 (2)	0.072 (2)	0.055 (2)	0.026 (2)	0.0196 (19)	0.0287 (19)
C5	0.066 (3)	0.061 (2)	0.088 (3)	0.024 (2)	0.030 (2)	0.038 (2)
C6	0.065 (3)	0.0468 (19)	0.093 (3)	0.0146 (19)	0.019 (2)	0.0109 (19)
C7	0.057 (2)	0.0533 (18)	0.0544 (19)	0.0098 (18)	0.0087 (18)	-0.0001 (15)
C8	0.0424 (17)	0.0504 (16)	0.0423 (15)	0.0161 (14)	0.0055 (14)	0.0084 (13)
C9	0.0431 (18)	0.0526 (16)	0.0426 (15)	0.0169 (15)	0.0089 (14)	0.0118 (13)
Co2	0.0625 (3)	0.0495 (2)	0.0336 (2)	0.0159 (2)	0.0033 (2)	0.0060 (2)
Cl4	0.1100 (10)	0.0605 (5)	0.0572 (5)	0.0355 (6)	-0.0038 (6)	0.0083 (4)
Cl5	0.1390 (14)	0.0598 (6)	0.0697 (7)	0.0027 (8)	-0.0386 (8)	-0.0015 (5)
Cl6	0.0837 (9)	0.1346 (11)	0.0667 (6)	0.0474 (9)	0.0314 (6)	0.0256 (7)
N11	0.0564 (17)	0.0426 (13)	0.0350 (11)	0.0145 (13)	0.0044 (12)	0.0038 (10)
N12	0.069 (2)	0.0467 (14)	0.0382 (13)	0.0191 (15)	0.0004 (14)	0.0000 (11)
N13	0.069 (2)	0.0590 (16)	0.0355 (12)	0.0233 (16)	0.0014 (14)	-0.0013 (11)
C14	0.061 (2)	0.077 (2)	0.0418 (17)	0.027 (2)	0.0099 (17)	0.0230 (17)
C15	0.067 (3)	0.073 (3)	0.062 (2)	0.029 (2)	0.017 (2)	0.035 (2)
C16	0.068 (3)	0.0489 (18)	0.069 (2)	0.0182 (19)	0.010 (2)	0.0174 (17)
C17	0.061 (2)	0.0456 (16)	0.0485 (18)	0.0161 (17)	0.0085 (17)	0.0057 (14)
C18	0.0440 (17)	0.0461 (15)	0.0353 (14)	0.0155 (14)	0.0053 (13)	0.0083 (11)
C19	0.048 (2)	0.0568 (18)	0.0362 (14)	0.0211 (16)	0.0094 (14)	0.0070 (13)
N21	0.065 (2)	0.0601 (17)	0.0392 (14)	0.0147 (16)	-0.0007 (15)	0.0087 (12)
N22	0.069 (2)	0.0514 (16)	0.0605 (19)	0.0148 (16)	-0.0018 (18)	0.0115 (14)
N23	0.0579 (19)	0.0524 (15)	0.0470 (15)	0.0168 (15)	-0.0020 (14)	-0.0018 (12)
C24	0.058 (2)	0.078 (2)	0.0472 (18)	0.028 (2)	0.0105 (18)	0.0202 (17)
C25	0.056 (2)	0.073 (3)	0.084 (3)	0.024 (2)	0.019 (2)	0.040 (2)
C26	0.055 (2)	0.0473 (19)	0.105 (4)	0.0122 (19)	0.005 (2)	0.011 (2)
C27	0.057 (2)	0.0491 (19)	0.071 (3)	0.0143 (18)	-0.003 (2)	-0.0047 (17)
C28	0.0429 (18)	0.0501 (17)	0.0452 (16)	0.0122 (15)	0.0032 (14)	0.0064 (13)
C29	0.0408 (18)	0.0527 (17)	0.0400 (15)	0.0162 (15)	0.0043 (14)	0.0044 (13)
N31	0.072 (3)	0.0657 (19)	0.0476 (16)	0.0224 (19)	0.0014 (17)	0.0150 (14)
N32	0.079 (3)	0.063 (2)	0.063 (2)	0.0251 (19)	0.0019 (19)	0.0135 (16)
N33	0.060 (2)	0.0635 (19)	0.0504 (17)	0.0182 (17)	0.0014 (15)	0.0054 (14)
C34	0.053 (2)	0.088 (3)	0.069 (3)	0.023 (2)	0.013 (2)	0.034 (2)
C35	0.069 (3)	0.066 (3)	0.109 (4)	0.019 (2)	0.019 (3)	0.036 (3)
C36	0.068 (3)	0.056 (2)	0.116 (4)	0.013 (2)	0.017 (3)	0.005 (2)
C37	0.069 (3)	0.070 (2)	0.065 (2)	0.016 (2)	0.011 (2)	-0.0019 (19)
C38	0.052 (2)	0.063 (2)	0.0476 (17)	0.0197 (18)	0.0068 (16)	0.0093 (15)
C39	0.045 (2)	0.0578 (18)	0.0526 (19)	0.0151 (17)	0.0035 (16)	0.0111 (15)
O1	0.161 (5)	0.126 (4)	0.074 (2)	0.015 (4)	0.002 (3)	-0.019 (2)
O2	0.187 (6)	0.111 (3)	0.065 (2)	0.033 (4)	0.016 (3)	-0.014 (2)

Geometric parameters (Å, °)

Co1—Cl1	2.2649 (13)	C6—C7	1.371 (6)
Co1—Cl2	2.2616 (13)	C7—C8	1.403 (5)
Co1—Cl3	2.2195 (16)	C8—C9	1.392 (5)
Co1—N1	2.027 (3)	C4—H4	0.93
Co2—Cl4	2.2441 (13)	C5—H5	0.93
Co2—Cl5	2.2450 (15)	C6—H6	0.93

Co2—C16	2.2584 (16)	C7—H7	0.93
Co2—N11	2.022 (3)	C14—C15	1.355 (6)
O1—H1B	0.82	C14—C19	1.404 (5)
O1—H1A	0.82	C15—C16	1.419 (7)
N1—C8	1.371 (4)	C16—C17	1.361 (6)
N1—N2	1.320 (4)	C17—C18	1.399 (5)
N2—N3	1.332 (4)	C18—C19	1.397 (4)
N3—C9	1.360 (5)	C14—H14	0.93
O2—H2B	0.82	C15—H15	0.93
O2—H2A	0.82	C16—H16	0.93
N3—H3	0.86	C17—H17	0.93
N11—N12	1.320 (4)	C24—C25	1.378 (6)
N11—C18	1.373 (4)	C24—C29	1.393 (5)
N12—N13	1.332 (4)	C25—C26	1.409 (7)
N13—C19	1.354 (5)	C26—C27	1.352 (7)
N13—H13	0.86	C27—C28	1.407 (5)
N21—C28	1.359 (5)	C28—C29	1.386 (5)
N21—N22	1.311 (4)	C24—H24	0.93
N22—N23	1.309 (4)	C25—H25	0.93
N23—C29	1.363 (5)	C26—H26	0.93
N21—H21	0.86	C27—H27	0.93
N23—H23	0.86	C34—C39	1.407 (6)
N31—C38	1.366 (5)	C34—C35	1.350 (8)
N31—N32	1.310 (5)	C35—C36	1.403 (9)
N32—N33	1.316 (5)	C36—C37	1.358 (7)
N33—C39	1.349 (5)	C37—C38	1.388 (6)
N31—H31	0.86	C38—C39	1.388 (5)
N33—H33	0.86	C34—H34	0.93
C4—C5	1.365 (7)	C35—H35	0.93
C4—C9	1.398 (6)	C36—H36	0.93
C5—C6	1.405 (7)	C37—H37	0.93
C11—Co1—C12	106.81 (5)	C5—C6—H6	119
C11—Co1—C13	113.73 (7)	C8—C7—H7	122
C11—Co1—N1	106.87 (10)	C6—C7—H7	122
C12—Co1—C13	116.75 (6)	C15—C14—C19	115.9 (3)
C12—Co1—N1	105.84 (10)	C14—C15—C16	122.6 (4)
C13—Co1—N1	106.16 (11)	C15—C16—C17	121.8 (4)
C16—Co2—N11	102.70 (11)	C16—C17—C18	116.4 (3)
C14—Co2—C16	114.49 (6)	C17—C18—C19	121.4 (3)
C14—Co2—N11	110.22 (10)	N11—C18—C19	107.4 (3)
C14—Co2—C15	108.33 (6)	N11—C18—C17	131.2 (3)
C15—Co2—N11	109.56 (10)	N13—C19—C14	134.5 (3)
C15—Co2—C16	111.40 (7)	C14—C19—C18	121.8 (3)
H1A—O1—H1B	125	N13—C19—C18	103.7 (3)
N2—N1—C8	109.5 (3)	C15—C14—H14	122
Co1—N1—N2	118.5 (2)	C19—C14—H14	122
Co1—N1—C8	131.9 (2)	C16—C15—H15	119

N1—N2—N3	107.4 (3)	C14—C15—H15	119
N2—N3—C9	111.5 (3)	C15—C16—H16	119
H2A—O2—H2B	124	C17—C16—H16	119
C9—N3—H3	124	C18—C17—H17	122
N2—N3—H3	124	C16—C17—H17	122
N12—N11—C18	109.6 (3)	C25—C24—C29	115.7 (3)
Co2—N11—N12	120.6 (2)	C24—C25—C26	121.3 (4)
Co2—N11—C18	129.7 (2)	C25—C26—C27	123.6 (4)
N11—N12—N13	106.7 (3)	C26—C27—C28	115.4 (4)
N12—N13—C19	112.7 (3)	N21—C28—C29	104.7 (3)
N12—N13—H13	124	N21—C28—C27	133.7 (3)
C19—N13—H13	124	C27—C28—C29	121.6 (4)
N22—N21—C28	113.5 (3)	C24—C29—C28	122.5 (4)
N21—N22—N23	104.1 (3)	N23—C29—C24	133.3 (3)
N22—N23—C29	113.6 (3)	N23—C29—C28	104.3 (3)
C28—N21—H21	123	C29—C24—H24	122
N22—N21—H21	123	C25—C24—H24	122
N22—N23—H23	123	C24—C25—H25	119
C29—N23—H23	123	C26—C25—H25	119
N32—N31—C38	112.6 (3)	C25—C26—H26	118
N31—N32—N33	105.2 (3)	C27—C26—H26	118
N32—N33—C39	112.3 (3)	C28—C27—H27	122
C38—N31—H31	124	C26—C27—H27	122
N32—N31—H31	124	C35—C34—C39	115.8 (5)
C39—N33—H33	124	C34—C35—C36	122.5 (5)
N32—N33—H33	124	C35—C36—C37	122.5 (5)
C5—C4—C9	116.0 (4)	C36—C37—C38	115.8 (5)
C4—C5—C6	122.1 (4)	N31—C38—C39	104.2 (3)
C5—C6—C7	122.2 (4)	N31—C38—C37	133.7 (4)
C6—C7—C8	116.4 (3)	C37—C38—C39	122.1 (4)
N1—C8—C9	107.2 (3)	C34—C39—C38	121.3 (4)
N1—C8—C7	131.9 (3)	N33—C39—C34	133.0 (4)
C7—C8—C9	120.9 (3)	N33—C39—C38	105.7 (3)
N3—C9—C8	104.5 (3)	C35—C34—H34	122
N3—C9—C4	133.1 (3)	C39—C34—H34	122
C4—C9—C8	122.4 (4)	C36—C35—H35	119
C5—C4—H4	122	C34—C35—H35	119
C9—C4—H4	122	C35—C36—H36	119
C6—C5—H5	119	C37—C36—H36	119
C4—C5—H5	119	C38—C37—H37	122
C7—C6—H6	119	C36—C37—H37	122

Hydrogen-bond geometry (Å, °)

<i>D</i> —H \cdots <i>A</i>	<i>D</i> —H	H \cdots <i>A</i>	<i>D</i> \cdots <i>A</i>	<i>D</i> —H \cdots <i>A</i>
O1—H1B \cdots Cl6	0.82	2.79	3.279 (5)	120
O2—H2A \cdots O1 ⁱ	0.82	2.41	3.142 (7)	148
O2—H2B \cdots Cl6	0.82	2.52	3.295 (5)	159

N3—H3···O2 ⁱⁱ	0.86	1.87	2.721 (4)	173
N13—H13···C12 ⁱⁱⁱ	0.86	2.28	3.127 (3)	169
N21—H21···C14	0.86	2.80	3.458 (4)	134
N21—H21···C16 ^{iv}	0.86	2.75	3.316 (4)	125
N23—H23···C11 ^v	0.86	2.30	3.123 (3)	160
N31—H31···C13 ^{vi}	0.86	2.61	3.214 (4)	128
N33—H33···O1 ⁱ	0.86	1.79	2.646 (5)	170

Symmetry codes: (i) $x, y+1, z$; (ii) $x, y-1, z$; (iii) $x, y, z-1$; (iv) $x+1, y, z$; (v) $x+1, y+1, z$; (vi) $x-1, y+1, z-1$.