

## Bis(methanol- $\kappa$ O)bis(quinoline-2-carboxylato- $\kappa^2$ N,O)nickel(II)

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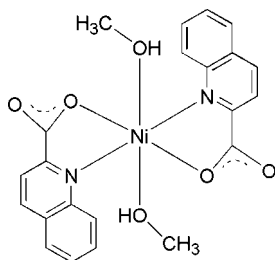
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Key indicators: single-crystal X-ray study;  $T = 293$  K; mean  $\sigma(\text{C}-\text{C}) = 0.003$  Å;  $R$  factor = 0.027;  $wR$  factor = 0.076; data-to-parameter ratio = 13.2.

In the title complex,  $[\text{Ni}(\text{C}_{10}\text{H}_6\text{NO}_2)_2(\text{CH}_3\text{OH})_2]$ , the  $\text{Ni}^{\text{II}}$  ion lies on an inversion center and is coordinated by two quinoline-2-carboxylate ligands in the equatorial sites and two axial methanol ligands, forming a distorted octahedral environment. In the crystal, molecules are linked *via*  $\text{O}-\text{H}\cdots\text{O}$  hydrogen bonds into a two-dimensional network parallel to  $(10\bar{1})$ .

### Related literature

For interactions of metal ions with amino acids, see: Daniele *et al.* (2008); Parkin (2004); Tshuva & Lippard (2004); Stoumpos *et al.* (2009). For related structures, see: Lee *et al.* (2008); Park *et al.* (2008); Shin *et al.* (2009); Song *et al.* (2009); Yu *et al.* (2008, 2009, 2010); Kim *et al.* (2011).



### Experimental

#### Crystal data

$[\text{Ni}(\text{C}_{10}\text{H}_6\text{NO}_2)_2(\text{CH}_3\text{O})_2]$	$V = 988.8$ (3) Å <sup>3</sup>
$M_r = 467.11$	$Z = 2$
Monoclinic, $P2_1/n$	Mo $K\alpha$ radiation
$a = 10.411$ (2) Å	$\mu = 1.03$ mm <sup>-1</sup>
$b = 7.3910$ (15) Å	$T = 293$ K
$c = 13.556$ (3) Å	$0.40 \times 0.10 \times 0.10$ mm
$\beta = 108.57$ (3)°	

#### Data collection

Bruker SMART CCD area-detector diffractometer	5292 measured reflections
Absorption correction: multi-scan (SADABS; Bruker, 1997)	1929 independent reflections
$T_{\text{min}} = 0.884$ , $T_{\text{max}} = 0.903$	1666 reflections with $I > 2\sigma(I)$
	$R_{\text{int}} = 0.018$

#### Refinement

$R[F^2 > 2\sigma(F^2)] = 0.027$	H atoms treated by a mixture of independent and constrained refinement
$wR(F^2) = 0.076$	$\Delta\rho_{\text{max}} = 0.22$ e Å <sup>-3</sup>
$S = 1.07$	$\Delta\rho_{\text{min}} = -0.31$ e Å <sup>-3</sup>
1929 reflections	
146 parameters	
1 restraint	

Table 1

Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{O}3-\text{H}3\text{O}\cdots\text{O}2^i$	0.86 (1)	1.81 (1)	2.655 (2)	167 (2)

Symmetry code: (i)  $-x + \frac{3}{2}, y - \frac{1}{2}, -z + \frac{1}{2}$ .

Data collection: SMART (Bruker, 1997); cell refinement: SAINT (Bruker, 1997); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: SHELXTL (Sheldrick, 2008); software used to prepare material for publication: SHELXTL.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: LH5343).

### References

- Bruker (1997). SMART, SAINT and SADABS. Bruker AXS Inc., Madison, Wisconsin, USA.
- Daniele, P. G., Foti, C., Gianguzza, A., Prenesti, E. & Sammartano, S. (2008). *Coord. Chem. Rev.* **252**, 1093–1107.
- Kim, J. H., Kim, C. & Kim, Y. (2011). *Acta Cryst.* **E67**, m3–m4.
- Lee, E. Y., Park, B. K., Kim, C., Kim, S.-J. & Kim, Y. (2008). *Acta Cryst.* **E64**, m286.
- Park, B. K., Jang, K.-H., Kim, P.-G., Kim, C. & Kim, Y. (2008). *Acta Cryst.* **E64**, m1141.
- Parkin, G. (2004). *Chem. Rev.* **104**, 699–767.
- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
- Shin, D. H., Han, S.-H., Kim, P.-G., Kim, C. & Kim, Y. (2009). *Acta Cryst.* **E65**, m658–m659.
- Song, Y. J., Lee, S.-W., Jang, K. H., Kim, C. & Kim, Y. (2009). *Acta Cryst.* **E65**, m1495–m1496.
- Stoumpos, C. C., Gass, I. A., Milios, C. J., Lalioti, N., Terzis, A., Aromi, G., Teat, S. J., Brechin, E. K. & Perlepes, S. P. (2009). *Dalton Trans.* pp. 307–317.
- Tshuva, E. Y. & Lippard, S. J. (2004). *Chem. Rev.* **104**, 987–1012.
- Yu, S. M., Koo, K., Kim, P.-G., Kim, C. & Kim, Y. (2010). *Acta Cryst.* **E66**, m61–m62.
- Yu, S. M., Park, C.-H., Kim, P.-G., Kim, C. & Kim, Y. (2008). *Acta Cryst.* **E64**, m881–m882.
- Yu, S. M., Shin, D. H., Kim, P.-G., Kim, C. & Kim, Y. (2009). *Acta Cryst.* **E65**, m1045–m1046.

## supporting information

*Acta Cryst.* (2011). E67, m1511 [doi:10.1107/S1600536811041134]

**Bis(methanol- $\kappa$ O)bis(quinoline-2-carboxylato- $\kappa^2$ N,O)nickel(II)**

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**S1. Comment**

The interaction of transition metal ions with biologically active molecules such as amino acids, proteins, sugars, and various acids is of great importance in biological systems (Daniele, *et al.*, 2008; Parkin, 2004; Tshuva & Lippard, 2004; Stoumpos, *et al.*, 2009). As models to examine the interaction, we have intensively studied the interaction of transition metal ions with various acids such as benzoic acid, fulvic acids and humic acids and have reported a variety of structures of copper(II), cadmium(II), and zinc(II) benzoates with quinoxaline,6-methylquinoline, 3-methylquinoline, *trans*-1-(2-pyridyl)-2-(4-pyridyl)ethylene, and di-2-pyridyl ketone (Lee, *et al.*, 2008; Yu, *et al.*, 2008; Park, *et al.*, 2008; Shin, *et al.*, 2009; Song, *et al.*, 2009; Yu, *et al.*, 2008,2009,2010; Kim, *et al.*, 2011). In this work, we have employed nickel(II) chloride as a building block and quinaldic acid as a ligand. We report herein the structure of the title complex.

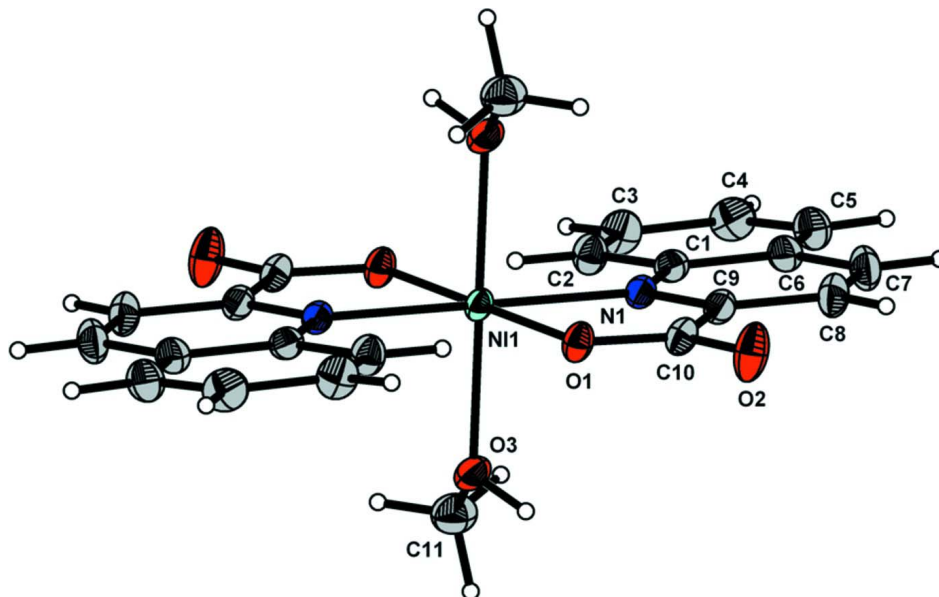
In the crystal structure of the title compound,  $[\text{Ni}(\text{C}_{10}\text{H}_6\text{NO}_2)_2(\text{CH}_3\text{OH})_2]$ , the  $\text{Ni}^{\text{II}}$  ion occupies a crystallographic inversion center. Two quinoline-2-carboxylate ligands coordinate the  $\text{Ni}^{\text{II}}$  ion in the equatorial sites and two methanol ligands coordinate the  $\text{Ni}^{\text{II}}$  ion in axial sites to form a distorted octahedral environment (Fig. 1). In the crystal, molecules are linked via  $\text{O}—\text{H}\cdots\text{O}$  hydrogen bonds to form a two-dimensional network parallel to  $[1\ 0\ -1]$ .

**S2. Experimental**

Quinaldic acid (17.7 mg, 0.1 mmol) and  $\text{NH}_4\text{OH}$  (13.9 ml, 0.1 mmol) were dissolved in 4 ml methanol and carefully layered with 4 ml methanol solution of nickel(II) chloride hexahydrate (11.9 mg, 0.05 mmol). Suitable crystals of the title compound for X-ray analysis were obtained in two weeks.

**S3. Refinement**

H atoms bonded to C atoms were placed in calculated positions with C—H distances of 0.93–0.96 Å. They were included in the refinement in a riding-motion approximation with  $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$  or  $U_{\text{iso}}(\text{H}) = 1.5U_{\text{eq}}(\text{C}_{\text{methyl}})$ . The positions of O—H atoms of the methanol ligands were refined with O—H restraints (0.86 Å) and  $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{O})$ .

**Figure 1**

The molecular structure of the title compound with displacement ellipsoids shown at the 30% probability level. Unlabeled atoms are related by the symmetry operator  $(-x+2, -y, -z+1)$ .

### Bis(methanol- $\kappa$ O)bis(quinoline-2-carboxylato- $\kappa^2$ N,O)nickel(II)

#### Crystal data

$[\text{Ni}(\text{C}_{10}\text{H}_6\text{NO}_2)_2(\text{CH}_4\text{O})_2]$

$M_r = 467.11$

Monoclinic,  $P2_1/n$

Hall symbol:  $-P\ 2_1n$

$a = 10.411\ (2)\ \text{\AA}$

$b = 7.3910\ (15)\ \text{\AA}$

$c = 13.556\ (3)\ \text{\AA}$

$\beta = 108.57\ (3)^\circ$

$V = 988.8\ (3)\ \text{\AA}^3$

$Z = 2$

$F(000) = 484$

$D_x = 1.569\ \text{Mg m}^{-3}$

Mo  $K\alpha$  radiation,  $\lambda = 0.71073\ \text{\AA}$

Cell parameters from 3640 reflections

$\theta = 2.6\text{--}28.1^\circ$

$\mu = 1.03\ \text{mm}^{-1}$

$T = 293\ \text{K}$

Rod, colorless

$0.40 \times 0.10 \times 0.10\ \text{mm}$

#### Data collection

Bruker SMART CCD area-detector  
diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

$\varphi$  and  $\omega$  scans

Absorption correction: multi-scan

(*SADABS*; Bruker, 1997)

$T_{\min} = 0.884$ ,  $T_{\max} = 0.903$

5292 measured reflections

1929 independent reflections

1666 reflections with  $I > 2\sigma(I)$

$R_{\text{int}} = 0.018$

$\theta_{\max} = 26.0^\circ$ ,  $\theta_{\min} = 2.2^\circ$

$h = -12 \rightarrow 12$

$k = -9 \rightarrow 9$

$l = -9 \rightarrow 16$

#### Refinement

Refinement on  $F^2$

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.027$

$wR(F^2) = 0.076$

$S = 1.07$

1929 reflections

146 parameters

1 restraint

Primary atom site location: structure-invariant  
direct methods

Secondary atom site location: difference Fourier map  
 Hydrogen site location: inferred from neighbouring sites  
 H atoms treated by a mixture of independent and constrained refinement

$$w = 1/[\sigma^2(F_o^2) + (0.0451P)^2 + 0.2296P]$$

where  $P = (F_o^2 + 2F_c^2)/3$   
 $(\Delta/\sigma)_{\max} = 0.001$   
 $\Delta\rho_{\max} = 0.22 \text{ e } \text{\AA}^{-3}$   
 $\Delta\rho_{\min} = -0.31 \text{ e } \text{\AA}^{-3}$

*Special details*

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
Ni1	1.0000	0.0000	0.5000	0.02463 (12)
N1	0.79816 (13)	0.10618 (19)	0.48009 (11)	0.0270 (3)
O1	0.94209 (12)	0.07175 (19)	0.35013 (9)	0.0314 (3)
O2	0.78467 (15)	0.2090 (2)	0.22280 (11)	0.0530 (4)
O3	0.91781 (13)	-0.25715 (18)	0.45429 (10)	0.0362 (3)
H3O	0.8599 (16)	-0.259 (3)	0.3928 (7)	0.043*
C1	0.72490 (17)	0.1223 (2)	0.54860 (13)	0.0296 (4)
C2	0.77500 (19)	0.0457 (3)	0.64897 (15)	0.0360 (4)
H2	0.8589	-0.0116	0.6698	0.043*
C3	0.7003 (2)	0.0554 (3)	0.71576 (16)	0.0427 (5)
H3	0.7331	0.0020	0.7811	0.051*
C4	0.5748 (2)	0.1451 (3)	0.68685 (17)	0.0453 (5)
H4	0.5257	0.1514	0.7333	0.054*
C5	0.5250 (2)	0.2221 (3)	0.59184 (17)	0.0429 (5)
H5	0.4425	0.2826	0.5740	0.051*
C6	0.59717 (18)	0.2118 (2)	0.51872 (15)	0.0347 (4)
C7	0.54761 (19)	0.2841 (3)	0.41796 (17)	0.0414 (5)
H7	0.4653	0.3455	0.3971	0.050*
C8	0.61990 (19)	0.2645 (3)	0.35068 (15)	0.0385 (4)
H8	0.5871	0.3101	0.2833	0.046*
C9	0.74535 (17)	0.1737 (2)	0.38513 (14)	0.0300 (4)
C10	0.82890 (17)	0.1504 (3)	0.31228 (14)	0.0315 (4)
C11	0.8677 (2)	-0.3714 (3)	0.51879 (17)	0.0481 (5)
H11A	0.9424	-0.4196	0.5740	0.072*
H11B	0.8167	-0.4690	0.4779	0.072*
H11C	0.8102	-0.3025	0.5478	0.072*

Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
Ni1	0.02223 (18)	0.03058 (19)	0.01837 (18)	0.00163 (12)	0.00265 (12)	-0.00029 (12)
N1	0.0236 (7)	0.0301 (8)	0.0250 (7)	0.0008 (6)	0.0045 (6)	-0.0006 (6)
O1	0.0268 (6)	0.0448 (7)	0.0206 (6)	0.0047 (6)	0.0045 (5)	0.0017 (6)
O2	0.0390 (7)	0.0871 (12)	0.0295 (7)	0.0141 (8)	0.0060 (6)	0.0189 (8)
O3	0.0352 (7)	0.0369 (7)	0.0291 (7)	-0.0050 (6)	0.0000 (5)	-0.0025 (6)
C1	0.0267 (8)	0.0299 (9)	0.0320 (10)	-0.0015 (7)	0.0089 (7)	-0.0048 (7)
C2	0.0307 (9)	0.0458 (11)	0.0317 (10)	0.0052 (8)	0.0105 (8)	-0.0009 (8)
C3	0.0430 (11)	0.0562 (12)	0.0317 (10)	0.0024 (10)	0.0158 (9)	-0.0019 (9)
C4	0.0414 (11)	0.0561 (13)	0.0459 (12)	-0.0013 (10)	0.0246 (10)	-0.0108 (10)
C5	0.0323 (10)	0.0456 (12)	0.0539 (13)	0.0064 (9)	0.0182 (9)	-0.0057 (10)
C6	0.0292 (9)	0.0328 (10)	0.0419 (11)	0.0019 (8)	0.0110 (8)	-0.0035 (8)
C7	0.0288 (9)	0.0418 (11)	0.0511 (12)	0.0112 (8)	0.0094 (9)	0.0063 (9)
C8	0.0304 (9)	0.0429 (11)	0.0366 (10)	0.0061 (8)	0.0027 (8)	0.0114 (9)
C9	0.0249 (8)	0.0317 (9)	0.0302 (9)	-0.0013 (7)	0.0043 (7)	0.0015 (8)
C10	0.0275 (8)	0.0384 (10)	0.0241 (9)	-0.0002 (8)	0.0017 (7)	0.0025 (8)
C11	0.0544 (13)	0.0428 (12)	0.0447 (12)	-0.0066 (10)	0.0124 (10)	0.0000 (10)

Geometric parameters ( $\text{\AA}$ ,  $^\circ$ )

Ni1—O1 <sup>i</sup>	1.9979 (12)	C3—C4	1.405 (3)
Ni1—O1	1.9980 (12)	C3—H3	0.9300
Ni1—O3 <sup>i</sup>	2.0954 (13)	C4—C5	1.351 (3)
Ni1—O3	2.0954 (13)	C4—H4	0.9300
Ni1—N1	2.1779 (14)	C5—C6	1.423 (3)
Ni1—N1 <sup>i</sup>	2.1779 (14)	C5—H5	0.9300
N1—C9	1.326 (2)	C6—C7	1.403 (3)
N1—C1	1.382 (2)	C7—C8	1.363 (3)
O1—C10	1.267 (2)	C7—H7	0.9300
O2—C10	1.231 (2)	C8—C9	1.409 (3)
O3—C11	1.429 (3)	C8—H8	0.9300
O3—H3O	0.859 (2)	C9—C10	1.519 (3)
C1—C2	1.411 (3)	C11—H11A	0.9600
C1—C6	1.424 (2)	C11—H11B	0.9600
C2—C3	1.371 (3)	C11—H11C	0.9600
C2—H2	0.9300		
O1 <sup>i</sup> —Ni1—O1	180.0	C2—C3—H3	119.5
O1 <sup>i</sup> —Ni1—O3 <sup>i</sup>	88.71 (6)	C4—C3—H3	119.5
O1—Ni1—O3 <sup>i</sup>	91.29 (6)	C5—C4—C3	120.31 (19)
O1 <sup>i</sup> —Ni1—O3	91.29 (6)	C5—C4—H4	119.8
O1—Ni1—O3	88.71 (6)	C3—C4—H4	119.8
O3 <sup>i</sup> —Ni1—O3	180.00 (7)	C4—C5—C6	120.93 (18)
O1 <sup>i</sup> —Ni1—N1	100.86 (6)	C4—C5—H5	119.5
O1—Ni1—N1	79.14 (6)	C6—C5—H5	119.5
O3 <sup>i</sup> —Ni1—N1	89.82 (5)	C7—C6—C5	123.07 (17)

O3—Ni1—N1	90.18 (5)	C7—C6—C1	118.33 (17)
O1 <sup>i</sup> —Ni1—N1 <sup>i</sup>	79.14 (6)	C5—C6—C1	118.60 (17)
O1—Ni1—N1 <sup>i</sup>	100.86 (6)	C8—C7—C6	120.06 (17)
O3 <sup>i</sup> —Ni1—N1 <sup>i</sup>	90.18 (5)	C8—C7—H7	120.0
O3—Ni1—N1 <sup>i</sup>	89.82 (5)	C6—C7—H7	120.0
N1—Ni1—N1 <sup>i</sup>	180.0	C7—C8—C9	118.69 (18)
C9—N1—C1	118.24 (14)	C7—C8—H8	120.7
C9—N1—Ni1	109.96 (11)	C9—C8—H8	120.7
C1—N1—Ni1	131.69 (11)	N1—C9—C8	123.71 (17)
C10—O1—Ni1	118.31 (11)	N1—C9—C10	116.25 (15)
C11—O3—Ni1	123.34 (12)	C8—C9—C10	120.04 (16)
C11—O3—H3O	107.6 (15)	O2—C10—O1	124.64 (18)
Ni1—O3—H3O	113.5 (15)	O2—C10—C9	119.26 (16)
N1—C1—C2	119.98 (15)	O1—C10—C9	116.09 (15)
N1—C1—C6	120.95 (16)	O3—C11—H11A	109.5
C2—C1—C6	119.06 (17)	O3—C11—H11B	109.5
C3—C2—C1	120.16 (18)	H11A—C11—H11B	109.5
C3—C2—H2	119.9	O3—C11—H11C	109.5
C1—C2—H2	119.9	H11A—C11—H11C	109.5
C2—C3—C4	120.9 (2)	H11B—C11—H11C	109.5
O1 <sup>i</sup> —Ni1—N1—C9	175.65 (12)	C2—C3—C4—C5	0.6 (3)
O1—Ni1—N1—C9	-4.35 (12)	C3—C4—C5—C6	1.0 (3)
O3 <sup>i</sup> —Ni1—N1—C9	87.00 (12)	C4—C5—C6—C7	177.7 (2)
O3—Ni1—N1—C9	-93.00 (12)	C4—C5—C6—C1	-1.5 (3)
O1 <sup>i</sup> —Ni1—N1—C1	-0.27 (16)	N1—C1—C6—C7	-0.3 (3)
O1—Ni1—N1—C1	179.73 (16)	C2—C1—C6—C7	-178.75 (18)
O3 <sup>i</sup> —Ni1—N1—C1	-88.92 (15)	N1—C1—C6—C5	179.02 (16)
O3—Ni1—N1—C1	91.08 (15)	C2—C1—C6—C5	0.5 (3)
O3 <sup>i</sup> —Ni1—O1—C10	-85.58 (14)	C5—C6—C7—C8	-177.88 (19)
O3—Ni1—O1—C10	94.42 (14)	C1—C6—C7—C8	1.4 (3)
N1—Ni1—O1—C10	3.98 (13)	C6—C7—C8—C9	-1.1 (3)
N1 <sup>i</sup> —Ni1—O1—C10	-176.02 (13)	C1—N1—C9—C8	1.4 (3)
O1 <sup>i</sup> —Ni1—O3—C11	27.49 (15)	Ni1—N1—C9—C8	-175.15 (15)
O1—Ni1—O3—C11	-152.51 (15)	C1—N1—C9—C10	-179.23 (14)
N1—Ni1—O3—C11	-73.38 (15)	Ni1—N1—C9—C10	4.23 (18)
N1 <sup>i</sup> —Ni1—O3—C11	106.62 (15)	C7—C8—C9—N1	-0.3 (3)
C9—N1—C1—C2	177.39 (17)	C7—C8—C9—C10	-179.64 (18)
Ni1—N1—C1—C2	-7.0 (2)	Ni1—O1—C10—O2	176.36 (16)
C9—N1—C1—C6	-1.1 (2)	Ni1—O1—C10—C9	-2.9 (2)
Ni1—N1—C1—C6	174.56 (12)	N1—C9—C10—O2	179.42 (18)
N1—C1—C2—C3	-177.54 (18)	C8—C9—C10—O2	-1.2 (3)
C6—C1—C2—C3	1.0 (3)	N1—C9—C10—O1	-1.3 (2)
C1—C2—C3—C4	-1.5 (3)	C8—C9—C10—O1	178.12 (17)

Symmetry code: (i)  $-x+2, -y, -z+1$ .

*Hydrogen-bond geometry (Å, °)*

<i>D</i> —H $\cdots$ <i>A</i>	<i>D</i> —H	H $\cdots$ <i>A</i>	<i>D</i> $\cdots$ <i>A</i>	<i>D</i> —H $\cdots$ <i>A</i>
O3—H3O $\cdots$ O2 <sup>ii</sup>	0.86 (1)	1.81 (1)	2.655 (2)	167 (2)

Symmetry code: (ii)  $-x+3/2, y-1/2, -z+1/2$ .