

2-[2-(Trifluoromethyl)phenyl]-2H-1-benzopyran-4(3H)-one

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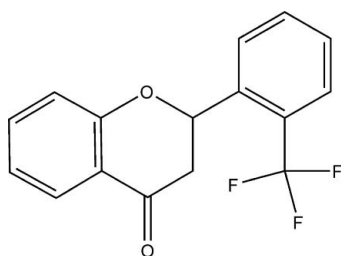
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Key indicators: single-crystal X-ray study; $T = 293$ K; mean $\sigma(\text{C}-\text{C}) = 0.006$ Å; R factor = 0.087; wR factor = 0.109; data-to-parameter ratio = 13.5.

In the title compound, $\text{C}_{16}\text{H}_{11}\text{F}_3\text{O}_2$, the γ -pyranone ring adopts an envelope conformation with the chiral C atom standing out of the ring plane. In the crystal, molecules are linked by $\text{C}-\text{H}\cdots\text{O}$ and $\text{C}-\text{H}\cdots\text{F}$ interactions.

Related literature

For general background to flavones, see: Harborne & Williams (2000). For related flavonoids, see: Benavente-García & Castillo (2008); Rodeiro *et al.* (2006). For related structures, see: Wera *et al.* (2012); Białońska *et al.* (2007); Krishnaiah *et al.* (2005); Wu *et al.* (2005). For van der Waals radii, see: Bondi (1964).



Experimental

Crystal data

$\text{C}_{16}\text{H}_{11}\text{F}_3\text{O}_2$	$V = 1329.2(3) \text{ Å}^3$
$M_r = 292.25$	$Z = 4$
Orthorhombic, $Pna2_1$	Mo $K\alpha$ radiation
$a = 8.2291(9) \text{ Å}$	$\mu = 0.12 \text{ mm}^{-1}$
$b = 22.020(3) \text{ Å}$	$T = 293 \text{ K}$
$c = 7.3355(11) \text{ Å}$	$0.18 \times 0.02 \times 0.02 \text{ mm}$

Data collection

Oxford Diffraction Xcalibur	2558 independent reflections
Sapphire3 Gemini diffractometer	1462 reflections with $I > 2\sigma(I)$
4585 measured reflections	$R_{\text{int}} = 0.049$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.087$	1 restraint
$wR(F^2) = 0.109$	H-atom parameters constrained
$S = 1.12$	$\Delta\rho_{\text{max}} = 0.16 \text{ e Å}^{-3}$
2558 reflections	$\Delta\rho_{\text{min}} = -0.15 \text{ e Å}^{-3}$
190 parameters	

Table 1

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{C3}-\text{H3}\cdots\text{O2}^{\text{i}}$	0.93	2.54	3.311 (5)	140
$\text{C5}-\text{H5}\cdots\text{F3}^{\text{ii}}$	0.93	2.54	3.313 (5)	141

Symmetry codes: (i) $x+1, y, z$; (ii) $-x, -y+2, z+\frac{1}{2}$.

Data collection: *CrysAlis PRO* (Oxford Diffraction, 2010); cell refinement: *CrysAlis PRO*; data reduction: *CrysAlis PRO*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEP-3* (Farrugia, 1997); software used to prepare material for publication: *WinGX* (Farrugia, 1999), *PLATON* (Spek, 2009) and *PARST* (Nardelli, 1995).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: KJ2198).

References

- Benavente-García, O. & Castillo, J. (2008). *J. Agric. Food Chem.* **56**, 6185–6205.
- Białońska, A., Ciunik, Z., Kostrzewa-Susłow, E. & Dmochowska-Gładysz, J. (2007). *Acta Cryst.* **E63**, o432–o433.
- Bondi, A. (1964). *J. Phys. Chem.* **68**, 441–451.
- Farrugia, L. J. (1997). *J. Appl. Cryst.* **30**, 565.
- Farrugia, L. J. (1999). *J. Appl. Cryst.* **32**, 837–838.
- Harborne, J. B. & Williams, C. A. (2000). *Phytochemistry*, **55**, 481–504.
- Krishnaiah, M., Ravi Kumar, R., Jagadeesh Kumar, N., Gunasekar, D. & Jayaprakasam, B. (2005). *Acta Cryst.* **E61**, o3565–o3567.
- Nardelli, M. (1995). *J. Appl. Cryst.* **28**, 659.
- Oxford Diffraction (2010). *CrysAlis PRO*. Oxford Diffraction Ltd., Yarnton, Oxfordshire, England.
- Rodeiro, I., Cancino, L., González, J. E., Morffi, J., Garrido, G., González, R. M., Nuñez, A. & Delgado, R. (2006). *Food Chem. Toxicol.* **44**, 1707–1713.
- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
- Spek, A. L. (2009). *Acta Cryst.* **D65**, 148–155.
- Wera, M., Chalyi, A. G., Roshal, A. D. & Błażejowski, J. (2012). *Acta Cryst.* **E68**, o253–o254.
- Wu, H., Xu, Z., Wan, Y., Liang, Y.-M. & Yu, K.-B. (2005). *Acta Cryst.* **E61**, o1692–o1693.

supporting information

Acta Cryst. (2012). E68, o1522 [https://doi.org/10.1107/S160053681201687X]

2-[2-(Trifluoromethyl)phenyl]-2*H*-1-benzopyran-4(3*H*)-one

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S1. Comment

The title compound belongs to the group of flavanones which occur predominantly in citrus fruits. Citrus flavonoids were reported (Benavente-García & Castillo, 2008) as having antimicrobial, antifungal, antiviral, anti-allergenic, anti-inflammatory (Harborne & Williams, 2000) and antioxidant (Rodeiro *et al.*, 2006) properties. Case control studies suggest that flavonoids may reduce the risk of cardiovascular disease and stroke. Considering the wide spectrum of activities of natural flavonoids further structure modification of these molecules is aimed to enhance their interaction with the target sites in cells.

The title compound adopts a typical conformation of flavanones with the γ -pyranone ring adopting the envelope conformation. In this conformation the carbon atoms C1, C6, C7, C8 and oxygen O1 are nearly coplanar with a root mean square deviation from the mean plane of 0.035 Å, while the C9 carbon atom is standing out from this plane with an atom-to-plane distance of 0.652 (6) Å. All bond lengths and angles in the title compound show usual values for this type of compounds (Wera *et al.*, 2012; Białońska *et al.*, 2007; Krishnaiah *et al.*, 2005; Wu *et al.*, 2005). In this crystal packing of the title compound there are 4 intramolecular, 2 intermolecular (C—H $\cdots$ O and C—H $\cdots$ F) and one π – π interaction. All of the fluoride atoms participate in weak intramolecular C—H $\cdots$ F interactions (Fig. 1) with H $\cdots$ F distances equal or shorter than 2.50 Å, which is shorter than the sum of their Van der Waals radii (Bondi, 1964). Two of the fluoride atoms interact with the hydrogen atom connected to the chiral carbon atom C9 while the third fluoride atom interacts with one of the phenyl hydrogen atoms (Table 1). The dihedral angle between Cg1 and Cg2 rings in previously published structures of flavanones (Wera *et al.*, 2012; Białońska *et al.*, 2007; Krishnaiah *et al.*, 2005; Wu *et al.*, 2005) is in the range from 55 to 75° while the corresponding dihedral angle in the case of the title compound is 66.06 (15)°.

The flavanone molecules are connected into rows by the C3—H3 $\cdots$ O2 intermolecular interaction forming chains down the crystallographic *a* axis (Fig. 2). The second intermolecular interaction, C5—H5 $\cdots$ F3, connects the molecules into another chain in the direction of the screw axis following the crystallographic *c* axis, thus forming a two dimensional net of molecules (Fig. 3). Strings of molecules along the crystallographic *c* axis are further connected by π – π interactions (Fig. 4). Perpendicular distance from the centroid of one Cg1 ring, molecule at (*x*, *y*, *z*), to the plane of the second Cg1 ring, molecule at (1 - *x*, 2 - *y*, 1/2 + *z*), and *vice versa* are 3.70 and 3.62 Å respectively, while the distance between the ring centroids measures 4.101 (3) Å. The Cg1 ring planes of molecules at (*x*, *y*, *z*) and (1 - *x*, 2 - *y*, 1/2 + *z*) are nearly parallel with the dihedral angle being 5.59 (3)°.

S2. Experimental

The proposed compound was prepared in two steps. The first step in the synthetic route consisted of the condensation of 2-hydroxy acetophenone with 2-trifluoromethylbenzaldehyde to give an α,β -unsaturated ketone (chalcone). In the second step the obtained chalcone was dissolved under stirring in a 50:50 water/ethanol mixture. The pH was set to 9.0 with 0.1

M NaOH and the reaction mixture was refluxed for 2 h after which it was cooled over night to room temperature. Small crystals of the title compound formed in the reaction vessel and, after draining excess fluid, the crystals were dried at room temperature.

S3. Refinement

The H atoms bonded to C atoms were placed at geometrically calculated positions and refined using a riding model. C—H distances were fixed to 0.93 Å for aromatic C atoms, 0.97 Å for the secondary CH₂ group and 0.98 Å for the tertiary CH group. Their $U_{\text{iso}}(\text{H})$ values were equal to 1.2 times U_{eq} of the corresponding C atom.

In the absence of significant anomalous scattering, the absolute configuration could not be reliably determined so the Friedel pairs were merged and any references to the Flack parameter were removed.

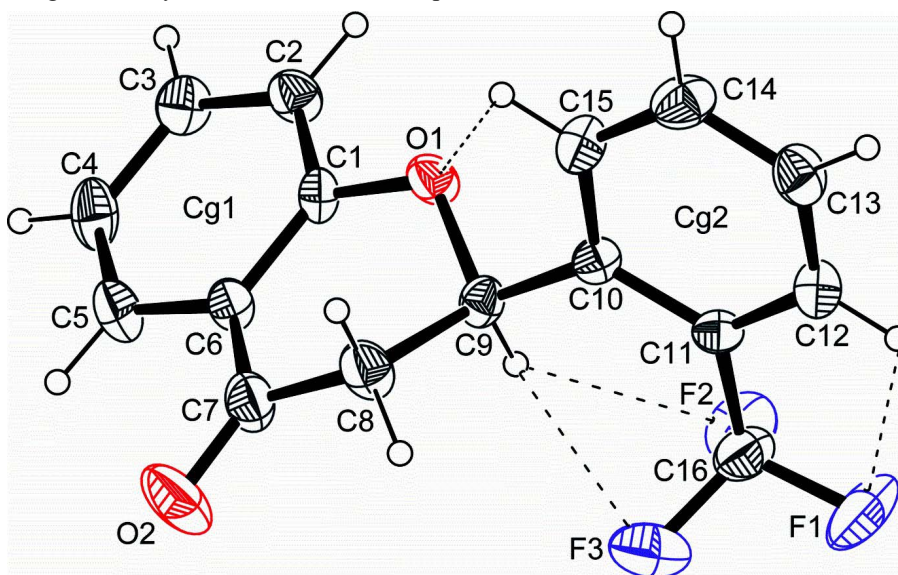


Figure 1

The molecular structure of (I), with atom labels and 30% probability displacement ellipsoids for non-H atoms. H atoms are represented as small spheres of arbitrary radii. Intramolecular interactions are shown as dashed lines.

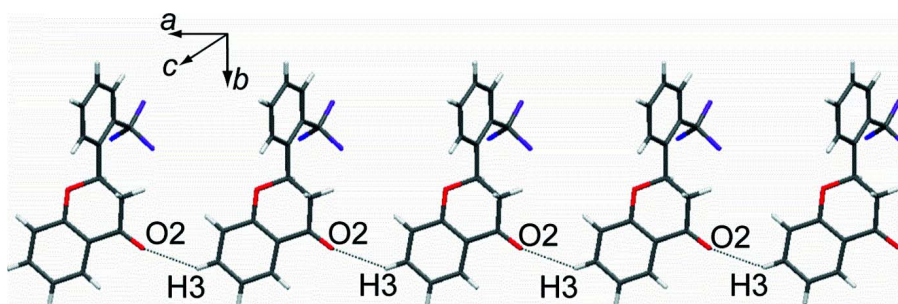


Figure 2

Molecular packing of the title compound and intermolecular interactions along the crystallographic *a* axis.

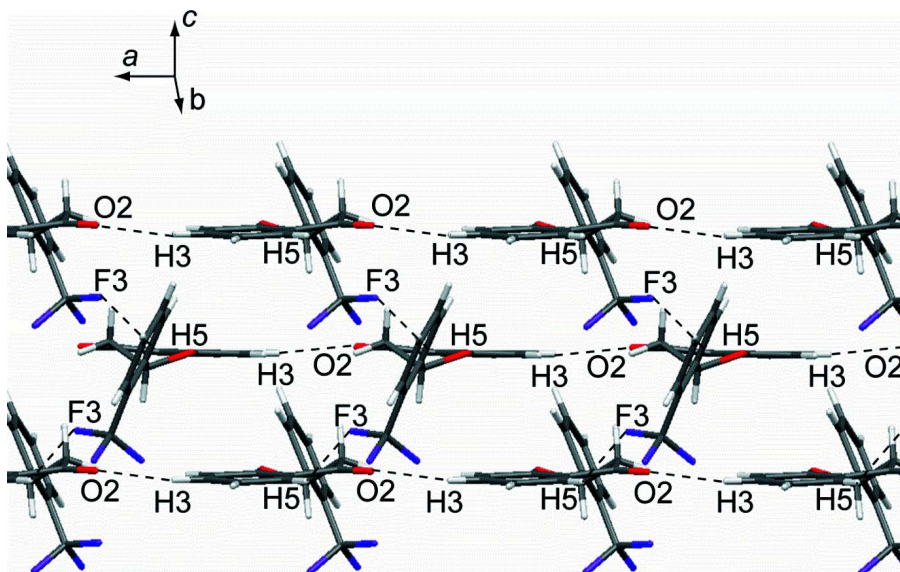


Figure 3

Molecular packing of the title compound and intermolecular interactions along the crystallographic *c* axis.

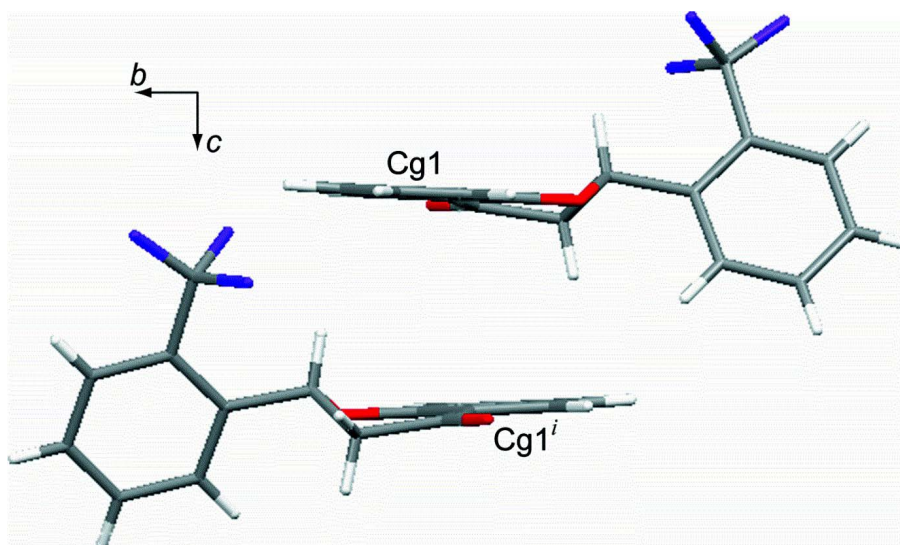


Figure 4

π - π interactions of two neighbouring Cg1 rings. Symmetry code: (i) $1 - x, 2 - y, 1/2 + z$.

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Crystal data

$C_{16}H_{11}F_3O_2$
 $M_r = 292.25$
 Orthorhombic, $Pna2_1$
 Hall symbol: $P\ 2c\ -2n$
 $a = 8.2291\ (9)\ \text{\AA}$
 $b = 22.020\ (3)\ \text{\AA}$
 $c = 7.3355\ (11)\ \text{\AA}$
 $V = 1329.2\ (3)\ \text{\AA}^3$
 $Z = 4$

$F(000) = 600$
 $D_x = 1.460\ \text{Mg m}^{-3}$
 Mo $K\alpha$ radiation, $\lambda = 0.71073\ \text{\AA}$
 Cell parameters from 966 reflections
 $\theta = 3.3\text{--}28.9^\circ$
 $\mu = 0.12\ \text{mm}^{-1}$
 $T = 293\ \text{K}$
 Needle, colourless
 $0.18 \times 0.02 \times 0.02\ \text{mm}$

Data collection

Oxford Diffraction Xcalibur Sapphire3 Gemini
diffractometer
Radiation source: Enhance (Mo) X-ray Source
Graphite monochromator
Detector resolution: 16.3280 pixels mm⁻¹
 ω scans
4585 measured reflections

2558 independent reflections
1462 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.049$
 $\theta_{\text{max}} = 28.9^\circ$, $\theta_{\text{min}} = 3.7^\circ$
 $h = -11 \rightarrow 10$
 $k = -28 \rightarrow 29$
 $l = -9 \rightarrow 9$

Refinement

Refinement on F^2
Least-squares matrix: full
 $R[F^2 > 2\sigma(F^2)] = 0.087$
 $wR(F^2) = 0.109$
 $S = 1.12$
2558 reflections
190 parameters
1 restraint
Primary atom site location: structure-invariant
direct methods

Secondary atom site location: difference Fourier
map
Hydrogen site location: inferred from
neighbouring sites
H-atom parameters constrained
 $w = 1/[\sigma^2(F_o^2) + (0.025P)^2]$
where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\text{max}} < 0.001$
 $\Delta\rho_{\text{max}} = 0.16 \text{ e } \text{\AA}^{-3}$
 $\Delta\rho_{\text{min}} = -0.15 \text{ e } \text{\AA}^{-3}$

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ($\text{\AA}^2$)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
O1	0.3661 (3)	0.90660 (11)	0.7245 (5)	0.0481 (8)
O2	-0.0253 (3)	1.02162 (14)	0.7338 (6)	0.0901 (14)
F1	0.0147 (4)	0.74844 (13)	0.3065 (5)	0.1002 (11)
F2	0.1886 (3)	0.82036 (13)	0.2974 (4)	0.0759 (8)
F3	-0.0440 (3)	0.83559 (15)	0.4155 (5)	0.0867 (10)
C1	0.3889 (4)	0.96841 (17)	0.7160 (7)	0.0394 (10)
C2	0.5488 (4)	0.98855 (19)	0.7098 (7)	0.0475 (12)
H2	0.6339	0.9608	0.7105	0.057*
C3	0.5801 (5)	1.04910 (19)	0.7029 (7)	0.0496 (12)
H3	0.6872	1.0626	0.7023	0.060*
C4	0.4543 (5)	1.09121 (19)	0.6967 (7)	0.0581 (13)
H4	0.4767	1.1325	0.6878	0.070*
C5	0.2973 (5)	1.07110 (18)	0.7040 (7)	0.0557 (13)
H5	0.2132	1.0993	0.7028	0.067*
C6	0.2595 (4)	1.00912 (18)	0.7131 (7)	0.0416 (11)
C7	0.0917 (5)	0.98787 (19)	0.7367 (8)	0.0563 (14)
C8	0.0756 (4)	0.92045 (17)	0.7675 (7)	0.0555 (15)

H8A	−0.0302	0.9070	0.7254	0.067*
H8B	0.0831	0.9119	0.8969	0.067*
C9	0.2082 (4)	0.88567 (18)	0.6662 (6)	0.0393 (11)
H9	0.1964	0.8925	0.5349	0.047*
C10	0.2017 (4)	0.81852 (19)	0.7043 (6)	0.0365 (10)
C11	0.1381 (5)	0.7759 (2)	0.5836 (6)	0.0387 (11)
C12	0.1326 (5)	0.7148 (2)	0.6313 (7)	0.0478 (13)
H12	0.0877	0.6868	0.5509	0.057*
C13	0.1925 (5)	0.6955 (2)	0.7952 (8)	0.0540 (12)
H13	0.1884	0.6545	0.8253	0.065*
C14	0.2581 (5)	0.7360 (2)	0.9145 (7)	0.0583 (14)
H14	0.2998	0.7227	1.0254	0.070*
C15	0.2626 (5)	0.7976 (2)	0.8697 (7)	0.0505 (13)
H15	0.3071	0.8252	0.9518	0.061*
C16	0.0756 (6)	0.7942 (2)	0.4026 (8)	0.0594 (15)

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
O1	0.0319 (14)	0.0360 (17)	0.076 (2)	−0.0008 (13)	−0.0020 (16)	0.001 (2)
O2	0.0468 (16)	0.064 (2)	0.160 (4)	0.0188 (17)	0.013 (2)	0.002 (3)
F1	0.149 (3)	0.079 (2)	0.072 (2)	−0.041 (2)	−0.051 (2)	−0.002 (2)
F2	0.0938 (18)	0.087 (2)	0.0464 (17)	−0.0169 (17)	0.0070 (17)	0.0086 (19)
F3	0.0697 (18)	0.097 (2)	0.093 (2)	0.0126 (19)	−0.0212 (18)	0.022 (2)
C1	0.044 (2)	0.032 (2)	0.042 (3)	−0.002 (2)	0.007 (2)	−0.009 (3)
C2	0.035 (2)	0.046 (3)	0.062 (3)	0.000 (2)	0.007 (3)	0.001 (3)
C3	0.050 (2)	0.045 (3)	0.055 (3)	−0.011 (2)	0.006 (3)	0.005 (3)
C4	0.070 (3)	0.031 (2)	0.073 (4)	−0.011 (3)	0.018 (3)	−0.005 (3)
C5	0.060 (3)	0.036 (3)	0.071 (3)	0.009 (2)	0.010 (3)	−0.001 (3)
C6	0.041 (2)	0.036 (3)	0.049 (3)	0.002 (2)	0.005 (3)	−0.002 (3)
C7	0.047 (2)	0.047 (3)	0.076 (4)	0.002 (2)	0.006 (3)	−0.013 (4)
C8	0.037 (2)	0.045 (3)	0.085 (4)	−0.002 (2)	0.009 (3)	−0.004 (3)
C9	0.040 (2)	0.042 (2)	0.036 (3)	−0.003 (2)	0.005 (2)	−0.007 (2)
C10	0.0344 (19)	0.037 (3)	0.038 (3)	−0.005 (2)	0.001 (2)	0.002 (3)
C11	0.030 (2)	0.044 (3)	0.042 (3)	−0.001 (2)	−0.002 (2)	0.002 (3)
C12	0.056 (3)	0.041 (3)	0.046 (3)	−0.001 (3)	0.004 (2)	−0.005 (3)
C13	0.060 (3)	0.035 (3)	0.067 (3)	−0.004 (3)	0.005 (3)	0.006 (3)
C14	0.066 (3)	0.060 (4)	0.050 (3)	−0.009 (3)	−0.014 (3)	0.015 (3)
C15	0.061 (3)	0.046 (3)	0.045 (3)	−0.008 (2)	−0.001 (3)	−0.006 (3)
C16	0.060 (3)	0.058 (4)	0.060 (4)	−0.011 (3)	−0.016 (3)	0.001 (4)

Geometric parameters (Å, °)

O1—C1	1.376 (4)	C7—C8	1.508 (5)
O1—C9	1.443 (4)	C8—C9	1.526 (5)
O2—C7	1.217 (4)	C8—H8A	0.9700
F1—C16	1.328 (5)	C8—H8B	0.9700
F2—C16	1.339 (5)	C9—C10	1.506 (5)

F3—C16	1.345 (5)	C9—H9	0.9800
C1—C2	1.389 (5)	C10—C15	1.391 (5)
C1—C6	1.392 (5)	C10—C11	1.392 (5)
C2—C3	1.359 (5)	C11—C12	1.390 (6)
C2—H2	0.9300	C11—C16	1.480 (7)
C3—C4	1.390 (5)	C12—C13	1.367 (6)
C3—H3	0.9300	C12—H12	0.9300
C4—C5	1.367 (5)	C13—C14	1.361 (6)
C4—H4	0.9300	C13—H13	0.9300
C5—C6	1.401 (5)	C14—C15	1.397 (6)
C5—H5	0.9300	C14—H14	0.9300
C6—C7	1.468 (5)	C15—H15	0.9300
C1—O1—C9	115.2 (3)	O1—C9—C8	109.8 (3)
O1—C1—C2	116.5 (3)	C10—C9—C8	112.2 (3)
O1—C1—C6	122.3 (3)	O1—C9—H9	109.3
C2—C1—C6	121.2 (3)	C10—C9—H9	109.3
C3—C2—C1	119.6 (4)	C8—C9—H9	109.3
C3—C2—H2	120.2	C15—C10—C11	117.9 (4)
C1—C2—H2	120.2	C15—C10—C9	118.3 (4)
C2—C3—C4	121.0 (4)	C11—C10—C9	123.9 (4)
C2—C3—H3	119.5	C12—C11—C10	120.3 (4)
C4—C3—H3	119.5	C12—C11—C16	118.5 (4)
C5—C4—C3	119.1 (4)	C10—C11—C16	121.2 (4)
C5—C4—H4	120.4	C13—C12—C11	120.7 (5)
C3—C4—H4	120.4	C13—C12—H12	119.7
C4—C5—C6	121.8 (4)	C11—C12—H12	119.7
C4—C5—H5	119.1	C14—C13—C12	120.3 (5)
C6—C5—H5	119.1	C14—C13—H13	119.8
C1—C6—C5	117.3 (3)	C12—C13—H13	119.8
C1—C6—C7	120.8 (4)	C13—C14—C15	119.7 (5)
C5—C6—C7	121.6 (4)	C13—C14—H14	120.1
O2—C7—C6	123.2 (4)	C15—C14—H14	120.1
O2—C7—C8	122.3 (4)	C10—C15—C14	121.1 (4)
C6—C7—C8	114.5 (4)	C10—C15—H15	119.4
C7—C8—C9	111.0 (3)	C14—C15—H15	119.4
C7—C8—H8A	109.4	F1—C16—F2	106.4 (5)
C9—C8—H8A	109.4	F1—C16—F3	106.0 (4)
C7—C8—H8B	109.4	F2—C16—F3	104.9 (4)
C9—C8—H8B	109.4	F1—C16—C11	113.7 (4)
H8A—C8—H8B	108.0	F2—C16—C11	113.2 (4)
O1—C9—C10	106.9 (3)	F3—C16—C11	112.0 (5)
C9—O1—C1—C2	−158.8 (4)	C7—C8—C9—C10	176.3 (4)
C9—O1—C1—C6	21.0 (7)	O1—C9—C10—C15	43.2 (5)
O1—C1—C2—C3	−179.4 (5)	C8—C9—C10—C15	−77.3 (5)
C6—C1—C2—C3	0.7 (8)	O1—C9—C10—C11	−136.9 (4)
C1—C2—C3—C4	−1.8 (8)	C8—C9—C10—C11	102.7 (4)

C2—C3—C4—C5	2.2 (8)	C15—C10—C11—C12	1.7 (6)
C3—C4—C5—C6	−1.6 (8)	C9—C10—C11—C12	−178.2 (4)
O1—C1—C6—C5	−179.9 (4)	C15—C10—C11—C16	−178.1 (4)
C2—C1—C6—C5	−0.1 (7)	C9—C10—C11—C16	2.0 (6)
O1—C1—C6—C7	5.9 (7)	C10—C11—C12—C13	−1.4 (6)
C2—C1—C6—C7	−174.3 (5)	C16—C11—C12—C13	178.4 (4)
C4—C5—C6—C1	0.5 (8)	C11—C12—C13—C14	0.2 (7)
C4—C5—C6—C7	174.7 (5)	C12—C13—C14—C15	0.7 (7)
C1—C6—C7—O2	−179.4 (6)	C11—C10—C15—C14	−0.9 (6)
C5—C6—C7—O2	6.6 (9)	C9—C10—C15—C14	179.1 (4)
C1—C6—C7—C8	1.3 (7)	C13—C14—C15—C10	−0.3 (7)
C5—C6—C7—C8	−172.6 (5)	C12—C11—C16—F1	2.2 (7)
O2—C7—C8—C9	148.5 (5)	C10—C11—C16—F1	−178.0 (4)
C6—C7—C8—C9	−32.3 (6)	C12—C11—C16—F2	−119.4 (4)
C1—O1—C9—C10	−174.2 (4)	C10—C11—C16—F2	60.4 (6)
C1—O1—C9—C8	−52.2 (5)	C12—C11—C16—F3	122.3 (4)
C7—C8—C9—O1	57.6 (5)	C10—C11—C16—F3	−57.9 (6)

Hydrogen-bond geometry (Å, °)

<i>D</i> —H $\cdots$ <i>A</i>	<i>D</i> —H	H $\cdots$ <i>A</i>	<i>D</i> $\cdots$ <i>A</i>	<i>D</i> —H $\cdots$ <i>A</i>
C9—H9 $\cdots$ F2	0.98	2.36	3.068 (5)	129
C9—H9 $\cdots$ F3	0.98	2.50	2.984 (5)	110
C12—H12 $\cdots$ F1	0.93	2.33	2.677 (6)	102
C15—H15 $\cdots$ O1	0.93	2.50	2.760 (5)	96
C3—H3 $\cdots$ O2 ⁱ	0.93	2.54	3.311 (5)	140
C5—H5 $\cdots$ F3 ⁱⁱ	0.93	2.54	3.313 (5)	141

Symmetry codes: (i) $x+1, y, z$; (ii) $-x, -y+2, z+1/2$.