

2,2,2-Tris(pyrazol-1-yl)ethanol

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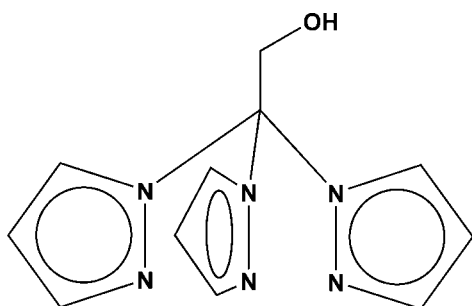
Received 31 March 2011; accepted 6 April 2011

Key indicators: single-crystal X-ray study; $T = 93$ K; mean $\sigma(\text{C}-\text{C}) = 0.002$ Å; R factor = 0.034; wR factor = 0.087; data-to-parameter ratio = 14.0.

The title compound TPE, $\text{C}_{11}\text{H}_{12}\text{N}_6\text{O}$, was prepared by slow evaporation from diethyl ether. In the crystal, there is a hydrogen bond between the alcohol H atom and an N in the pyrazole ring of a neighboring molecule.

Related literature

For the original preparation, see: Reger *et al.* (2000). The title compound was prepared as part of our efforts to study tridentate scorpionate and pseudo-scorpionate ligands for coordination to vanadium, see: McLauchlan *et al.* (2004, 2009); McLauchlan & McDonald (2005, 2006). For coordination complexes with TPE, see: Sánchez-Méndez *et al.* (2004), García-Orozco *et al.* (2006); Silva *et al.* (2009). For applications following substitution of the alcohol, see: Reger, Wright *et al.* (2001); Reger, Semeniuc *et al.* (2001); Reger & Grattan (2003); Pettinari & Pettinari (2005); Silva *et al.* (2009).



Experimental

Crystal data

$\text{C}_{11}\text{H}_{12}\text{N}_6\text{O}$
 $M_r = 244.27$
 Monoclinic, $C2/c$
 $a = 19.6589$ (14) Å
 $b = 11.5155$ (8) Å
 $c = 12.4185$ (18) Å
 $\beta = 125.740$ (1)°

$V = 2281.9$ (4) Å³
 $Z = 8$
 Mo $K\alpha$ radiation
 $\mu = 0.10$ mm⁻¹
 $T = 93$ K
 $0.40 \times 0.39 \times 0.33$ mm

Data collection

Bruker SMART APEX CCD diffractometer
 Absorption correction: multi-scan (SADABS; Bruker, 2008)
 $T_{\min} = 0.699$, $T_{\max} = 0.746$

9505 measured reflections
 2334 independent reflections
 2187 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.018$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.034$
 $wR(F^2) = 0.087$
 $S = 1.05$
 2334 reflections
 167 parameters

H atoms treated by a mixture of independent and constrained refinement
 $\Delta\rho_{\max} = 0.27$ e Å⁻³
 $\Delta\rho_{\min} = -0.25$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$O1-H1O\cdots N5^i$	0.851 (18)	2.003 (18)	2.8494 (13)	172.9 (16)

Symmetry code: (i) $x, -y + 2, z - \frac{1}{2}$.

Data collection: APEX2 (Bruker, 2008); cell refinement: SAINT (Bruker, 2008); data reduction: SAINT; program(s) used to solve structure: SHELXTL (Sheldrick, 2008); program(s) used to refine structure: SHELXTL; molecular graphics: ORTEP-III (Burnett & Johnson, 1996) and Mercury (Macrae *et al.*, 2008); software used to prepare material for publication: SHELXTL (Sheldrick, 2008).

This work was supported by Illinois State University, the American Chemical Society Petroleum Research Fund (46064-B3) and the National Science Foundation (US, CHE-0645081). RSP also acknowledges the support of the American Chemical Society Project SEED. The authors also thank the STaRBURST Cyberdiffraction Consortium and Dr M. Zeller (Youngstown State University) for the data collection.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: DN2673).

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supporting information

Acta Cryst. (2011). E67, o1133–o1134 [https://doi.org/10.1107/S1600536811012931]

2,2,2-Tris(pyrazol-1-yl)ethanol

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S1. Comment

In our efforts to study tridentate scorpionate and psuedo-scorpionate ligands for coordination to vanadium (McLauchlan *et al.*, 2004; McLauchlan & McDonald, 2005; McLauchlan & McDonald, 2006; McLauchlan *et al.*, 2009), we have prepared and crystallized the title compound, tris-2,2,2-(1-pyrazolyl)ethanol, TPE, Figure 1. TPE had been reported by Reger *et al.* (2000) and used as a valuable tool for coordination chemistry owing to the ability to functionalize the alcohol (inter alia Reger, Wright *et al.*, 2001; Reger, Semeniuc *et al.*, 2001; Reger & Grattan, 2003; Pettinari & Pettinari, 2005; Silva *et al.*, 2009). Others have reported coordination complexes with TPE, including Sánchez-Méndez *et al.* (2004), Garcia-Orozco *et al.* (2006), and Silva *et al.* (2009). The structure of TPE shows evidence of hydrogen bonding between the alcoholic H and an N in the pyrazole ring of a neighboring molecule (Table 1) to form chains of TPE parallel to the *c* axis (Figure 2). The structural parameters are similar to those of that seen in the structure of TPE bound to Cu(II) by Silva *et al.* (2009).

S2. Experimental

The title compound was prepared using the method of Reger *et al.* (2000). Use of freshly sublimed KO^tBu was critical for reasonable yield (Silva *et al.*, 2009). Crystals suitable for diffraction study were obtained from slow evaporation of a diethyl ether solution of the product. Characterization data are in line with published data (Reger *et al.*, 2000).

S3. Refinement

The H atoms were geometrically placed (C—H = 0.93–0.97 Å) and refined as riding with $U_{\text{iso}}(\text{H}) = 1.2U_{\text{iso}}(\text{C})$ or $1.5U_{\text{eq}}(\text{methylene C})$. H atoms for the pyrazoles and the methylene unit were modeled in fixed positions whereas the alcoholic H (on O1) was allowed to refine freely.

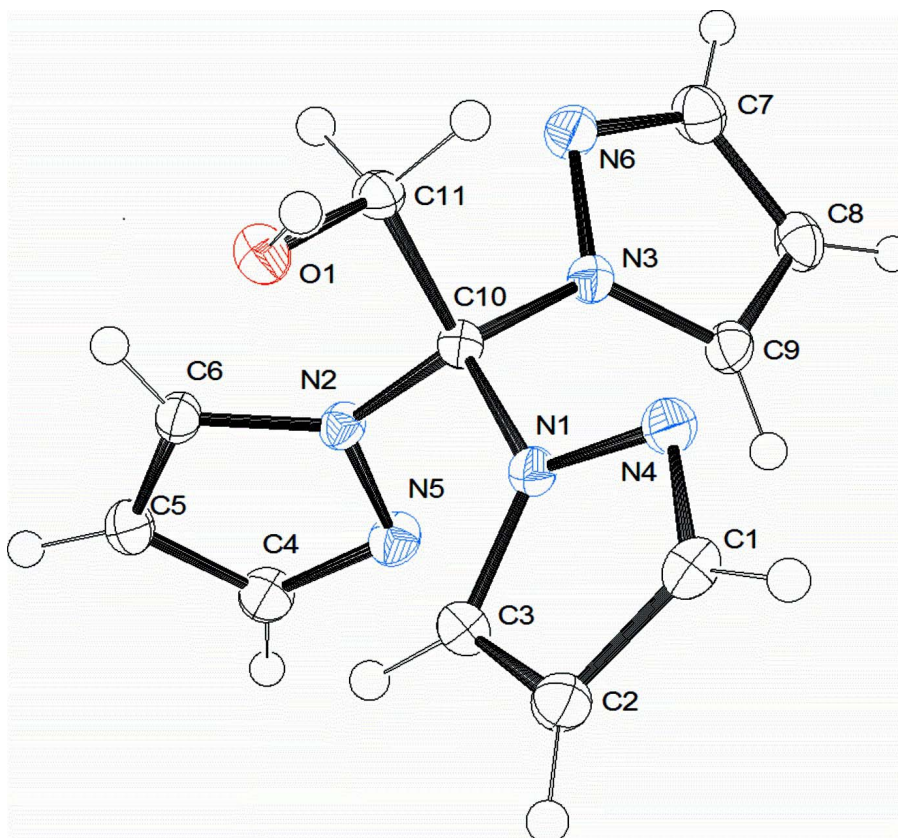


Figure 1

View of the molecular structure of the title compound showing 50% displacement ellipsoids (arbitrary spheres for the H atoms).

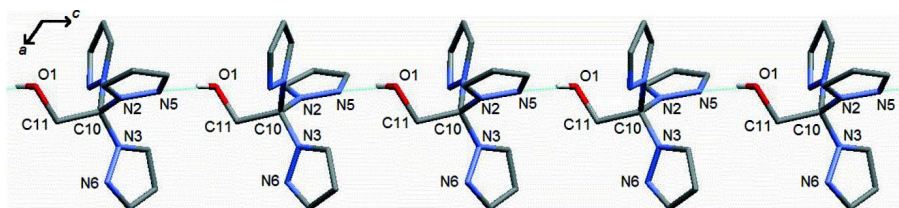


Figure 2

Several symmetry equivalent molecules of the title compound viewed along the *b* axis. Hydrogen bonds between the alcoholic O and the neighboring pyrazole N are indicated in light blue and are parallel to the *c* axis. The neighboring molecule is related by the symmetry code *x*, *-y* + 2, *z* - 1/2. Non-H-bonding hydrogen atoms have been removed for clarity.

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Crystal data

$C_{11}H_{12}N_6O$
 $M_r = 244.27$
 Monoclinic, *C*2/*c*
 Hall symbol: -C 2yc
 $a = 19.6589(14) \text{ \AA}$

$b = 11.5155(8) \text{ \AA}$
 $c = 12.4185(18) \text{ \AA}$
 $\beta = 125.740(1)^\circ$
 $V = 2281.9(4) \text{ \AA}^3$
 $Z = 8$

$F(000) = 1024$
 $D_x = 1.422 \text{ Mg m}^{-3}$
 Melting point: 387(2) K
 Mo $K\alpha$ radiation, $\lambda = 0.71073 \text{ \AA}$
 Cell parameters from 6831 reflections

$\theta = 2.2\text{--}31.4^\circ$
 $\mu = 0.10 \text{ mm}^{-1}$
 $T = 93 \text{ K}$
 Block, colorless
 $0.40 \times 0.39 \times 0.33 \text{ mm}$

Data collection

Bruker SMART APEX CCD
 diffractometer
 Radiation source: fine-focus sealed tube
 Graphite monochromator
 ω scans
 Absorption correction: multi-scan
 (SADABS; Bruker, 2008)
 $T_{\min} = 0.699$, $T_{\max} = 0.746$

9505 measured reflections
 2334 independent reflections
 2187 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.018$
 $\theta_{\max} = 26.4^\circ$, $\theta_{\min} = 2.2^\circ$
 $h = -24 \rightarrow 23$
 $k = -14 \rightarrow 14$
 $l = -15 \rightarrow 15$

Refinement

Refinement on F^2
 Least-squares matrix: full
 $R[F^2 > 2\sigma(F^2)] = 0.034$
 $wR(F^2) = 0.087$
 $S = 1.05$
 2334 reflections
 167 parameters
 0 restraints
 Primary atom site location: structure-invariant
 direct methods

Secondary atom site location: difference Fourier
 map
 Hydrogen site location: inferred from
 neighbouring sites
 H atoms treated by a mixture of independent
 and constrained refinement
 $w = 1/[\sigma^2(F_o^2) + (0.0454P)^2 + 1.8611P]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\max} = 0.001$
 $\Delta\rho_{\max} = 0.27 \text{ e \AA}^{-3}$
 $\Delta\rho_{\min} = -0.25 \text{ e \AA}^{-3}$

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) etc. and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger. H atoms for the pyrazoles and the methylene unit were modeled in fixed positions whereas the alcoholic H (on O1) was allowed to refine freely.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ($\text{\AA}^2$)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
O1	0.16215 (5)	0.93915 (7)	0.05201 (8)	0.01904 (19)
N1	0.14888 (6)	1.06979 (8)	0.22818 (9)	0.0163 (2)
N2	0.18638 (6)	0.87863 (8)	0.30788 (9)	0.0161 (2)
N3	0.29052 (6)	1.02830 (8)	0.40204 (9)	0.0169 (2)
N4	0.16341 (6)	1.17862 (8)	0.20317 (10)	0.0198 (2)
N5	0.17186 (6)	0.88180 (8)	0.40312 (9)	0.0207 (2)
N6	0.36644 (6)	0.98761 (9)	0.43963 (10)	0.0225 (2)
C1	0.08823 (7)	1.22859 (10)	0.13521 (11)	0.0196 (2)
H1A	0.0782	1.3066	0.1046	0.024*

C2	0.02568 (7)	1.15312 (10)	0.11394 (11)	0.0214 (2)
H2A	−0.0324	1.1691	0.0684	0.026*
C3	0.06644 (7)	1.05131 (10)	0.17310 (11)	0.0198 (2)
H3A	0.0418	0.9812	0.1753	0.024*
C4	0.13735 (8)	0.77969 (10)	0.39302 (12)	0.0231 (3)
H4A	0.1210	0.7557	0.4480	0.028*
C5	0.12786 (7)	0.71147 (10)	0.29158 (12)	0.0221 (2)
H5A	0.1045	0.6358	0.2652	0.027*
C6	0.15956 (7)	0.77737 (9)	0.23876 (11)	0.0180 (2)
H6A	0.1623	0.7563	0.1674	0.022*
C7	0.42096 (8)	1.04416 (11)	0.55168 (12)	0.0242 (3)
H7A	0.4800	1.0341	0.6022	0.029*
C8	0.38124 (8)	1.12033 (10)	0.58653 (11)	0.0231 (3)
H8A	0.4068	1.1699	0.6617	0.028*
C9	0.29742 (7)	1.10762 (10)	0.48837 (11)	0.0199 (2)
H9A	0.2527	1.1469	0.4821	0.024*
C10	0.21515 (7)	0.98322 (9)	0.28018 (10)	0.0156 (2)
C11	0.23604 (7)	0.95939 (10)	0.17984 (10)	0.0168 (2)
H11A	0.2662	1.0269	0.1771	0.020*
H11B	0.2732	0.8908	0.2089	0.020*
H1O	0.1605 (10)	0.9921 (15)	0.0027 (17)	0.033 (4)*

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
O1	0.0198 (4)	0.0210 (4)	0.0151 (4)	−0.0026 (3)	0.0094 (3)	0.0001 (3)
N1	0.0177 (5)	0.0141 (4)	0.0187 (5)	−0.0005 (3)	0.0116 (4)	−0.0007 (3)
N2	0.0192 (5)	0.0152 (4)	0.0161 (4)	0.0000 (3)	0.0115 (4)	−0.0003 (3)
N3	0.0167 (5)	0.0166 (4)	0.0171 (5)	−0.0006 (3)	0.0096 (4)	−0.0015 (4)
N4	0.0228 (5)	0.0154 (5)	0.0249 (5)	0.0002 (4)	0.0160 (4)	0.0017 (4)
N5	0.0278 (5)	0.0207 (5)	0.0192 (5)	−0.0005 (4)	0.0169 (4)	0.0002 (4)
N6	0.0168 (5)	0.0238 (5)	0.0220 (5)	0.0014 (4)	0.0087 (4)	−0.0031 (4)
C1	0.0235 (6)	0.0177 (5)	0.0190 (5)	0.0032 (4)	0.0132 (5)	0.0008 (4)
C2	0.0176 (5)	0.0227 (6)	0.0208 (6)	0.0010 (4)	0.0095 (5)	−0.0018 (4)
C3	0.0175 (5)	0.0199 (5)	0.0215 (5)	−0.0029 (4)	0.0111 (5)	−0.0029 (4)
C4	0.0292 (6)	0.0211 (6)	0.0255 (6)	−0.0015 (5)	0.0196 (5)	0.0024 (4)
C5	0.0246 (6)	0.0168 (5)	0.0266 (6)	−0.0023 (4)	0.0159 (5)	−0.0010 (4)
C6	0.0189 (5)	0.0157 (5)	0.0187 (5)	0.0000 (4)	0.0106 (5)	−0.0019 (4)
C7	0.0197 (6)	0.0265 (6)	0.0203 (6)	−0.0025 (5)	0.0081 (5)	−0.0023 (5)
C8	0.0267 (6)	0.0234 (6)	0.0184 (6)	−0.0069 (5)	0.0128 (5)	−0.0045 (4)
C9	0.0254 (6)	0.0189 (5)	0.0205 (5)	−0.0032 (4)	0.0163 (5)	−0.0031 (4)
C10	0.0158 (5)	0.0147 (5)	0.0162 (5)	−0.0004 (4)	0.0093 (4)	−0.0010 (4)
C11	0.0166 (5)	0.0188 (5)	0.0161 (5)	−0.0012 (4)	0.0102 (4)	−0.0009 (4)

Geometric parameters (Å, °)

O1—C11	1.4108 (13)	C2—C3	1.3669 (17)
O1—H1O	0.851 (18)	C2—H2A	0.9500

N1—C3	1.3582 (14)	C3—H3A	0.9500
N1—N4	1.3610 (13)	C4—C5	1.4015 (17)
N1—C10	1.4578 (13)	C4—H4A	0.9500
N2—C6	1.3586 (14)	C5—C6	1.3683 (16)
N2—N5	1.3698 (13)	C5—H5A	0.9500
N2—C10	1.4546 (14)	C6—H6A	0.9500
N3—C9	1.3535 (14)	C7—C8	1.4005 (18)
N3—N6	1.3607 (13)	C7—H7A	0.9500
N3—C10	1.4613 (14)	C8—C9	1.3696 (17)
N4—C1	1.3307 (15)	C8—H8A	0.9500
N5—C4	1.3263 (15)	C9—H9A	0.9500
N6—C7	1.3292 (16)	C10—C11	1.5475 (15)
C1—C2	1.3992 (16)	C11—H11A	0.9900
C1—H1A	0.9500	C11—H11B	0.9900
C11—O1—H1O	105.9 (11)	C6—C5—H5A	127.4
C3—N1—N4	112.15 (9)	C4—C5—H5A	127.4
C3—N1—C10	127.64 (9)	N2—C6—C5	106.90 (10)
N4—N1—C10	118.58 (9)	N2—C6—H6A	126.6
C6—N2—N5	111.62 (9)	C5—C6—H6A	126.6
C6—N2—C10	128.49 (9)	N6—C7—C8	112.10 (11)
N5—N2—C10	119.22 (9)	N6—C7—H7A	124.0
C9—N3—N6	112.23 (9)	C8—C7—H7A	124.0
C9—N3—C10	129.19 (9)	C9—C8—C7	104.84 (10)
N6—N3—C10	118.58 (9)	C9—C8—H8A	127.6
C1—N4—N1	103.86 (9)	C7—C8—H8A	127.6
C4—N5—N2	104.35 (9)	N3—C9—C8	106.80 (10)
C7—N6—N3	104.04 (9)	N3—C9—H9A	126.6
N4—C1—C2	112.28 (10)	C8—C9—H9A	126.6
N4—C1—H1A	123.9	N2—C10—N1	107.50 (8)
C2—C1—H1A	123.9	N2—C10—N3	110.38 (9)
C3—C2—C1	104.87 (10)	N1—C10—N3	108.60 (8)
C3—C2—H2A	127.6	N2—C10—C11	111.49 (8)
C1—C2—H2A	127.6	N1—C10—C11	110.70 (9)
N1—C3—C2	106.78 (10)	N3—C10—C11	108.15 (9)
N1—C3—H3A	126.6	O1—C11—C10	110.62 (9)
C2—C3—H3A	126.6	O1—C11—H11A	109.5
N5—C4—C5	111.97 (10)	C10—C11—H11A	109.5
N5—C4—H4A	124.0	O1—C11—H11B	109.5
C5—C4—H4A	124.0	C10—C11—H11B	109.5
C6—C5—C4	105.14 (10)	H11A—C11—H11B	108.1
C3—N1—N4—C1	2.18 (12)	C6—N2—C10—N1	105.80 (12)
C10—N1—N4—C1	168.76 (9)	N5—N2—C10—N1	−63.96 (12)
C6—N2—N5—C4	1.32 (12)	C6—N2—C10—N3	−135.91 (11)
C10—N2—N5—C4	172.70 (9)	N5—N2—C10—N3	54.33 (12)
C9—N3—N6—C7	−0.22 (13)	C6—N2—C10—C11	−15.69 (15)
C10—N3—N6—C7	−179.92 (10)	N5—N2—C10—C11	174.56 (9)

N1—N4—C1—C2	−1.23 (12)	C3—N1—C10—N2	−20.24 (14)
N4—C1—C2—C3	−0.10 (13)	N4—N1—C10—N2	175.52 (9)
N4—N1—C3—C2	−2.31 (13)	C3—N1—C10—N3	−139.67 (11)
C10—N1—C3—C2	−167.39 (10)	N4—N1—C10—N3	56.09 (12)
C1—C2—C3—N1	1.41 (12)	C3—N1—C10—C11	101.74 (12)
N2—N5—C4—C5	−1.10 (13)	N4—N1—C10—C11	−62.50 (12)
N5—C4—C5—C6	0.51 (14)	C9—N3—C10—N2	−91.92 (13)
N5—N2—C6—C5	−1.04 (12)	N6—N3—C10—N2	87.72 (11)
C10—N2—C6—C5	−171.42 (10)	C9—N3—C10—N1	25.69 (15)
C4—C5—C6—N2	0.32 (13)	N6—N3—C10—N1	−154.67 (9)
N3—N6—C7—C8	0.03 (14)	C9—N3—C10—C11	145.87 (11)
N6—C7—C8—C9	0.15 (14)	N6—N3—C10—C11	−34.49 (13)
N6—N3—C9—C8	0.32 (13)	N2—C10—C11—O1	70.77 (11)
C10—N3—C9—C8	179.98 (10)	N1—C10—C11—O1	−48.84 (12)
C7—C8—C9—N3	−0.28 (13)	N3—C10—C11—O1	−167.70 (8)

Hydrogen-bond geometry (Å, °)

<i>D</i> —H $\cdots$ <i>A</i>	<i>D</i> —H	H $\cdots$ <i>A</i>	<i>D</i> $\cdots$ <i>A</i>	<i>D</i> —H $\cdots$ <i>A</i>
O1—H1O $\cdots$ N5 ⁱ	0.851 (18)	2.003 (18)	2.8494 (13)	172.9 (16)

Symmetry code: (i) $x, -y+2, z-1/2$.